

## Supplement

Unveiling the multiphase fate of 2,4-dinitrophenol on aerosols: Interfacial hydration governs competing oxidation pathways and unexpected toxicity amplification

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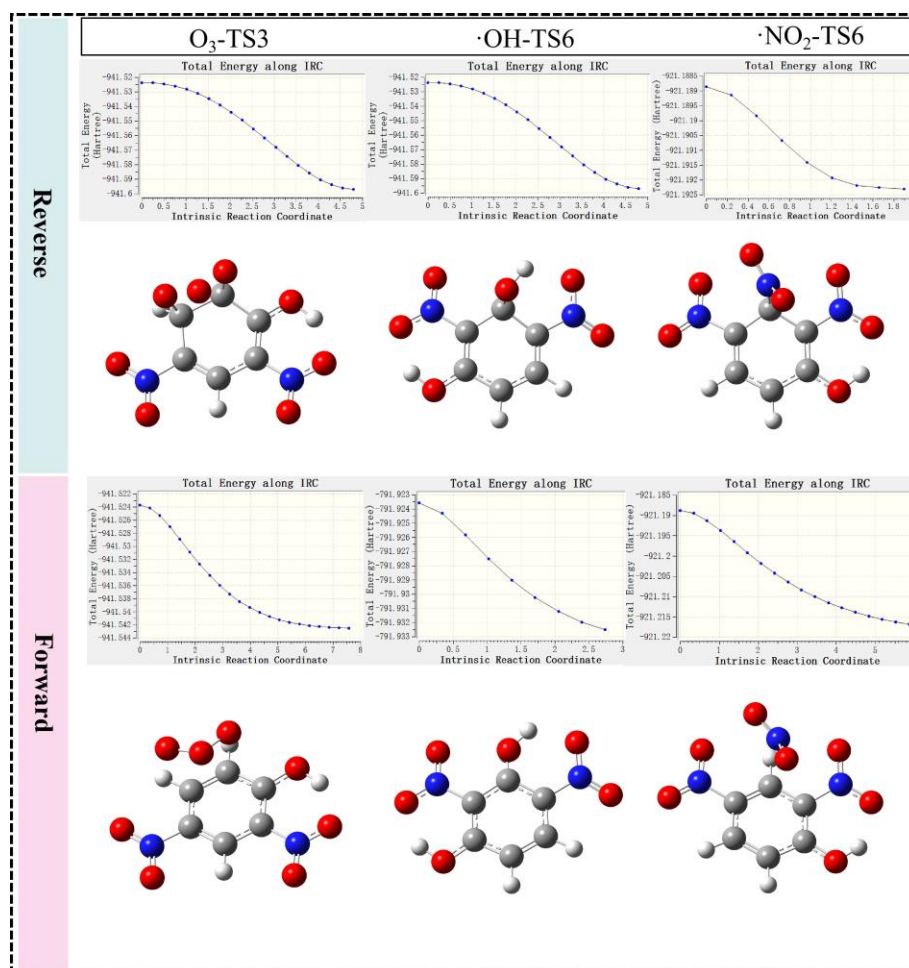
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**Table S1** The reaction energies ( $\Delta_r E$ ) of 2,4-DNP +  $\cdot\text{OH} \rightarrow \text{IM3}$  at various levels of theory. Units of is kcal mol<sup>-1</sup>.

Methods	$\Delta_r E(\text{kcal mol}^{-1})$	Note
CCSD(T)/6-31+g(d,p)	-17.11	
M062x/6-311++g(3df,2p)	-18.28	ZPEs from
M062X/def2-TZVP	-18.73	M06-2X/6-31+g(d,p)
M062X/aug-cc-pVDZ	-19.50	

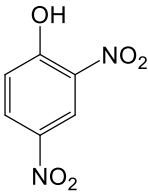
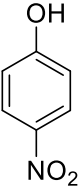
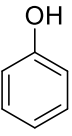
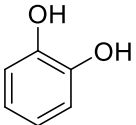


**Figure S1** The IRC analysis of important transition states.

**Table S2** The total rate constants ( $\text{cm}^3 \text{ molecule}^{-1} \text{ s}^{-1}$ ) of 2,4-DNP with  $\text{O}_3$ ,  $\bullet\text{OH}$ ,  $\bullet\text{NO}_2$  in the gas phase and aqueous phase

T(K)	Phase	$k\text{-O}_3$	$k\text{-}\bullet\text{OH}$	$k\text{-}\bullet\text{NO}_2$
200	Gas	$1.78 \times 10^{-26}$	$2.87 \times 10^{-16}$	$5.37 \times 10^{-35}$
225		$2.36 \times 10^{-25}$	$6.13 \times 10^{-16}$	$8.12 \times 10^{-33}$
250		$1.92 \times 10^{-24}$	$1.16 \times 10^{-15}$	$4.65 \times 10^{-31}$
275		$1.09 \times 10^{-23}$	$2.02 \times 10^{-15}$	$1.31 \times 10^{-29}$
300		$4.72 \times 10^{-23}$	$3.30 \times 10^{-15}$	$2.16 \times 10^{-28}$
325		$1.66 \times 10^{-22}$	$5.10 \times 10^{-15}$	$2.36 \times 10^{-27}$
350		$4.97 \times 10^{-22}$	$7.55 \times 10^{-15}$	$1.87 \times 10^{-26}$
375		$1.30 \times 10^{-21}$	$1.08 \times 10^{-14}$	$1.14 \times 10^{-25}$
400		$3.07 \times 10^{-21}$	$1.50 \times 10^{-14}$	$6.99 \times 10^{-38}$
200		Aqueous	$1.38 \times 10^{-21}$	$6.26 \times 10^{-22}$
225	$4.02 \times 10^{-20}$		$2.09 \times 10^{-20}$	$5.76 \times 10^{-32}$
250	$6.05 \times 10^{-19}$		$3.54 \times 10^{-19}$	$8.35 \times 10^{-30}$
275	$5.63 \times 10^{-18}$		$3.65 \times 10^{-18}$	$5.35 \times 10^{-28}$
300	$3.65 \times 10^{-17}$		$2.60 \times 10^{-17}$	$1.82 \times 10^{-26}$
325	$1.79 \times 10^{-16}$		$1.39 \times 10^{-16}$	$3.79 \times 10^{-25}$
350	$7.05 \times 10^{-16}$		$5.90 \times 10^{-16}$	$5.31 \times 10^{-24}$
375	$2.33 \times 10^{-15}$		$2.09 \times 10^{-15}$	$5.38 \times 10^{-23}$
400	$6.70 \times 10^{-15}$		$6.39 \times 10^{-15}$	$6.99 \times 10^{-38}$

**Table S3** The experimental rate constant ( $\text{cm}^3 \text{ molecule}^{-1} \text{ s}^{-1}$ ) values of various phenols with  $\bullet\text{OH}$  in the gas phase.

Compound	Structure formula	Rate constants
2,4DNP		$2.91 \times 10^{-12}$
Nitrophenol		$1.99 \times 10^{-11}$ <small>错误!未定义书签。</small>
Phenol		$2.30 \times 10^{-11}$
Catechol		$1.04 \times 10^{-10}$



**Table S4** The Ames mutagenicity, development toxicity for 2,4-DNP and the intermediates (IMs series) and final products (Ps series).

Compound	Ames muagenicity	Developmental toxicity
2,4DNP	0.13	0.55
IM17	0	0.67
IM25	0	0.85
IM44	0	0.84
IM54	0	0.94
P1	0.53	0.58
P2	0.21	0.56
P3	0.47	0.65
P4	0.21	0.56

Note: Ames mutagenicity negative: Ames mutagenicity  $\leq 0.50$ , Ames mutagenicity positive: Ames mutagenicity  $> 0.50$ . Development non-toxicant: developmental toxicity  $< 0.50$ , development toxicant: development toxicity  $> 0.50$ .