Mechanistic Insights into I_2O_5 Heterogeneous Hydrolysis and Its Role in Iodine Aerosol Growth in Pristine and Polluted Atmospheres

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Abstract. Higher-order iodine oxides are intricately linked to marine aerosol formation; however, the underlying physicochemical mechanisms remain poorly constrained, particularly for I_2O_5, which is stable yet conspicuously absent in the atmosphere. While reactivity with water has been implicated, the direct hydrolysis of I_2O_5 ($I_2O_5 + H_2O \rightarrow 2HIO_3$) fails to account for this discrepancy due to its high activation barrier (21.8 kcal mol^{-1}). Herein, we have probed heterogeneous hydrolysis of I_2O_5 mediated by prevalent chemicals over oceans through Born-Oppenheimer molecular dynamics and well-tempered metadynamics simulations. Our results demonstrate that self-catalyzed pathways involving I_2O_5 and its hydrolysis product HIO_3 substantially reduce the reaction barrier, thereby accelerating the conversion of I_2O_5 to HIO_3 in pristine marine environments. In polluted regions, interfacial hydrolysis of I_2O_5 mediated by acidic or basic pollutants (e.g., H_2SO_4 or amines) proceeds with even greater efficiency, characterized by remarkably low barriers ($\leq$1.3 kcal mol^{-1}). Collectively, these proposed heterogeneous reactions of I_2O_5 are relatively effective, acting as a hitherto unrecognized sink for I_2O_5 and a source of HIO_3—processes that facilitate marine aerosol growth and rationalize the high iodate abundances detected in aerosols. These findings provide mechanistic insight into the elusive I_2O_5-to-HIO_3 conversion, offering an unheeded step toward improving the representation of iodine chemistry and marine aerosol formation in atmospheric models, with implications for climate prediction and environmental impact assessment.

1 Introduction

Given the vast coverage (71%) of oceans on Earth's surface, marine aerosols exert substantial impact on the global atmosphere, altering cloud microphysical properties, radiative balance, and climate change (Kaloshin, 2021; Mahowald et al., 2018). Hence, understanding how aerosols form is fundamental for the climate system, especially ruinous extreme weather, as it introduces the greatest uncertainty in climate forcing (Gettelman and Kahn, 2025). Earlier studies have confirmed a solid association between marine aerosols and sulfur chemicals derived from the oxidation of dimethyl sulfide (DMS) (Barnes et al., 2006; Chen et al., 2018; Li et al., 2024; Ning and Zhang, 2022; Shen et al., 2020; Zhang et al., 2022a, 2024).

While the recent evidence has shown that iodine-bearing precursors of oceanic origin such as iodine oxides (I_2O_{2-5}) and iodine oxyacids (HIO_{2-3}) significantly contribute to marine aerosol formation, owing to their high efficiency and rising levels (Allan et al., 2015; Carpenter et al., 2021; Droste et al., 2021; Gómez Martín et al., 2021, 2022b; Huang et al., 2022; Li et al.,

35 2022; McFiggans et al., 2010; O'Dowd et al., 2002; O'Dowd and De Leeuw, 2007; Roscoe et al., 2015; Saiz-Lopez et al., 2012, 2014). Despite the experimental and theoretical studies uncovering the role of HIO_{2-3} in aerosol formation (He et al., 2021, 2023; Liu et al., 2023; Zhang et al., 2022b, 2024; Zu et al., 2024), the atmospheric fate and impacts of I_2O_{2-5} are yet to be fully established (Finkenzeller et al., 2023; Gómez Martín et al., 2013, 2020; Leroy and Bosland, 2023; Lewis et al., 2020; Liang et al., 2022), limiting the accuracy of atmospheric models in simulating iodine-driven climate effects.

40 As one of the highest iodine oxides, iodine pentoxide (I_2O_5) serves as the anhydride of iodic acid (HIO_3), playing a vital role in bridging iodine oxides and iodic acid in iodine chemical cycle, however, its true fate is still elusive, with the conflicting findings reported. Specifically, I_2O_5 is highly thermally stable (Kaltsoyannis and Plane, 2008), yet it is not evidently present   in the gas phase, implying its unknown sinks. Sipilä et al. (2016) hypothesized that I_2O_5 may participate in the formation of marine aerosol particles through the dehydration reaction (R1: $2HIO_3 \rightarrow I_2O_5 + H_2O$) based on the

45 identified oxygen-to-iodine ratios of 2.4 in field-collected particles. Nevertheless, no direct detection of I_2O_5, such as mass spectral signals, was provided. And the theoretical study has shown that reaction R1 can hardly occur due to its high energy barrier (27.8 kcal mol^{-1}) and endothermic nature (Khanniche et al., 2016), suggesting that this sink pathway seems not well-supported. A most recent experimental evidence indicates that I_2O_5 can indeed participate in aerosol particle formation, but under the unrealistically low-humidity conditions (Rörup et al., 2024). Upon the addition of water, I_2O_5 is markedly depleted,

50 suggesting its reaction with water, accompanied by the formation of HIO_3 (Gómez Martín et al., 2022a). This finding seems to support well the absence of I_2O_5 in high-humidity marine atmospheres. While the theoretical study of Xia et al. (2020) found that the direct reaction between I_2O_5 and H_2O (R2: $I_2O_5 + H_2O \rightarrow 2HIO_3$) in gas phase also needs to cross a high energy barrier of 21.8 kcal mol^{-1}, suggesting that this process is unlikely to occur effectively, even with another H_2O or I_2O_5 acting as a catalyst (Kumar et al., 2018). As a result, this gas-phase mechanism alone appears insufficient to explain the

55 observed absence of I_2O_5 in experimental or clean marine environment, with the I_2O_5 sink remaining unrevealed. Furthermore, Kumar et al. (2018) proposed the heterogeneous reactions of I_2O_5 at the air-water interface, which is ubiquitous over oceans like aerosol surface (Zhong et al., 2018). Unfortunately, this interfacial mechanism still fails to account for the observed well-established link between I_2O_5 depletion and HIO_3 formation, as it points to the formation of other novel product $H_2I_2O_6$. Despite its dynamic stability within the simulation scale (34 ps) (Kumar et al., 2018), $H_2I_2O_6$ has not been

60 detected in the prior iodine-related experimental and field studies (Gómez Martín et al., 2020, 2022a; He et al., 2021, 2023; Rörup et al., 2024; Sipilä et al., 2016), highlighting the poor understanding of the reaction mechanism of I_2O_5 under pristine conditions.

The mystery of I_2O_5 chemistry likely deepens in polluted coastal regions, e.g., Zhejiang, China (Yu et al., 2019), where the evident burst of iodine aerosols has been also observed. Similarly, I_2O_5 is also evidently absent but HIO_3 (detected as IO_3^-)

65 is abundant in the field-collected aerosols, indicating I_2O_5 may undergo an I_2O_5-to-HIO_3 conversion. Yet the underlying

mechanisms in polluted regions remain unclear, which is likely more complex than that in pristine regions due to the pollutant-mediated impacts from typical amine or sulfuric acid that can facilitate the heterogeneous reactions of other iodine oxides (Ning et al., 2023, 2024). Taken together, these unresolved discrepancies and mechanistic uncertainties render the fate of I_2O_5 elusive in both pristine and polluted marine environment, which limits our understanding of its true role in atmospheric iodine chemistry and aerosol formation.

In this study, we have performed the Born-Oppenheimer molecular dynamics (BOMD) simulations to elucidate the potential hydrolysis mechanism of I_2O_5 at the air-water interface, as represented by aqueous aerosol surfaces, mediated by different kinds of prevalent chemicals. We first investigate the heterogenous hydrolysis of I_2O_5 mediated by I_2O_5 and HIO_3, which potentially occur in pristine marine regions; while the reactions of I_2O_5 mediated by pollutants such as acid (sulfuric acid, SA) and alkaline (methylamine (MA), dimethylamine (DMA), trimethylamine (TMA)) were also probed to comprehend the loss path of I_2O_5 in polluted marine environment. Alongside mechanistic insights, we further quantified the reaction barriers by metadynamics simulations (MetaD) and block error analysis. By comparing the resulting energy barriers, the dominant mechanisms of I_2O_5 hydrolysis under different environment were identified. Furthermore, we performed wavefunction analysis to examine the reactive sites of I_2O_5 hydrolysis products at the air-water interface, which may facilitate the condensation of gaseous precursors to promote aerosol growth. This work provides mechanistic insights into heterogeneous hydrolysis of I_2O_5 that potentially occur under marine conditions, enhancing our understanding of iodine chemistry and marine aerosol formation.

2 Methods

2.1 Quantum Chemistry Calculations

Structural optimizations for gas-phase molecular geometries were performed using the Gaussian 16 software package (Frisch et al., 2016) with tight convergence criteria. The I_2O_5 molecule with lowest Gibbs free energy has been selected from isomers (Kaltsoyannis and Plane, 2008; Khanniche et al., 2016; Kim and Yoo, 2016), and details of the structures and coordinates are provided in SI (Figure S1 and Table S3). During the geometry optimization, the M06-2X functional was selected for gas-phase molecules due to its excellent performance in calculating main group elements (Zhao and Truhlar, 2008), and the adopted basis set was aug-cc-pVTZ (Kendall et al., 1992) (for C, H, O, N, and S atoms) + aug-cc-pVTZ-PP with ECP28MDF (for I atom) (Peterson et al., 2003). The frequency analysis was performed at the same level of theory to ensure that the optimized structures have no imaginary frequencies. The geometries and coordinates of gas-phase molecules (i.e. I_2O_5, HIO_3, H_2SO_4, MA, DMA, and TMA) are provided in Figure S2 and Table S2 in the supporting information (SI), respectively.

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2.2 Classical MD Simulations

Classical molecular dynamics (MD) simulations were carried out by the GROMACS 2022 package (Van Der Spoel et al., 2005) to probe the interfacial propensity of I_2O_5 at air-water interface. Regarding the initial configurations for MD simulations, we used the PACKMOL code (Martínez et al., 2009) to build a cubic water slab with 3.3 nm edge, the simulation box, consisting of 1000 water molecules with a tolerance of 2.0 Å (minimum interatomic distance). And the simulation box was expanded to $3.3 \times 3.3 \times 16$ nm^3 along the Z-axis to build the vacuum layers, forming the air-water interface. Umbrella sampling was employed, and a harmonic bias of 5000 kJ mol^{-1} nm^{-2} was applied to I_2O_5 molecule along the Z-axis to investigate its free energy profile crossing the air-water interface from bulk water into gas phase. Furthermore, the generalized AMBER force field (GAFF) (Wang et al., 2004) with parameters obtained from the Sobtop 1.0 code (Lu, 2024) was combined with the TIP4P water model (Jorgensen et al., 1983) to describe the driving forces in the MD simulations. The restrained electrostatic potential (RESP) method (Bayly et al., 1993) was used to fit the atomic charges using the Multiwfn 3.7 code (Lu and Chen, 2012). Prior to MD simulations, the energy minimization of the system was performed by the steepest descent algorithm. The MD simulations were carried out in the constant volume and temperature (*NVT*) ensemble, with the temperature controlled at 300 K by the Nosé−Hoover thermostat (Evans and Holian, 1985). A time step of 2.0 fs was employed for all MD simulations. Lennard−Jones (LJ) and Coulomb potentials were used to simulate nonbonded interactions, with the electrostatic interactions calculated using the particle mesh Ewald (PME) summation method (Darden et al., 1993) and a 10 Å real-space cutoff for nonbonded interactions. The hydrogen-involved bonds were treated via the LINCS algorithm (Hess et al., 1997). In each umbrella sampling window, the system was equilibrated for 5 ns. The free-energy profile was calculated by the weighted histogram analysis method (WHAM) (Kumar et al., 1992).

2.3 BOMD Simulations

To explore the interfacial reaction of I_2O_5, the underlying mechanism was studied by BOMD and the stepwise multi-subphase space metadynamics (SMS-MetaD) (Fang et al., 2022) simulations using the CP2K 2022 program (Kühne et al., 2020) combined with PLUMED software (Anon, 2019). The SMS-MetaD method is well-suited for effectively modeling chemical systems with complex potential energy surface, especially at the air-water interface (Fang et al., 2022), which has already been successfully employed in the studies of heterogeneous reactions (Fang et al., 2024a, b; Tang et al., 2024; Wan et al., 2023). The QUICKSTEP module was applied with the Gaussian and plane wave (GPW) method, and the electronic exchange−correlation term was described by the BLYP-D3 functional (Perdew and Wang, 1992). The DZVP-MOLOPT-SR-GTH basis set and Goedecker−Teter−Hutter (GTH) pseudopotentials (Goedecker et al., 1996; Hartwigsen et al., 1998) were adopted here. The plane-wave cutoff was set to 280 Ry, and that for Gaussian was 40 Ry. All BOMD simulations were performed in the *NVT* ensemble, with the temperature controlled at 300 K by the Nosé−Hoover thermostat (Evans and Holian, 1985), and the time step was set to 1.0 fs. The dimensions of 3-D periodic cell are $x = 18$ Å, $y = 18$ Å, and $z = 30$ Å, in which the water slab containing 128 H_2O was fully relaxed to reach a statistical equilibrium for temperature and potential

energy (see Figure S4 in SI). To probe the reaction barriers of I_2O_5 hydrolysis mediated by different chemicals, a series of SMS-MetaD simulations were carried out. In the SMS-MetaD simulations, collective variable (CV), as the key parameter, was set to effectively differentiate distinct states (Figures S3 and S8-9), with its upper and lower limits referring to the interfacial behavior observed in unbiased simulations (see more details in SI). Gaussian hills with adaptive heights and sigma widths of 0.1 Å were deposited every 50 steps to efficiently explore the free energy landscape and accelerate the convergence (Figures S10-S13). And block average analysis was performed to assess the error associated with the calculated reaction barriers (Figure S14). The wave function analysis and visualization of key structures from the BOMD simulations were carried out using VMD 1.9.3 (Humphrey et al., 1996) and Multiwfn 3.7 software (Lu and Chen, 2012).

3 Results and Discussion

3.1 Surface Preference

The interfacial preference of I_2O_5 influences whether its heterogeneous reactions can effectively occur. Herein, umbrella sampling was applied to obtain the free energy profile of I_2O_5 from the bulk water to the air along the Z direction, where I_2O_5 was initially placed in the center of water layer. As shown in Figure 1, the air-water interface is defined as the region with $10\% - 90\%$ of the bulk density, delineated by the two dashed lines. As I_2O_5 moves from the bulk water into air across the interface, the free energy first decreases and then increases, reaching the lowest values near the Gibbs dividing surface (GDS, the blue dash line). The results show that I_2O_5 has an evident preference for aqueous surface, suggesting its potential to react with other chemicals at the air-water interface.
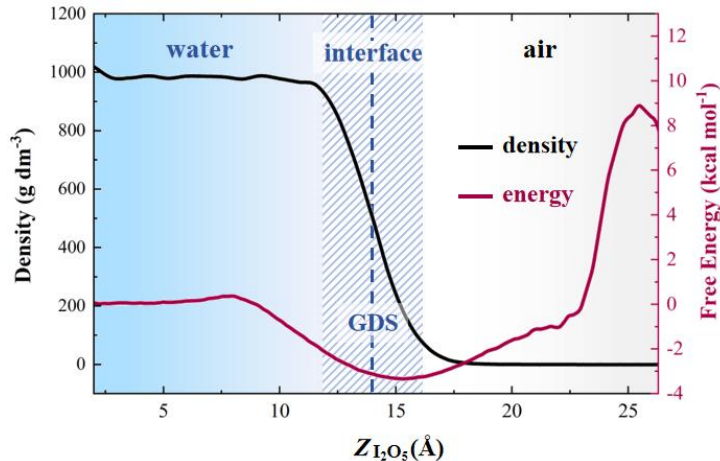

Figure 1. Interfacial propensity of I_2O_5. Mass density (g dm^{-3}) of the water (black line); the free energy of I_2O_5 (purple line) along the Z axis (Å). The vertical dashed line indicates the Gibbs dividing surface (GDS), which corresponds to a location with half of the bulk density. The shaded region represents the air-water interface ($10\% - 90\%$ of bulk density).

3.2 H₂O-Mediated Interfacial Hydrolysis of I₂O₅

150 Experimental evidence (Gómez Martín et al., 2022a) has confirmed a high reactivity of I_2O_5 in the presence of H_2O, which stands in contrast to theoretical predictions based on gas-phase reactions (Khanniche et al., 2016; Xia et al., 2020). To address this gap, we first probed unexplored heterogeneous hydrolysis of I_2O_5 at the air-water interface by the BOMD simulations. The results indicate no apparent tendency for I_2O_5 to react with interfacial H_2O to form HIO_3 in the unbiased BOMD simulation within 50 ps (Figure S7), implying this process may involve an energy barrier. Accordingly, the enhanced

155 sampling (i.e., SMS-MetaD) simulations were further performed to examine the interfacial reactions of I_2O_5 and to quantify the associated reaction barrier (see CV settings in SI). Figure 2(a)-(b) illustrates the process of I_2O_5 hydrolysis, where the reaction initiates with proton (H^+) transfer from H_2O to I_2O_5 facilitated by another H_2O as a 'bridge'. Then I_2O_5 binds with H^+ to form the intermediate $HI_2O_5^+$, with the system existing in a quasi-transition state (TS). Subsequently the I–O bond in the $HI_2O_5^+$ breaks, and the OIO^+ motif quickly combines with the OH^- to form a HIO_3, leaving another HIO_3. Figure 2(c)

160 presents the free energy profile for the interfacial hydrolysis of I_2O_5, showing a reaction barrier of ~9.4 kcal mol^{-1} upon convergence of the results (Figure S10), with ± 0.1 kcal mol^{-1} error as indicated by the block analysis (see more details in SI), which is lower than that of the reported corresponding gas-phase reaction (13.6 kcal mol^{-1}) (Xia et al., 2020). However, this barrier remains considerable, limiting the direct heterogeneous hydrolysis from proceeding rapidly. The previous study (Kumar et al., 2018) reported the formation of $H_2I_2O_6$ from the reaction of I_2O_5 with interfacial water and indicated its

165 dynamic stability over a 35 ps simulation period. However, $H_2I_2O_6$ was unfortunately not observed in the reported iodine-related experimental and field studies (Gómez Martín et al., 2020, 2022a; He et al., 2021, 2023; Rörup et al., 2024; Sipilä et al., 2016), nor was it observed in our unbiased BOMD simulations. Thus, H_2O-mediated heterogeneous mechanism still faces challenges in representing the identified I_2O_5-to-HIO_3 conversion, pointing to the faster pathways mediated by other chemicals responsible for the absence of I_2O_5 in the experimental or field observations.
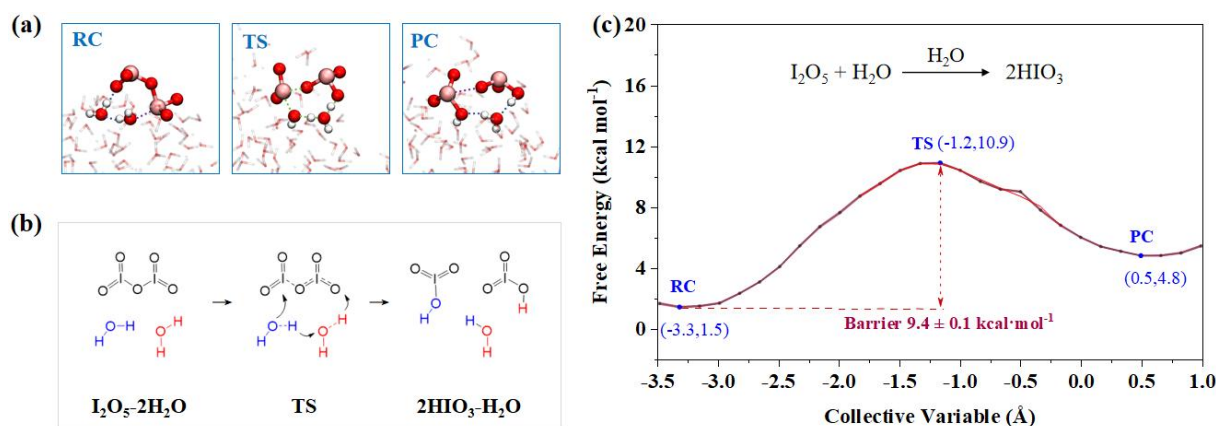
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Figure 2. Details of H_2O-mediated hydrolysis of I_2O_5. (a) Snapshot structures of SMS-MetaD simulation representing reactant complex (RC), transition state (TS), and product complex (PC). The pink, red and white spheres represent I, O, and H atoms, respectively (the same

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applies in Figures 3-6 below). (b) Reaction mechanism for the H_2O-mediated hydrolysis of I_2O_5. Blue molecular indicates the water involved in hydrolysis, black one indicates the I_2O_5, red one indicates the catalyst. The arrows indicate the direction of atom transfer (the same below). (c) Free energy profile for I_2O_5 hydrolysis at the air-water interface mediated by H_2O (black line) with the error band (pink region).

3.3 Self-Mediated Hydrolysis of I_2O_5

Given that H_2O-mediated hydrolysis is inefficient, we next explored whether leftover I_2O_5 can self-catalyze a faster reaction—a critical question for explaining its atmospheric depletion. During the unbiased BOMD simulation, two I_2O_5 molecules combined together to form iodine oxide complex ($I_2O_5\cdots I_2O_5$) via intermolecular halogen bond (XB), without undergoing hydrolysis within 50 ps (Figure S7). Further SMS-MetaD simulations, shown in Figure 3(a)-(b), present the mechanistic details of this hydrolysis. The $I_2O_5\cdots I_2O_5$ complex initially binds a H_2O molecule via hydrogen bond (HB) and XB, forming a three-membered ring. Next, the H_2O molecule dissociates into OH^- and H^+, each approaching a distinct I_2O_5. Meanwhile, the H_2O-involved HB (O–H$\cdots$O) and XB (O–I$\cdots$O) generally converted into new H–O and I–O covalent bonds, resulting in the formation of two HIO_3 molecules. And the residual OIO and IO_3 of the two I_2O_5 molecules combined to form a new I_2O_5, finalizing the self-catalyzed reaction. As shown in Figure 3(c), the resulting reaction barrier is 3.5 kcal mol^{-1}, which is 5.9 kcal mol^{-1} lower than that of H_2O-catalyzed mechanism. The results suggest that I_2O_5 facilitates its own hydrolysis more pronouncedly than H_2O. Thus, the I_2O_5-mediated mechanism efficiently drives I_2O_5-to-HIO_3 conversion, resolving the earlier paradox of I_2O_5's stability yet absence in the atmosphere—and explaining the high HIO_3 levels in marine aerosols.
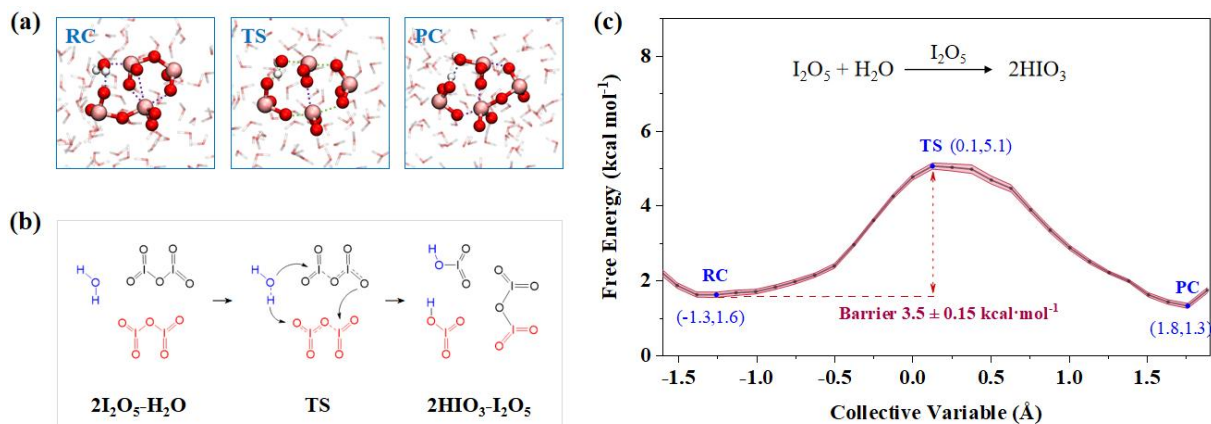

Figure 3. Details of I_2O_5-mediated hydrolysis of I_2O_5. (a) Snapshot structures of SMS-MetaD simulation representing RC, TS, and PC. (b) Reaction mechanism for the I_2O_5-mediated hydrolysis of I_2O_5. (c) Free energy profile for I_2O_5 hydrolysis at the air-water interface mediated by I_2O_5 (black line) with the error band (pink region).

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3.4 HIO$_3$-Mediated Interfacial Hydrolysis of I$_2$O$_5$

We further performed SMS-MetaD simulations to explore the interfacial reaction mediated by the product HIO$_3$, due to its widespread presence in gas and aerosol phase. And HIO$_3$ possesses both proton donor and acceptor sites, potentially facilitating proton transfer, which is a key step in I$_2$O$_5$ hydrolysis. As shown in Figure 4(a)-(b), H$^+$ transfers from H$_2$O to HIO$_3$, then HIO$_3$ donates H$^+$ to I$_2$O$_5$, finally yielding two new HIO$_3$. The resulting energy barrier of this process is ~4.3 kcal mol^{-1} (Figure 4(c)), slightly higher than that mediated by I$_2$O$_5$ (3.5 kcal mol^{-1}). However, given the high concentration and widespread distribution of HIO$_3$ over oceans, its role in mediating hydrolysis of I$_2$O$_5$ should be fully considered. Moreover, as this reaction proceeds, the product HIO$_3$ accumulates to further catalyze the reaction, establishing positive feedback. Thus, in the pristine iodine-rich region such as Mace Head, with HIO$_3$ concentrations as high as 10^8 molec. cm^{-3}, (Sipilä et al., 2016) the HIO$_3$-catalyzed heterogeneous hydrolysis of I$_2$O$_5$ may represent its primary loss pathway. This may help explain why I$_2$O$_5$, though hypothesized to be involved in aerosol formation, is not observed in mass spectra, where iodate signals appear instead.
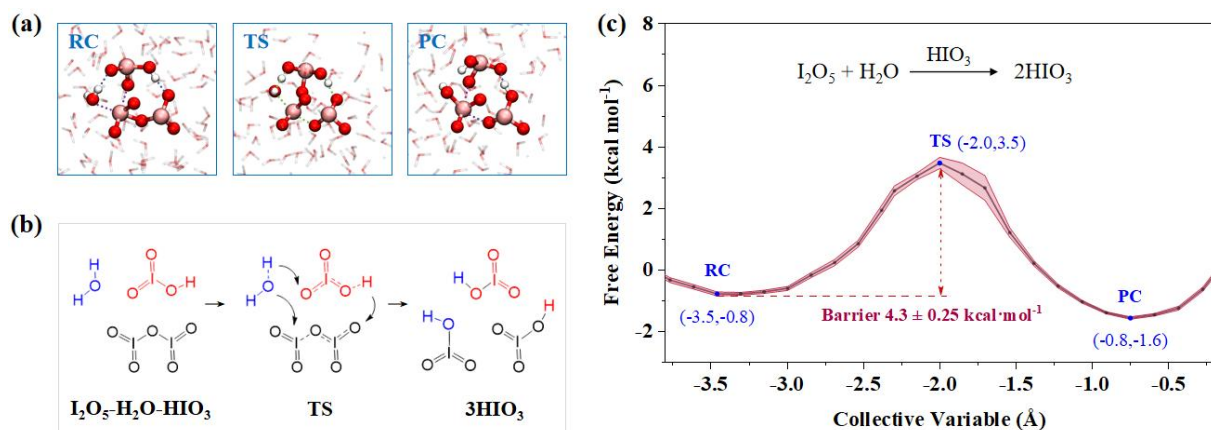

Figure 4. Details of HIO$_3$-mediated hydrolysis of I$_2$O$_5$. (a) Snapshot structures of SMS-MetaD simulation representing RC, TS, and PC. (b) Reaction mechanism for the HIO$_3$-mediated hydrolysis of I$_2$O$_5$. (c) Free energy profile for I$_2$O$_5$ hydrolysis at the air-water interface mediated by HIO$_3$ (black line) with the error band (pink region).

3.5 Sulfuric Acid-Mediated Interfacial Hydrolysis of I$_2$O$_5$

In fact, I$_2$O$_5$ seems not only absent over pristine oceans but also in polluted marine environment. For example, along the eastern coast of China (e.g., Zhejiang), bursts of iodine-bearing aerosols were observed, featuring high levels of IO$_3^-$, without I$_2$O$_5$ signal detected (Yu et al., 2019). Despite the proposed conversion of hydrated I$_2$O$_5$ to HIO$_3$ (detected as IO$_3^-$) (Gómez Martín et al., 2022a), the underlying mechanism remains unclear, especially under the impacts of pollutants. Notably, HSO$_4^-$ (ionization of typical acidic pollutant H$_2$SO$_4$) was also detected at considerable concentrations within iodine aerosols. Accordingly, we explored the H$_2$SO$_4$-mediated hydrolysis of I$_2$O$_5$ at the molecular level. As a strong acid (pK$_{a1}$ = -3.0),

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H_2SO_4 can rapidly lose a proton to form HSO_4^- at the air-water interface, which has a surface preference (Hua et al., 2015). At this stage, HSO_4^- loses strong acidity and shows dissociation equilibrium, thus serving as a proton-transfer bridge. As shown in Figure 5(b), H^+ initially moves from HSO_4^- to I_2O_5, meanwhile, the H atom of H_2O is attracted by SO_4^{2-} and remains the OH^-, leading to I–O bond breaking, forming two HIO_3 molecules. This process couples the self-dissociation of HSO_4^- and the hydrolysis of I_2O_5, expected to occur efficiently on aqueous surface, owing to its lower barrier (~1.3 kcal mol^{-1}) than that mediated by I_2O_5 (3.5 kcal mol^{-1}). The findings indicate that HSO_4^--mediated pathway may play an important role in I_2O_5 depletion under polluted conditions, thereby providing mechanistic insights into the absence of I_2O_5 and the coexistence of HSO_4^- and IO_3^- in field-collected aerosols (Yu et al., 2019). To further quantify the efficiency of these reactions, we calculated the reaction rate constants through Transition State Theory (Table S1). Notably, the hydrolysis of I_2O_5 mediated by I_2O_5 or HIO_3 proceeds 4–5 orders of magnitude faster than that mediated by H_2O, but still more slowly (1-2 orders of magnitude) than that mediated by H_2SO_4. These results highlight the significance of intercomponent coupling in enhancing heterogeneous iodine chemistry in marine atmospheres, from pristine to polluted.
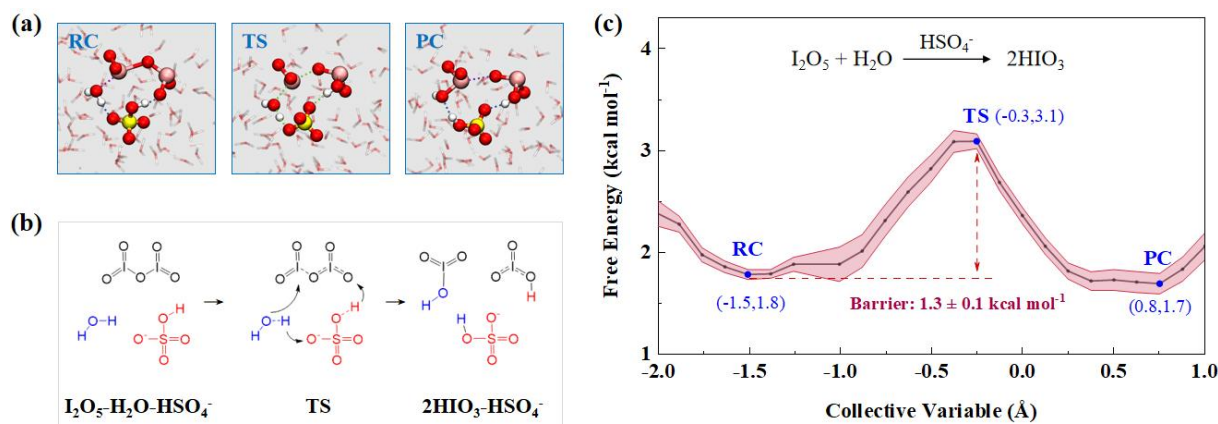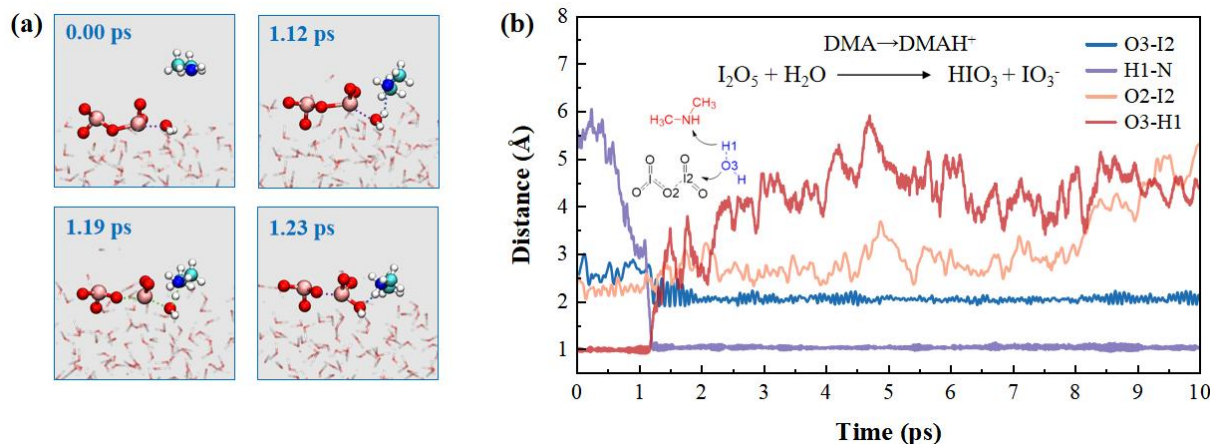

Figure 5. Details of SA-mediated hydrolysis of I_2O_5. (a) Snapshot structures of SMS-MetaD simulation representing RC, TS, and PC. Yellow atoms represent S. (b) Reaction mechanism for the SA-mediated hydrolysis of I_2O_5. (c) Free energy profile for I_2O_5 hydrolysis at the air-water interface mediated by SA (black line) with the error band (pink region).

3.6 Amine-Mediated Interfacial Hydrolysis of I_2O_5

Aside from acid pollutants, atmospheric bases may also affect interfacial chemistry under polluted conditions. Amines, as the typical bases, exhibit evident surface preference and the potential to affect hydrolysis of iodine oxides at the air-water interface (Ning et al., 2024). Accordingly, we investigated the amine-mediated hydrolysis of I_2O_5. Figure 6 shows the DMA-involved key structural snapshots and time-dependent evolution of key bond distances. Initially, DMA is positioned in the gas phase, without notable interaction with aqueous surface. Over time, DMA gradually approaches interfacial water to form HB (O3–H1···N). At 1.12 ps, interfacial water bridges I_2O_5 and DMA via HB (O3–H1···N) and XB (O2–I2···O3), forming
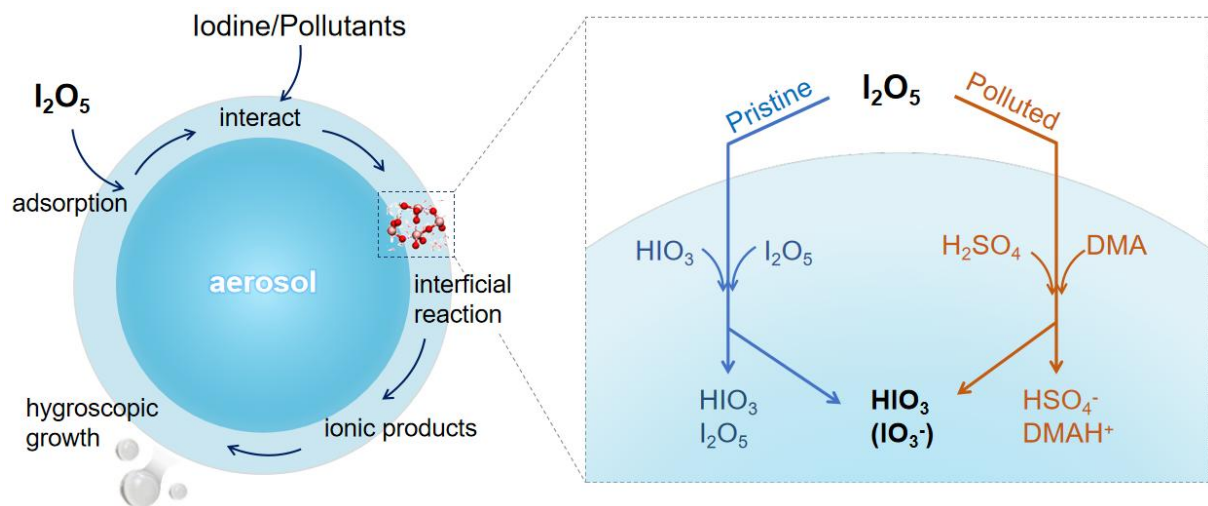
245   a pre-reaction complex. By 1.19 ps, the covalent bond between the H1 and O3 elongated, forming a TS-like structure, then proton H1 transfers from H_2O to DMA. The deprotonated H_2O molecule, generating OH^-, immediately approaches the I2 atom of I_2O_5. The $I-O$ bond in I_2O_5 gradually changes from covalent bond into XB, ultimately yielding HIO_3 and IO_3^- (Figure S18). By 1.23 ps, the reaction ends to form HIO_3, IO_3^-, and alkylammonium salts ($DMAH^+$). The BOMD results indicate that the DMA-mediated hydrolysis of I_2O_5 occurs rapidly on a picosecond timescale, thus attesting to a barrierless

250 pathway. Also, we studied the MA- and TMA-mediated hydrolysis of I_2O_5, which shows similar capabilities of other amines like MA and TMA to accelerate this process (see Figures S5 and S6 in SI). Accordingly, in polluted environments, the role of pollutants in activating iodine chemistry should not be overlooked, as their mediation of rapid heterogeneous processes likely affects aerosol composition and growth dynamics.

255 **Figure 6.** Details of DMA-mediated hydrolysis of I_2O_5. (a) Snapshot structures of BOMD simulation, which illustrate the mechanism for this process. The dashed lines indicate the intermolecular interactions (hydrogen or halogen bonding). Blue and cyan atoms represent N and C atoms, respectively. (b) Time evolution of key bond distances (O3–I2, H1–N, O2–I2, and O3–H1). The arrows indicate the direction of atom transfer.

     In addition to the above analysis of the reaction mechanism and energy barriers, we further explored how interfacial

260 hydrolysis of I_2O_5 affects aerosol growth (Scheme 1) through analysis of dynamic trajectories and wave function of key structures. Here, the DMA-mediated process is taken as an example (see details in SI). The resulting products (HIO_3, IO_3^-, and $DMAH^+$) are partially solvated, as they are surrounded by interfacial water (Figures S15 and S16) via HBs and XBs but still exposed outside to air with unoccupied HB or XB sites (Figure S17). These HB and XB interactions serve the following roles: i) stabilizing the product complex by preventing its decomposition via intermolecular interactions within complex; ii)

265 inhibiting evaporation of product via interactions with interfacial water; iii) promoting aerosol growth by unoccupied HB or XB sites for uptake of gaseous species, in particular, the hydrophilic ionic products (e.g., MAH^+, $DMAH^+$, and $TMAH^+$) accelerate water condensation. Thus, the chemisorption of I_2O_5 contributes to aerosol growth not only directly, by yielding low-volatility products, but also indirectly by enhancing hygroscopic growth. Understanding the impacts of these

heterogeneous reactions and the resulting products may advance our knowledge of how iodine oxides influence iodate aerosol formation.

Scheme 1. Illustration of aerosol growth driven by I_2O_5 hydrolysis at the air-water interface, highlighting the potential reaction pathways and resulting products in pristine and polluted environments.

4 Conclusion

Iodine chemicals are closely linked to marine aerosol formation; however, the underlying physicochemical process, especially involving higher-order iodine oxides (e.g., I_2O_x, $x \geq 2$), remains highly uncertain. HIO_3 is widely detected in marine aerosols, while I_2O_5 is absent, pointing to a potential I_2O_5-to-HIO_3 conversion, although the mechanism is still elusive. To address this, we have investigated interfacial hydrolysis of I_2O_5, the typical I_2O_x, on aqueous aerosol surfaces at the molecular level by BOMD and SMS-MetaD simulations. The results show that I_2O_5 uptakes directly onto the pure water surface, where the interfacial hydrolysis proceeds slowly ($I_2O_5 + H_2O \xrightarrow{H_2O} HIO_3$ in Figure 2). And the self-catalyzed pathways involving I_2O_5 (reactant) and HIO_3 (product) notably lower the energy barrier and facilitate the reaction. Thus, in pristine marine environments, the proposed I_2O_5- and HIO_3-mediated heterogeneous mechanisms (Figures 3 and 4) potentially serve as both an important sink for I_2O_5 and a source of HIO_3, which helps clarify the chemical speciation of field-collected iodine aerosols. In contrast, in polluted regions, pollutant-mediated pathways—with barriers as low as 1.3 kcal mol^{-1} (H_2SO_4-mediated) or even barrierless (amine-mediated)—dominate over self-catalyzed processes, driving faster I_2O_5 interfacial hydrolysis and yielding HIO_3, which helps to elucidate the rapid formation of iodate aerosols observed in polluted eastern coast of China (Yu et al., 2019). It also highlights the role of air pollutants in activating iodine chemistry. Accordingly, these heterogeneous mechanisms can advance our understanding of the roles of I_2O_5 in iodine chemistry, which may help explain the well-established but mechanistically elusive I_2O_5-to-HIO_3 conversion. Notably, the hydrolysis products

(i.e. HIO_3, IO_3^-, and ionic species) enhance aerosol growth through their low volatility, unoccupied HB/XB sites, and hygroscopicity—linking I_2O_5 hydrolysis directly to aerosol growth. These insights deepen our understanding of how iodine oxide affects aerosol formation—both directly and indirectly—in both pristine and polluted marine environments.

Altogether, this study highlights the critical role of heterogeneous hydrolysis of I_2O_5 mediated by some important chemicals in iodine chemistry and aerosol growth. Here, we find that intercomponent coupling-i.e., interactions of iodine oxides with themselves, iodic acids, and atmospheric pollutants-is a key, yet frequently overlooked factor that can reshape heterogenous iodine chemistry across environments, where the reactive pollutant-mediated processes outcompete self-catalyzed pathways. These findings can help explain the observed but poorly understood I_2O_5 and HIO_3 distribution, as well as their chemical conversion, which sheds light on iodine aerosol growth. The iodine-driven impacts are likely global and expected to grow, due to rapidly increasing oceanic iodine emissions. Accordingly, integrating these proposed mechanisms into atmospheric models is expected to reduce uncertainties (unrecognized sink or source for iodine oxides and iodine oxyacids) arising from iodine oxides, which can refine the representation of iodine cycling and further improve the assessment of globally environmental and climatic effects (like aerosol-driven radiative forcing) driven by atmospheric iodine aerosols. The real atmosphere is chemically complex, including iodine species (e.g. HOI, HIO_2, HIO_3, I_2O_3, I_2O_4, and I_2O_5) and atmospheric pollutants (e.g. H_2SO_4, HNO_3, organic acids, and ammonia), which are likely to influence the heterogeneous hydrolysis of I_2O_5. In future work, we intend to confirm the impacts from other atmospheric components.

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