



1 CO₂ deviation in a cylinder due to consumption of a 2 standard gas mixture

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14

15 **Abstract:** The CO₂ molar fraction in standard gas mixtures is known to deviate as a result of
16 adsorption/desorption to/from the inner surface of a high-pressure cylinder and thermal diffusion
17 fractionation caused by the temperature distribution in the cylinder. This deviation reduces the consistency
18 of atmospheric CO₂ observations, because the standard gas mixtures are used to calibrate all measurement
19 systems for precise CO₂ observations. To maintain the consistency of CO₂ values over the long term, a
20 quantitative understanding of the deviations in the CO₂ molar fraction in a standard gas mixture is needed.
21 Thus far, this understanding has not been achieved sufficiently well, because the contribution of thermal
22 diffusion fractionation is less well understood than that of adsorption/desorption. In this study, offsets of
23 $0.013 \pm 0.015 \mu\text{mol mol}^{-1}$ and $-0.014 \pm 0.011 \mu\text{mol mol}^{-1}$ were observed in the outflowing gas from
24 horizontally and vertically positioned cylinders, respectively, at a flow rate of 0.080 L min^{-1} . These offsets
25 are attributed to thermal diffusion effects, which diluted and enriched the CO₂ mole fraction by -0.045
26 $\mu\text{mol mol}^{-1}$ (horizontal cylinder) and $0.048 \mu\text{mol mol}^{-1}$ (vertical cylinder) as the relative pressure dropped
27 to 0.03. In the experiments at same flow rate, the adsorption/desorption effect enriched the CO₂ mole



1 fraction by $0.06 \mu\text{mol mol}^{-1}$ (horizontal cylinder) and $0.10 \mu\text{mol mol}^{-1}$ (vertical cylinder). Therefore,
2 attention should be paid to both thermal diffusion fractionation and adsorption/desorption effects for precise
3 calibration of long-term observations of CO_2 molar fractions, although past studies have ignored the
4 contribution of thermal diffusion fractionation at the low flow rates ($<0.3 \text{ L min}^{-1}$) examined in this study.
5 Furthermore, the deviation of the CO_2 molar fraction depends only on the pressure relative to the initial
6 pressure of the cylinder. This result suggests that the recommendation by the World Meteorological
7 Organization (WMO) to replace the standard gas mixture once the cylinder pressure drops to 2 MPa needs
8 to be revised.

9 **Keywords:** standard gas mixture, atmospheric CO_2 , adsorption/desorption, thermal diffusion fractionation

10 1 Introduction

11 Carbon dioxide (CO_2) is an important greenhouse gas that contributes markedly to the radiative forcing of
12 the atmosphere. Systematic observations of atmospheric CO_2 have been conducted by numerous
13 laboratories around the world to better understand its sources and sinks. By determining the CO_2 molar
14 fraction in the atmosphere based on a scale established on the basis of primary standard gas mixtures in
15 high-pressure aluminum cylinders, the laboratories ensure consistency of the observed values over the long
16 term. Because deviations of the CO_2 molar fractions in the cylinders lead to over- or underestimation of the
17 measured CO_2 molar fraction and reduce the comparability of worldwide CO_2 observations, deviations of
18 the CO_2 molar fractions in the cylinders should be a focus of attention.

19 Langenfelds et al. (2005) reported that the air composition of a standard gas mixture in a high-pressure
20 cylinder could be modified by diffusive and surface processes. Subsequently, Leuenberger et al. (2015) and
21 Schibig et al. (2018) conducted “decanting experiments”, in which a CO_2 -in-air mixture leaving a cylinder
22 was measured continuously, and found that the deviation of the CO_2 molar fraction in the cylinder could be
23 explained by adsorption/desorption phenomena to/from the cylinder inner surface. In the studies of
24 Leuenberger et al. (2015) and Schibig et al. (2018), the amounts of CO_2 adsorbed on the inner surface of



1 the cylinder, expressed as a fraction of the total gas in the cylinder, were estimated to be $0.028 \mu\text{mol mol}^{-1}$
2 and $0.0165 \pm 0.0016 \mu\text{mol mol}^{-1}$, respectively, in decanting experiments using 29.5 L aluminum cylinders.
3 Aoki et al. (2022) estimated the adsorbed CO_2 molar fraction to be $0.027 \pm 0.004 \mu\text{mol mol}^{-1}$ using 10 L
4 aluminum cylinders. Moreover, Schibig et al. (2018) reported that other effects such as thermal diffusion
5 fractionation became more pronounced than adsorption/desorption effects when the flow rate of the
6 outflowing gas from the cylinder was increased. Aoki et al. (2022) also suggested that thermal diffusion
7 fractionation was the main contributor to the “other effects” in their mother–daughter transfer experiments.
8 Aoki et al. (2022) and Schibig et al. (2018) pointed out that thermal diffusion fractionation depended on
9 the position of the cylinder: CO_2 molar fractions were enriched in vertically positioned cylinders but diluted
10 in horizontally positioned cylinders. Thermal diffusion fractionation is driven by the difference in the
11 diffusion velocity between CO_2 and air caused by the temperature gradient in the cylinder, with heavier
12 molecules preferentially accumulating in colder regions. Therefore, these results suggest that colder air
13 leaves from horizontally positioned cylinders and warmer air leaves from vertically positioned cylinders.
14 The same series of primary standard gas mixtures should be used for as long a time as possible to maintain
15 consistency of the CO_2 molar fractions. However, it is not possible to use standard gas mixtures down to
16 lower pressure because the CO_2 molar fraction in the cylinder deviates as the pressure drops as a result of
17 adsorption/desorption and thermal diffusion effects. Therefore, the World Meteorological Organization
18 (WMO) recommends that the standard gas mixtures should be replaced once the cylinder pressure has
19 decreased to 2 MPa. Leuenberger et al. (2015) and Schibig et al. (2018) recommended that the usage of
20 standard gas mixtures in aluminum cylinders should be restricted to pressures above 3 MPa to remain within
21 the WMO’s compatibility goal of $0.1 \mu\text{mol mol}^{-1}$ for the northern hemisphere and $0.05 \mu\text{mol mol}^{-1}$ for the
22 southern hemisphere. If the deviation of the CO_2 molar fraction could be corrected, standard gas mixtures
23 could be used down to lower pressure than the recommended value. However, currently it is difficult to
24 apply this correction because the magnitude of thermal diffusion fractionation has not been sufficiently
25 evaluated, in contrast to the considerable work on adsorption/desorption in previous studies.



1 In this study, we quantitatively estimated the deviation of the CO₂ molar fraction in 10 L aluminum
2 cylinders as the pressure dropped. First, CO₂ deviations were evaluated by means of decanting experiments
3 with different flow rates of the outflowing gas. Second, the fractionation factors of CO₂ resulting from
4 thermal diffusion fractionation were determined by subtracting the adsorption/desorption effect from the
5 deviation in the CO₂ molar fraction measured in the decanting experiment. Last, the actual offsets of the
6 CO₂ values caused by thermal diffusion effect were compared with the offset values calculated based on
7 the fractionation factors. In addition, we discussed how the standard gas mixture in the cylinder should be
8 operated based on the results obtained in this study.

9 **2 Methods**

10 **2.1 Experiment**

11 **2.1.1 Sample gas mixtures**

12 CO₂-in-air mixtures were used as a sample gas to measure the deviations of CO₂ molar fractions. The
13 mixtures were prepared by mixing pure CO₂ (>99.995 %, Nippon Ekitan Corp., Japan) with purified air
14 (G1-grade, <0.1 μmol mol⁻¹ for CO, CO₂, THC, <0.01 μmol mol⁻¹ for NO_x, SO₂, <-80 °C for H₂O, Japan
15 Fine Products, Japan) into a 10 L aluminum cylinder (Luxfer Gas Cylinders, UK;). The CO₂ mole fractions
16 in the CO₂-in-air mixtures were adjusted to an atmospheric level.

17 **2.1.2 Decanting experiment**

18 The CO₂-in-air mixtures in 10 L aluminum cylinders positioned horizontally and vertically were decanted
19 from 10.0 MPa to 0.3 MPa at outflowing gas rates of 0.080 L min⁻¹, 0.15 L min⁻¹, 0.30 L min⁻¹, 1.2 L
20 min⁻¹, and 6.0 L min⁻¹. A schematic diagram of the decanting experiment is shown in Fig. 1. The mixture
21 leaving the cylinder via a single-stage regulator (Torr 1300, NISSAN TANAKA Co., Japan) was divided
22 into two by means of T-pieces. The branched flows were controlled using two mass flow controllers, one
23 of which (SEC-Z512MGX 100 SCCM, Horiba STEC Co., Ltd., Japan) was introduced into a Picarro G2301
24 gas analyzer (Picarro, Inc., California, USA) at a flow rate of 0.080 L min⁻¹, and the other (SEC-Z512MGX

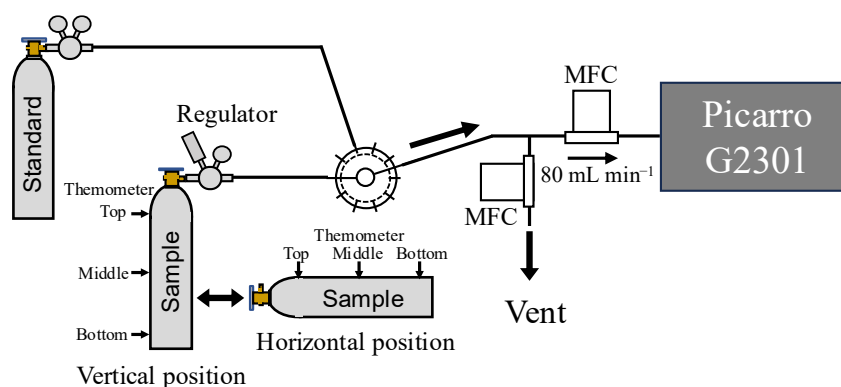


Figure 1. Schematic diagram of the piping used to introduce the CO₂-in-air mixture in a cylinder to a Picarro G2301 in the decanting experiment. MFC, mass flow controller.

1 1 SLM or 10 SLM, Horiba STEC Co., Ltd., Japan) was exhausted to the surroundings at flow rates of 0.0
2 L min⁻¹, 0.070 L min⁻¹, 0.22 L min⁻¹, 1.12 L min⁻¹, and 5.92 L min⁻¹. An absolute pressure gauge of flush
3 diaphragm type (PPA-33X, KELLER AG, Switzerland) attached to the regulator was used to measure
4 pressures in the cylinders. The output values obtained from the Picarro G2301 were linearly calibrated
5 using one of nine standard gas mixtures with CO₂ molar fractions from 337 μmol mol⁻¹ to 452 μmol mol⁻¹
6 and standard uncertainties of less than 0.05 μmol mol⁻¹. Here the output signal was assumed to be zero
7 when the CO₂ molar fraction was zero. The standard gas mixtures were gravimetrically prepared by mixing
8 pure CO₂ and purified air (Aoki et al., 2022). A 0.9 L aluminum cylinder (Luxfer Gas Cylinders, UK) was
9 filled with pure CO₂ and weighed using a balance (AX2005, Mettler Toledo, Switzerland) with a resolution
10 of 0.01 mg and a maximum load of 2 kg. A 10 L aluminum cylinder was filled with purified air and weighed
11 using another balance (XP26003L, Mettler Toledo, Switzerland) with a resolution of 1 mg and a maximum
12 load of 26 kg (Matsumoto et al., 2004; Aoki et al., 2019). The outflowing standard gas mixture from a
13 cylinder with a flow rate of 0.080 L min⁻¹ was introduced directly into the Picarro G2301. After measuring
14 the outflowing standard gas mixture cylinder for 20 min to calibrate the Picarro G2301, the outflowing gas
15 from horizontally or vertically positioned cylinders were measured continuously for 100 min. This cycle
16 was repeated until the pressure dropped to 0.3 MPa. In the decanting experiment at an outflowing gas rate



1 of 6.0 L min⁻¹, the temperatures at the top, middle, and bottom of the cylinders were measured by using a
2 thermocouple-type thermometer that consisted of an insulated thermocouple wire (TT-K-36-SLE-100,
3 OMEGA, Norwalk, California, USA) and a digital multimeter (DMM6500, KEITHLEY, Ohio, USA) with
4 a scanner card (Model 2000-SCAN, KEITHLEY, Ohio, USA) as shown in Fig.1. To investigate the
5 dependence on initial pressure, some decanting experiments were also performed at an outflowing gas flow
6 rate of 0.15 L min⁻¹ and initial pressures of 2.1 MPa, 6.5 MPa, and 11.0 MPa.

7 **2.1.3 Measurement for validation**

8 Three experiments were conducted to validate the fractionation factors obtained by the decanting
9 experiments. The first experiment was measurement of the deviation of the CO₂ value using the Picaro
10 G2301 when the flow rate of gas leaving a cylinder was changed at 20 min intervals. Flow rates of 0.080 L
11 min⁻¹, 0.15 L min⁻¹, 0.30 L min⁻¹, 1.2 L min⁻¹, and 6.0 L min⁻¹ were used in this experiment. The second
12 experiment was measurement of the CO₂ mole fraction in outflowing gas from cylinders positioned
13 vertically and horizontally using the Picaro G2301 and evaluation of the difference in the CO₂ mole fraction
14 between the two positions. An outflowing gas flow rate of 0.080 L min⁻¹ was used in this experiment. The
15 third experiment was measurement of the $\delta(^{29}\text{N}_2/^{28}\text{N}_2)$, $\delta(^{34}\text{O}_2/^{32}\text{O}_2)$, $\delta(^{40}\text{Ar}/^{36}\text{Ar})$, $\delta(^{32}\text{O}_2/^{28}\text{N}_2)$, and
16 $\delta(^{40}\text{Ar}/^{28}\text{N}_2)$ at the start and end of the decanting experiment using a mass spectrometer (Delta-V, Thermo
17 Fisher Scientific Inc., Massachusetts, USA) to clarify the contribution of thermal fractionation during the
18 decanting experiment based on the relationship between the measured elemental and isotopic ratios (e.g.,
19 Langenfels et al., 2003; Ishidoya et al., 2013). The details of the measurement technique using the mass
20 spectrometer have been provided by Ishidoya and Murayama (2014).

21 **2.2 Analytical method for the decanting experiments**

22 **2.2.1 Langmuir adsorption/desorption model**

23 To evaluate the deviation of the CO₂ molar fraction in the CO₂-in-air mixture caused by
24 adsorption/desorption effects, the decanting experiments were repeated using vertically positioned



1 cylinders with low flow rates ($<0.30 \text{ L min}^{-1}$). Each measurement run of every cylinder was used to
2 individually fit a function based on the Langmuir adsorption/desorption model (Langmuir, 1916, 1918) as
3 derived by Leuenberger et al. (2015):

4

$$5 \quad X_{\text{CO}_2, \text{meas}} = X_{\text{CO}_2, \text{ad}} \cdot \left(\frac{K \cdot (P - P_0)}{1 + K \cdot P} + (1 + K \cdot P_0) \cdot \ln \left(\frac{P_0 \cdot (1 + K \cdot P)}{P \cdot (1 + K \cdot P_0)} \right) \right) + X_{\text{CO}_2, \text{initial}}, \quad (1)$$

6

7 Where P is the actual pressure of the cylinder (MPa), P_0 is the initial pressure of the cylinder (MPa) before
8 the decanting experiment, $X_{\text{CO}_2, \text{meas}}$ is the measured CO_2 molar fraction in the outflowing gas, $X_{\text{CO}_2, \text{ad}}$ is
9 the CO_2 molar fraction multiplied by the occupied adsorption sites at pressure P_0 , $X_{\text{CO}_2, \text{initial}}$ is the CO_2
10 molar fraction measured in the outflowing gas at pressure P_0 , and K is the ratio of the adsorption rate
11 constant to the desorption rate constant (unit MPa^{-1}). $X_{\text{CO}_2, \text{ad}}$, $X_{\text{CO}_2, \text{initial}}$, and K were obtained from the
12 nonlinear least-squares fit to the measurement results.

13 **2.2.1 Rayleigh distillation model and its combination with the Langmuir adsorption/desorption** 14 **model**

15 The offset of the CO_2 molar fraction in the outflowing gas caused by thermal diffusion fractionation can be
16 represented using a Rayleigh distillation model (Rayleigh, 1902; Matsubaya and Matsuo, 1982;
17 Langenfelds et al., 2005) according to the following equation:

18

$$19 \quad \frac{X}{X_0} = \left(\frac{P}{P_0} \right)^{\alpha - 1}, \quad (2)$$

20

21 where X corresponds to the measured CO_2 mole fraction; X_0 corresponds to the initial CO_2 mole fraction in
22 the outflowing gas; and α is the fractionation factor of CO_2 when the CO_2 -in-air mixture leaves the cylinder.
23 The CO_2 molar fraction in the outflowing gas is depleted if $\alpha < 1$, which increases the CO_2 molar fraction
24 in the remaining CO_2 -in-air mixture in the cylinder (and vice versa). It is possible to obtain reasonable fits



1 to the measured CO₂ molar fraction data by the Langmuir adsorption/desorption model (Eq. (1)) or Rayleigh
2 distillation function (Eq. (2)); in other words, it is difficult to separate the contributions of
3 adsorption/desorption and thermal diffusion fractionation. Therefore, the Langmuir–Rayleigh model,
4 which integrates the Langmuir model and the Rayleigh function, is required to evaluate
5 adsorption/desorption and thermal diffusion effects. The Langmuir–Rayleigh model was proposed by
6 Schibig et al. (2018) to analyze the results of decanting experiments as follows:

$$8 \quad X_{\text{CO}_2, \text{meas}} = X_{\text{CO}_2, \text{ad,ave}} \cdot \left(\frac{K_{\text{ave}} \cdot (P - P_0)}{1 + K_{\text{ave}} \cdot P} \right) + (1 + K_{\text{ave}} \cdot P_0) \cdot \ln \left(\frac{P_0 \cdot (1 + K_{\text{ave}} \cdot P)}{P \cdot (1 + K_{\text{ave}} \cdot P_0)} \right) + X_0 \cdot \left(\frac{P}{P_0} \right)^{\alpha-1} \quad (3)$$

9
10 where $X_{\text{CO}_2, \text{ad,ave}}$ is the average $X_{\text{CO}_2, \text{ad}}$ coefficient of the low-flow experiments, and K_{ave} is the
11 average ratio of the adsorption and desorption rate constants of the low-flow experiments. The value of α
12 can be obtained by fitting Eq. (3) to the results of the decanting experiments, with the values of $X_{\text{CO}_2, \text{ad,ave}}$
13 and K_{ave} determined in advance.

15 3 Results

16 3.1 Decanting experiments

17 The decanting experiments were performed to evaluate the deviation of the CO₂ molar fraction in the 10 L
18 aluminum cylinders resulting from thermal diffusion fractionation as the pressure dropped. Decanting the
19 CO₂-in-air mixtures from the 10 L aluminum cylinders reduced cylinder temperatures by a maximum of ~6
20 K depending on the outflowing gas flow rate. The temperature distribution in the cylinder depends on the
21 outflowing gas flow rate and the cylinder position (Schibig et al., 2018; Aoki et al., 2022). The temperature
22 reduction could also alter the amount of CO₂ adsorbed on the inner surface of the cylinder, because the
23 adsorption energy changes depending on the cylinder temperature. However, the change of the adsorbed
24 CO₂ amount resulting from temperature variation is estimated to be less than 0.002 μmol mol⁻¹ because the



1 temperature dependence that was observed for aluminum cylinders by Leuenberger et al. (2015) was
 2 between -0.0002 and $-0.0003 \mu\text{mol mol}^{-1} \text{K}^{-1}$. The change is negligible because the contribution is below
 3 the CO_2 value reproducibility of $0.005 \mu\text{mol mol}^{-1}$. Therefore, CO_2 dilution and enrichment in cylinders
 4 with different flow rates, which ranged from -0.08 to $0.31 \mu\text{mol mol}^{-1}$ (Fig. 2), depends on thermal
 5 diffusion fractionation rather than adsorption/desorption.

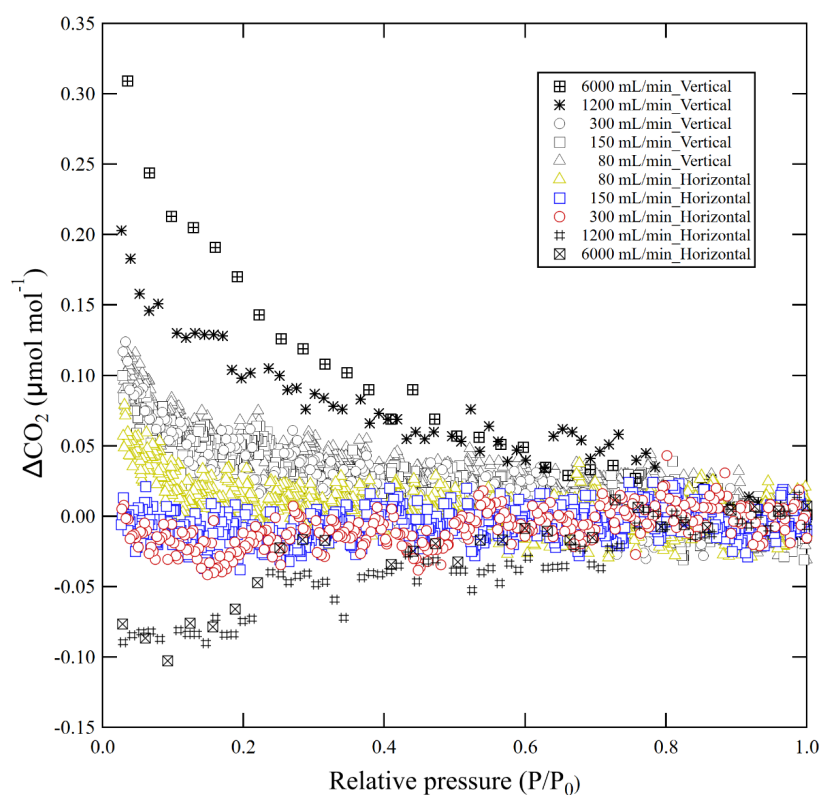


Figure 2. Plot showing deviation of the CO_2 molar fraction from the initial value against relative pressure. These results were obtained by decanting experiments at outflowing gas flow rates between 0.080 L min^{-1} , 0.15 L min^{-1} , 0.30 L min^{-1} , 1.2 L min^{-1} , and 6.0 L min^{-1} with vertically positioned cylinders and horizontally positioned cylinders



3.1.1 Flow rate dependency

The decanting experiments were performed at outflowing gas flow rates of 0.080 L min⁻¹, 0.15 L min⁻¹, 0.30 L min⁻¹, 1.2 L min⁻¹ and 6.0 L min⁻¹ for cylinders positioned horizontally and vertically until the pressure dropped from 10 MPa to 0.3 MPa. Figure 2 shows the deviations of the CO₂ molar fraction in the outflowing gas as the relative pressure (P/P_0) in the cylinders dropped. For a horizontally positioned cylinder, the deviations of CO₂ molar fraction at a relative pressure of 0.03 were between 0.06 μmol mol⁻¹ to -0.08 μmol mol⁻¹ relative to the initial CO₂ molar fractions from the initial value as summarized in Table 1. The deviation decreased as the flow rate increased, indicating that thermal diffusion fractionation acted to dilute the CO₂ molar fraction in the horizontally positioned cylinder because adsorption/desorption acted to enrich the CO₂ molar fraction (Leuenberger et al., 2015; Schibig et al., 2018; Aoki et al., 2022). These results also mean that the contribution of thermal diffusion fractionation increased at higher flow rates. At a flow rate of 0.080 L min⁻¹, the CO₂ molar fraction was enriched as the relative pressure dropped, indicating that the effect of adsorption/desorption was larger than that of thermal diffusion fractionation. At flow rates of 0.15 L min⁻¹ and 0.30 L min⁻¹, the CO₂ molar fractions were almost constant, indicating that the increase due to adsorption/desorption was cancelled out by the decrease due to thermal diffusion fractionation. At flow rates of 1.2 L min⁻¹ and 6.0 L min⁻¹, the CO₂ molar fractions decreased as the pressure dropped, indicating that the thermal diffusion effect was larger than the adsorption/desorption effect.

Table 1 Deviations of CO₂ molar fraction in the outflowing gas from initial values measured by decanting experiments at flow rates of 0.080 L min⁻¹, 0.15 L min⁻¹, 0.30 L min⁻¹, 1.2 L min⁻¹, and 6.0 L min⁻¹.

Flow rate	Deviations at a relative pressure of 0.03 (μmol mol ⁻¹)	
	Horizontally positioned cylinder	Vertically positioned cylinder
0.080 L min ⁻¹	0.06	0.12
0.15 L min ⁻¹	-0.002	0.11
0.30 L min ⁻¹	-0.005	0.12
1.2 L min ⁻¹	-0.08	0.20
6.0 L min ⁻¹	-0.08	0.31



1 For vertically positioned cylinders, at all outflow rates the CO₂ molar fraction in the outflowing gas
2 increased from the initial value as the pressure dropped. The increases in the CO₂ molar fractions at a
3 relative pressure of 0.03 were between 0.12 μmol mol⁻¹ and 0.31 μmol mol⁻¹ relative to the initial values
4 (Table 1). The increases were larger at higher flow rates, indicating that thermal diffusion fractionation
5 acted to enrich the CO₂ molar fraction and its contribution was greater with increased flow rate. However,
6 there was little difference in the CO₂ enrichment for flow rates less than 0.30 L min⁻¹, suggesting that the
7 contribution of thermal diffusion fractionation was minimal at these rates, and the CO₂ enrichment can
8 mainly be attributed to adsorption/desorption effects.

9 To understand the mechanism of thermal diffusion fractionation, the temperatures at the top, middle, and
10 bottom of the cylinders were measured using a thermocouple-type thermometer (Fig.1). Figure 3a shows
11 the relationship of pressure and temperatures at the top, middle, and bottom of a horizontally positioned
12 cylinder when decanting the CO₂-in-air mixture from 10 MPa to 0.3 MPa at a flow rate of 6.0 L min⁻¹. The
13 temperatures at the top, middle, and bottom of the cylinder decreased as the pressure dropped, while the
14 temperatures at the top, middle, and bottom were almost equivalent at all pressures. These results do not
15 provide insights into the thermal distribution that drives thermal diffusion fractionation; thus, further study
16 of the mechanism of thermal diffusion fractionation in a horizontally positioned cylinder is required. Figure
17 3b shows the relationship between pressure and the temperatures at the top, middle, and bottom of a
18 vertically positioned cylinder during decanting of the CO₂-in-air mixture at a flow rate of 6.0 L min⁻¹ from
19 10 MPa to 0.3 MPa. The temperatures at the top, middle, and bottom of the cylinder decreased as the
20 pressure dropped, reaching stable values below 2 MPa, while the temperature differences between the
21 different parts of the cylinder increased as the pressure dropped. The temperature difference between the
22 cylinder top and bottom was about 0.7 °C at pressures below 2 MPa, indicating that thermal diffusion
23 fractionation was caused by the temperature difference between the upper and lower parts of the cylinder.

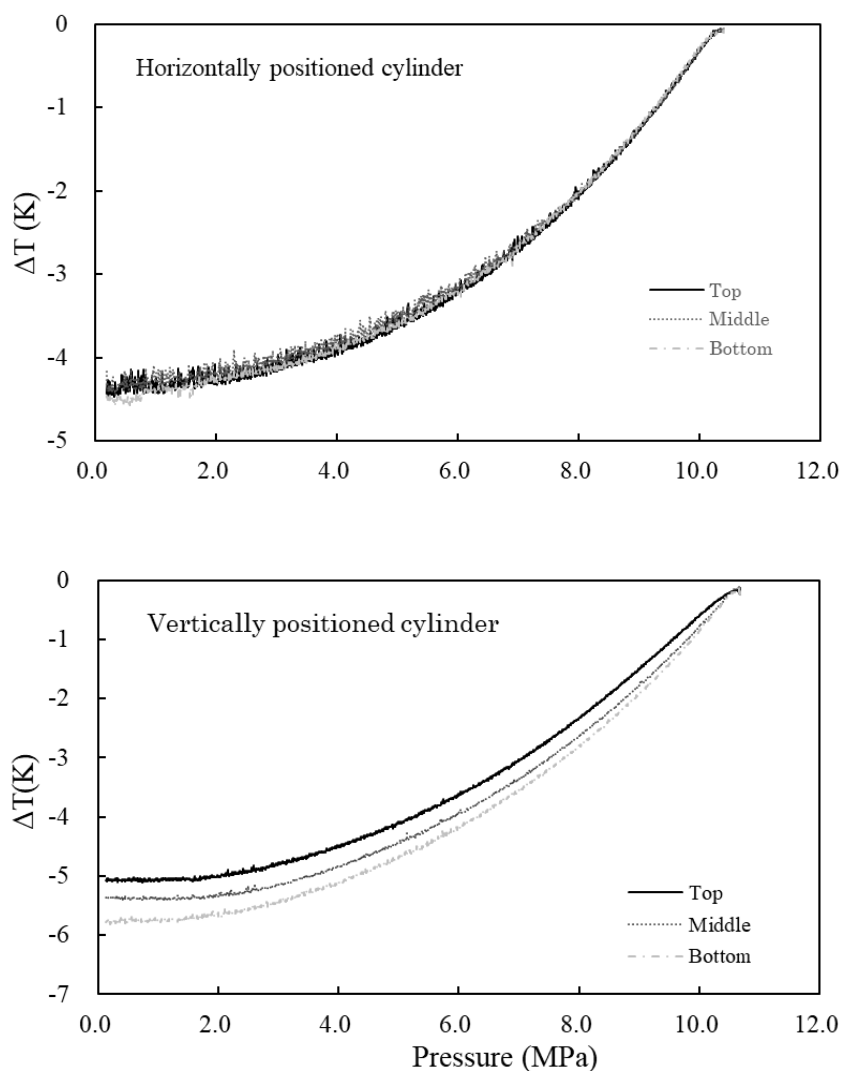


Figure 3. Temperature changes from the initial values of at the top, middle, and bottom of the cylinder when the CO₂-in-air mixture was decanted at a flow rate of 6 L min⁻¹.

1 3.1.2 Evaluation of adsorption/desorption effect

- 2 CO₂ enrichment in a vertically positioned cylinder is considered to be mainly due to adsorption/desorption
- 3 in the decanting experiment performed at an outflowing gas flow rate of less than 0.30 L min⁻¹, as described



- 1 in Sect. 3.1.1. In this section, the adsorption/desorption effect was quantitatively evaluated from the results
- 2 obtained by repeating the decanting experiment at a flow rate of less than 0.30 L min^{-1} with a vertically

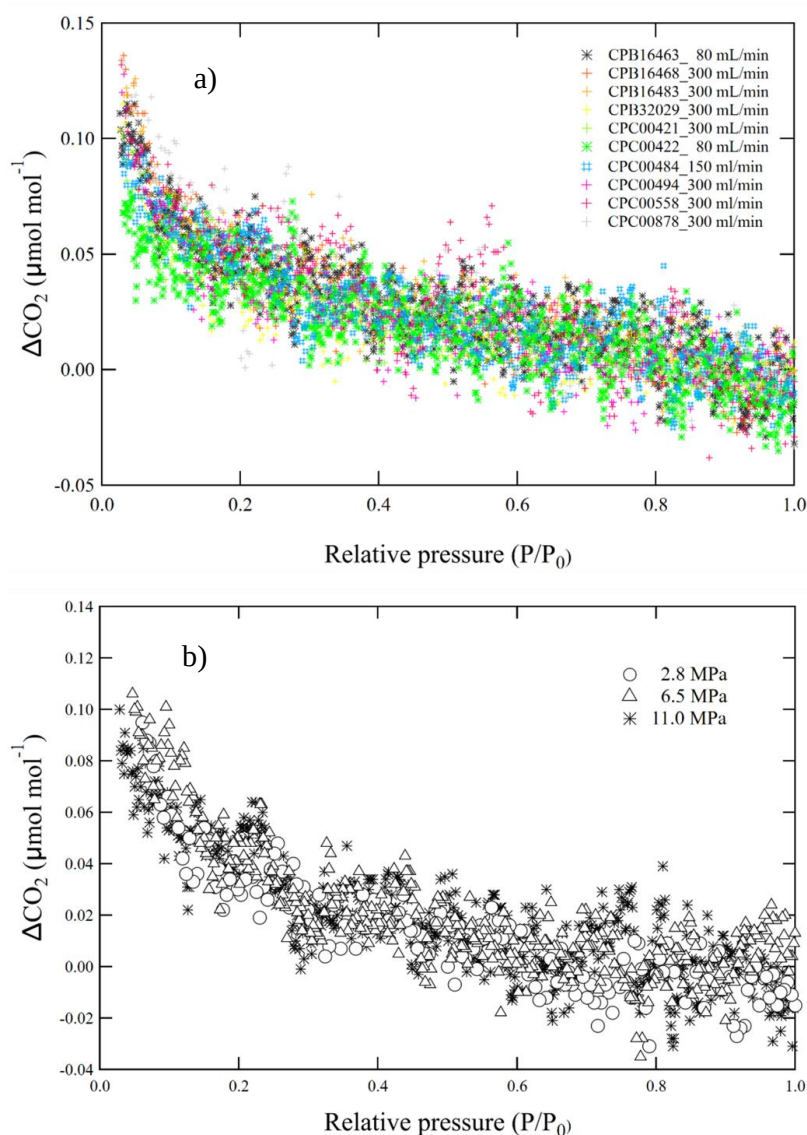


Figure 4. (a) Deviations of the CO₂ molar fraction from the initial value versus relative pressure at flow rates of less than 0.30 L min^{-1} in vertically positioned cylinders. (b) Deviations of the CO₂ molar fraction from the initial value versus relative pressure for initial pressures of 2.1 MPa, 6.5 MPa, and 11.0 MPa.



1 positioned cylinder. In this experiment, the CO₂ enrichment was assumed to be caused by only
2 adsorption/desorption effects.

3 The decanting experiments were initially repeated seven times with a CPC00494 cylinder to determine the
4 measurement uncertainty of CO₂ enrichment as the pressure dropped. The Langmuir model was fitted to
5 each measurement result. The average values of K and $X_{\text{CO}_2,\text{ad}}$ were $0.020 \pm 0.036 \text{ MPa}^{-1}$ and $0.027 \pm$
6 $0.002 \mu\text{mol mol}^{-1}$, respectively. Here, the number following the symbol represents the standard deviation.

7 The decanting experiments were then repeated 10 times, each with a different cylinder with the same types
8 of internal surface treatment and diaphragm valve, to determine the adsorption/desorption effect
9 quantitatively. Figure 4a shows the deviations of the CO₂ molar fraction in the outflowing gas with decreasing
10 P/P_0 , obtained from the decanting experiments with 10 replicates. The CO₂ molar fraction increased from
11 $0.08 \mu\text{mol mol}^{-1}$ to $0.15 \mu\text{mol mol}^{-1}$ from initial values as P/P_0 dropped to 0.03. The average K and $X_{\text{CO}_2,\text{ad}}$
12 values were $0.024 \pm 0.035 \text{ MPa}^{-1}$ and $0.028 \pm 0.005 \mu\text{mol mol}^{-1}$, respectively, when fitting a function based
13 on the Langmuir model. The averages were consistent with that for the CPC00494 cylinder within
14 uncertainty, demonstrating that K and $X_{\text{CO}_2,\text{ad}}$ do not differ in different cylinders.

15 In addition, decanting experiments were performed at initial pressures P_0 of 2.1 MPa, 6.5 MPa, and 11.0
16 MPa with a vertically positioned cylinder. A flow rate of 0.15 L min^{-1} was used for the outflowing gas.
17 Figure 4b shows the deviations of the CO₂ molar fraction in the outflowing gas from the CO₂-in-air mixture
18 with decreasing P/P_0 . The deviations obtained from the three experiments agreed well with each other,
19 indicating that the adsorption/desorption effect in the vertically positioned cylinder depends on P/P_0 rather
20 than P .

21 3.1.3 Estimation of thermal diffusion fractionation

22 Fractionation factors for the CO₂-in-air mixture leaving the cylinders were obtained by fitting a function
23 based on the Langmuir–Rayleigh model (Eq. (3)) to the results described in Sect. 3.1.1; the functions are
24 shown in Fig. 5. The constant coefficient K_{ave} and $X_{\text{CO}_2,\text{ad,ave}}$ were $0.024 \pm 0.035 \text{ MPa}^{-1}$ and $0.028 \pm$
25 $0.005 \mu\text{mol mol}^{-1}$, respectively, as determined in the previous section. The fractionation factors (α)



- 1 obtained from the fit functions following the Langmuir–Rayleigh model and the deviation of the CO₂ molar
- 2 fraction calculated based on the α values are summarized in Table 2.

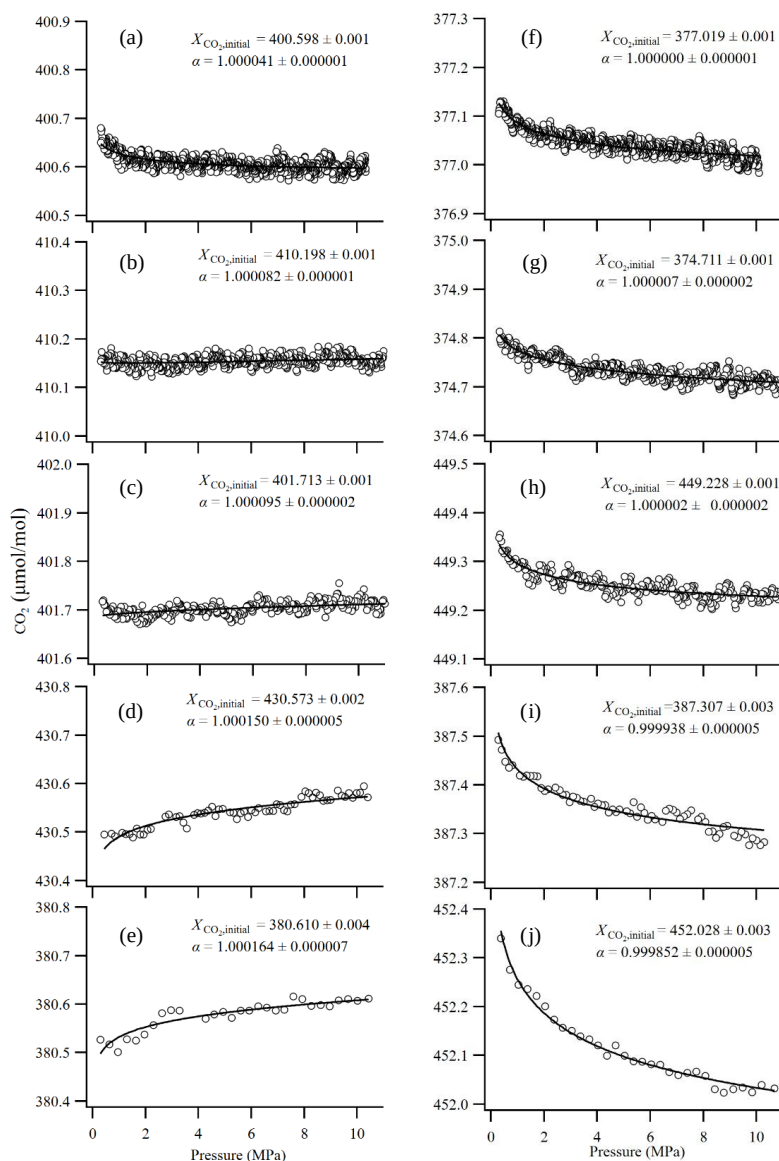


Figure 5. Results from fitting the equation combining the Langmuir and Rayleigh distillation functions to the deviations of the CO₂ molar fraction versus relative pressure. (a–e) Results for horizontally positioned cylinders. (f–j) Results for vertically positioned cylinders.



- 1 For horizontally positioned cylinders, the fractionation factor α was between 1.000041 ± 0.000001 and
- 2 1.000164 ± 0.000007 ; the outflowing gas had offsets between $0.017 \pm 0.000 \mu\text{mol mol}^{-1}$ and 0.069 ± 0.002

Table 2 Fractionation factors for CO₂-in-air mixtures leaving 10 L aluminum cylinders obtained by fitting the Langmuir–Rayleigh model to the decanting measurements. Offsets and differences are from the original values in the cylinders and from the values for 0.080 L min⁻¹, and they were calculated from the fractionation factors.

Flow rate	Fractionation factor ^a	Offsets calculated from the original values ^b ($\mu\text{mol mol}^{-1}$)	Measured differences from the 0.080 L min ⁻¹ values ^c ($\mu\text{mol mol}^{-1}$)	Theoretical differences from the 0.080 L min ⁻¹ values ^d ($\mu\text{mol mol}^{-1}$)
Horizontally positioned cylinder				
0.080 L min ⁻¹	1.000041 ± 0.000001	0.017 ± 0.000	—	—
0.15 L min ⁻¹	1.000082 ± 0.000001	0.034 ± 0.000	0.009 ± 0.018	0.017 ± 0.001
0.30 L min ⁻¹	1.000095 ± 0.000002	0.040 ± 0.001	0.025 ± 0.018	0.023 ± 0.001
1.2 L min ⁻¹	1.000150 ± 0.000005	0.063 ± 0.002	0.049 ± 0.018	0.046 ± 0.002
6.0 L min ⁻¹	1.000164 ± 0.000007	0.069 ± 0.003	0.050 ± 0.018	0.052 ± 0.003
Vertically positioned cylinder				
0.080 L min ⁻¹	1.000000 ± 0.000001	0.000 ± 0.001	—	—
0.15 L min ⁻¹	1.000007 ± 0.000002	0.003 ± 0.001	0.010 ± 0.018	0.003 ± 0.001
0.30 L min ⁻¹	1.000002 ± 0.000002	0.001 ± 0.001	0.011 ± 0.018	0.001 ± 0.001
1.2 L min ⁻¹	0.999938 ± 0.000005	-0.026 ± 0.002	-0.016 ± 0.018	-0.025 ± 0.002
6.0 L min ⁻¹	0.999852 ± 0.000005	-0.062 ± 0.002	-0.074 ± 0.018	-0.060 ± 0.002

The number following the symbol represents the standard uncertainty.

^a These values were calculated by fitting Eq. (3) to the results of the decanting experiments described in Sect. 3.1.1. The standard uncertainty represents the standard deviation obtained from the fitting.

^b Offsets of the CO₂ molar fraction in the outflowing gas for a CO₂-in-air mixture with an original molar fraction of $420 \mu\text{mol mol}^{-1}$. These offsets represent the differences between the original values and the values obtained by multiplying the original values by the fractionation factors.

^c Differences from the 0.080 L min⁻¹ value when changing the outflowing gas flow rate. These differences were determined by measuring CO₂-in-air mixtures with CO₂ molar fractions of $421.2 \mu\text{mol mol}^{-1}$ and $406.5 \mu\text{mol mol}^{-1}$ for horizontally and vertically positioned cylinders, respectively.

^d Differences from the 0.080 L min⁻¹ value when changing the outflowing gas flow rate when decanting CO₂-in-air mixtures with CO₂ molar fractions of $421.2 \mu\text{mol mol}^{-1}$ and $406.5 \mu\text{mol mol}^{-1}$ for horizontally and vertically positioned cylinders, respectively. These differences were calculated based on the fractionation factors.



1 $\mu\text{mol mol}^{-1}$ from the original values in the cylinders (Table 2) for a CO_2 molar fraction of $420 \mu\text{mol mol}^{-1}$
2 (the atmospheric level).

3 For vertically positioned cylinders, α was between 1.000000 ± 0.000001 and 0.999852 ± 0.000005 ; the
4 outflowing gas had offsets between $0.000 \pm 0.001 \mu\text{mol mol}^{-1}$ and $-0.062 \pm 0.002 \mu\text{mol mol}^{-1}$ from the
5 original values, respectively (Table 2) for a CO_2 molar fraction of $420 \mu\text{mol mol}^{-1}$.

6 **3.2 Validation of thermal diffusion fractionation**

7 The fractionation factors determined in the previous section were validated in three ways: first, by
8 measuring the offset of the CO_2 molar fraction corresponding to the fractionation factors when changing
9 the flow rate of the outflowing gas (see Sect. 3.2.1); second, by measuring the CO_2 molar fraction in the
10 outflowing gas from the same horizontally or vertically positioned cylinder at a flow rate of 0.080 L min^{-1}
11 and comparing the difference between values (see Sect. 3.2.2); and third, by measuring $\delta(^{29}\text{N}_2/^{28}\text{N}_2)$,
12 $\delta(^{34}\text{O}_2/^{32}\text{O}_2)$, $\delta(^{32}\text{O}_2/^{28}\text{N}_2)$, $\delta(^{40}\text{Ar}/^{28}\text{N}_2)$, and $\delta(^{40}\text{Ar}/^{36}\text{Ar})$ by mass spectrometry before and after the
13 decanting experiment (see Sect. 3.2.3).

14 **3.2.1 Deviations of CO_2 molar fractions at different flow rates**

15 The fractionation factors determined in Sect. 3.1.3 suggest that the CO_2 molar fractions in the outflowing
16 gas have the offsets from the original values depending on the flow rate. The outflowing gas from
17 horizontally and vertically positioned cylinders with CO_2 molar fractions of $421.2 \mu\text{mol mol}^{-1}$ and 406.6
18 $\mu\text{mol mol}^{-1}$ were continuously measured as the outflowing gas flow rate was varied from 0.080 L min^{-1} to
19 6.0 L min^{-1} at 20 min intervals.

20 The differences from the CO_2 value of the 0.080 L min^{-1} flow rate were between $0.009 \pm 0.018 \mu\text{mol mol}^{-1}$
21 and $0.050 \pm 0.018 \mu\text{mol mol}^{-1}$ (Table 2). Here, the number following the symbol represents the standard
22 uncertainty ($\sqrt{0.013^2 + 0.013^2} = 0.018$), which was calculated by combining the measurement repeatability
23 of the CO_2 values ($0.013 \mu\text{mol mol}^{-1}$) at each flow rate and at 0.080 L min^{-1} . The theoretical differences
24 from the CO_2 value at 0.080 L min^{-1} were calculated based on the fractionation factors to be between 0.017



1 $\pm 0.001 \mu\text{mol mol}^{-1}$ and $0.052 \pm 0.003 \mu\text{mol mol}^{-1}$ (Table 2). The measured difference values agreed with
2 the theoretical values within the uncertainties, suggesting that the differences between the fractionation
3 factors are valid for the horizontally positioned cylinders.

4 The differences from the CO_2 value at 0.080 L min^{-1} were between $0.010 \pm 0.018 \mu\text{mol mol}^{-1}$ and -0.074
5 $\pm 0.018 \mu\text{mol mol}^{-1}$. Theoretical differences from the CO_2 value for 0.080 L min^{-1} were calculated based
6 on the fractionation factors to be between $0.003 \pm 0.001 \mu\text{mol mol}^{-1}$ and $-0.060 \pm 0.002 \mu\text{mol mol}^{-1}$ (Table
7 2). All of the measured difference values also agreed with the theoretical values, suggesting that the
8 differences in fractionation factor are valid for vertically positioned cylinders.

9 **3.2.2 Difference in CO_2 molar fractions for vertically and horizontally positioned cylinders**

10 The dependence of the fractionation factor on the outflowing gas flow rates in each cylinder position was
11 verified in the previous section; however, the difference between vertically and horizontally positioned
12 cylinders was not verified. In this section, the CO_2 differences of a cylinder containing the same CO_2 -in-air
13 mixture with a CO_2 molar fraction of $391.9 \mu\text{mol mol}^{-1}$ was measured in both horizontal and vertical
14 positions to evaluate whether an offset of the CO_2 molar fraction corresponding to the fractionation factors
15 could be detected between the positions.

16 The CO_2 offsets at an outflow rate of 0.080 L min^{-1} were calculated to be $0.017 \pm 0.001 \mu\text{mol mol}^{-1}$
17 (horizontal cylinder) and $0.000 \pm 0.001 \mu\text{mol mol}^{-1}$ (vertical cylinder), based on fractionation factors of
18 1.000041 ± 0.000001 (horizontal cylinder) and 1.000000 ± 0.000001 (vertical cylinder). The
19 difference of the CO_2 molar fraction between the horizontal and vertical positions is estimated to be 0.017
20 $\pm 0.001 \mu\text{mol mol}^{-1}$. Here, the number following the symbol is the standard uncertainty obtained by
21 combining the uncertainties of both offsets. To detect the difference, the cylinder was left in a horizontal
22 position overnight and measured once, then left in a vertical position overnight and measured once, and the
23 measurement sequence was performed four times. The average value of the measured difference between
24 the two positions was $0.011 \pm 0.004 \mu\text{mol mol}^{-1}$ (Fig. 6). The number following the symbol represents the
25 standard uncertainty, which was calculated by combining the standard error of the CO_2 molar fraction for



1 each cylinder position. The expanded uncertainties ($k = 2$) of the measured and estimated differences were
2 $0.008 \mu\text{mol mol}^{-1}$ and $0.002 \mu\text{mol mol}^{-1}$, respectively. These measured and estimated differences of 0.011
3 $\pm 0.008 \mu\text{mol mol}^{-1}$ and $0.016 \pm 0.002 \mu\text{mol mol}^{-1}$ are in agreement within uncertainty, suggesting that the
4 difference in the fractionation factors is valid between horizontal and vertical cylinder positions.

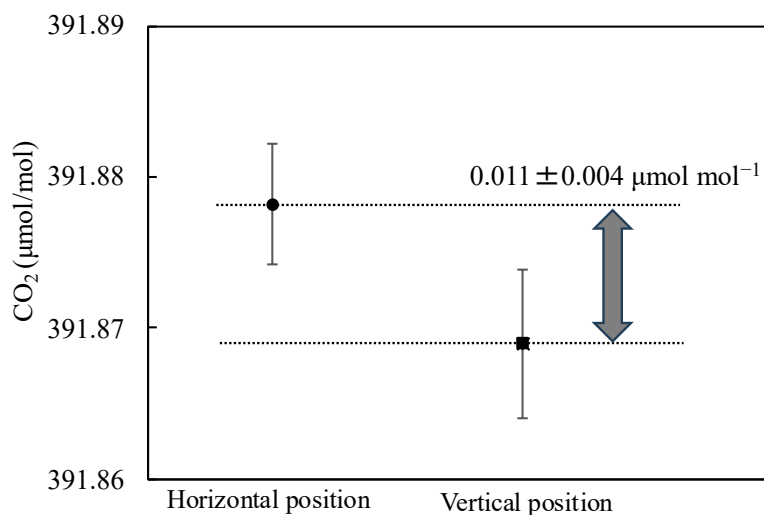


Figure 6. CO₂ molar fraction in a cylinder measured in both vertical and horizontal positions. Error bars represent standard errors.

5 3.2.3 Contribution of thermal diffusion fractionation at 0.080 L min^{-1} flow rate

6 As discussed above, the relationship of the fractionation factors between the different outflowing gas rates
7 summarized in Table 2 is relatively valid. However, the fractionation factors were calculated by assuming
8 that thermal diffusion fractionation was negligible for gas flowing out from a vertically positioned cylinder
9 at a flow rate of less than 0.30 L min^{-1} (Fig. 2). To validate this assumption, we measured $\delta(^{34}\text{O}_2/^{32}\text{O}_2)$,
10 $\delta(^{40}\text{Ar}/^{36}\text{Ar})$, $\delta(^{32}\text{O}_2/^{28}\text{N}_2)$, $\delta(^{40}\text{Ar}/^{28}\text{N}_2)$, and $\delta(^{29}\text{N}_2/^{28}\text{N}_2)$ in the outflowing gas before and after decanting
11 from 8 MPa to below 0.9 MPa. The experiments were carried out by using a vertically positioned cylinder



1 with flow rates of 0.080 L min⁻¹, 0.15 L min⁻¹, and 0.30 L min⁻¹, and a horizontally positioned cylinder
2 with a flow rate of 0.080 L min⁻¹.

3 Figure 7 shows the relationship of the deviations of $\delta(^{34}\text{O}_2/^{32}\text{O}_2)$, $\delta(^{40}\text{Ar}/^{36}\text{Ar})$, $\delta(^{32}\text{O}_2/^{28}\text{N}_2)$, and $\delta(^{40}\text{Ar}/^{28}\text{N}_2)$
4 values against that of $\delta(^{29}\text{N}_2/^{28}\text{N}_2)$. Most of the deviations of $\delta(^{34}\text{O}_2/^{32}\text{O}_2)$, $\delta(^{40}\text{Ar}/^{36}\text{Ar})$, $\delta(^{32}\text{O}_2/^{28}\text{N}_2)$, and
5 $\delta(^{40}\text{Ar}/^{28}\text{N}_2)$ values against $\delta(^{29}\text{N}_2/^{28}\text{N}_2)$ fall on the dotted lines within uncertainties, suggesting that the
6 deviations were caused by thermal diffusion fractionation. Thus, thermal diffusion fractionation occurs
7 even at low flow rates whether the cylinders are positioned horizontally or vertically.

8 The deviations of $\delta(^{29}\text{N}_2/^{28}\text{N}_2)$ at a flow rate of 0.080 L min⁻¹ were -2.7 ± 1.4 per meg in the depletion from
9 8.3 to 0.6 MPa for the horizontally positioned cylinder and 3.9 ± 1.4 per meg in the depletion from 8.5 MPa
10 to 0.2 MPa for the vertically positioned cylinder. These values correspond to CO₂ deviations of $-0.032 \pm$
11 $0.017 \mu\text{mol mol}^{-1}$ and $0.047 \pm 0.017 \mu\text{mol mol}^{-1}$ for horizontally and vertically positioned cylinders,
12 respectively, based on the relationship between the $\delta(\text{CO}_2/\text{N}_2)$ deviations and those of $\delta(^{29}\text{N}_2/^{28}\text{N}_2)$ of Aoki
13 et al. (2022; their fig. 5). The number following the symbol indicates the standard uncertainties of the
14 deviations, which were based on the uncertainties of the deviations of $\delta(^{29}\text{N}_2/^{28}\text{N}_2)$. When substituting the
15 CO₂ molar fractions and the pressures before and after each decanting experiments into the function based
16 on the Rayleigh distillation model (Eq. (2)), the fractionation factors were calculated to be $1.000030 \pm$
17 0.000037 for the horizontally positioned cylinder and 0.999968 ± 0.000027 for the vertically positioned
18 cylinder with the atmospheric CO₂ level of $420 \mu\text{mol mol}^{-1}$. The fractionation factors correspond to offsets
19 in the outflowing gas of $0.013 \pm 0.015 \mu\text{mol mol}^{-1}$ (horizontal cylinder) and $-0.014 \pm 0.011 \mu\text{mol mol}^{-1}$
20 (vertical cylinder), meaning that the CO₂ mole fraction in the horizontally and vertically positioned cylinder
21 deviated by $-0.045 \mu\text{mol mol}^{-1}$ and $0.048 \mu\text{mol mol}^{-1}$, respectively, as the relative pressure dropped to 0.03.
22 The difference in the CO₂ molar fraction between outflowing gases for both cylinder positions was
23 calculated to be $0.027 \pm 0.038 \mu\text{mol mol}^{-1}$, consistent with the difference of $0.011 \pm 0.008 \mu\text{mol mol}^{-1}$
24 between the horizontally and vertically positioned cylinders obtained in the previous section. The numbers

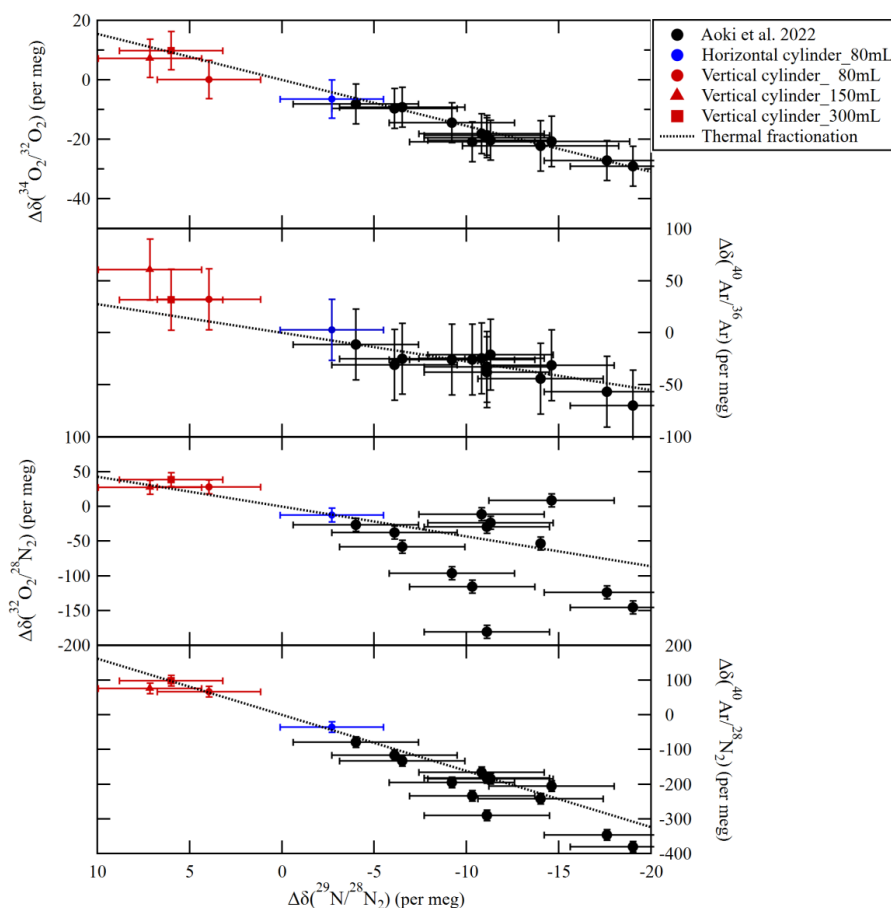


Figure 7. Relationship between the deviations of $\delta(^{34}\text{O}_2/^{32}\text{O}_2)$, $\delta(^{40}\text{Ar}/^{36}\text{Ar})$, $\delta(^{32}\text{O}_2/^{28}\text{N}_2)$, and $\delta(^{40}\text{Ar}/^{28}\text{N}_2)$ and those of $\delta(^{29}\text{N}_2/^{28}\text{N}_2)$ in daughter cylinders relative to their initial value when CO_2/air mixtures with an atmospheric CO_2 level were decanted from the cylinder. The error bars indicate the expanded uncertainties ($k = 2$) of the deviations. The dotted lines represent the deviations due to thermal diffusion, which were experimentally estimated by Ishidoya et al. (2013, 2014). The black closed circles represent the deviations in daughter cylinders relative to their mother cylinders obtained by mother–daughter experiments (Aoki et al. 2022).

- 1 after the symbol represent the expanded uncertainties ($k = 2$), which were calculated by combining the
- 2 standard uncertainties for both cylinder positions. This finding indicates that the fractionation factors



1 obtained using the mass spectrometer are reasonable and the assumption that thermal diffusion fractionation
2 is negligible in the vertical position was not correct. The difference from the fractionation factor of less
3 than 0.080 L min^{-1} in the vertical position is reasonable, although the absolute fractionation factors need to
4 be revised based on the fractionation factors obtained using the mass spectrometer.

5 **4 Discussion**

6 In actual atmospheric observation, the standard gas mixture is used intermittently rather than continuously,
7 whereas the results in this study are based on decanting experiments in which the CO_2 -in-air mixture was
8 used continuously. Therefore, it is necessary to confirm that adsorption and thermal diffusion effects are
9 equivalent between continuous and intermittent use of standard gas mixtures, to be able to discuss how to
10 operate the standard gas mixtures taking into account the results from this study. Schibig et al. (2018)
11 reported that the CO_2 desorption energy (E_d) from an aluminum cylinder inner surface was 10 kJ mol^{-1} ,
12 meaning that the only adsorption mechanism for CO_2 on the inner wall of the cylinder is physisorption. The
13 desorption lifetime τ on the inner surface of the cylinder is expressed by the following Arrhenius-type
14 equation (Arrhenius, 1889a, b; Laidler, 1949; Frenkel, 1924; Laidler et al., 1940):

$$15 \quad \tau = \frac{1}{A_{des}} \times e^{E_d/RT} \quad (4)$$

16 where A_{des} is a pre-exponential factor (10^{12} s) (Knopf et al., 2024), and R and T represent the gas constant
17 ($8.314 \text{ J K}^{-1} \text{ mol}^{-1}$) and room temperature (298 K), respectively. Using these values, τ is calculated to be
18 $6 \times 10^{-11} \text{ s}$. Because the desorption lifetime is sufficiently shorter than the pressure change rate of $1.4 \times$
19 $10^{-5} \text{ MPa s}^{-1}$, the CO_2 on the inner surface and in the standard gas mixture is estimated to have always been
20 in equilibrium over the experiments in this study. The adsorption/desorption effect would be comparable
21 for intermittent and continuous use. However, thermal diffusion fractionation could differ between
22 intermittent and continuous use if the thermal distribution in the cylinder takes a long time to reach
23 equilibrium. The equilibrium time for the temperature distribution can be estimated from the time it takes
24 for the CO_2 value to stabilize; in the experiment in Sect. 3.2.1, the temperature distribution reaches



1 equilibrium within a few minutes even when the flow rate of the outflowing gas is changed. Because actual
2 measurements of standard gas mixtures are carried out continuously over several tens of minutes, which is
3 longer than the equilibrium time for the thermal distribution, it can be estimated that even intermittent use
4 is not markedly different from continuous measurements. Hence, we discuss how to operate the standard
5 gas mixtures based on the results of this study.

6 Thermal diffusion fractionation has been demonstrated to have diluted the CO₂ molar fraction in the
7 horizontal cylinder and to have enriched the molar fraction in the vertical cylinder as the pressure dropped.
8 This effect also increased as the outflowing gas flow rate increased, although the adsorption/desorption
9 effect was constant. These findings are consistent with the results of previous studies (Schibig et al., 2018;
10 Aoki et al., 2022). Furthermore, information on the mechanism of thermal diffusion fractionation was
11 obtained from the temperature changes at the top, middle, and bottom of the cylinder monitored in the
12 decanting experiments performed with horizontally and vertically positioned cylinders at a flow rate of 6 L
13 min⁻¹. The temperature difference between the top, middle, and bottom of the cylinder was negligible when
14 the cylinder was horizontal, but the temperatures at the top and bottom were 0.3 K and -0.4 K higher than
15 that in the middle when the cylinder was vertical (Fig. 3). The offset of the outflowing gas was 0.069 μmol
16 mol⁻¹ for the horizontal cylinder and -0.062 μmol mol⁻¹ for the vertical cylinder (Table 2). The detected
17 offset was estimated to be driven by a temperature difference of 0.9 K, computed using the thermal diffusion
18 coefficient reported by Severinghaus et al. (1996). Because the thermal conductivity of the aluminum
19 cylinder is higher than that of the internal gas, the measured temperature difference of the cylinder is
20 expected to be smaller than the actual temperature difference of the gas mixture. The temperature difference
21 of 0.3 K between the top and middle of the vertical cylinder appears to support the validity of the calculated
22 temperature difference of 0.9 K. These results mean that the outflowing gas would have been taken out
23 from the warmer gas at the cylinder top of the vertical cylinder, although the temperature distribution
24 causing thermal diffusion fractionation could not be determined for the horizontal cylinder. However, the
25 contribution of thermal diffusion fractionation has been understood to be negligible at low flow rates in



1 previous studies (Schibig et al., 2018; Hall et al., 2019; Aoki et al., 2022). Notwithstanding, even at a flow
2 rate of 0.080 L min^{-1} , which is within the usual range of flow rates used by observation laboratories, we
3 found that thermal diffusion fractionation produced offsets of the CO_2 molar fractions of 0.013 ± 0.015
4 $\mu\text{mol mol}^{-1}$ (horizontal cylinder) and $-0.014 \pm 0.011 \mu\text{mol mol}^{-1}$ (vertical cylinder) in the outflowing gases.
5 These offsets are driven by temperature difference in the cylinder as small as 0.18 K and indicate that a
6 difference of $0.027 \mu\text{mol mol}^{-1}$ can be produced simply by changing the cylinder from horizontal to vertical.
7 Measuring standard gas mixtures while keeping the cylinder in the same position will be an effective means
8 of maintaining the long-term consistency of observed values, because it is difficult to completely suppress
9 the occurrence of such small temperature differences.

10 Furthermore, the CO_2 deviation resulting from the adsorption/desorption effect and thermal diffusion
11 effects as pressure dropped were verified using the results of the decanting experiment at a flow rate of
12 0.080 L min^{-1} in Fig. 2. The CO_2 deviation due to thermal diffusion fractionation as the pressure dropped
13 was calculated by substituting fractionation factors of 1.000030 ± 0.000037 for a horizontally positioned
14 cylinder and 0.999968 ± 0.000027 for a vertically positioned cylinder into the Rayleigh function (Eq. (2))
15 (orange shading in Fig. 8). The total CO_2 deviation was estimated by fitting the Langmuir–Rayleigh model
16 (Eq. (3)) to the results of the decanting experiment at a flow rate of 0.080 L min^{-1} . The CO_2 deviation due
17 to adsorption/desorption was calculated by subtracting the thermal diffusion fractionation deviation from
18 the total CO_2 deviation. The contributions to the total change were 60% (vertical cylinder) and 70%
19 (horizontal cylinder) for adsorption/desorption, and those of thermal diffusion fractionation were 40%
20 (vertical cylinder) and 30% (horizontal cylinder) (Fig. 8). Here, the CO_2 deviation from the initial value
21 due to adsorption/desorption at a relative pressure of 0.03 was $\sim 0.1 \mu\text{mol mol}^{-1}$ for a horizontally positioned
22 cylinder and $\sim 0.06 \mu\text{mol mol}^{-1}$ for a vertically positioned cylinder. This difference is assumed to be
23 uncertainty because the contribution of adsorption/desorption should be constant regardless of the cylinder
24 position. That is, it is necessary to understand that this estimated contribution contains a large uncertainty.



- 1 The contribution of adsorption/desorption is larger than that of thermal diffusion, but fractionation due to
- 2 thermal diffusion is not negligible.

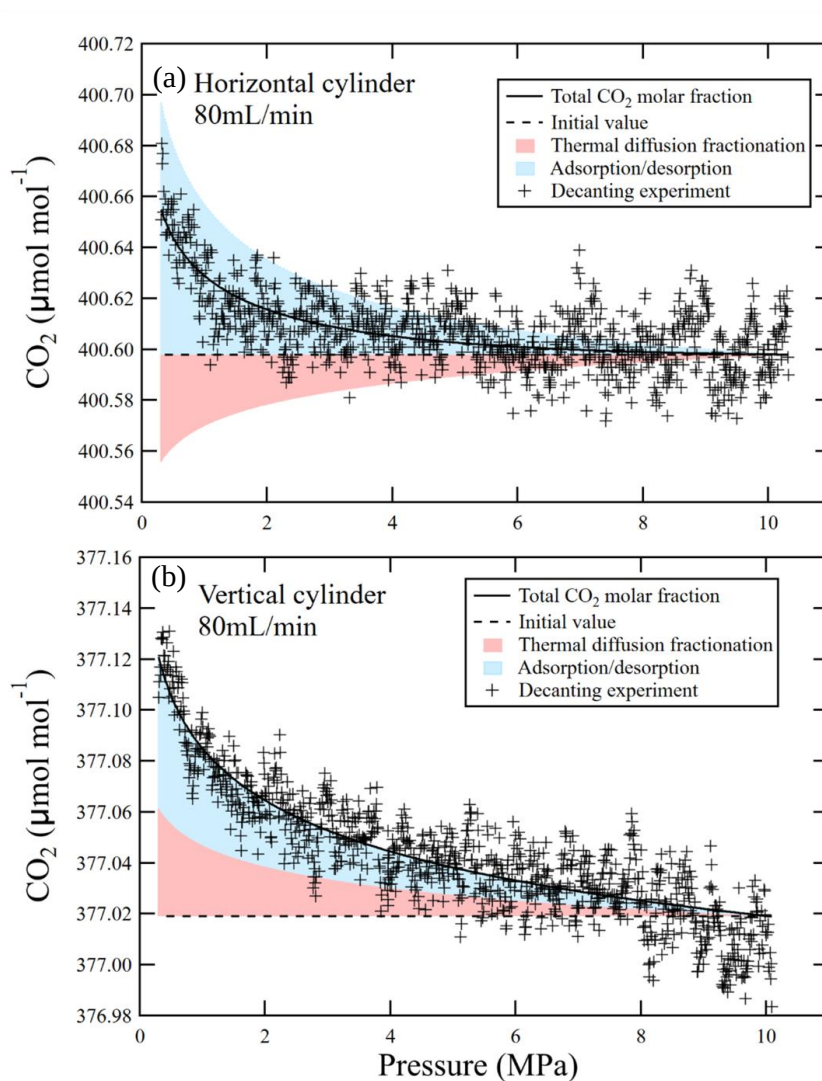


Figure 8. Total CO₂ deviation, deviation due to adsorption/desorption, and deviation due to thermal diffusion fractionation relative to the initial value for a CO₂-in-air mixture leaving (a) a horizontally positioned cylinder and (b) a vertically positioned cylinder at an outflowing gas flow rate of 0.080 L



1 The WMO recommends that calibration standard gas mixtures of CO₂ should be replaced once the cylinder
2 pressure has dropped to 2 MPa (WMO report No. 292). Leuenberger et al. (2015) and Schibig et al. (2018)
3 recommended that the usage of standard gas mixtures should be restricted to pressures above 3 MPa to
4 remain within the WMO's compatibility goal of 0.1 $\mu\text{mol mol}^{-1}$ for the northern hemisphere and 0.05 μmol
5 mol^{-1} for the southern hemisphere. However, the CO₂ enrichment shown in Fig. 4b depends only on relative
6 pressure, not absolute pressure, suggesting that determining the minimum operating pressure by considering
7 the absolute pressure is not efficient. For example, if the initial pressure is low, the standard gas mixture
8 will be replaced at a pressure at which it should have been usable, resulting in waste of the standard gas
9 mixture. If the initial pressure is high, the standard gas mixture will not be replaced at the pressure at which
10 it should be replaced, leading to poor consistency because of overestimation or underestimation of the
11 observed values. Therefore, we recommend that the WMO's compatibility goal should be modified so that
12 laboratories use the relative pressure as a criterion. If the CO₂ mole fraction is allowed to increase to 0.05
13 mol mol^{-1} , the standard gas mixture should be replaced when the cylinder pressure drops to 3 MPa, 2 MPa,
14 or 1 MPa for initial pressures of 15 MPa, 10 MPa, and 5 MPa, respectively. In this way, the standard gas
15 mixture can be used efficiently without waste.

16 The question arises as to whether the cylinder should be positioned horizontally or vertically during
17 measurement. From Fig. 5, it appears to be best to operate a horizontally positioned cylinder with an
18 outflowing gas flow rate of between 0.15 L min⁻¹ and 0.30 L min⁻¹, because it may not be necessary to pay
19 attention to deviations of CO₂ levels in the cylinders when taking out the standard gas mixture. However,
20 a lower flow rate such as 0.080 L min⁻¹ may be desirable if the same set of standard gas mixtures is used
21 for a long time. In that case, a correction would be necessary to ensure long-term consistency of the CO₂
22 molar fraction because the CO₂ deviation as the pressure drops cannot be ignored. Our results showed that
23 the relative pressure determines the amount of CO₂ deviation, provided that the cylinder position and the
24 gas outflow rate are constant. Therefore, CO₂ deviation could be corrected for by determining in advance
25 the relationship between CO₂ deviation and relative pressure and the flow rate of the outflowing gas. It



1 should be noted, however, that this method is for correction of the CO₂ molar fraction in the cylinder, not
2 correction of the positive and negative offsets in outflowing gases. The offset values should be corrected
3 using a fractionation factor as determined in Sect. 3.2.3. However, this correction may not be very useful,
4 because as the offsets would be at the same level as the measurement uncertainty. Indeed, it is important to
5 note that the atmospheric CO₂ molar fraction is difficult to determine with an uncertainty of less than 0.01
6 $\mu\text{mol mol}^{-1}$ due to the thermal diffusion effect.

7 From the above discussion, the standard gas mixture should be operated during observation as follows.

- 8 1. The flow rate of outflowing gas from the cylinders should be as low as possible to reduce the
9 contribution of thermal diffusion fractionation.
- 10 2. Throughout the observation, cylinders should be used in either a horizontal or a vertical position, and
11 the position of the cylinders should not be altered.
- 12 3. The cylinder pressure at which a standard gas mixture should be changed should be determined based
13 on the relative pressure.

14

15 5 Conclusions

16 We attempted to quantitatively estimate the factors that cause the CO₂ molar fraction in a cylinder to deviate
17 as the pressure drops, to facilitate a shift from the use of standard gas mixtures based on empirical
18 knowledge to use based on theoretical understanding. We found that the CO₂ mole fraction in the cylinder
19 changes from the initial value as a result of thermal diffusion fractionation as well as adsorption and
20 desorption. We found that thermal diffusion fractionation operates even at low gas outflow rates, for which
21 adsorption/desorption effects had been considered to be the main cause. A further important finding was
22 that this deviation of the CO₂ molar fraction is independent of the initial pressure and depends on the relative
23 pressure rather than the absolute pressure. Our results demonstrate the necessity for a new way of operating
24 cylinders that is different from the conventional empirical knowledge of the use of standard gases.



1 Furthermore, long-term consistency of values will be ensured by correcting for deviations in the CO₂ molar
2 fraction due to the pressure drop. Thus, this study is an important contribution to ensuring the consistency
3 of observed values, which has been a concern in long-term CO₂ observations.

4 **Data availability.** The data presented in this article are available upon request to Nobuyuki Aoki (aoki-
5 nobu@aist.go.jp).

6 **Author contributions.** NA designed the study. NA performed the experiment and prepared the first draft.
7 SI performed mass spectrometry measurements. Both authors contributed to the preparation of the final
8 version of the manuscript.

9 **Competing interests**

10 The authors declare that they have no conflict of interest.

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