TEOS-10 and the Climatic Relevance of Ocean-Atmosphere Interaction

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Abstract: Unpredicted observations in the climate system, such as recently an excessive ocean warming, are often lacking immediate causal explanations and are challenging the numerical models. As a highly advanced mathematical tool, the Thermodynamic Equation of Seawater – 2010 (TEOS-10) had been established by international bodies as an interdisciplinary standard and is recommended for use in geophysics, such as especially in climate research. From its very beginning, the development of TEOS-10 was supported by *Ocean Science* through publishing successive stages and results. Here, the history and properties of TEOS-10 are briefly reviewed. With focus on the air-sea interface, selected current problems of climate research are discussed and tutorial examples for the possible use of TEOS-10 in the associated context are presented, such as related to ocean heat content, latent heat and rate of marine evaporation, properties of sea spray aerosol, or climatic effects of low-level clouds. Appended to this article, a list of publications and their metrics is provided for illustrating the uptake of TEOS-10 by the scientific community, along with some continued activities, addressing still pending, connected issues such as uniform standard definitions of uncertainties, of relative humidity, seawater salinity or pH.

This article is dedicated to the Jubilee celebrating 20 years of Ocean Science.

This article is also dedicated to the memory of Wolfgang Wagner who sadly and unexpectedly has passed away on 12 August 2024. His contributions to TEOS-10 are truly indispensable constituents; Wolfgang was an essential co-author of various related documents and articles. He will deeply be missed.

All the rivers run into the sea; yet the sea is not full;
unto the place from whence the rivers come, thither they return again.
The King James Bible: Ecclesiastes, 450 – 150 BCE

He wraps up the waters in his clouds,
yet the clouds do not burst under their weight.
Holy Bible: New International Version, Job 26:8

Of the air, the part receiving heat is rising higher.
So, evaporated water is lifted above the lower air.
Leonardo da Vinci: Primo libro delle acque, Arundel Codex, ca. 1508

Two-thirds of the Sun's energy falling on the Earth's surface is needed
to vaporize ... water ... as a heat source for a gigantic steam engine.
Heinrich Hertz: Energiehaushalt der Erde, 1885

The sea surface interaction is obviously
a highly significant quantity in simulating climate.
Andrew Gilchrist, Klaus Hasselmann: Climate Modelling, 1986

The climate of the Earth is ultimately determined
by the temperatures of the oceans.
Donald Rapp: Assessing Climate Change, 2014

1 Introduction

Quite recently in 2024, climate research has published alarming news: “The world’s oceans absorbed more heat in 2023 than in any other year since records began in the 1950s. … Data show that the heat stored in the upper 2,000 metres of oceans increased by 15 zettajoules (1 zettajoule is 10^{21} joules) in 2023 compared with that stored in 2022. This is an enormous amount of energy — for comparison, the world’s total energy consumption in 2022 was roughly 0.6 zettajoules” (You 2024: p. 434). Dividing this value by the global ocean surface area and by the duration of a year, the reported ocean’s average warming rate amounts to 1.3 W m^{-2}, and is apparently even increasing. “Earth’s net global energy imbalance (12 months up to September 2023) amounts to +1.9 W m^{-2}, … ensuring further heating of the ocean” (Kuhlbrodt et al. 2024: p. E474). „Climate models struggle to explain why planetary temperatures spiked suddenly. … No year has confounded climate scientists’ predictive capabilities more than 2023. … This sudden heat spike greatly exceeds predictions made by statistical climate models” (Schmidt 2024: p. 467).

The currently observed *ocean heat content* (OHC) represents a merely transient maximum after a decade-long systematic warming process in the past, see Fig. 18 in **Section 6**, which may proceed to even higher values in the future. In **Section 3**, thermodynamic aspects of related OHC definitions will be considered. Regarding the long-term period since 1971, “the drivers of a larger Earth energy imbalance in the 2000s than [before] are still unclear. … Future studies are needed to further explain the drivers of this change” (von Schuckmann et al. 2023: p. 1694). Laterally, the observed heat excess is unevenly distributed over the world ocean (Fig. 1), in contrast to what naively may be expected from rising atmospheric CO_2 concentrations. Rather, warming seems to be most pronounced in the cloudy austral and boreal west-wind belts. Selected thermodynamic relations between OHC and cloudiness are briefly discussed in **Section 6**.

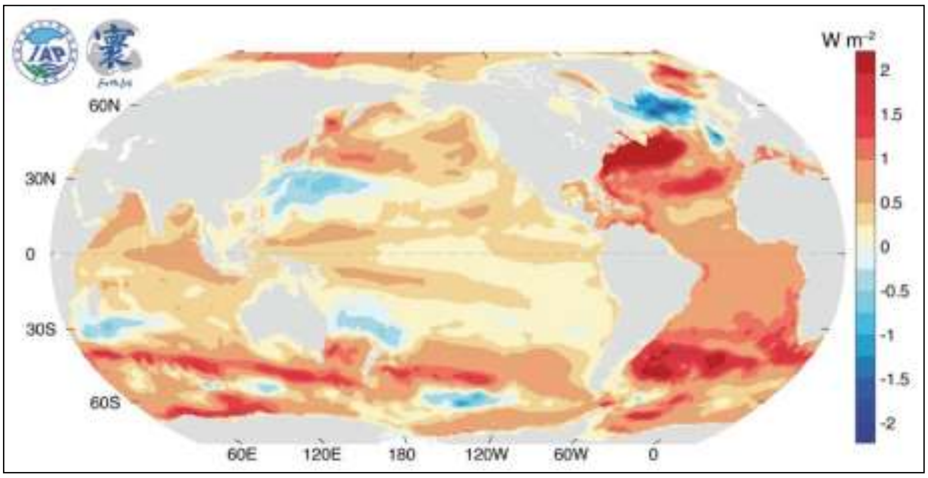

Fig. 1: Observed trend 1958 through 2022 of the upper 2000 m ocean heat content (WMO 2024). Image reproduction permitted by WMO Copyright.

Sunlight is the only available heat source of sufficient power to cause the observed warming, while the globally averaged geothermal heat flux is estimated to be just 0.087 W m^{-2} (Pollack et al. 1993), and is not expected to suddenly rise recently due to human impact. Irradiation is hampered by clouds, dust and absorbing gases, and water surface reflection such as by whitecaps, waves or plankton layers (Cahill et al. 2023). Heat absorbed in the water column may effectively exit the ocean again only across the air-sea interface via sensible, radiative and latent heat flux. All these effects may vary in the climate system in a complicated, mutually interacting manner. Typically, present numerical climate models suffer from an "ocean heat budget closure problem" (Josey et al. 1999) and describe the ocean-atmosphere heat flux only to within uncertainties between 10 W m^{-2} and 30 W m^{-2} (Josey et al. 2013). According to recent model comparison studies, many of those "models fail to provide as much heat into the ocean as observed" (Weller et al. 2022: p. E1968). Dynamical models, rather than observed correlations, are the most reliable tools for the detection and verification of causal relations (Feistel 2023), however, such as in this case of air-sea interaction, large uncertainties may prevent any significant conclusions to be drawn regarding the causes of the observed ocean warming rate of 1.3 W m^{-2}.

Of the increasing amount of water vapour contained in the global troposphere, 85 % results from ocean evaporation (Gimeno et al. 2013). Corresponding to 1200 mm annual evaporation (Budyko 1963, 1984, Baumgartner and Reichel 1975, Peters-Lidard et al. 2019), the associated latent heat flux of about 95 W m^{-2} per ocean surface area represents the strongest energy supply for the atmospheric dynamics (Albrecht 1940) and at the same time the strongest cooling process of the sea. This flux depends sensitively on the relative humidity (RH) at the water surface; an RH increase by 1 %rh can be estimated to reduce evaporation by 5 W m^{-2} (Feistel 2015, 2024, Feistel and Hellmuth 2021, 2023), so that minor additional 0.2 %rh may already suffice to warm up the ocean by the observed 1.3 W m^{-2}. Unfortunately, marine RH is observed only with uncertainties between 1 and 5 %rh (Lovell-Smith et al. 2016), or, accordingly, between 5 and 25 W m^{-2} of latent heat flux, which is roughly corresponding to unknown variations ranging up to 50 … 250 mm evaporation per year. It remains unclear to what extent minor, yet unnoticed changes in marine RH may be responsible for the recent ocean warming.

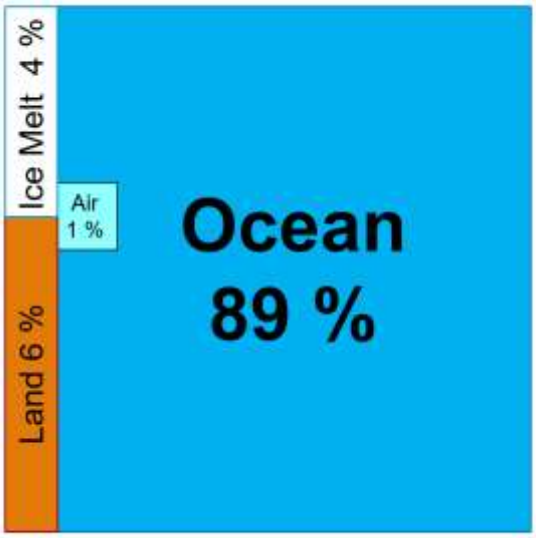

Fig. 2: Heat fractions stored additionally in the different parts of the Earth system 1960–2020 (values from von Schuckmann et al. 2023), represented graphically by partial areas. Obviously, the oceans dominate global warming.

According to Fig. 2, a paramount share of 94 % of global warming occurs in different phases and geophysical mixtures of water, in particular in seawater. Considering this situation, the *Scientific Committee on Oceanic Research* (SCOR) in cooperation with the *International Association for the*

Physical Sciences of the Oceans (IAPSO) decided at its 2005 Cairns meeting the establishment of the *SCOR/IAPSO Working Group 127 on Thermodynamics of Seawater* (WG127) (Millero 2010, Pawlowicz et al. 2012, Smythe-Wright et al. 2019), which held its inaugural meeting in 2006 at Warnemünde (Fig. 3). It had been recognised that “modelling of the global heat engine needs accurate expressions for the entropy, enthalpy, and internal energy of seawater so that heat fluxes can be more accurately determined in the ocean” (Millero 2010: p. 28) while such properties were not available from the thermodynamic seawater standard at that time, the 1980 Equation of State of Seawater (EOS-80) (Fofonoff and Millard Jr. 1983).

Fig. 3: Participants of the 2006 kick-off meeting of SCOR/IAPSO WG127 at the Leibniz Institute for Baltic Sea Research (IOW) in Warnemünde, Germany. From left to right: Chen-Tung Arthur Chen (Taiwan), Frank Millero (USA), Brian King (UK), Rainer Feistel (WG vice chair, Germany), Daniel Wright (Canada, deceased 2010), Trevor McDougall (WG chair, Australia) and Giles Marion (USA).

Fig. 4: Participants of the BIPM-IAPWS meeting in February 2012 at the Pavillon de Breteuil, Sèvres. From left to right: Dan Friend (IAPWS), Karol Daučik (IAPWS president), Jeff Cooper (IAPWS), Alain Picard (BIPM, deceased 2015), Petra Spitzer (WG127), Rainer Feistel (WG127), Michael Kühne (director BIPM), Andy Henson (BIPM) and Robert Wielgosz (BIPM).

The foundation of WG127 happened almost coincidently with the establishment of the *Ocean Science* journal of the *European Geosciences Union* (EGU) in 2004/05. The development of the new standard by WG127, the Thermodynamic Equation of Seawater – 2010 (TEOS-10) was very successfully supported by Ocean Science, publishing the Special Issue #14 on “Thermophysical Properties of Seawater” with 16 articles between 2008 and 2012 (Feistel et al. 2008a). **Appendix A** reports the current metrics of this Special Issue. Also in 2008, at its conference in Berlin, Germany, the *International Association for the Properties of Water and Steam* (IAPWS) established a new *Subcommittee on Seawater* (SCSW) that cooperated closely with WG127. In the form of carefully verified mathematical formulations for properties of water, ice, seawater and humid air, IAPWS adopted nine fundamental documents related to TEOS-10 (IAPWS AN6-16 2016), see **Appendix A**.

With respect to problems yet pending after the official adoption of TEOS-10, especially for the preparation of future novel international definitions of seawater salinity, seawater pH and atmospheric relative humidity (Feistel et al. 2016, Pawlowicz et al. 2016, Dickson et al. 2016, Lovell-Smith et al. 2016), the standing *IAPSO/SCOR/IAPWS Joint Committee on the Properties of Seawater* (JCS) was established in 2012. In 2011, IAPWS also extended its cooperation with the *International Bureau for Weights and Measures* (BIPM), see Fig. 4. Further details on TEOS-10 (IOC et al. 2010, McDougall et al. 2013, Feistel 2018, Wikipedia 2024) are available from the TEOS-10 homepage, www.teos-10.org, and are briefly reviewed in **Section 2** and **Appendix B**.

In the context of the predecessor EOS-80, the *ocean heat content* (OHC) was defined in terms of *potential temperature* (Abraham et al. 2013). Improving this method, TEOS-10 entropy and enthalpy of seawater provided a proper quantitative basis for a novel, thermodynamically rigorous definition of the OHC in the form of seawater *potential enthalpy* (McDougall 2003, McDougall et al. 2013, Graham and McDougall 2013, McDougall et al. 2021), equivalently defined as *Conservative Temperature* and briefly discussed in **Section 3**.

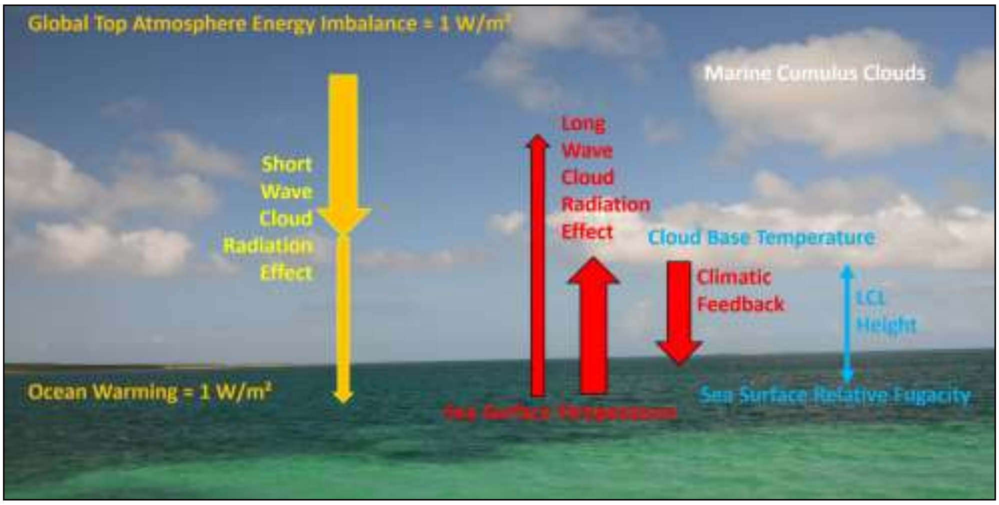

Fig. 5: Schematic of *cloud radiation effects* (CRE). The *short-wave effect* (SW CRE) controls the downward flux of solar irradiation while the *long-wave effect* governs the infrared radiation balance between water surface and cloud base. By thermal convection, cumulus clouds emerge at the isentropic *lifted condensation level* (LCL). Figure from Feistel and Hellmuth (2024b)

Currently implemented parameterisations of marine evaporation rates in the form of historical *Dalton equations* (Stewart 2008, Josey et al. 1999, 2013) may be replaced by TEOS-10 chemical potentials which provide the proper quantitative basis for a thermodynamically rigorous formulation of non-equilibrium Onsager forces and fluxes in terms of *relative fugacity* (RF) of humid air (Kraus and Businger 1994, Feistel and Lovell-Smith 2017, Feistel and Hellmuth 2023, 2024a), as described in **Section 4.** *Relative humidity* (RH) is defined relative to the saturation state of moist air, which in turn is controlled by the chemical potentials of water in the gas and liquid phase. It is only natural, therefore, to define RH in terms of chemical potentials, which in fact is performed by RF. The uncertainty of latent heat flux with respect to the uncertainty of surface RH observation is shown to be significantly larger than the observed warming of $1.3\ \mathrm{W\ m^{-2}}$, so that this unpredicted warming may or may not be caused by so-far ignored minor RH increase.

The conceptual model of sea air as a two-phase composite thermodynamic system is outlined in **Section 5**. The roles of enthalpy, chemical potential and entropy are explained by means of explicit theoretical descriptions of three simplified tutorial examples, (i) for the latent heat of evaporation, (ii) for the heat capacity of humid air containing sea spray, and (iii) for the entropy production of irreversible evaporation.

Clouds do not only release the latent heat which water vapour has carried away from the ocean, they also interfere substantially in the global radiation balance, cooling the surface by reflecting short-wave solar irradiation, and warming the surface by sending back down long-wave thermal radiation, see Fig. 5. In the course of global warming, cloudiness has been found to exhibit a systematic trend of reduction, see **Section 6**, which affects the ocean heat content in a non-trivial, non-uniform manner. Marine cumulus clouds arise by isentropic uplift of thermal convection. Their height controls their temperature and their thermal downward radiation, affecting the ocean's energy balance. Updating previous results (Romps 2017) for the *lifted condensation level* (LCL) of marine cumulus clouds to thermodynamically rigorous TEOS-10 standard equations (Feistel and Hellmuth 2024b), the radiative effect of those clouds can be estimated from *sea-surface temperature* (SST) and surface relative humidity. This effect turns out to be weakly cooling and cannot provide a reasonable explanation for the so-far unclear strong ocean warming. The effect of increasing SST in the past decades turns out to be minor in comparison to that caused by RH uncertainty.

Section 7 provides a summery of this paper, **Appendix A** reports collections of publications with respect to TEOS-10 as well as their metrics, and **Appendix B** gives a short introduction into the concept of thermodynamic potentials.

2 Thermodynamic Equation of Seawater – 2010 (TEOS-10)

In the climate system, the omnipresent and dominant substance is water in various phases and mixtures. For example, "water vapor is by far the most important greenhouse gas, in the sense that it absorbs more irradiance from the Earth than all other greenhouse gases combined" (Rapp 2014: p. 381). Textbooks and other publications offer numerous collections of various different property equations for water, ice, seawater or moist air, but uncertainties and mutual consistencies of those equations are often unclear. To improve this situation, a novel *Thermodynamic Equation of Seawater – 2010* (TEOS-10) was developed by the members of the SCOR/IAPSO Working Group 127 (WG 127) in close cooperation with the International Association for the Properties of Water and Steam (IAPWS). TEOS-10 is described in a detailed Manual (IOC et al. 2010) and has been adopted and recommended by IOC-UNESCO (2009) in Paris and by the IUGG (2011) in Melbourne, see also Feistel (2008b, 2012, 2018), Valladares et al. (2011) and Pawlowicz et al. (2012). Starting in 2008 with a Special Issue of *Ocean Science* (Feistel et al. 2008a), a large number of scientific publications has appeared in the meantime, supporting, extending or exploiting TEOS-10. A collection of selected papers related to TEOS-10 is summarised in Appendix A together with metrics that illustrate the growing uptake of TEOS-10 by the scientific community.

The development of the first numerical thermodynamic Gibbs potentials (see Appendix B) for seawater (Feistel 1991, 1993, Feistel and Hagen 1995) was based on the works of Millero and Leung (1976) and Millero (1982, 1983), together with high-pressure background data of the previous EOS-80 standard (Unesco 1981). Independently of that, a Helmholtz potential for pure fluid water had been adopted by IAPWS in 1996 at Fredericia (Harvey 1998, Wagner and Pruß 2002). These were the key activities which eventually culminated in the formulation of TEOS-10 about two decades later. By combining those equations for pure and seawater, some known pending problems of EOS-80 (Fofonoff and Millard Jr. 1983) could incidentally be resolved (Feistel 2003). In the end, TEOS-10 has been assembled from four basic thermodynamic potentials derived from mutually consistent, most comprehensive and accurate datasets of measured properties available at that time. Those potentials are (IAPWS R6-95 2016, IAPWS R10-06 2009, IAPWS R13-08 2008, IAPWS G8-10 2010, respectively):

(i) A Helmholtz function of fluid water, $f^{\mathrm{F}}(T,\rho) \equiv f^{\mathrm{W}}(T,\rho) \equiv f^{\mathrm{V}}(T,\rho)$, known as the IAPWS-95 formulation (Wagner and Pruß 2002), which is identical for liquid water, $f^{\mathrm{W}}(T,\rho)$ and for water vapour, $f^{\mathrm{V}}(T,\rho)$. It describes de-aerated water of a fixed isotopic composition, termed *Standard Mean Ocean Water* (SMOW), with density ρ and temperature T.

(ii) A Gibbs function of ambient hexagonal ice I, $g^{\mathrm{Ih}}(T,p)$, or IAPWS-06 formulation (Feistel and Wagner 2006), see Tables A2 and A3 of Appendix A, depending on pressure p.

(iii) A Gibbs function of *IAPSO Standard Seawater*, $g^{\mathrm{SW}}(S,T,p)$, or IAPWS-08 formulation (Feistel 2008a), see Tables A2 and A3 of Appendix A. The variable S, at which a subscript A is omitted here for simplicity, is the specific or *Absolute Salinity*, the mass fraction of dissolved salt in seawater, which differs from *Practical Salinity*, S_{P}, measured by present-day oceanographic instruments, as well as from various other obsolete salinity scales (Millero et al. 2008). Throughout this paper, the term “salinity” is exclusively short hand for TEOS-10 Absolute Salinity. Sea salt is assumed to have stoichiometric *Reference Composition*. The pure-water limit, $g^{\mathrm{SW}}(0,T,p) = g^{\mathrm{W}}(T,p)$, is the Gibbs function of liquid water computed from the IAPWS-95 Helmholtz function $f^{\mathrm{W}}(T,\rho)$. For brackish seawater, g^{SW} has implemented Debye’s root law of dilute electrolyte solutions (Landau and Lifschitz 1966, Falkenhagen et al. 1971).

(iv) A Helmholtz function of humid air, $f^{\mathrm{AV}}(A,T,\rho)$, or IAPWS-10 formulation (Feistel et al. 2010a), see Tables A1 and A2 of Appendix A. The variable A is the mass fraction of dry air admixed with water vapour, so that $q = 1 - A$ is the *specific humidity*. The dry-air limit $f^{\mathrm{AV}}(1,T,\rho) = f^{\mathrm{A}}(T,\rho)$ equals, up to modified reference-state conditions, the equation of state of Lemmon et al. (2000). The air-free limit $f^{\mathrm{AV}}(0,T,\rho) = f^{\mathrm{V}}(T,\rho)$ equals the IAPWS-95 Helmholtz function of water vapour. In f^{AV}, the interaction of water vapour with dry air is described by 2$^{\mathrm{nd}}$ and 3$^{\mathrm{rd}}$ virial coefficients.

Thermodynamic potentials include certain adjustable constants expressing the absolute energies and entropies of the particular substances, which are not available from measurement (Planck 1906, Feistel 2019b) and have, in turn, no effect on measurable properties derived from those potentials. In fact, among the comprehensive experimental data sets from which the TEOS-10 equations were derived, none of those are suitable for fitting the empirical coefficients that represent absolute energies and entropies of those equations. For this reason, the International Conference on the Properties of Steam at London defined in 1967 the common triple point of water as the reference state at which those absolute values were arbitrarily set. Since then, no evidence has appeared for putative conflicts caused by such settings with any technical or scientific applications of the equations. Despite this, Feistel and Wagner (2006) and Feistel et al. (2008b) discuss the implementation of alternative residual entropies of water, if that should be of interest in exceptional applications of TEOS-10. For recent discussions of Pauling's absolute "residual" entropy at zero kelvin and Nernst's Third Law of thermodynamics, see Kozliak and Lambert (2008), Gutzow and Schmelzer (2011), Takada et al. (2015), Schmelzer and Tropin (2018), Feistel (2019b), or Shirai (2023).

The TEOS-10 reference states (Feistel et al. 2008b, 2010a) are the triple point of water, $T_{\mathrm{TP}} = 273.16\ \mathrm{K}$, $p_{\mathrm{TP}} = 611.654\ 771\ \mathrm{Pa}$, where the conditions

$$\eta_{\mathrm{TP}}^{\mathrm{W}} = 0, \qquad e_{\mathrm{TP}}^{\mathrm{W}} = 0, \qquad (1)$$

are imposed, and the standard ocean state at Absolute Salinity, $S_{\mathrm{SO}} = 35.165\ 04\ \mathrm{g\ kg^{-1}}$, absolute temperature, $T_{\mathrm{SO}} = 273.15\ \mathrm{K}$, and absolute pressure, $p_{\mathrm{SO}} = 101\ 325\ \mathrm{Pa}$, with the conditions for sea salt,

$$\eta_{\mathrm{SO}}^{\mathrm{SW}} = 0, \qquad h_{\mathrm{SO}}^{\mathrm{SW}} = 0, \qquad (2)$$

and for dry air,

$$\eta_{\mathrm{SO}}^{\mathrm{A}} = 0, \qquad h_{\mathrm{SO}}^{\mathrm{A}} = 0. \qquad (3)$$

Here, η, e and h, respectively, are specific entropy, internal energy and enthalpy of water (superscript W), seawater (superscript SW) and dry air (superscript A). The TEOS-10 potential functions and properties derived thereby are numerically implemented in two different libraries, the Sea-Ice-Air (SIA) and the Gibbs-Seawater (GSW) libraries, see Table A4 in Appendix A.

The SIA library includes empirical coefficients only in the four fundamental potentials (Feistel 2010d, Wright et al. 2010). All other potential functions and properties are derived strictly by mathematical operations to ensure consistent results, even at the cost of low computation speeds as a result of stacked iteration procedures. All quantities are exclusively expressed in basic SI units such as kg, m, J or Pa. A more recent extension of SIA code is reported in Feistel et al. (2022) for the computation of relative fugacity.

The GSW library is tailored for oceanographic models, optimised in computation speed (Roquet et al 2015). For fast numerical evaluation, GSW procedures contain new empirical coefficients determined from the SIA library functions by regression. Units and variables are adjusted to common oceanographic practice such as pressure in decibars relative to surface pressure, or temperatures in °C. *Conservative Temperature* (CT) is used as a new preferred thermal variable. An additional thermodynamic potential has been constructed (McDougall et al. 2023) that supports the use of CT universally as an independent variable.

3 Potential Enthalpy and Ocean Heat Content (OHC)

Thermodynamically, the term "Ocean Heat Content" (OHC) is a sloppy wording. "Content" means a state quantity of a body or volume while, by contrast, "heat" is an exchange quantity rather than a state quantity. "We have … a right to speak of heat as a *measurable quantity*, … however, … we have no right to treat heat as a *substance*" (Maxwell 1888: p. 7). "The obsolete hypothesis of heat being a substance is excluded" (Sommerfeld 1988: p. 6). "Heat is not a substance! More formally: Heat is not a thermodynamic function of state" (Romer 2001: p. 107). This distinction is qualitatively fundamental (Feistel 2023). Physical conservation quantities such as energy or mass have the key property that the change of that quantity in a volume equals the flux of that quantity across the boundary (Landau and Lifschitz 1966, Glansdorff and Prigogine 1971), but this does not apply to "heat". For example, a heat engine receives a permanent net heat flux without getting permanently hotter. While asking how much "heat" is contained in the ocean may find ambiguous answers, it is well defined to say how much heat has entered or left the ocean across its boundary by a specified process that transfers the ocean from a certain state of reference to the current state of interest. In this section, based upon TEOS-10, related states and processes are described which may properly specify what is commonly termed OHC. This consideration intrinsically connects OHC with ocean-atmosphere exchange processes relevant for climate change.

Since a long time, measuring and calculating the ocean's "heat" has been a question of central interest to oceanography. Recently, this issue has become even more important and urgent in the context of climate change. "The total energy imbalance at the top of atmosphere is best assessed by taking an inventory of changes in energy storage. The main storage is in the ocean" (Abraham et al. 2013: p. 450). The conventional approach is a formally defined mathematical procedure based on potential temperatures. "Changes to ocean heat content (OHC) can be calculated from measurements of the temperature evolution of the ocean. The OHC is attained from the difference of the measured potential temperature profile and the potential temperature climatology. This difference is integrated over a particular reference depth (for instance, 700 m) and is multiplied by a constant ocean density reference and heat capacity" (Abraham et al. 2013: p. 468). However, in representing a kind of "heat substance", this OHC definition has no rigorous thermodynamic justification, and the relation to processes of ocean-atmosphere heat fluxes is not entirely clear. If a sea-air heat flux of 1 W m^{-2} warms up the atmosphere, by what rate exactly will that OHC decrease?

Making the seawater properties entropy and enthalpy quantitatively available, TEOS-10 has offered a thermodynamically improved option for defining OHC (McDougall et al. 2021), in the form of the integral over the ocean volume,

$$OHC = \int h^{\mathrm{SW}}(S,\eta,p_0)\,\rho^{\mathrm{SW}}(S,\eta,p)\mathrm{d}V\,. \tag{4}$$

Here, $h^{\mathrm{SW}}(S,\eta,p_0)$ is the *potential enthalpy* (McDougall 2003) relative to the surface pressure, p_0, and $\rho^{\mathrm{SW}}(S,\eta,p)$ is the in-situ mass density at the pressure p of a parcel with salinity and entropy equal to those before. This definition can be understood in terms of both, a specified process of heat exchange, and a reference state relative to which OHC is counted, as follows (Feistel 2024):

(i) A virtual **heat exchange process** supporting the definition (4) is sketched in Fig. 6. In turn, each ocean parcel with in-situ properties (S,η,p) is lifted to the surface pressure p_0, keeping its salinity and entropy constant. There, it reversibly exchanges heat, $\mathrm{d}h = T\mathrm{d}\eta$, with a measuring device until the parcel's entropy has reached a certain reference value, η_{ref}, while the parcel's salinity remains unchanged. Subsequently, the heat is reversibly put back to the parcel which is then returned to its original location. The work required to lift and lower the parcel is balanced because the parcel's thermodynamic state is exactly the same before and after the balanced reversible heat exchange across the surface. The "heat content" defined this way for a single parcel is added up then over all ocean parcels to result in its total OHC value.

(ii) The **reference state** relative to which OHC is measured may freely be specified at will, but beneficially be chosen with respect to its convenience or usefulness. In the case of eq. (4), the OHC reference state is zero potential enthalpy (or zero Conservative Temperature, McDougall 2003) of all ocean parcels.

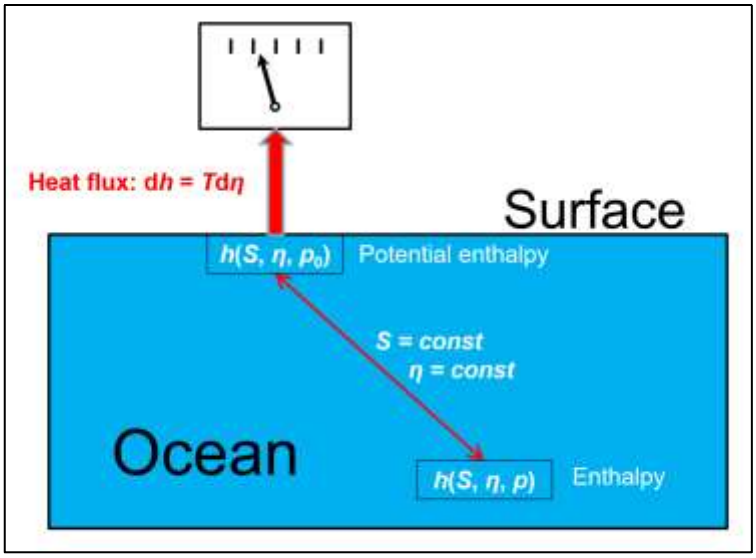

Fig. 6: Schematic of a conceptual process defining the ocean heat "content" (OHC) by measuring heat flux across the ocean boundary according to eqs. (4) and (5).

The process depicted in Fig. 6 measures the total heat flux $\int \mathrm{d}h = \int T\mathrm{d}\eta$ which changes the entropy of the given sample from the current value, η, to some arbitrary reference value, η_{ref}, and this way, the process also changes the parcel's enthalpy from $h^{\mathrm{SW}}(S,\eta,p_0)$ to $h^{\mathrm{SW}}(S,\eta_{\mathrm{ref}},p_0)$. Integration over all ocean samples results in an OHC value of

$$OHC^* = \int\left[h^{\mathrm{SW}}(S,\eta,p_0) - h^{\mathrm{SW}}(S,\eta_{\mathrm{ref}},p_0)\right]\rho^{\mathrm{SW}}(S,\eta,p)\mathrm{d}V. \quad (5)$$

While the choice of the OHC reference state is - in principle - entirely arbitrary, such as simply putting $\eta_{\mathrm{ref}} = 0$, it is reasonable to better adapt this selection to the purpose of the OHC definition. The main purpose of estimating OHC is keeping track of the ocean's long-term energy balance, in particular of the ocean's share of global warming. Three conditions appear immediately plausible in order to achieve this goal,

(i) *The OHC definition should ensure that OHC differences represent a suitable spatial integral over the heat fluxes crossing the ocean's boundaries*. As discussed in more detail in Section 5.3, production of entropy, $\mathrm{d_i}\eta$, caused by irreversible processes between different parcels within the ocean, does not affect the ocean's total enthalpy budget. This is quite in contrast to entropy exchange, $\mathrm{d_e}\eta$, of the given sample in the form of reversible heat flux across its boundary. Such irreversible processes affect the ocean's total potential enthalpy much less than its total entropy (McDougall et al. 2021). For this reason the OHC reference state should explicitly be defined in terms of potential enthalpy, $h^{\mathrm{SW}}(S,\eta_{\mathrm{ref}},p_0)$, and this way only implicitly in terms of entropy by specifying $\eta_{\mathrm{ref}}(S)$.

(ii) *Provided that the ocean's mass remains the same between any two ocean states (1) and (2), the difference OHC(1) – OHC(2) should depend only on the surface heat flux balance during the time in between*. *In particular, differences OHC(1) – OHC(2) should not depend on the OHC reference state.* For this reason, the OHC reference value should be independent of changes occurring in the density distribution, $\rho^{\mathrm{SW}}(S,\eta,p)$. This can be achieved by assigning to each ocean parcel the same reference potential enthalpy,

$h^{\mathrm{SW}}(S, \eta_{\mathrm{ref}}, p_0) = \mathrm{const}$, even though such a state may hardly ever be observed in the real ocean.

(iii) *Quantitatively, OHC values estimated at different times or places should be mutually comparable without estimation bias resulting from possibly changing methods of OHC calculation*. For this reason, resulting OHC values should be independent of the inevitable arbitrary, physically irrelevant reference-state conditions imposed on energy and entropy, such as eqs. (1)-(3). This can be achieved by assigning to each ocean parcel the same standard-ocean enthalpy as its reference potential enthalpy, $h^{\mathrm{SW}}(S, \eta_{\mathrm{ref}}, p_0) = h_{\mathrm{SO}}$. In the special case of TEOS-10 enthalpy, this value is defined by eq. (2), $h_{\mathrm{SO}} = 0$. This choice is implicitly made by the definition (4) but needed to be considered explicitly as soon as alternative equations for seawater enthalpy or entropy are employed, such as those of Millero and Leung (1976) and Millero (1982, 1983).

In a sense consistent with the previous OHC definition (Abraham et al. 2013), also a climatological average state could in principle be chosen as the OHC reference. However, this option includes the problem that the salinity distribution of the current ocean may differ from the reference ocean, and that thermodynamically properly treating the required salt exchange processes at the surface may turn the issue unnecessarily complicated. A detailed comparison of the OHC definition (4) with its precursor prior to TEOS-10 is provided by McDougall et al. (2021). OHC as a part of the total energy balance of the ocean is analysed by Tailleux (2010, 2018) and Tailleux and Dubos (2024).

4 Relative Fugacity and Ocean Evaporation Rate

"The global water cycle and the exchange of freshwater between the atmosphere and ocean is poorly understood. … It has been predicted that increasing global temperatures will lead to an enhanced global water cycle" (Holliday et al. 2011: p. 34). In the past, several climate researchers have argued that along with global warming the marine evaporation has or will be "amplified" or "intensified" (Feistel and Hellmuth 2021). However, it was not always made clear whether this may mean that (a) in the course of a year, more water vapour is transferred from the global ocean to the atmosphere, or (b) that the global mean evaporation rate remains unchanged while locally or temporally, evaporation is more intense, or (c) any combination of the two variants. Conclusions of kind (a) were drawn by renowned climatologists such as Budyko (1984), Flohn et al. (1992), Yu (2007), Randall (2012), Francis (2021) or Zhang et al. (2021).

By contrast, in favour of option (b), the currently observed ocean warming at a rate about 1 W m^{-2} does not support assumptions of an enhanced hydrological cycle with related latent-heat cooling, rather, it more likely suggests a slight reduction of evaporation. Two decades ago, Held and Soden (2006: p. 5687-5689) had already clearly stated that "it is important that the global-mean precipitation or evaporation, commonly referred to as the strength of the hydrological cycle, does not scale with Clausius–Clapeyron. … We can, alternatively, speak of the mean residence time of water vapor in the troposphere as increasing with increasing temperature." Subsequent observations have underpinned their statement.

Between 1979 and 2022, annual mean global precipitation values, see Fig. 7, fluctuated by about $\pm 10\ \mathrm{mm\ yr^{-1}}$, in particular due to La Niña events, but do not exhibit a significant long-term trend (Vose et al. 2023). Under the common assumption that global precipitation is balanced against evaporation, no substantial strengthening of the hydrological cycle may be observed yet.

Probably, the minor trend of 0.06 mm yr^{-2} of the data displayed in Fig. 7 is statistically insignificant. Associated with this apparent trend, the latent heat transferred to the troposphere can be estimated to a negligible putative warming rate of additional 0.5 mW m^{-2} per year, which could explain only 10 % of observed atmospheric warming by 1.7 °C per century (Morice et al. 2012, Feistel and Hellmuth 2021).

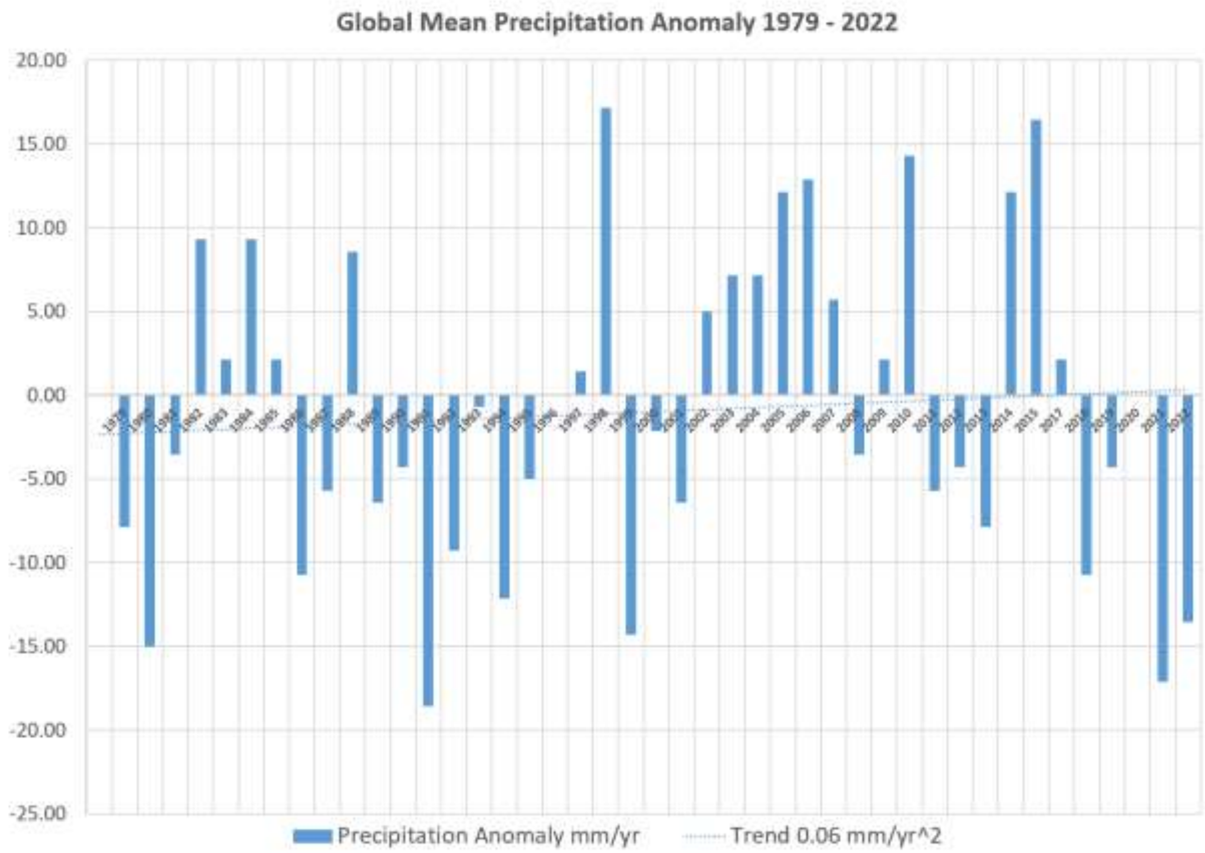

Fig. 7: Global mean precipitation anomaly 1979-2022 in mm yr^{-1}. The values displayed exhibit a minor increasing trend (dotted line) of 0.06 mm yr^{-2}. Data from Vose et al. (2023)

The thermodynamic driving force for evaporation is the difference between the chemical potentials of water in humid air and in seawater at the two sides of the sea-air interface (Kraus and Businger 1994). TEOS-10 has made this difference numerically available in the form of the water mass evaporation rate (Feistel and Hellmuth 2022, 2023)

$$J_{\mathrm{W}} = -D_f(u)\ln\frac{\psi_f}{x_{\mathrm{W}}}\,. \qquad (6)$$

Here, x_{W} is the mole fraction of water in seawater. Consistent with Wüst (1920), for the standard ocean with Reference Composition, this fraction is (Millero et al. 2008: Table 4),

$$x_{\mathrm{W}} = \frac{53.556\,514\,4}{54.676\,283\,8} = 0.979\,52, \qquad \ln x_{\mathrm{W}} = -0.0206926. \qquad (7)$$

In eq. (6), $D_f(u)$, the *Dalton coefficient*, is an empirical transfer coefficient as a function of the wind speed, u, as a parameterisation of the turbulent transport processes of water in the vicinity of the interface. Applications of the *Monin-Obukhov Similarity Theory* (MOST) in order to estimate the Dalton coefficient are reviewed by Liu et al. (1979), Foken and Richter (1991), Foken (2004, 2016) and in the Digital Supplement of Feistel and Hellmuth (2024). A review of empirical Dalton coefficients is given by Debski (1966); historical evaporation experiments are summarised by Biswas (1969).

In eq. (6), the sea-surface humidity is expressed by the *relative fugacity* (RF), ψ_f, defined by the ratio of the water-vapour fugacity in humid air, f_V, to that fugacity at saturation, $f_\mathrm{V}^\mathrm{sat}$ (Feistel and Lovell-Smith 2017), see eq. (49). In ideal-gas approximation, RF equals conventional RH (Lovell-Smith et al. 2016)

$$\psi_f \equiv \frac{f_\mathrm{V}}{f_\mathrm{V}^\mathrm{sat}} \approx \psi_x \equiv \frac{x}{x^\mathrm{sat}}\,. \tag{8}$$

Here, the mole fraction of water vapour in humid air is x, and its value at saturation is x^sat. Further, ψ_x is the conventional definition of RF in metrology and meteorology which, however, is inconsistent with alternative definitions such as the one employed in climatology (Lovell-Smith et al. 2016). Independent of ideal-gas conditions, but sufficiently close to saturation, such as near the sea surface, RF can be estimated in excellent approximation from the Clausius-Clapeyron formula (Feistel et al. 2022),

$$\psi_f \approx \exp\left\{\frac{L(T_\mathrm{dp},p)}{R_\mathrm{W}}\left(\frac{1}{T}-\frac{1}{T_\mathrm{dp}}\right)\right\}. \tag{9}$$

The evaporation enthalpy of pure water (IAPWS SR1-86 1992) at the dewpoint T_dp is L, and $R_\mathrm{W} = 461.523\ \mathrm{J\ kg^{-1}\ K^{-1}}$ is the specific gas constant of water. The typical marine RF is

$$\psi_f \approx 80\ \%\mathrm{rh}, \tag{10}$$

and is fairly independent of region, season or global warming (Dai 2006, Randall 2012, Rapp 2014, MetOffice 2020). Indeed, observed ocean surface RH has no significant climatological trend (Willett et al. 2023). Similarly, observed ocean wind speeds seem to be unaffected by global warming (Azorin-Molina et al. 2023). Eq. (6) for the evaporation rate depends only on wind speed and RF, so that it may be concluded that also the global mean evaporation rate has no significant climatic trend. In turn, as far as the release of latent heat is the main driving force of marine tropospheric dynamics, without increase of that release the mean wind speed is not expected to grow. “Latent heat is the main fuel that powers hurricanes, thunderstorms and normal bouts of lousy weather” (Francis 2021). Hence, the TEOS-10 approach in the form of eq. (6) appears to be consistent with the prediction of Held and Soden (2006) that the global evaporation does not increase along with temperature.

Various empirical evaporation equations, commonly known as *Dalton equations*, are found in the literature (Wüst 1920, Sverdrup 1936, 1937, Montgomery 1940, Debski 1966, Biswas 1969, Baumgartner and Reichel 1975). Several numerical climate models estimate evaporation from the formula (Stewart 2008, Pinker et al. 2014),

$$J_\mathrm{W} = D_q(u)(q_0 - q_{10}), \tag{11}$$

where q_0 is the specific humidity at the sea surface and q_{10} is that at 10 m height, or from (Josey et al. 1999, 2013)

$$J_\mathrm{W} = D_q(u)(0.98\, q^\mathrm{sat} - q). \tag{12}$$

Here, q is the near-surface specific humidity, and q^sat is the saturation value at the same temperature and pressure. The factor 0.98 accounts for the salinity, see eq. (7). After a few approximation steps (Feistel and Hellmuth 2023), these Dalton equations can be derived from the TEOS-version, eq. (6), however, there is an important qualitative difference. At constant RH, due to global warming, specific humidities such as q and q^sat, as well as their difference, are increasing following the Clausius-Clapeyron saturation formula. Accordingly, eq. (12) implies that also the evaporation rate J_{W} is growing this way, by contrast to eq. (6). This virtual acceleration of the hydrological cycle is evidently inconsistent with the prediction of Held and Soden (2006). This parameterisation-caused additional latent heat flux implies a spurious ocean cooling that may contribute to the finding that many numerical climate models tend to underestimate the observed ocean warming (Weller et al. 2022).

From eq. (6), the sensitivity of the latent heat flux, LJ_{W}, with respect to RH variations is easily estimated. For a mean evaporation rate of 1200 mm per year, the corresponding mass flux is about $J_{\mathrm{W}} \approx 3.8 \times 10^{-5}\ \mathrm{kg\ m^{-2}s^{-1}}$ and the related heat flux is $LJ_{\mathrm{W}} \approx 95\ \mathrm{W\ m^{-2}}$ with respect to the ocean surface area and a specific evaporation enthalpy of $L = 2\,501\ \mathrm{kJ\ kg^{-1}}$. At a surface humidity of $\psi_f = 0.8$, a value of $D_f(u) \approx 1.87 \times 10^{-4}\ \mathrm{kg\ m^{-2}s^{-1}}$ can be concluded for the mass transfer coefficient, and of $LD_f(u) \approx 468\ \mathrm{W\ m^{-2}}$ for that of latent heat. Then, from

$$\Delta(LJ_{\mathrm{W}}) = L\frac{\partial J_{\mathrm{W}}}{\partial \psi_f}\Delta\psi_f = -LD_f(u)\frac{\Delta\psi_f}{\psi_f} \tag{13}$$

it follows that an increase by $\Delta\psi_f = 1\ \%\mathrm{rh}$ results in a heat flux reduction by $\Delta(LJ_{\mathrm{W}}) = 5.85\ \mathrm{W\ m^{-2}}$. So, the currently observed ocean warming (Cheng et al. 2024) of $1.3\ \mathrm{W\ m^{-2}}$ could theoretically be caused already by a minor marine humidity increase of $\Delta\psi_f = 0.2\ \%\mathrm{rh}$, a value far below the present measurement uncertainty between 1 and 5 %rh of relative humidity. The resolution of climate models and observation seems to be insufficient yet to identify the possible role of RH for the unclear explanation of the warming ocean.

5 Sea Air as a Two-Phase Composite

Gibbs' (1873) method of using potential functions can be applied to any systems possessing stable thermodynamic equilibria and obeying energy conservation, without being restricted to merely homogeneous or single-phase samples. The intentionally strict mutual consistency of the different TEOS-10 potential functions permits a mathematical description of multi-phase composites such as sea ice, consisting of ice with included brine pockets (Feistel and Hagen 1998, Feistel and Wagner 2005), or clouds, where liquid water or ice is floating in saturated humid air (Hellmuth et al. 2021). Another important model is that of *sea air*, a sample consisting of a mass m^{SW} of seawater in thermodynamic equilibrium with a mass m^{AV} of humid air (Feistel et al. 2010d, Feistel and Hellmuth 2023). Such a model may serve as a mathematical description for certain thermodynamic properties of ocean-atmosphere interaction.

Extensive thermodynamic functions such as Gibbs energy or enthalpy are additive with respect to the two separate phases of the sample. Equilibrium between those parts requires equal temperatures and pressures. For this reason, a Gibbs function of sea air is an appropriate potential for the composite system with the TEOS-10 Gibbs functions $g^{\mathrm{SW}}(S,T,p)$ describing the liquid part and $g^{\mathrm{AV}}(A,T,p)$ the gas part. Let the masses of the substances in the parts be m^{W} of liquid water, m^{S} of dissolved salt, m^{A} of dry air and m^{V} of water vapour. Note that TEOS-10 neglects solubility of dry air constituents in liquid water. From combinations of the partial masses follow the liquid mass, $m^{\mathrm{SW}} = m^{\mathrm{S}} + m^{\mathrm{W}}$, the gas mass, $m^{\mathrm{AV}} = m^{\mathrm{A}} + m^{\mathrm{V}}$, the total mass $m = m^{\mathrm{SW}} + m^{\mathrm{AV}}$, the total water mass $m^{\mathrm{WV}} = m^{\mathrm{W}} + m^{\mathrm{V}}$, the salinity $S = m^{\mathrm{S}}/m^{\mathrm{SW}}$ and the dry-air fraction $A = 1 - q = m^{\mathrm{A}}/m^{\mathrm{AV}}$.

The Gibbs energies of the two phases of sea air are additive,

$$G^{\mathrm{SA}} = G^{\mathrm{SW}} + G^{\mathrm{AV}} = mg^{\mathrm{SA}}\,, \tag{14}$$

and, accordingly, the Gibbs function of sea air, g^{SA}, may be constructed from that of seawater, $g^{\mathrm{SW}}(S,T,p)$, with a liquid mass fraction of $w^{\mathrm{SW}} = m^{\mathrm{SW}}/m$ and that of humid air, $g^{\mathrm{AV}}(A,T,p)$, with a gaseous mass fraction of $w^{\mathrm{AV}} = m^{\mathrm{AV}}/m = 1 - w^{\mathrm{SW}}$,

$$g^{\mathrm{SA}}\left(S,A,w^{\mathrm{SW}},T,p\right) = w^{\mathrm{SW}} g^{\mathrm{SW}}(S,T,p) + \left(1 - w^{\mathrm{SW}}\right) g^{\mathrm{AV}}(A,T,p). \qquad (15)$$

If the two phases are assumed to be at mutual equilibrium, they possess the same temperature, pressure and chemical potentials, see eq. (B.11) in Appendix B, $\mu_{\mathrm{W}}^{\mathrm{SW}} = \mu_{\mathrm{V}}^{\mathrm{AV}}$, namely that of water in seawater,

$$\mu_{\mathrm{W}}^{\mathrm{SW}}(S,T,p) = g^{\mathrm{SW}} - S\left(\frac{\partial g^{\mathrm{SW}}}{\partial S}\right)_{T,p}, \qquad (15)$$

equalling that of water vapour in humid air,

$$\mu_{\mathrm{V}}^{\mathrm{AV}}(A,T,p) = g^{\mathrm{AV}} - A\left(\frac{\partial g^{\mathrm{AV}}}{\partial A}\right)_{T,p}. \qquad (16)$$

5.1 Sea Air as a Model for Latent Heat of Evaporation

Water evaporated from the ocean surface drives the climate system. “The by far largest part of heat conveyed to the air is in the form of latent heat during subsequent condensation along with cloud formation. The heat budget over the sea is mainly controlled by the latent heat released to the air” (Albrecht 1940). It is the “*heat source for a gigantic steam engine*”, as Heinrich Hertz had put it in his 1885 inaugural lecture at Karlsruhe (Mulligan and Hertz 1997). The latent heat of evaporation of pure liquid water into pure water vapour is numerically well known from experiments (IAPWS SR1-86 1992, Harvey 1998, Wagner and Pruß 2002). Slightly differing values are reported in various textbooks on hydrology (Debski 1966: p. 332), meteorology (Linke and Baur 1970) or geophysics (Gill 1982, Kraus and Businger 1994). TEOS-10, however, permits the computation of evaporation properties from seawater into humid air, based on the first-time availability of standard equations for enthalpies and chemical potentials of those non-ideal mixtures.

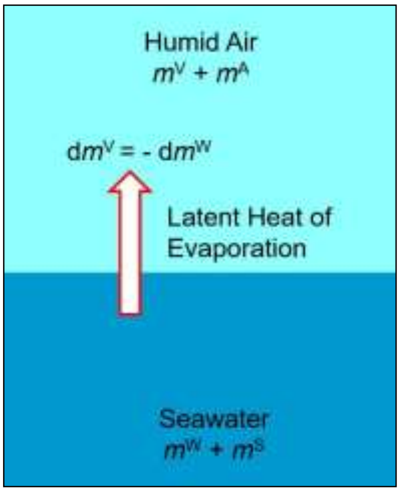

Fig. 8: Conceptual thermodynamic “sea air” model of ocean-atmosphere interaction as a two-phase composite of seawater and humid air

“Latent heat is the quantity of heat which must be communicated to a body in a given state in order to convert it into another state without changing its temperature” (Maxwell 1888: p.73). If an infinitesimal amount of water is transferred from the liquid to the gas phase (Fig. 8), while temperature and pressure remain at their equilibrium values, and the total masses of salt, m^{S}, dry air, m^{A}, and water, m^{WV}, are not affected, the isobaric-isothermal latent heat of evaporation may be defined by

$$L^{\mathrm{SA}} \equiv \left(\frac{\partial H^{\mathrm{SA}}}{\partial m^{\mathrm{V}}}\right)_{T,p,m^{\mathrm{S}},m^{\mathrm{A}},m^{\mathrm{WV}}}. \qquad (17)$$

This latent heat accounts for the loss of total heat of the sea-air sample associated with the loss of liquid water and equal gain of water vapour,

$$\frac{\partial m^{\mathrm{V}}}{\partial T} = -\frac{\partial m^{\mathrm{W}}}{\partial T}. \qquad (18)$$

Here, H^{SA} is the enthalpy of sea air, available from the Gibbs function (15) through the sum

$$H^{\mathrm{SA}} \equiv m^{\mathrm{SW}} h^{\mathrm{SW}} + m^{\mathrm{AV}} h^{\mathrm{AV}}. \qquad (19)$$

Here, the specific enthalpies of seawater,

$$h^{\mathrm{SW}} = g^{\mathrm{SW}} - T\left(\frac{\partial g^{\mathrm{SW}}}{\partial T}\right)_{S,p}, \qquad (20)$$

and of humid air,

$$h^{\mathrm{AV}} = g^{\mathrm{AV}} - T\left(\frac{\partial g^{\mathrm{AV}}}{\partial T}\right)_{A,p}, \qquad (21)$$

are defined in terms of the related Gibbs functions.

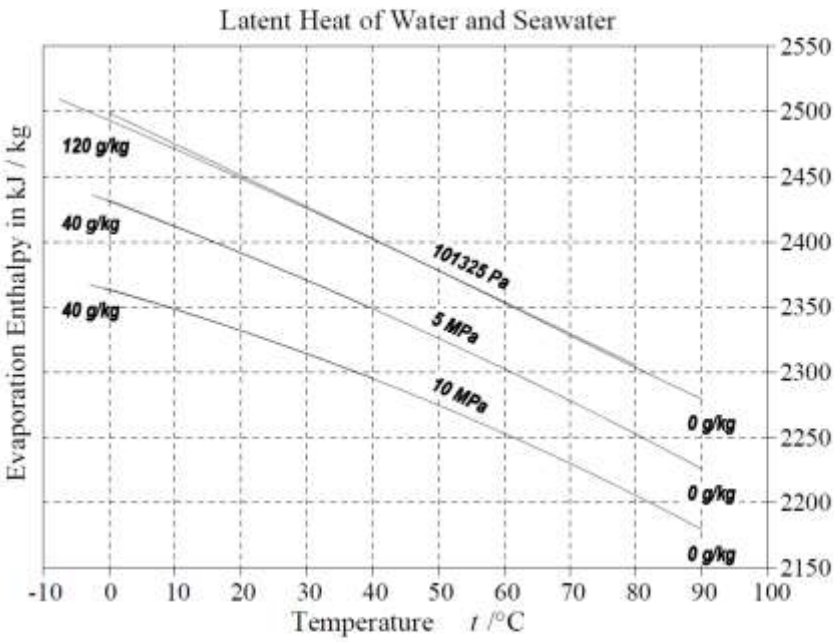

Fig. 9: Evaporation enthalpy, eq. (23), of seawater in equilibrium with humid air at different temperatures, pressures and salinities. The dependence on salinity is very weak; graphically, the related curves are hardly distinguishable. The nonlinear dependence on temperature is more pronounced at elevated pressures. Figure from Feistel et al. (2010a: p. 105)

The derivative (17) is carried out in the form

$$L^{\mathrm{SA}} = -h^{\mathrm{SW}} - m^{\mathrm{SW}} \left(\frac{\partial h^{\mathrm{SW}}}{\partial S}\right)_{T,p} \left(\frac{\partial S}{\partial m^{\mathrm{V}}}\right)_{m^{\mathrm{S}},m^{\mathrm{WV}}} + h^{\mathrm{AV}} + m^{\mathrm{AV}} \left(\frac{\partial h^{\mathrm{AV}}}{\partial A}\right)_{T,p} \left(\frac{\partial A}{\partial m^{\mathrm{V}}}\right)_{m^{\mathrm{A}}}, \qquad (22)$$

which results in the TEOS-10 latent-heat equation (Feistel et al. 2010a, Feistel and Hellmuth 2023),

$$L^{\mathrm{SA}} = h^{\mathrm{AV}} - A \left(\frac{\partial h^{\mathrm{AV}}}{\partial A}\right)_{T,p} - h^{\mathrm{SW}} + S \left(\frac{\partial h^{\mathrm{SW}}}{\partial S}\right)_{T,p}, \qquad (23)$$

with typical values shown in Fig. 9. If seawater is in mutual equilibrium with humid air at given temperature and pressure, salinity and humidity of the parts of sea air satisfy the condition $\mu_{\mathrm{W}}^{\mathrm{SW}} = \mu_{\mathrm{V}}^{\mathrm{AV}}$, given by eqs. (15) and (16),

$$\Delta\mu \equiv g^{\mathrm{SW}} - S \left(\frac{\partial g^{\mathrm{SW}}}{\partial S}\right)_{T,p} - g^{\mathrm{AV}} + A \left(\frac{\partial g^{\mathrm{AV}}}{\partial A}\right)_{T,p} = 0 \qquad (24)$$

At given masses of salt, m^{S}, of dry air, m^{A}, and of total water, $m^{\mathrm{WV}} = m^{\mathrm{W}} + m^{\mathrm{V}}$, eq. (24) controls the value of either m^{W} or m^{V}, and this way also of S and A as functions of T, p, m^{S}, m^{A} and m^{WV}. Related numerical solutions are readily implemented in the TEOS-10 SIA library; the latent heat of sea air can be computed by calling the function sea_air_enthalpy_evap_si(), see Wright et al. (2010).

Latent heat of eq. (23) is valid regardless of the equilibrium condition, eq. (24), being satisfied or not. The non-equilibrium case is considered separately in Section 5.3.

5.2 Sea Air as a Model of Sea Spray

As a special form of air-sea interaction, sea spray is typically ejected from the crest of a breaking wave, which may happen all along oceanic coasts but also wherever whitecaps are produced from swell or stormy sea state, see Fig. 10. In contrast to fresh-water haze, droplets of sea spray cannot completely evaporate for the salt they contain, and rather develop into a floating persistent Köhler (1936) equilibrium between droplet size, droplet salinity and ambient relative fugacity (Hellmuth and Shchekin 2015, Pöhlker et al. 2023). This equilibrium can be described by the TEOS-10 model of sea air if the additional Kelvin pressure caused by the surface tension is allowed for.

In the infrared spectral range, sea spray as well as other aerosols (Carlon 1970, 1980) may be considered as a black absorber and emitter of thermal radiation. The resulting “gray atmosphere” is a conveniently simple conceptual model for the long-wave radiative effects of dust or haze in the climate system (Emden 1913). When heated from below, as in the case of the clear-sky marine troposphere, a theoretical finding is that the thermally stratified gray troposphere exhibits a special critical value of the isobaric heat capacity at $c_p = 4R$ (Pierrehumbert 2010: p. 201), R being the molar gas constant. Vertical stability may be lost at $c_p > 4R$ and turbulent mixing is expected to commence (Feistel 2011b: eq. 58 therein). Such a kinetic phase transition could substantially modify the thermal radiation balance between troposphere and ocean surface.

The terrestrial atmosphere is dominated by the two-atomic gases N_2 and O_2 with heat capacities about 3.5 R which prevent the putative radiative vertical instability to occur. This situation may change, however, in the presence of haze or sea spray. To investigate this effect theoretically, in this section a TEOS-10 equation for the heat capacity of equilibrium sea air is derived from the definition

$$c_p^{\mathrm{SA}} \equiv \frac{1}{m} \left(\frac{\partial H^{\mathrm{SA}}}{\partial T}\right)_{p,m^{\mathrm{S}},m^{\mathrm{A}},m^{\mathrm{WV}}}. \qquad (25)$$

The enthalpy of sea air is given by eq. (19). Taking into account water conservation upon evaporation, $m^{\mathrm{WV}} = \mathrm{const}$, that is,

$$\frac{\partial m^{\mathrm{V}}}{\partial T} = -\frac{\partial m^{\mathrm{W}}}{\partial T}, \quad (26)$$

and of eq. (23), the isobaric heat capacity of sea air is concluded to be

$$c_p^{\mathrm{SA}} = w^{\mathrm{SW}} c_p^{\mathrm{SW}} + w^{\mathrm{AV}} c_p^{\mathrm{AV}} + L^{\mathrm{SA}} \frac{1}{m} \frac{\partial m^{\mathrm{V}}}{\partial T}. \quad (27)$$

To the additive contributions of the partial heat capacities of the liquid and the gas part, there appears the latent heat of the water mass that evaporates from the liquid as vapour. This evaporation rate is governed by the mutual equilibrium between seawater and humid air.

Fig. 10: Sea spray ejection from a breaking wave crest of Atlantic swell. Photo taken at Cabo Trafalgar in March 2011.

During the temperature change, sea-air equilibrium, eq. (24), is assumed to be maintained by water transfer between the phases, changing S and A along with T,

$$\left(\frac{\partial \Delta\mu}{\partial T}\right)_{p,m^{\mathrm{S}},m^{\mathrm{A}},m^{\mathrm{WV}}} = 0. \quad (28)$$

Carrying out the derivative, this condition reads

$$\left(\frac{\partial g^{\mathrm{SW}}}{\partial T}\right)_{S,p} - S\left(\frac{\partial^2 g^{\mathrm{SW}}}{\partial S \partial T}\right)_p - S\left(\frac{\partial^2 g^{\mathrm{SW}}}{\partial S^2}\right)_{T,p}\left(\frac{\partial S}{\partial T}\right)_{m^{\mathrm{S}}}$$

$$= \left(\frac{\partial g^{\mathrm{AV}}}{\partial T}\right)_{A,p} - A\left(\frac{\partial^2 g^{\mathrm{AV}}}{\partial A \partial T}\right)_p - A\left(\frac{\partial^2 g^{\mathrm{AV}}}{\partial A^2}\right)_{T,p}\left(\frac{\partial A}{\partial T}\right)_{m^{\mathrm{A}}}. \quad (29)$$

On the other hand, from combining eq. (23) with eq. (24) it follows that the latent heat may be expressed by,

$$L^{\mathrm{SA}} = T\left\{\left(\frac{\partial g^{\mathrm{SW}}}{\partial T}\right)_{S,p} - S\left(\frac{\partial^2 g^{\mathrm{SW}}}{\partial S \partial T}\right)_p - \left(\frac{\partial g^{\mathrm{AV}}}{\partial T}\right)_{A,p} + A\left(\frac{\partial^2 g^{\mathrm{AV}}}{\partial A \partial T}\right)_p\right\}, \quad (30)$$

so that eq. (29) may be written as

$$L^{\mathrm{SA}} = T\left\{S\left(\frac{\partial^2 g^{\mathrm{SW}}}{\partial S^2}\right)_{T,p}\left(\frac{\partial S}{\partial T}\right)_{m^{\mathrm{S}}} - A\left(\frac{\partial^2 g^{\mathrm{AV}}}{\partial A^2}\right)_{T,p}\left(\frac{\partial A}{\partial T}\right)_{m^{\mathrm{A}}}\right\}. \quad (31)$$

Further, the total water mass balance, eq. (26), implies that

$$\left(\frac{\partial S}{\partial T}\right)_{m^{\mathrm{S}}} = \left(\frac{\partial S}{\partial m^{\mathrm{W}}}\right)_{m^{\mathrm{S}}}\frac{\partial m^{\mathrm{W}}}{\partial T} = \frac{S}{m^{\mathrm{SW}}}\frac{\partial m^{\mathrm{V}}}{\partial T}, \quad (32)$$

and similarly,

$$\left(\frac{\partial A}{\partial T}\right)_{m^{\mathrm{A}}} = \left(\frac{\partial A}{\partial m^{\mathrm{V}}}\right)_{m^{\mathrm{A}}}\frac{\partial m^{\mathrm{V}}}{\partial T} = -\frac{A}{m^{\mathrm{AV}}}\frac{\partial m^{\mathrm{V}}}{\partial T}. \quad (33)$$

Inserting those expressions into eq. (31), the equation for the isobaric evaporation rate of sea air is

$$\frac{\partial m^{\mathrm{V}}}{\partial T} = \frac{L^{\mathrm{evap}}}{T}\left\{\frac{S^2}{m^{\mathrm{SW}}}\left(\frac{\partial^2 g^{\mathrm{SW}}}{\partial S^2}\right)_{T,p} + \frac{A^2}{m^{\mathrm{AV}}}\left(\frac{\partial^2 g^{\mathrm{AV}}}{\partial A^2}\right)_{T,p}\right\}^{-1}. \quad (34)$$

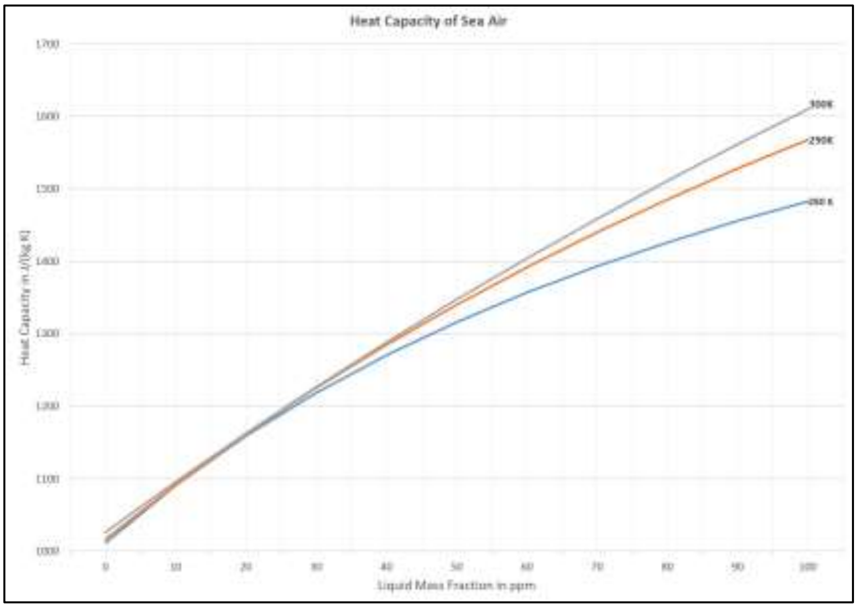

Fig. 11: TEOS-10 values for the isobaric specific heat capacity, eq. (35), of sea air at atmospheric pressure and sea-spray standard-ocean salinity, S = 35.165 04 g kg^{-1}, at temperatures of 280 K (lower curve), 290 K (middle curve) and 300 K (upper curve) as functions of the liquid mass fraction, w^{SW}, up to 100 ppm.

Together with eq. (34), the desired formula for the isobaric heat capacity (27) of sea air finally becomes (Feistel et al. 2010a: eq. 6.22 therein),

$$c_p^{\mathrm{SA}} = w^{\mathrm{SW}}c_p^{\mathrm{SW}} + \left(1 - w^{\mathrm{SW}}\right)c_p^{\mathrm{AV}} + \frac{\left(L^{\mathrm{SA}}\right)^2}{T}\left\{\frac{S^2}{w^{\mathrm{SW}}}\left(\frac{\partial^2 g^{\mathrm{SW}}}{\partial S^2}\right)_{T,p} + \frac{A^2}{(1-w^{\mathrm{SW}})}\left(\frac{\partial^2 g^{\mathrm{AV}}}{\partial A^2}\right)_{T,p}\right\}^{-1}. \quad (35)$$

Of the *liquid water content*, expressed in form of the liquid mass fraction, w^{SW}, realistic values may typically range between 10^{-6} and 10^{-4} in the troposphere. Growing along with this fraction, related heat capacities of sea air, eq. (35), may substantially exceed that of liquid-free humid air, c_p^{AV}, see Fig. 11.

5.3 Sea Air as a Model for Irreversible Evaporation

The climate system functions far from thermodynamic equilibrium, permanently producing and exporting entropy at an average rate about 1 W m^{-2} K^{-1} per global surface area (Ebeling and Feistel 1982, Feistel and Ebeling 2011). By contrast, TEOS-10 is a mathematical description of equilibrium properties (Appendix B). The latter is applicable to states away from thermodynamic equilibrium under the assumption of *local equilibrium* as introduced by Ilya Prigogine (1947, 1978). This assumption means that spatially extended substances such as ocean or atmosphere consist of sufficiently small volume elements that may reasonably be described as macroscopic equilibrium states, homogeneous in temperature, pressure and chemical potentials. TEOS-10 thermodynamic potentials can be used to describe those local states.

By definition, if a volume at equilibrium is divided into partial volumes, each of those parts is at equilibrium itself, and each pair of those is at mutual equilibrium also. By contrast, the combination of several local-equilibrium elements forms a non-equilibrium state if pairs of elements exist that are out of mutual equilibrium. Extensive properties such as mass, energy, entropy or enthalpy can be added up to give correct values of the entire system. When exchange processes between those elements occur, gains and losses of masses, energies or enthalpies are mutually balanced by conservation laws, however, this is not the case for entropy.

A tutorial case of a local equilibrium system may be the model of sea air (Feistel and Hellmuth 2024a) depicted in Fig. 8. It consists of a mass $m^{\mathrm{SW}} = m^{\mathrm{S}} + m^{\mathrm{W}}$ of seawater in contact with a mass $m^{\mathrm{AV}} = m^{\mathrm{A}} + m^{\mathrm{A}}$ of humid air. Both fluids are assumed to be at internal equilibrium themselves but not necessarily in mutual equilibrium with one another. This is a natural geophysical situation – marine RH has typical values of 80 %rh while the equilibrium of humid air with seawater, eq. (24), is established at about 98 %rh. For simplicity, let all parts have equal temperatures and pressures.

If evaporation takes place, the partial water masses involved will change by a mass flux across the sea surface,

$$J_m \equiv \frac{\mathrm{d}m^{\mathrm{AV}}}{\mathrm{d}t} = \frac{\mathrm{d}m^{\mathrm{V}}}{\mathrm{d}t} = -\frac{\mathrm{d}m^{\mathrm{SW}}}{\mathrm{d}t} = -\frac{\mathrm{d}m^{\mathrm{W}}}{\mathrm{d}t}. \tag{36}$$

The change of the total enthalpy of the sea-air sample is available from eqs. (17) and (23),

$$\frac{\mathrm{d}H^{\mathrm{SA}}}{\mathrm{d}t} = \left(\frac{\partial H^{\mathrm{SA}}}{\partial m^{\mathrm{V}}}\right)_{T,p,m^{\mathrm{S}},m^{\mathrm{A}},m^{\mathrm{WV}}} \frac{\mathrm{d}m^{\mathrm{V}}}{\mathrm{d}t} = L^{\mathrm{SA}} J_m \,. \tag{37}$$

This expression of energy conservation, the 1st law of thermodynamics, is similarly valid for equilibrium and non-equilibrium conditions of the sample. For comparison, of the total entropy defined by,

$$N^{\mathrm{SA}} \equiv m^{\mathrm{SW}}\eta^{\mathrm{SW}} + m^{\mathrm{AV}}\eta^{\mathrm{AV}}, \tag{38}$$

the change is given by

$$\frac{\mathrm{d}N^{\mathrm{SA}}}{\mathrm{d}t} = \left(\frac{\partial N^{\mathrm{SA}}}{\partial m^{\mathrm{V}}}\right)_{T,p,m^{\mathrm{S}},m^{\mathrm{A}},m^{\mathrm{WV}}} \frac{\mathrm{d}m^{\mathrm{V}}}{\mathrm{d}t} \,. \tag{39}$$

In terms of its two parts, eq. (38), this change takes the form,

$$\frac{\mathrm{d}N^{\mathrm{SA}}}{\mathrm{d}t} = \left[\eta^{\mathrm{AV}} - A\left(\frac{\partial \eta^{\mathrm{AV}}}{\partial A}\right)_{T,p} - \eta^{\mathrm{SW}} + S\left(\frac{\partial \eta^{\mathrm{SW}}}{\partial S}\right)_{T,p}\right] J_m. \tag{40}$$

In oceanography, the symbol N for entropy was suggested by Fofonoff (1962) to avoid confusion with salinity S. Making use of their local equilibria, specific entropy of each part can be expressed by the difference, eq. (B.6),

$$\eta = \frac{h-g}{T}\,, \tag{41}$$

between specific enthalpy, h, and specific Gibbs energy, g, so that the entropy change (40) becomes

$$T\frac{\mathrm{d}N^{\mathrm{SA}}}{\mathrm{d}t} = \left(L^{\mathrm{SA}} + \Delta\mu\right)J_m. \tag{42}$$

Here, the latent heat, L^{SA}, is given by eq. (23), and the distance from mutual equilibrium, $\Delta\mu$, by eq. (24).

The first term,

$$T\frac{\mathrm{d_e}N^{\mathrm{SA}}}{\mathrm{d}t} \equiv L^{\mathrm{SA}}J_m, \tag{43}$$

is the *external* entropy change (subscript e) in the form of the heat flux required to maintain the sample's temperature, in the sense of Maxwell's (1888) definition of latent heat, compensating the storage of latent heat by emitting water vapour.

The second term,

$$T\frac{\mathrm{d_i}N^{\mathrm{SA}}}{\mathrm{d}t} \equiv J_m\Delta\mu. \tag{44}$$

is the *internal* entropy change (subscript i), or *entropy production*, of the non-equilibrium sea-air sample. It represents the additional entropy gain of humid air compared to the entropy loss of seawater. This production happens at the air-sea interface and disappears as soon as mutual equilibrium, $\Delta\mu = 0$, is approached.

It is important to be aware that the external part, $\frac{\mathrm{d_e}N^{\mathrm{SA}}}{\mathrm{d}t}$, *always* constitutes a contribution to the system's energy balance while, by contrast, the internal part, $\frac{\mathrm{d_i}N^{\mathrm{SA}}}{\mathrm{d}t}$, is *never* any such contribution. The irreversible production of entropy is an internal conversion or redistribution of energy rather than a change of it. This implies that irreversible processes violate Gibbs' fundamental equation (B.8) in the sense that

$$\frac{\mathrm{d}H^{\mathrm{SA}}}{\mathrm{d}t} = -T\frac{\mathrm{d_e}N^{\mathrm{SA}}}{\mathrm{d}t} + V^{\mathrm{SA}}\frac{\mathrm{d}p}{\mathrm{d}t} + \sum_i \mu_i\,\frac{\mathrm{d}m_i}{\mathrm{d}t} > -T\frac{\mathrm{d}N^{\mathrm{SA}}}{\mathrm{d}t} + V^{\mathrm{SA}}\frac{\mathrm{d}p}{\mathrm{d}t} + \sum_i \mu_i\,\frac{\mathrm{d}m_i}{\mathrm{d}t}, \tag{45}$$

even though each of its local-equilibrium elements strictly satisfies the related fundamental equation (B.13), valid for reversible processes only,

$$dh = -Td\eta + vdp + \sum_{i=1}^{n-1}(\mu_i - \mu_0)\mathrm{d}w_i. \tag{46}$$

Entropy production appears wherever a flux is passing its driving gradient. Near equilibrium, this flux is proportional to its driving force (Glansdorff and Prigogine 1971, Landau and Lifschitz 1974, Kraus and Businger 1994, Feistel and Hellmuth 2024a), usually termed *Onsager force*. For example, the evaporation mass flux of water, eq. (6),

$$J_m = C\,\Delta\mu \tag{47}$$

may be assumed as being proportional to the difference between the chemical potentials of water across the air-sea interface. The related *Dalton equation* (6) was discussed in Section 4. The associated entropy production, eq. (44), obeys the 2nd law of thermodynamics by the inequality

$$\frac{\mathrm{d_i}N^{\mathrm{SA}}}{\mathrm{d}t} = C\,(\Delta\mu)^2 \geq 0, \qquad (48)$$

while the total entropy change, eq. (42) may possess any sign. In other words, the 2nd law forbids that *Onsager fluxes* may be directed against their causing Onsager forces. The *Prigogine Theorem* predicts that in linear irreversible thermodynamics, entropy production approaches minimum values at steady states (Glansdorff and Prigogine 1971).

Processes accompanied by entropy production are termed *irreversible* ones, since entropy once created may never be destroyed again. Related processes cannot be reversed unless lasting changes are left behind in the external world. By contrast, processes which transform an equilibrium state into another equilibrium state may *reversibly* be performed without producing entropy. Entropy production is possible only under non-equilibrium conditions.

Under typical marine circumstances, the entropy production density of ocean evaporation can be estimated to about 4 mW K^{-1} m^{-2}, contributing roughly 0.4 % to the global entropy production (Feistel and Ebeling 2011, Feistel and Hellmuth 2024a).

6 Cloudiness and Ocean Warming

"Cloud feedback on climate represents the largest uncertainty in our ability to understand the sensitivity of the planet to radiative forcing" (Gettelman and Sherwood 2016). On the long-term average, cloudiness is particularly strong in the low-pressure belts of the global tropospheric circulation, where air is ascending and its humidity is condensing, see Fig. 12. Except for the equatorial zone, those spatial cloudiness pattern correlate visibly with those of recent ocean warming, compare Fig. 1. It is a plausible working hypothesis that this correlation could also indicate a causal relation between the two phenomena. However, such correlations imply chicken-and-egg problems (Rapp 2014): putative causality relations between those trends cannot be derived from observation but only be concluded from reliable prediction models (Feistel 2023). May the observed systematic reduction of global cloudiness (Fasullo and Trenberth 2012) actually be responsible for the currently recorded excessive ocean warming (You 2024)? Unfortunately, and somewhat surprisingly, this assumption can apparently not be underpinned yet by closer investigation. Some related issues will be discussed in this section.

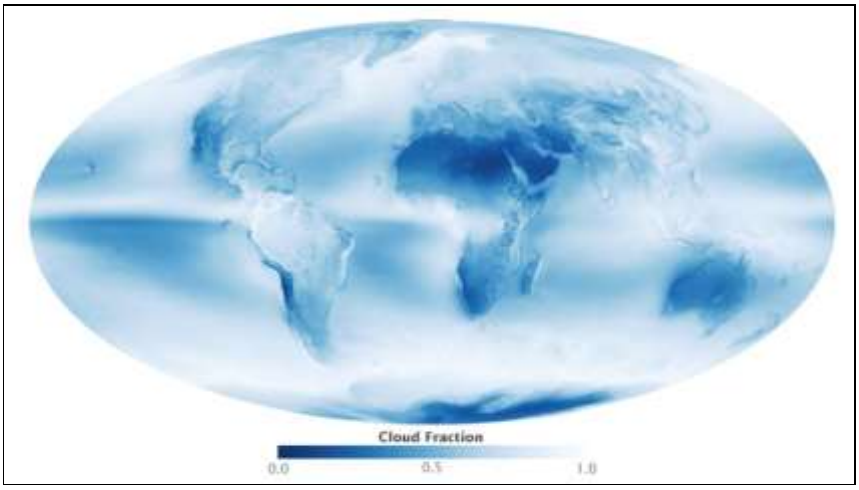

Fig. 12: Global distribution of cloudiness July 2002 – April 2015 (Allen and Ward 2015). Image reproduction permitted by NASA Copyright.

6.1 Cloudiness Trend

Global cloud-covered surface area fraction *C* has systematically been reducing by about 6 % per century, see Fig. 13, from $C \approx 67.5$ % in 1980 to $C \approx 65$ % in 2022 (Foster et al. 2023, Phillips and Foster 2023). Observed cloudiness values depend strongly on the way clouds are defined (Spänkuch et al. 2022) and on the measurement technology applied. For example, Rapp (2014: Fig. 6.20) reported a decrease in cloudiness in 30 years from 70 % in 1983 down to 63.5 % in, likely, 2013. This reduction rate of more than 20 % per century is three times as fast as that given in Fig. 13 and may result from different observation techniques.

Assuming that this shrinking occurred in a similar way above both land and sea, the ocean is expected to receive increasingly more solar irradiation. This phenomenon is known as the *short-wave cloud radiative effect* (SW CRE), see Fig. 14.

On the other hand, clouds are opaque with respect to oceanic upward thermal radiation and emit themselves downward infrared radiation. This phenomenon is known as the *long-wave cloud radiative effect* (LW CRE), see Fig. 15. Radiation models show that on the global average these two effects cancel each other almost completely up to minor residual of –1 mW m^{-2} yr^{-1}, so that the continuously shrinking cloudiness may be assumed to have practically no net effect on the ocean's radiation balance (Phillips and Foster 2023, Feistel and Hellmuth 2024b). However, more detailed investigations in the future may reveal more rigorous results for the ocean than this simplified picture.

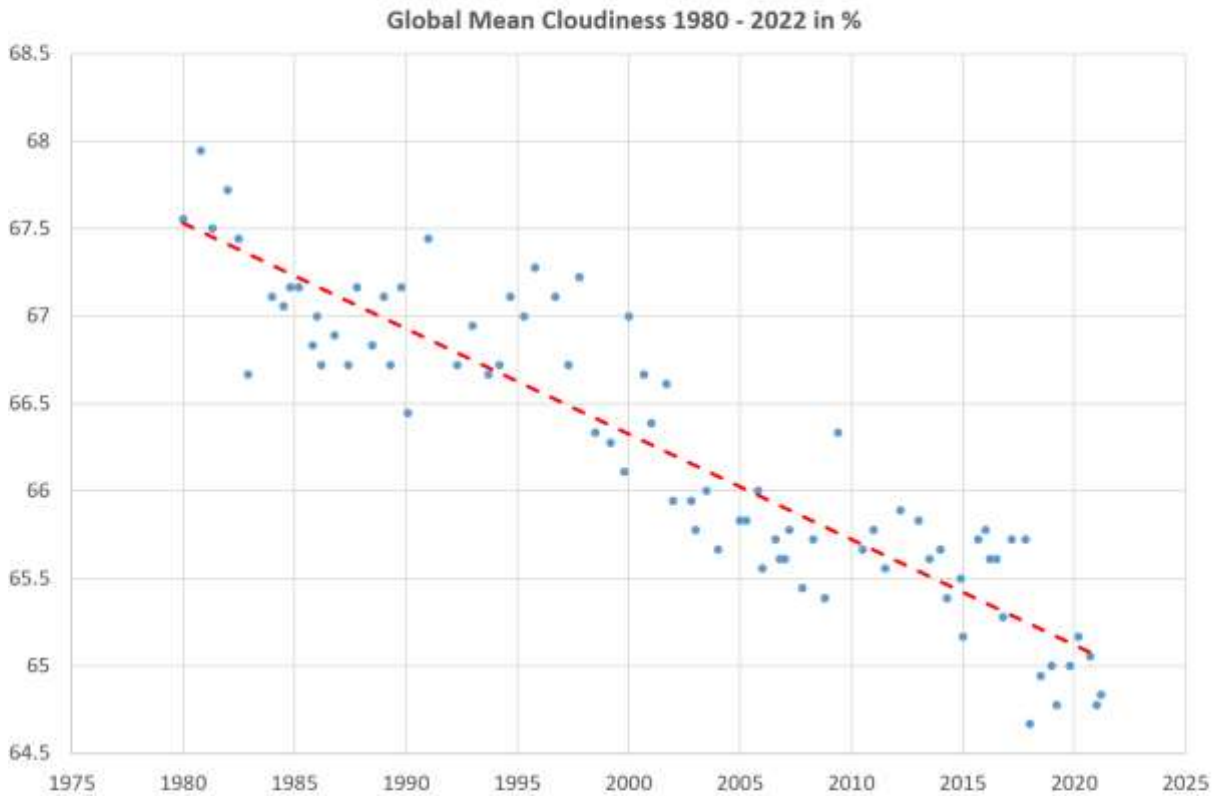

Fig. 13: Dots: satellite-derived global mean cloud area fractions 1980-2022 in percent. Data from Foster et al. (2023). Dashed line: present cloudiness is 65 % with a climatological linear shrinking trend of –6.2 % per century.

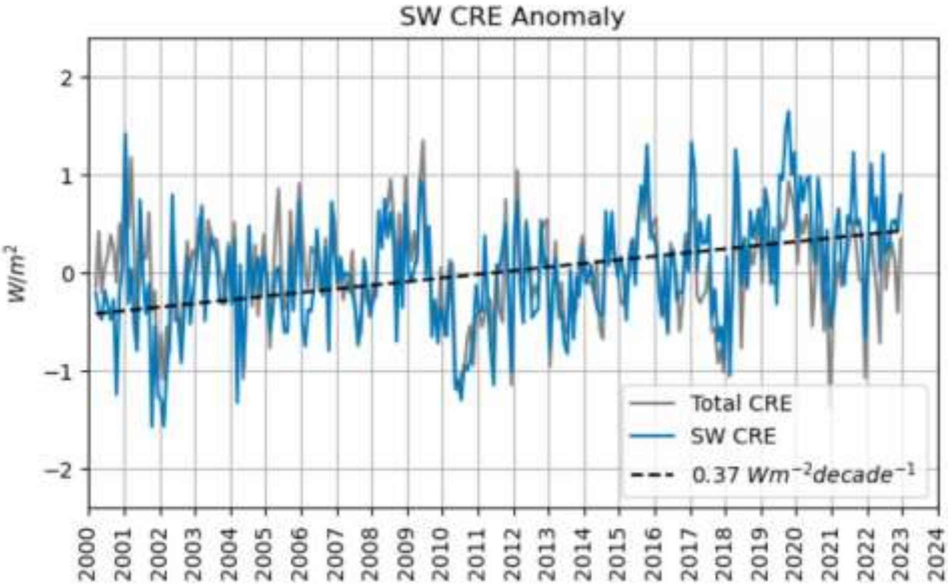

Fig. 14: Short-wave cloud radiative effect (SW CRE) of increasing solar irradiation. Image kindly provided by Coda Phillips (priv. comm.), with minor correction compared to the similar previous publication (Phillips and Foster 2023). Total CRE is the net effect of SW and LW CRE, see Fig. 15

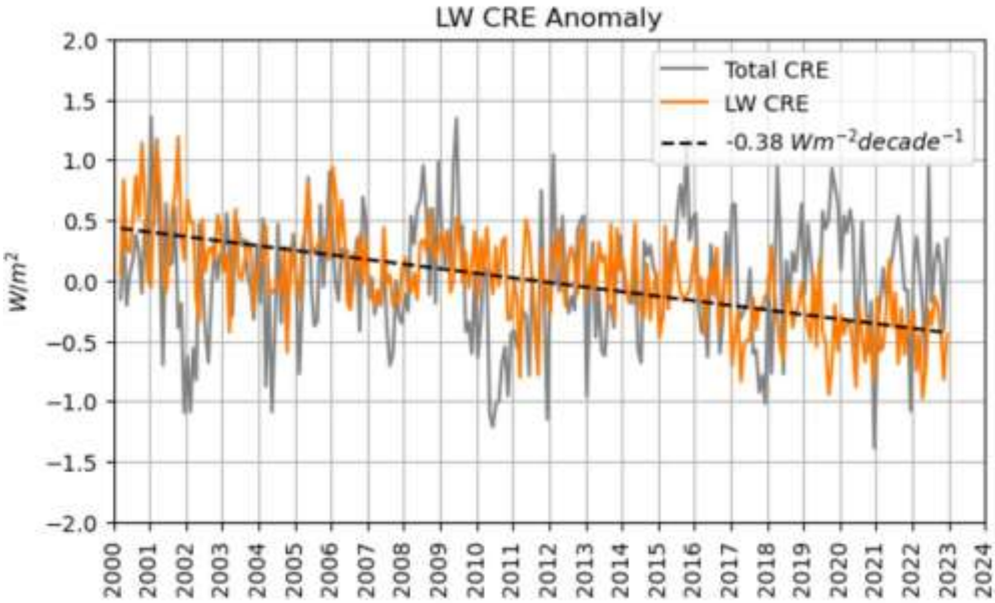

Fig. 15: Long-wave cloud radiative effect (LW CRE) of decreasing net thermal radiation. Image kindly provided by Coda Phillips (priv. comm.), with minor correction compared to the similar previous publication (Phillips and Foster 2023). Total CRE is the net effect of SW and LW CRE, see Fig. 14.

6.2 Cumulus Clouds

Cumulus clouds are often formed in the course of diurnal convection by isentropic uplift of humid air parcels from the sea surface to the condensation level, mostly located at low heights between 200 and 500 m. This process permits a thermodynamic description of such clouds (Romps 2014) by calculating the *lifted condensation level* (LCL) as the cumulus cloud base. In distinction to previous studies, as the first such international geophysical standard, TEOS-10 provides explicit equations for entropy, enthalpy and chemical potentials of humid air which may be used to derive reference equations and values of the LCL (Feistel and Hellmuth 2024b).

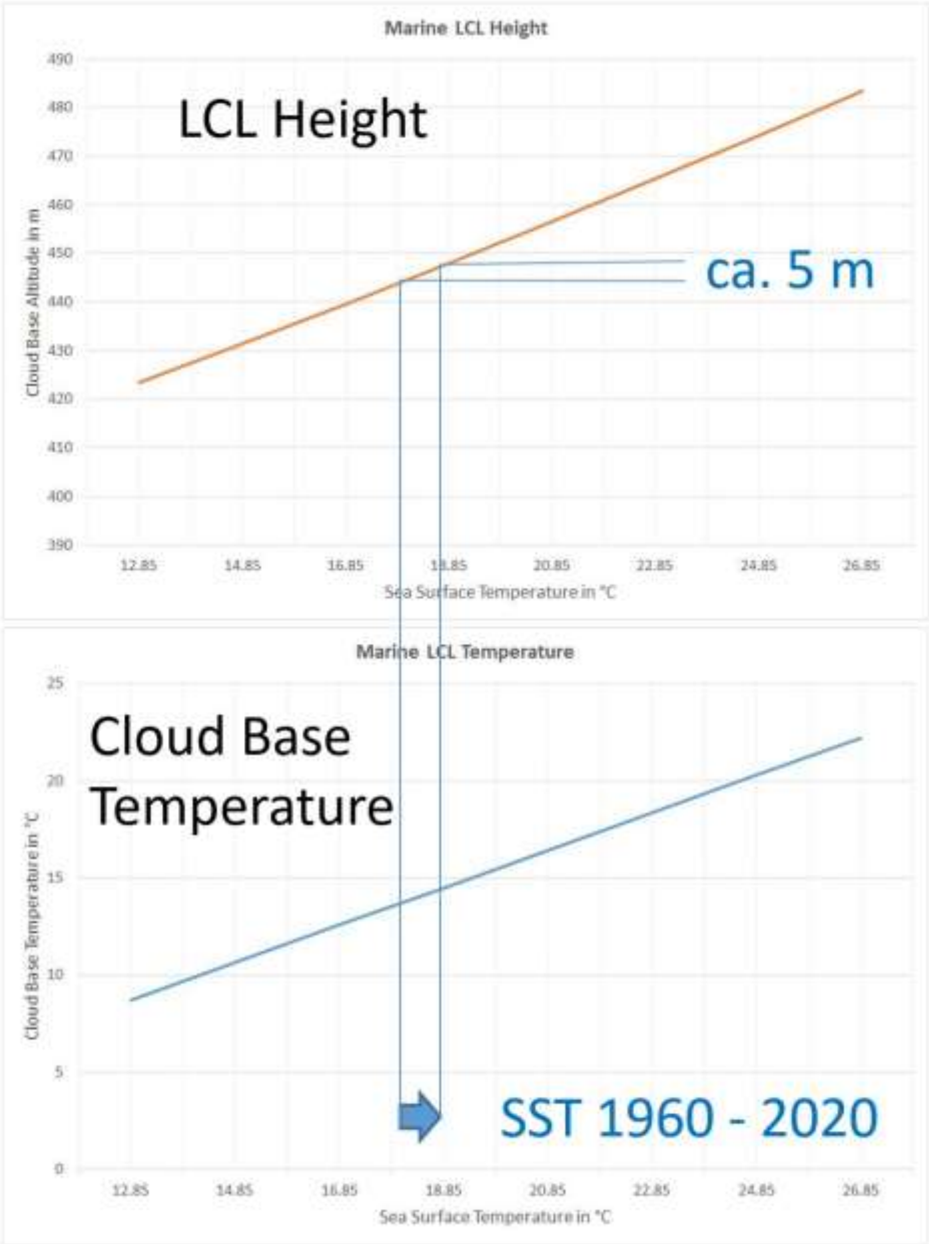

Fig. 16: As a function of typical low-latitude sea-surface temperatures, LCL height (top) and LCL temperature (bottom) are computed from the TEOS-10 equations (49) – (52) at a typical marine surface RH of 80 %rh. The added interval indicates the global mean SST change between 1960 and 2020 which has resulted in an increase of the cloud base altitude by about 5 m.

At the sea surface pressure, p_{SS}, the air parcel may possess the temperature T_{SS} and the relative fugacity ψ_f, which is a real-gas definition of relative humidity (Feistel and Lovell-Smith 2017) in terms of the chemical potential of water vapour in humid air, $\mu_{\mathrm{V}}^{\mathrm{AV}}$, and that of liquid water, μ_{W},

$$R_W T_{SS} \ln \psi_f = \mu_V^{AV}(A, T_{SS}, p_{SS}) - \mu_W(T_{SS}, p_{SS}). \tag{49}$$

Here, $R_W = 461.523\ \mathrm{J\ kg^{-1}\ K^{-1}}$ is the specific gas constant of water, and $A = 1 - q$ is the dry-air mass fraction of the parcel, to be determined from ψ_f by this condition.

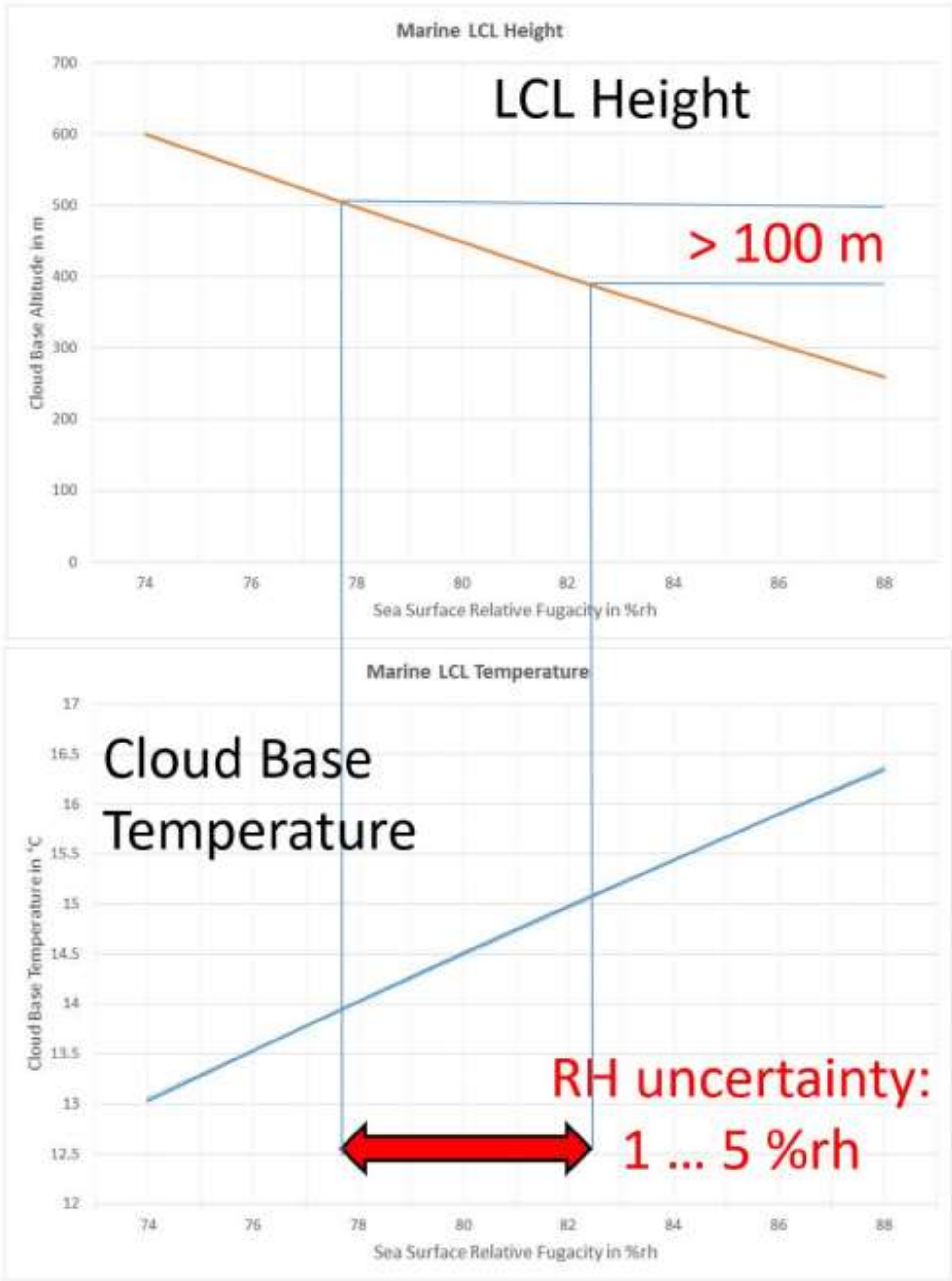

Fig. 17: As a function of typical marine RH values, LCL height (top) and LCL temperature (bottom) are computed from the TEOS-10 equations (49) – (52) at a sea surface temperature of 292 K, close to the current global mean SST of 18.8 °C, see Fig. 18. The added interval indicates the observation uncertainty of sea surface RH which corresponds to an uncertainty of the cloud base altitude of more than 100 m.

At the LCL, the parcel is saturated at $\psi_f = 1$, i.e.,

$$0 = \mu_{\mathrm{V}}^{\mathrm{AV}}(A, T_{\mathrm{LCL}}, p_{\mathrm{LCL}}) - \mu_{\mathrm{W}}(T_{\mathrm{LCL}}, p_{\mathrm{LCL}}). \tag{50}$$

During uplift, A is assumed to remain constant, as well as the parcel's entropy, η^{AV},

$$\eta^{\mathrm{AV}}(A, T_{\mathrm{SS}}, p_{\mathrm{SS}}) = \eta^{\mathrm{AV}}(A, T_{\mathrm{LCL}}, p_{\mathrm{LCL}}) \,. \tag{51}$$

Finally, the LCL altitude, z_{LCL}, above sea level follows from the isentropic integral of the hydrostatic equation in terms of the enthalpy, h^{AV}, of humid air,

$$z_{\mathrm{LCL}} = \frac{1}{g_{\mathrm{E}}}\left[h^{\mathrm{AV}}\left(A, \eta^{\mathrm{AV}}, p_{\mathrm{SS}}\right) - h^{\mathrm{AV}}\left(A, \eta^{\mathrm{AV}}, p_{\mathrm{LCL}}\right)\right]. \tag{52}$$

The gravity acceleration is $g_{\mathrm{E}} = 9.81\ \mathrm{m\ s^{-2}}$. The functions $\mu_{\mathrm{V}}^{\mathrm{AV}}$, η^{AV}, h^{AV} and μ_{W} can be expressed by partial derivatives of the TEOS-10 thermodynamic potentials of humid air and liquid water, and are numerically available from the *Sea-Ice-Air* (SIA) *library* (Feistel et al. 2010d, Wright et al. 2010). Solving eqs. (49) – (52) numerically, the LCL properties $(A, T_{\mathrm{LCL}}, p_{\mathrm{LCL}}, z_{\mathrm{LCL}})$ are obtained from the given surface properties, $\left(\psi_f, T_{\mathrm{SS}}, p_{\mathrm{SS}}\right)$.

Table 1: LCL cloud-base temperatures, T_{LCL}, pressures, p_{LCL}, and heights, z_{LCL}, as functions of the SST, T_{SS}, at marine surface relative fugacity of $\psi_f = 80\ \%\mathrm{rh}$, computed from TEOS-10 eqs. (49) – (52), as well as climatic LCL sensitivities, *α, β, γ,* eq. (53), with respect to increasing SST (Feistel and Hellmuth 2024). The row printed in bold approximates the current global mean SST, see Fig. 18.

T_{SS} K	T_{LCL} K	p_{LCL} hPa	z_{LCL} m	α % K^{-1}	β K K^{-1}	γ hPa K^{-1}
286	281.883	963.093	423.468	–0.0483	0.9634	–0.2742
288	283.810	962.542	431.481	–0.0542	0.9629	–0.2773
290	285.735	961.984	439.660	–0.0608	0.9624	–0.2806
292	**287.659**	**961.419**	**448.017**	**–0.0680**	**0.9619**	**–0.2841**
294	289.583	960.847	456.561	–0.0759	0.9614	–0.2878
296	291.505	960.268	465.305	–0.0846	0.9608	–0.2917
298	293.426	959.680	474.263	–0.0942	0.9603	–0.2959
300	295.346	959.084	483.449	–0.1047	0.9597	–0.3004

As solutions of eqs. (49) – (52), height and base temperature of marine cumulus clouds, as a function of the sea-surface temperature T_0 at a sea-surface relative fugacity of $\psi_f = 80\ \%\mathrm{rh}$, are displayed in Fig. 16. Similarly, height and base temperature, as a function of the sea-surface relative fugacity RF of ψ_f at a sea-surface temperature $T_{\mathrm{SS}} = 292\ \mathrm{K}$, close to the current global mean SST, are displayed in Fig. 17. It is obvious that the LCL effect of the range of RF uncertainty exceeds significantly the effect caused by global SST rise, so that unknown minor systematic RF changes may easily disguise the thermal effects on marine cumulus clouds.

Global mean sea-surface temperature has risen from about 17.9 °C in 1956 to 18.8 °C in 2023 (Cheng et al. 2024), see Fig. 18. This pronounced climatic trend is expected to let the cumulus cloud base lift up while at the same time warming it, see Fig. 16, but not as much as the SST itself is increasing. The related climatic sensitivities possess complicated dependencies but may directly be derived by taking the related derivatives

$$\alpha \equiv \left(\frac{\partial A}{\partial T_{\mathrm{SS}}}\right)_{p_{\mathrm{SS}}, \psi_f} = -\left(\frac{\partial q}{\partial T_{\mathrm{SS}}}\right)_{p_{\mathrm{SS}}, \psi_f}, \beta \equiv \left(\frac{\partial T_{\mathrm{LCL}}}{\partial T_{\mathrm{SS}}}\right)_{p_{\mathrm{SS}}, \psi_f} \text{ and } \gamma \equiv \left(\frac{\partial p_{\mathrm{LCL}}}{\partial T_{\mathrm{SS}}}\right)_{p_{\mathrm{SS}}, \psi_f}, \tag{53}$$

of the TEOS-10 LCL equations (49) – (52) with respect to the surface temperature while keeping surface RH fixed (Feistel and Hellmuth 2024). Selected results for those sensitivities are given in Table

1 relative to 1 °C rise of SST, similar to that in the past 70 years (Fig. 16). Here, $\alpha \approx -0.07\ \%\ \mathrm{K}^{-1}$ describes the rate of increase of specific humidity at the sea surface, often dubbed the "Clausius-Clapeyron effect". The value of $\beta \approx 0.96$ indicates that the cumulus cloud base warms up slower that the ocean by about 4 %, and $\gamma \approx -0.28\ \mathrm{hPa}\ \mathrm{K}^{-1}$ is the LCL pressure lowering caused by ocean warming, corresponding to ascending clouds. The value $\beta < 1$ implies that the thermal downward radiation from the cloud base does not keep pace with the ocean upward radiation, so that the net climatic feedback of cumulus clouds is negative and acts against ocean warming. These clouds do not provide a physical explanation for the observed enhanced ocean warming.

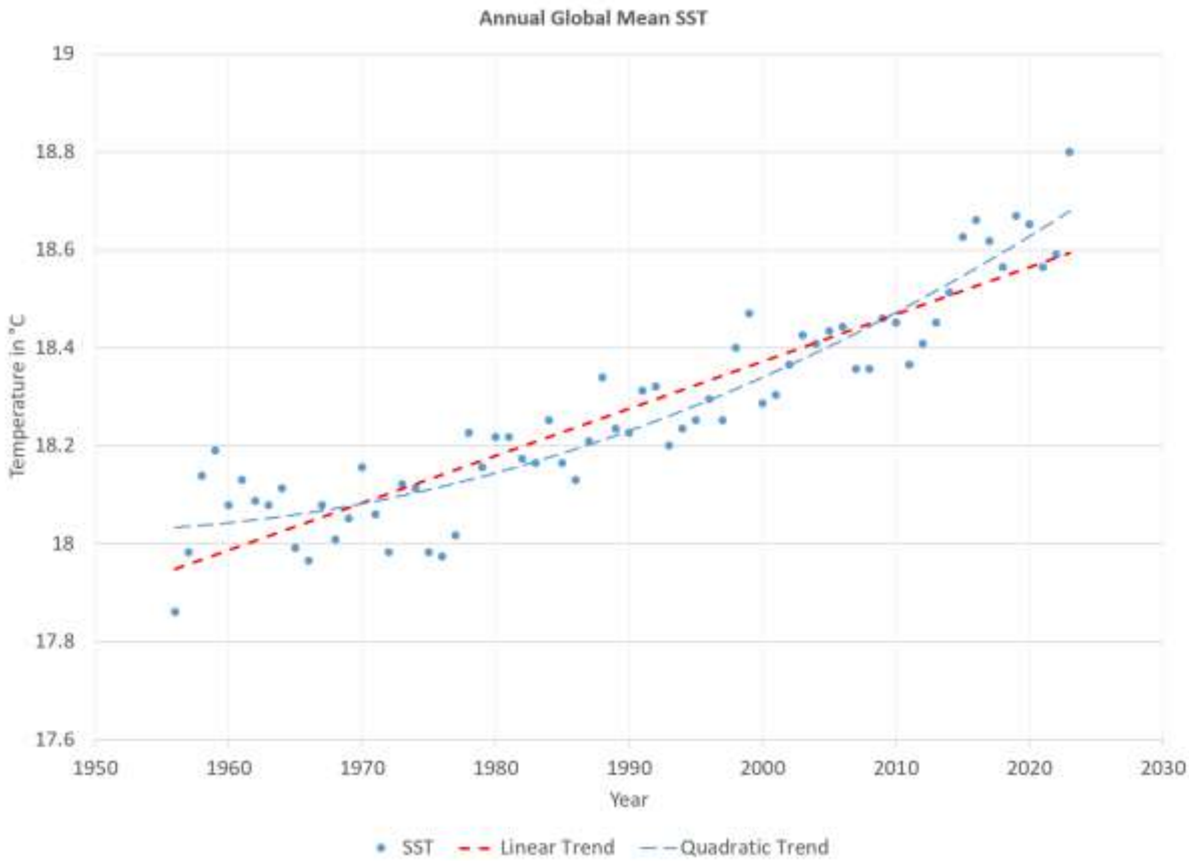

Fig. 18: Estimated increase 1957 – 2023 of global annual mean sea-surface temperatures (source: Cheng et al. 2024). The linear trend (red) is $t/°\mathrm{C} \approx 18 + 0.01 \times (yr - 1961)$. The quadratic trend curve (blue) suggests an acceleration of warming.

6.3 Stratocumulus and Other Clouds

"Marine low clouds strongly cool the planet" (Myers et al. 2021). Over the Atlantic, "the strongest surface longwave cloud effects were shown in the presence of low level clouds" (Kalisch and Macke 2012). "Low-cloud feedbacks are also a leading cause of uncertainty in future climate prediction because even small changes in cloud coverage and thickness have a major impact on the radiation budget" (Wood 2012: p. 2373).

Generally, however, the dominating cloud type over the ocean is stratocumulus (Eastman et al. 2011). "They are common over the cooler regions of subtropical and midlatitude oceans where their coverage can exceed 50% in the annual mean" (Wood 2012: p. 2373) with a typical thickness about 320 m and "a tendency for thicker clouds (median 420 m) in mid- and high latitudes" (Wood 2012: p. 2378). "Stratocumuli tend to form under statically stable lower-tropospheric conditions" (Wood

2012: p. 2374). On the annual average, stratocumulus is particularly frequent (up to 60 % coverage) at the subtropical coastal upwelling regions such as the cold Benguela, Humboldt and California Currents (Wood 2012: Fig. 4a, Muhlbauer et al. 2014: Fig. 2). However, in those areas there is no obvious correlation of cloud cover with ocean warming (Fig. 1). Stratocumulus also forms large cloud cover (about 20 % coverage) in the boreal and austral west-wind bands (Wood 2012: Fig. 4a) where the ocean is strongly warming up (Fig. 1).

"Only small changes in the coverage and thickness of stratocumulus clouds are required to produce a radiative effect comparable to those associated with increasing greenhouse gases" (Wood 2012: p. 2374). Marine stratocumulus cloud feedback is still a major challenge and source of uncertainty of climate models (Hirota et al. 2021). However, "similar to other low-cloud types in the marine boundary layer, the impact of stratocumulus clouds on the outgoing longwave radiation is marginal due to the lack of contrast between the temperature of stratocumulus cloud tops and the temperature of the sea surface over which they form. Thus, the net radiative effect of stratocumulus clouds is primarily controlled by factors influencing their shortwave cloud forcing such as the cloud albedo and the cloud coverage" (Muhlbauer et al. 2014: p. 6695).

Following this argumentation and assuming that the short-wave cloud effect of stratocumuli on the ocean radiation balance by far outweighs their long-wave effects, then the short-wave warming effect (Fig. 14) of decreasing cloudiness may dominate over the long-wave cooling (Fig. 15). Possibly, this could make stratocumulus a potential candidate for causing the unclear recent ocean warming.

Similarly, in the diurnal cycle, short-wave effects (Fig. 14) have an impact at daytime only, while long-wave effects (Fig. 15) are present all 24 hours. In this respect, Luo et al. (2024) report that low-level cloudiness has an asymmetric day/night trend which enhances ocean warming. Regionally, where in spring the days get longer, and the heavy cloudiness of the west-wind belt becomes replaced by fewer subtropical clouds (see Fig. 12), the systematic reduction of cloudiness may be expected to produce local excess warming such as near the subtropical fronts (see Fig. 1). Only dedicated future model studies, however, may reliably verify such speculations. As a recent example for the complexity of SST warming "by suppressing the evaporative cooling" of the ocean, Wang et al. (2024) explain dramatic but yet elusive warming events in the North-East Pacific by changes in ocean-atmosphere mechanisms caused by reduced Chinese aerosol emissions. Also, Berthou et al. (2024) describe cloud cover feedback over the sea during an unprecedented marine heatwave off northwest Europe in 2023. Excess water vapour emitted from a 2022 submarine volcano eruption may as well have contributed to the observed 2023 heat anomaly (Vömel et al. 2022, Jucker et al. 2024).

7 Summary

Substantial uncertainties of estimated heat fluxes at the ocean-atmosphere interface, such as the "ocean heat budget closure problem", prevent reliable model predictions and causal explanations of climate phenomena that may take place within the range of those uncertainties. Among such "surprises" is the currently registered excessive ocean warming, but are also the subsequent consequences of this warming, such as those expected for global weather processes.

Intending to reduce model uncertainties of thermal energies and heat fluxes in the climate system associated with the global circulation of water in its different phases and mixtures, the new geophysical thermodynamic standard TEOS-10 had been adopted internationally in 2009 and 2011. Meanwhile, the uptake of TEOS-10 by the scientific community is mainly focussed on ocean observations and modelling, as the related publication metrics are suggesting (Appendix A).

TEOS-10 is advanced over previous similar standards and various collections of tailored empirical property equations by (i) its completeness in describing all thermodynamic properties of seawater, humid air and ice, including their entropies, enthalpies and chemical potentials, by (ii) its perfect mutual consistency between different phases and mixtures, and by (iii) its minimum uncertainty over maximum ranges of validity. Among its particularly favourable fields of application are composite systems with internal phase boundaries such as air sea interaction or cloud formation.

In addition to entropies, enthalpies and chemical potentials, TEOS-10 has made available certain new quantities for the description and modelling of climate processes, such as (i) Absolute Salinity of the ocean with a specified Reference Composition, (ii) Conservative Temperature as a measure of Potential Enthalpy of seawater representing a definite heat content, and (iii) Relative Fugacity as the thermodynamic driving force of evaporation, suggesting an improved full-range definition of relative humidity as a substitute for mutually inconsistent and restricted such definitions in practical use in climatology, meteorology and physical chemistry.

This paper explains some tutorial examples for the application of TEOS-10 to selected current climate problems. There is (i) the two-phase conceptual model of "sea air" which provides rigorous equations for the latent heat of evaporation, for the heat capacity of humid air including salty aerosols (sea spray), and for the irreversible production of entropy by evaporation into the marine troposphere. There is also (ii) the formation of low marine cumulus clouds by isentropic thermal convection up to their condensation level, and their climatic feedback to surface temperature and humidity concerning their infrared radiation effects.

It is currently unclear why and how the ocean warming is intensifying, and when and how the related enormous amount of heat may transfer to the atmosphere. The observed systematic reduction of cloudiness may play an important role in this process, but responsible details and theoretical causes are unknown. Marine surface relative humidity is an important and rather sensitive "control valve" for the supply of the troposphere with latent heat, however, the common assumption of constant relative humidity during climate change lacks rigorous explanation and leaves open the question of its possible trends below the insufficiently high level of observational uncertainty. TEOS-10 may further assist climate modellers to address such issues.

Ocean Science has proved a scientifically well-reputable, reliable and successful partner journal for the publication of advanced results and methods in oceanography and geophysics. Cooperation with international bodies such as IUGG, UNESCO/IOC, IAPSO, SCOR, IAPWS and BIPM has made possible the development and international introduction of TEOS-10. The established standing committee JCS remains active with respect to related fundamental problems yet to be solved. It is hoped and expected that TEOS-10 may constitute a reliable long-term thermodynamic basis for interdisciplinary climate research.

Appendix A: Summary and Metrics of Selected Publications Related to TEOS-10

Between December 2008 and December 2012, supporting the activities of SCOR/IAPSO WG127, *Ocean Science* had published 16 articles open-access in its Special Issue #14, "Thermophysical properties of seawater" (Feistel et al. 2008a). From February 2013 on, monthly metrics have been recorded by the journal. Table A1 reports those metrics of the last decade.

For comparison, metrics – as far as published elsewhere by 04 April 2024 – of selected TEOS-10 articles listed at www.teos-10.org are reported in Table A2.

Table A1: Metrics of articles in the *Ocean Science* Special Issue #14, "Thermophysical properties of seawater" (Feistel et al. 2008a), from February 2013 till March 2024. "SIA" stands for the TEOS-10 Sea-Ice-Air open source code library.

Reference	Topic	Accessed	PDF Downloads	Cited
Millero and Huang (2009)	Seawater at High T,S	16 462	11 061	79
Feistel et al. (2010c)	Baltic Sea Density/Salinity	15 435	11 385	92
Pawlowicz et al. (2011)	Seawater Biogeochemistry	9 663	6 444	47
McDougall et al. (2012)	Global Absolute Salinity	9 290	5 489	116
Feistel et al. (2010a)	Humid Air Helmholtz Function	8 737	5 346	31
Safarov et al. (2009)	Seawater at High T,p	7 356	4 308	68
Wright et al. (2011)	Density Salinity	5 268	2 891	49
Marion et al. (2009)	$CaCO_3$ Solubility	5 169	3 170	36
Pawlowicz (2010)	Composition Variation	4 471	2 666	27
Feistel et al. (2010d)	SIA Library Equations	4 255	2 416	23
Wright et al. (2010)	SIA Library Routines	4 049	1 733	19
Feistel et al. (2008b)	Consistent New Potentials	3 585	1 527	27
Seitz et al. (2011)	Salinity Traceability	3 363	1 705	24
Feistel et al. (2010b)	Baltic Property Anomalies	3 183	1 500	12
Tailleux (2009)	Mixing Efficiency	2 752	1 303	11
Millero and Huang (2010)	Seawater at High T,S (corrig.)	2 189	909	1

Table A2: Metrics published by March 2024 of selected TEOS-10 related articles apart from *Ocean Science* Special Issue #14. "Ice Ih" is the ambient, hexagonal ice I phase of water.

Reference	Topic	Accessed	PDF Downloads	Cited
Wagner and Pruß (2002)	Water Helmholtz Function	7 516	7 516	3 457
Jackett et al. (2006)	Algorithms for Seawater	2 877	2 364	119
Feistel (2005)	Seawater Gibbs Function	2 584	1 126	10
Feistel et al. (2005)	Ice Ih Gibbs Function	2 288	1 015	5
Lemmon et al. (2000)	Dry Air Helmholtz Function	2 279	2 279	381
McDougall (2003)	Potential Enthalpy	1 970	1 367	50
Wagner et al. (2011)	Ice Melting/Sublimation	1 467	510	102
Seitz et al. (2010)	Salinity Determination	1 332		15
Feistel (2008b)	IAPWS-06 and IAPWS-08	1 279		4
Millero et al. (2008)	Seawater Composition	970		780
Feistel and Wagner (2006)	Ice Ih Gibbs Function	843	843	286
Feistel and Wagner (2005)	Ice Ih Gibbs Function	833		58
Graham and McDougall (2013)	Conservative Temperature	651	467	28
Feistel (2012)	New TEOS-10 Standard	436		27
Spall et al. (2013)	TEOS-10 for oceanography	230	128	3
Feistel (2008a)	Seawater Gibbs Function	134		133
Roquet et al. (2015)	TEOS-10 Polynomials	111		97
Feistel and Wagner (2007)	Ice Ih Sublimation >20 K	105		112
Feistel (2003)	Seawater Gibbs Function	100		105
McDougall et al. (2013)	Thermodynamics of Seawater	35		10
Feistel and Marion (2007)	Seawater Gibbs-Pitzer	25		32
Valladares et al. (2011)	Replacement of EOS-80	14+5		4+1
Feistel et al. (2006)	New Seawater Equation			

Table A3: IAPWS documents supporting TEOS-10, openly accessible at www.iapws.org. IAPWS documents are independently and painstakingly verified before they may become adopted at an annual meeting. No metrics available.

Document	Code	Topic	Meeting	Year
Release	R06-95	Water Helmholtz Function	Dresden	2016
Release	R10-06	Ice Ih Gibbs Function	Doorwerth	2009
Release	R13-08	Seawater Gibbs Function	Berlin	2008
Release	R14-08	Ice Melting/Sublimation	Pilsen	2011
Suppl. Release	SR1-86	Water Saturation Properties	St. Petersburg	1992
Suppl. Release	SR6-08	Liquid Water at 0.1 MPa	Pilsen	2011
Suppl. Release	SR7-09	Liquid Water Gibbs Function	Doorwerth	2009
Guideline	G05-01	Fundamental Constants	Virtual Online	2020
Guideline	G08-10	Humid Air Helmholtz Function	Niagara Falls	2010
Guideline	G09-12	Cold Water Vapour < 130 K	Boulder	2012
Guideline	G11-15	Fugacity Virial Equation	Stockholm	2015
Guideline	G12-15	Supercooled Water	Stockholm	2015
Advisory Note	AN4-09	IAPWS/CIPM Water Density	Doorwerth	2009
Advisory Note	AN5-13	Industrial Seawater	Dresden	2016
Advisory Note	AN6-16	IAPWS support for TEOS-10	Dresden	2016

Table A4: Numbers of unique internet downloads 2011-2023 of supporting material from the TEOS-10 homepage at www.teos-10.org. "GSW" stands for the TEOS-10 Gibbs Seawater open source code library. Data from Pawlowicz (2023)

Item	**2011-13**	**2013-14**	**2014-15**	**2015-16**	**2016-17**	**2017-18**	**2018-19**	**2019-20**	**2020-21**	**2021-22**	**2022-23**
TEOS-10 Manual	920	360	535	552	418	427	349	472	479	482	530
Getting Started	879	362	558	547	427	475	349	444	460	483	479
Lecture Slides	704	284	374	318	219	248	204	272	272	231	272
TEOS-10 Primer	584	197	289	297	222	217	187	253	260	226	268
GSW MATLAB	1920	1102	1485	1814	1235	1552	1233	1556	1504	1747	1897
GSW FORTRAN	366	222	171	162	127	116	82	98	83	92	87
GSW C	202	84	133	151	85	96	59	81	58	49	57
GSW PHP	-	55	61	43	29	60	28	52	22	22	21
SIA VB	72	100	46	45	45	48	43	47	47	38	30
SIA FORTRAN	59	118	58	44	36	42	37	42	31	33	31

Table A5: Selected additional TEOS-10 related readings, metrics by March 2024

Reference	Topic	Accessed	PDF Downloads	Cited
Turner et al. (2016)	Seawater Pitzer Model	13 780	1 175	21
Lovell-Smith et al. (2016)	Relative Humidity Challenges		6 502	27
Schmidt et al. (2018)	Density-Salinity Relation	9 421	5 481	28
Feistel et al. (2016a)	Challenges beyond TEOS-10		5 023	49
Dickson et al. (2016)	Seawater pH Challenges		2 818	43
Pawlowicz et al. (2016)	Seawater Salinity Challenges		2 738	40
Smythe-Wright et al. (2019)	IAPSO's history and roles	5 893	384	3
Foken et al. (2021)	Atmospheric Measurements	5 709		2
Feistel (2018)	TEOS-10 Review	5 441	1 632	38
Feistel and Hellmuth (2023)	Dalton Equation	5 068		1
McDougall et al. (2021)	Ocean Heat Flux and Content	4 993	1 425	5
Hellmuth et al. (2020)	Ice-Crystal Nucleation	4 811		6
Uchida et al. (2019)	Optical Density Sensor	3 513		19
Hellmuth et al. (2021)	Mass Density of Humid Air	2 643		4
Feistel and Lovell-Smith (2017)	Relative Fugacity Part 1		1 335	18
Le Menn et al. (2018)	Seawater Salinity Measurands		1 136	13
Schmidt et al. (2016)	Seawater Density up to 1 ppm		950	21
Ji et al. (2021)	Absolute Salinity off China	2 462	856	3
Von Rohden et al (2016)	Baltic Sound Speed	2 122	784	1
Feistel et al. (2016b)	Uncertainty of Correlation Eqs.		662	14
Martins and Cross (2022)	TEOS-10 Excel Code	2 087	542	2
Hellmuth and Feistel (2020)	Low-Density Subcooled Water	1 827		1
Feistel (2011a)	Stochastic Potential Functions	1 217		6
McDougall et al. (2014)	Sea Ice Formation	1 124	771	16
Feistel and Hellmuth (2024a)	Evaporation Entropy	1 038		0
Young (2010)	Boussinesq Approximation	928	724	56
Harvey et al. (2023)	Water Properties	874	369	9
Tailleux (2018)	Local Available Potential Energy	807	409	11
Uchida et al. (2020)	Seawater Intercomparison	707	764	6
Sharkawy et al. (2010)	Review of Seawater Correlations	701		946
Feistel (2019a)	Relative Fugacity Part 2		267	3
Feistel et al. (2022)	Relative Fugacity Part 3		252	4
McDougall et al. (2023)	Seawater Potential of (S, CT, p)	629	122	1
Feistel et al. (2015)	Virial Fugacity Equation	581		17
Nayar et al. (2016)	Seawater Property Review	553		366
Marion et al. (2011)	Seawater pH	491		170
Feistel and Lovell-Smith (2023)	Systematic Error in Regression	428	41	
Holzapfel and Klotz (2024)	H_2O and D_2O Ice Ih	285	58	
Ji et al. (2024)	Bohai Sea Salinity Anomaly	254	56	
Holzapfel and Klotz (2021)	Thermal Expansion of Ice Ih	245	77	1
Pawlowicz and Feistel (2012)	TEOS-10 in Limnology			22
Kretzschmar et al. (2015)	Industrial Seawater Equation	104		0
Ebeling et al. (2020)	Individual Ionic Activities	99		10
Marion et al. (2010)	FREZCHEM Solution Model	82		74
Sun et al. (2008)	Saline Thermal Fluid Equations	79		84
Almeida et al. (2018)	TEOS-10 Atlantic Impact	53		5
Safarov et al. (2012)	High-Salinity Seawater	42		21

Woosley et al. (2014)	World Ocean Absolute Salinity	39		16
Safarov et al. (2013)	Brackish Seawater Properties	35		15
Ulfsbo et al. (2015)	Seawater Activity Coefficients	34		11
Feistel and Hagen (1998)	Sea Ice Gibbs Function	24		31
Feistel (2010)	Seawater Gibbs Function	23		24
Tailleux (2010)	Buoyancy Power Input			20
Millero and Huang (2011)	Seawater Compressibility	19		19
Tchijov et al. (2008)	Ice at High *p* and Low *T*	19		6
Von Rohden et al. (2015)	Seawater Sound Speed 0.1 MPa			18
Budéus (2018)	TEOS-10 Density Bias ?	8		5
Lago et al. (2015)	Seawater Sound Speed < 70 MPa	8		4
Manaure et al. (2021)	Individual Ionic Activities	8		2
Weinreben and Feistel (2019)	Anomalous Salinity Density	8		1
Pawlowicz & Yerubandi (2024)	Water as a Substance	4		0
Ebeling et al. (2022)	Individual Ionic Activities	2		6
Waldmann et al. (2022)	Uncertainty of Ocean Variables	2		
Tailleux and Dubos (2024)	Seawater Static Energy	1		1
Pawlowicz (2013):	Physical Variables in the Ocean			
Laliberte (2015)	TEOS-10 Python Code			
Thol et al. (2024)	N_2-O_2-Ar Helmholtz Function			

Appendix B: Thermodynamic Potentials

This Appendix provides a short introduction to thermodynamic potentials, supporting the equations and topics discussed in this article. Alternative presentations from different perspectives are available from numerous textbooks such as Guggenheim (1949), Margenau and Murphy (1964), Landau and Lifschitz (1966) or Kittel (1969). For seawater, the use of a Gibbs thermodynamic potential was first suggested theoretically by Fofonoff (1958, 1962), see also Craig (1960).

A key theoretical tool for the physical investigation of the globally warming climate and the related energy balances is *thermodynamics*. It is known from experience that there exists a distinguished state of various ambient substances that is known as a *thermodynamic equilibrium state*. If a sample of matter is in this state, it may never spontaneously alter its measurable macroscopic properties unless it becomes disturbed by external contact and exchange of energy or matter with its surrounding. Typical properties which characterise a particular equilibrium state are the total mass of a sample, m, its volume, V, its temperature, T, or its pressure, p. Of a given sample, different equilibrium states may exist that differ in those quantities, but there exists a specific relation between those variables, known as an *equation of state*, which is characteristic for the given substance and remains universally valid at any of its possible equilibrium states. The most general and comprehensive equation of state of a given substance is a *thermodynamic potential* of that substance.

Thermodynamics is a mathematical theory for the construction and exploitation of equations of state and of properties derived therefrom for the prediction or verification of observations or experiments. Depending on the properties of interest, equations of state may be formulated in various different mathematical forms. It was discovered by J. Willard Gibbs (1873) that from a suitable thermodynamic potential all thermodynamic properties of a given substance at any of its equilibrium states can be derived by appropriate mathematical methods.

For theoretical reasons (namely, the statistical so-called *canonical ensemble*, Landau and Lifschitz 1966: §31; Kittel 1969: Ch. 18), a preferred thermodynamic potential of a pure substance is its

Helmholtz Energy, or *Free Energy*, $F(m,T,V)$, expressed in terms of the sample's mass, m, its temperature und volume. For mixtures, the single mass must be replaced by the set of partial masses of the species involved. Here, mass is used as a measure for the amount of substance, rather than particle or mole numbers, for the practical reason that in oceanography masses are easier measured than moles, and so TEOS-10 is following that tradition and is a mass-based description. Classical empirical thermodynamics of Clausius and Gibbs was formulated independently of the existence and properties of atoms or molecules which presently define the mole (BIPM 2019).

To the *Internal Energy* E of the sample, the Helmholtz energy is related by the Helmholtz Differential Equation,

$$E = F - T\left(\frac{\partial F}{\partial T}\right)_{m,V} \tag{B.1}$$

Note that IOC et al. (2010) uses the symbol U for the Internal Energy rather than E in eq. (B.1). This replacement is done here for denoting with u the wind speed, eq. (6), rather than specific internal energy, which is defined here by $e = E/m$, eqs. (1) and (B.3). The symbol E is frequently used in the thermodynamic literature, for example by Gibbs (1873a) and Landau and Lifschitz (1966).

The potential function F is extensive, which means that for instance $F(2\,m,T,2\,V) = 2\,F(m,T,V)$ is valid for an equilibrium sample of twice the mass. It follows that the mass-specific Helmholtz function, $F/m \equiv f(T,\rho)$, depends on two variables only, T and the mass density, $\rho \equiv m/V$, and is mathematically simpler and more convenient than F, which may always be retrieved from a given f by

$$F(m,T,V) = m \times f\left(T, \frac{m}{V}\right). \tag{B.2}$$

The quantitative description of a substance of interest in the form of a thermodynamic potential such as $f(T,\rho)$ has axiomatic properties. The description is *complete*, i.e., all thermodynamic properties of that substance are available, it is *consistent*, i.e., for any property one and only one result can be derived, and it is *independent*, i.e., no part of this description may be omitted without loosing the completeness. It is obvious that such axiomatic properties are very desirable for the description of geophysical substances, however, such thermodynamic potentials are rarely found in the corresponding literature. In particular in climate research which combines results and data from different disciplines, such as meteorology and oceanography, from research carried out all over the globe and over the years by subsequent generations of specialists, international binding standards such as the International System of Units (SI) are required that ensure mutual consistency and metrological comparability of any involved data produced from experiments, observations and models.

Gibbs' (1873a) original potential function was (internal) energy, $e = E/m$. It is known that a sample's energy can be increased by compression, $-p\mathrm{d}v$, where $v = 1/\rho$ is the specific volume, or by input of heat, $T\mathrm{d}\eta$, where $\eta = N/m$ is the specific entropy. As an extensive quantity, entropy introduced by Clausius (1865, 1976) is denoted here by N to avoid confusion with seawater salinity, S. Energy conservation implies that

$$\mathrm{d}e = T\mathrm{d}\eta - p\mathrm{d}v. \tag{B.3}$$

Any such change between different equilibrium states of the same sample takes place along a definite, substance-specific surface $e(\eta,v)$ so that $\mathrm{d}e$ in eq. (B.3) is mathematically an exact differential and the partial derivatives of e possess the physical meanings that

$$T = \left(\frac{\partial e}{\partial \eta}\right)_v, \qquad -p = \left(\frac{\partial e}{\partial v}\right)_\eta. \tag{B.4}$$

Gibbs (1873b) also demonstrated that for several equilibrium samples in contact with one another, in absence of gravity or accelerated motion, the samples are in mutual equilibrium only if they have equal values of the coefficients T and p of eq. (2.3).

In the geophysical practice, the quantities η and v are difficult to measure, in contrast to, say, T or p. Mathematically equivalent to $e(\eta, v)$, thermodynamic potentials in terms of the other three possible pairs of independent variables are formally obtained from so-called Legendre transforms (Alberty 2001), namely the *Helmholtz function* $f(T, v) \equiv e - T\eta$ with the differential

$$\mathrm{d}f = -\eta \mathrm{d}T - p\mathrm{d}v, \qquad \text{(B.5)}$$

the *Gibbs function* $g(T, p) \equiv f + pv = e - T\eta + pv$ with

$$\mathrm{d}g = -\eta \mathrm{d}T + v\mathrm{d}p, \qquad \text{(B.6)}$$

and the specific *enthalpy* $h(\eta, p) \equiv g + T\eta = f + T\eta + pv = e + pv$ with

$$\mathrm{d}h = T\mathrm{d}\eta + v\mathrm{d}p. \qquad \text{(B.7)}$$

Depending on the application purpose, each of these potential functions has certain advantages and disadvantages, and having all of them optionally at hand in mutually consistent versions is most useful.

Gibbs (1874-78) also considered a situation in which a given sample may exchange substance with its surrounding. If the exchanged mass of substance *i* is $\mathrm{d}m_i$, the related change of the sample's (extensive) energy E at constant entropy and volume is termed the *chemical potential* μ_i of that substance,

$$\mathrm{d}E = T\mathrm{d}N - p\mathrm{d}V + \textstyle\sum_i \mu_i \,\mathrm{d}m_i, \qquad \text{(B.8)}$$

so that this exact differential implies that the chemical potential is obtained from

$$\mu_i \equiv \left(\frac{\partial E}{\partial m_i}\right)_{N,V,m_{j\neq i}} = \left(\frac{\partial F}{\partial m_i}\right)_{T,V,m_{j\neq i}} = \left(\frac{\partial G}{\partial m_i}\right)_{T,p,m_{j\neq i}} = \left(\frac{\partial H}{\partial m_i}\right)_{N,p,m_{j\neq i}}. \qquad \text{(B.9)}$$

Equilibrium of a spatially extended substance, in absence of gravity or accelerated motion, requires that in addition to T and p, also the chemical potential μ_i separately for each present substance needs to possess the same value anywhere in the volume. "The potential for each component substance must be constant throughout the whole mass" (Gibbs 1874-78: p. 119).

As intensive properties, the specific energies cannot depend on the total mass but only on the mass fractions, $w_i \equiv m_i/m$. Because by definition $\sum w_i = 1$, only $(n-1)$ different fractions may be independent variables describing the n components of a mixture. For example, one of the components may be chosen as a master species, "0", such as a solvent, and the remaining ones, $i = 1, \dots, n-1$, may denote the solutes.

In terms of T and p, chemical potentials are computed from the Gibbs function, g, through the Gibbs energy, G, of eq. (B.9). Because the Gibbs function depends only on the independent intensive variables, $g(w_1, \dots, w_{n-1}, T, p)$, the solutes' chemical potentials, $i > 0$, are

$$\mu_i = \left(\frac{\partial G}{\partial m_i}\right)_{T,p,m_{j\neq i}} = \left(\frac{\partial (m\,g)}{\partial m_i}\right)_{T,p,m_{j\neq i}} = g + \left(\frac{\partial g}{\partial w_i}\right)_{T,p,w_{j\neq i}} - \sum_{j=1}^{n-1} w_j \left(\frac{\partial g}{\partial w_j}\right)_{T,p,w_{k\neq j}} \qquad \text{(B.10)}$$

Similarly, the solvent's chemical potential is

$$\mu_0 = \left(\frac{\partial G}{\partial m_0}\right)_{T,p,m_{j>0}} = \left(\frac{\partial (m\,g)}{\partial m_0}\right)_{T,p,m_{j>0}} = g - \sum_{j=1}^{n-1} w_j \left(\frac{\partial g}{\partial w_j}\right)_{T,p,w_{k\neq j}}. \qquad \text{(B.11)}$$

Therefore, the *relative chemical potentials* of the solutes are simply the partial derivatives,

$$\mu_i - \mu_0 = \left(\frac{\partial g}{\partial w_i}\right)_{T,p,w_{j\neq i}}. \qquad \text{(B.12)}$$

For mixtures, $n > 1$, the differential (B.6) of the Gibbs function takes the more general form

$$\mathrm{d}g = -\eta \mathrm{d}T + v\mathrm{d}p + \sum_{i=1}^{n-1}(\mu_i - \mu_0)\mathrm{d}w_i. \qquad \text{(B.13)}$$

It follows straightforwardly from (B.10), (B.11) that the sum,

$$\sum_{i=0}^{n-1} \mu_i\, m_i = m\, g = G\,, \qquad \text{(B.14)}$$

equals the Gibbs energy itself (Gibbs 1874-78: eq. (96) therein, Guggenheim 1949, Landau and Lifschitz 1966, Kittel 1969). In particular, if $n = 1$, the Gibbs function g of a pure substance represents its chemical potential,

$$g = \mu. \qquad \text{(B.15)}$$

Where two phases of a pure substance are in contact at mutual equilibrium, such as saturated water vapour at the liquid water surface, the mathematically distinct Gibbs functions of those phases take equal values. This indispensable condition for mutual consistency between the thermodynamic potentials of TEOS-10 is rigorously obeyed by virtue of appropriate reference-state conditions (Feistel et al. 2008b).

Code/Data availability. TEOS-10 library code used for this paper is available from www.teos-10.org

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