

## Review#2

This manuscript describes the first observation of BrCN in the atmosphere, during the global NASA ATom mission. Measurable BrCN was found within polar (Arctic & Antarctic) boundary layers, and was found in the presence of decreased O<sub>3</sub> (relative to background) and increased reactive Br species BrCl and BrO. These measurements are highly novel and provide new understanding of polar bromine chemistry. Detailed comments are provided below and primarily focus on data quantitation, additional experimental information needed, and needed clarifications.

We thank the reviewer for the positive comments about the novelty and importance of this work. We have the following responses to the questions and comments.

In Section 2.3, the authors state that only m/z 190 (and not the second isotope at m/z 192) was used to quantify BrCN (Line 193). They say that they divided the signal by its isotope abundance (0.51) "due to the calibration method employed". Please provide additional details about the calibration described on Lines 195-198. It is confusing then that the calibration was "compared to the sum of...m/z 190 and 192" (Page 6, Line 198). It is quite a coincidence that the "slope of the correlation of the CIT-CIMS vs I-CIMS is 0.52 +/- 0.01" (Line 207) (i.e. within error of the isotope ratio), when the CIT-CIMS m/z 190 signal was divided by 0.51. Perhaps there was a simple mistake in the data processing? This calibration is critical to the results presented because this slope led to correction factors being applied to the BrCN data from both the CIT-CIMS and I-CIMS.

We are sorry for the confusion concerning this point. The CIT-CIMS calibration is correct as written. The signals at m/z 190 and 192 were summed during calibration to improve the signal/noise of the analysis, so that the calibration factor included the sum of ion counts from both masses. Subsequent analysis revealed an intermittent interference at m/z 192 (the CIT CIMS had unit mass resolution), so that only m/z 190 could be used for quantitation. As a consequence, that signal at 190 was divided by its fractional isotope abundance before application of the calibration factor that included both ions. This procedure could not be responsible for the original CIT-CIMS data being lower by a factor of 0.52, since the signal that had not been corrected in this manner would have yielded even lower ambient mixing ratios.

Sections 2.2, 2.3, and 2.4.1: Describe the aircraft inlets (lengths, flows, materials, temperatures, etc) used for the two CIMS instruments and chemiluminescence instrument, as many of the compounds being measured can be impacted by line losses. I expect that BrO could have significant line losses, which is why the high flow Eisle-type inlet is typically used on the ground (Liao et al. 2011, JGR, A comparison of Arctic BrO measurements by CIMS and LP-DOAS); this, in particular, is not addressed.

The aircraft inlet for I-CIMS BrCN measurements is partially described by Veres et al., (Veres et al., 2020). Additional information is that the inlet consisted of a 0.75 m length of 0.75cm ID PFA

Teflon tubing thermostated at 35°C and had a flow rate of 6 slpm. The description in the paper will be amended to include this information. Several pieces of evidence argue against significant loss of BrCN inlets: BrCN equilibrates rapidly in the short PFA inlet, consistent with our understanding of the behavior of inorganic species on Teflon inlets as a function of Henry's Law. (see Liu et al., ) (Liu et al., 2019). BrCN hydrolysis is relatively slow and is base catalyzed (Roberts and Liu, 2019), and inlets used during ATom were acidic due to exposure to acidic aerosol (Nault et al., 2020), so we would not expect loss of BrCN due to hydrolysis.

The inlet of the CIT CIMS has been described previously (Crouse et al., 2006; Allen et al., 2022) and consisted of a fluoro-polymer (Fluoropol PFC 801A, Cytonix Corp) coated glass tube under high flow. Consequently, the same aspects of BrCN behavior on surfaces noted above pertain to this inlet as well. We will include this text and associated references in the revised manuscript.

The inlets for the O<sub>3</sub> and NO<sub>y</sub> chemiluminescence instruments have been described previously (Bourgeois et al., 2020) (Ryerson et al., 1999). The inlet for the NO measurement consisted of a 1.5 m long 4mm ID. PFA tube thermostated at 30°C at a flow rate of 1 slpm, coupled to a fused silica tube 12mm ID, 450mm length, used to match the residence time with the NO<sub>2</sub> channel of the instrument. The use of PFA Teflon for NO and O<sub>3</sub> sampling is well established and losses of NO or O<sub>3</sub> due to chemical reactions would only be expected under extreme circumstances, (sampling of highly polluted air with extremely high VOCs or NO). The NO<sub>y</sub> inlet consisted of a solid gold tube situated in the aircraft inlet and heated to 300°C. The loss or incomplete conversion of NO<sub>y</sub> species in the NO<sub>y</sub> convertor is routinely checked through the addition of NO<sub>2</sub> and HNO<sub>3</sub> standards, and found to be negligible. We do not use NO<sub>2</sub> data in this paper so will not discuss that inlet. We will put these details in the SI.

Losses of BrO on the actual I-ToF CIMS inlet were not apparent during calibrations that were performed after both ATom-3&4 deployments. The intercomparison of in situ BrO measurements by CIMS with Long Path Differential Absorption Spectroscopy (LP-DOAS) reported by Liao et al., (2011) (Liao et al., 2012) provides further evidence that our inlet did not have significant BrO losses. Aside from a large shroud that was required to establish flow from a static air mass, (not required for an aircraft inlet) the key feature of the Liao et al. CIMS inlet was a short section PFA Teflon tubing heated to 40°C and operated at high flow rates. Liao et al., (2011) report excellent agreement between the two methods for periods of steady wind flow without local NO pollution. Our inlet is slightly longer than that CIMS inlet, but was made of the same material and operated at a high flow rate. That study, combined with our laboratory experience indicate that we did not have significant losses of BrO. We will add a short description of this to the experimental methods section of our paper.

BrCN backgrounds were conducted by scrubbing the ambient air (Line 143). Please describe the scrubber (material, temp, etc) and its measured efficiency for removing BrCN and BrCl.

3. The scrubber system consisted of ambient air pumped from separate aircraft inlet, via a diaphragm pump, through a packed trap containing activated charcoal, Purafil, Puracarb AM, and Chlorosorb Ultra, and added to the main inlet periodically as an overflow. The scrubber material

was independently checked for removal of halogens ( $\text{Br}_2$ ,  $\text{Cl}_2$ , and  $\text{BrCl}$ ) and other analytes of interest. We have every reason to expect that this scrubber also removed  $\text{BrCN}$ , however, we have examined zeroing periods that occurred in the middle of  $\text{BrCN}$  measurements. An example of such a period is shown in Figure R1, from the Oct 19, 2018 flight, in which several of the zero periods happened during elevated  $\text{BrCN}$  periods. Also, all the zero periods exhibited non-varying signals at or lower than the ambient signals observed. We will add more details concerning the zeroing method and will show Figure R1 in the SI of the manuscript.

The authors state concern for sampling inlet reactions, previously reported by Neuman et al. (2010), as the reason why  $\text{Br}_2$  and  $\text{HOBr}$  were not quantified. However, the presence of  $\text{Br}_2/\text{HOBr}$ , while impacted by sampling inlet reactions, is another clear indicator of active bromine chemistry. Neuman et al. (2010) previously reported  $\text{Br}_2$  as a lower limit of the sum of  $\text{HOBr} + \text{Br}_2$ . I believe that it would be very useful for the authors to report their  $\text{Br}_2$  and  $\text{HOBr}$  data, even with high uncertainties, since observations of these species are limited and their presence would provide further support for the current work, especially since the authors propose that  $\text{HOBr}$  is a precursor to  $\text{BrCN}$  formation.

4. We will not report  $\text{HOBr}$  and  $\text{Br}_2$  data as mixing ratios when we have already explained that we don't have confidence in them. However, we will show a plot of the  $\text{HOBr}$  and  $\text{Br}_2$  signals observed during the 5/19/18 flight in the SI of revised paper, as shown in Figure R2 in this response. We note that they are very similar to the  $\text{BrCl}$  data (which are quantified in our work) in that these active Br compounds are present, but neither correlate closely with the  $\text{BrCN}$  values, nor anti-correlate closely with ozone during those periods. We think that this is due to the fact that  $\text{BrO}$ ,  $\text{Br}_2$  and  $\text{BrCl}$  are all relatively short-lived thus indicate that active Br chemistry is currently happening, while  $\text{BrCN}$  and  $\text{O}_3$  loss are more persistent and tend to integrate the amount of active Br chemistry that has happened in an airmass. We will add these points of discussion to the text of the paper in the appropriate section.

The authors discuss the potential for  $\text{BrCN}$  formation from reactions on the sampling lines. Is there any evidence of the presence of reduced nitrogen compounds on the sampling lines, or is there a simple test that could be done to evaluate this? Were the  $\text{Br}_2$  and  $\text{BrO}$  calibrations (noted on Lines 175-176) done using the aircraft inlet or just the instrument, as currently implied? If both  $\text{HOBr}$  and  $\text{HCN}$  flow through the inlet, is  $\text{BrCN}$  formed?

5. We do not have any direct evidence for the presence of reduced nitrogen compounds on the aircraft inlet. There are several lines of evidence that demonstrate that  $\text{BrCN}$  was not produced from reaction of active Br compounds with reduced nitrogen species on the aircraft inlet. First, a number of  $\text{BrO}$  and  $\text{Br}_2$  calibrations were conducted on the aircraft inlet that had been used during ATom-3 but not otherwise cleaned, and no production of  $\text{BrCN}$  was observed. Second, the NOAA group has conducted preliminary lab tests aimed at examining the temperature and pH dependence of the  $\text{Br}_2 + \text{HCN}$  chemistry on ice/water surfaces. These tests were conducted with the quadrupole I-CIMS that had lower sensitivity to  $\text{BrCN}$  than the ToF CIMS ( $\text{DL} = 30$  pptv for

10sec averages). In those tests, combining Br<sub>2</sub> (28ppbv) and HCN (5 ppbv) in the PFA Teflon inlet of that instrument (0.4cm ID, 2m length, 2.20 slpm flow of room air at 30% RH) produced no BrCN above detection limit. Substantial production of BrCN was apparent when Br<sub>2</sub> and HCN were reacted in a high-surface area reactor containing a small amount of liquid water at pH8, qualitatively confirming the chemistry proposed in this work. We will add a description of this observation in the SI as it pertains to possible instrument inlet effects, however any further description of the BrCN production experiments would be premature.

The methods section (Lines 163-165) states that “clear signals for ICN were observed in ambient air” but were not quantified. Yet, there is no mention of ICN in the Results. It would be useful to learn if ICN was observed concurrently with BrCN, or under different conditions, even it is not quantifiable.

6. We feel the presence of ICN is worth noting, however, we do not want to report observations that we do not have confidence in. Instead, we will add a brief summary to the Results section describing the nature of the ICN observations alluded to in the rest of the paper. We will present the ICN signals during the 5/19/18 flight as a figure in the SI, Figure R3 below, and note that it shows essentially what was described: ICN is sometimes observed when BrCN was observed, but they are not closely correlated. In addition, the instrument background for ICN is relatively high and variable, and the signal in scrubbed air is sometimes slightly higher than in ambient air, often leading to apparent negative ICN mixing ratios. We will show this figure in the SI to support these points. We believe these features, generally high background and apparent negative concentrations, are due to the fact that the reagent ion is iodide, made from the addition of substantial amounts of CH<sub>3</sub>I, hence the possibilities for ion exchange and surface artefact formation are numerous. We further recommend that alternate ionization schemes (e.g. Br<sup>-</sup> ions) would probably yield a better measurement of ICN.

For the ATom-2 Arctic flights on Feb. 18 & 19, 2017, the authors assert on Lines 293-294 that “There was essentially no photosynthetic activity at the Northern latitudes during this time”, and state Lines 296-298 that “there was Br activation initiated by O<sub>3</sub>-Br-chemistry...because there was insufficient photochemistry to carry the gas phase catalytic Br chemistry”. This is again repeated in the conclusions on Line 593. Yet, no observational support or reference is provided for these statements. These flights include MAs over BRW and SCC, where Polar sunrise occurs in late January, and by mid-February there are several hours of sunlight. Further, Raso et al. (2017, PNAS) and Custard et al. (2017, ACS Earth & Space Chem.) previously showed photochemical snowpack Br<sub>2</sub> and I<sub>2</sub> production at the beginning of February to mid-Feb in Utqiagvik (BRW). In addition, Pratt et al. (2013, Nat. Geosc.) showed that ozone reaction with bromide is far less efficient than snowpack photochemistry, and this is supported by the results presented by Custard et al. (2017). The authors should re-evaluate their explanation.

7. We thank the reviewer for making these points. We agree that we need to reevaluate our explanation for the ATom-2 observations. From the actinic flux measurements made on board the

aircraft, we can see that there was sunlight during the Feb 18 & 19 2017 flights, although it was much reduced (factors of 10 or more) relative to the ATom-4 periods. Also, temperatures were substantially lower in February relative to May. As a consequence, any photosynthetic sources of BrCN, and CHBr<sub>3</sub> (and subsequent biotic conversion to CH<sub>2</sub>Br<sub>2</sub>) were much less important during this ATom-2 period relative to ATom-4. We are also grateful for the reviewer pointing out that the snow pack photochemical source of Br<sub>2</sub> is operative during the late Winter period. We believe these features account for the BrCN, and high CHBr<sub>3</sub>/CH<sub>2</sub>Br<sub>2</sub> ratios that were observed during ATom-2. We will change our text in the Results and Conclusions sections to reflect this and will include the suggested references in the process.

In Figure S4, BrCN increases up to ~40 ppt below 600 m and O<sub>3</sub> decreases to ~15 ppb in this same altitude range, but this anti-correlation is not clearly described in the text on Line 297. The text also rules out a bromine explosion without O<sub>3</sub> below 10 ppbv; yet, BrCl and BrO data from ATom-3 and ATom-4 are discussed without this full ozone depletion, which is inconsistent. Please clarify the text. Also, is information about the boundary layer height available for these data points?

8. It is true that both high BrCN and lower O<sub>3</sub> values were observed at the low altitudes in the ATom-2 flights. However, unlike the May 19, 2018 flight there, there was not a close anticorrelation in the ATom-2 BrCN and O<sub>3</sub> data when the individual times are considered. We will make this clear in the text when these data are presented. We agree that the two aspects of our text were not in agreement. There is no requirement that O<sub>3</sub> be below 10ppbv for the Br explosion chemistry to be happening. We will change the way the ATom-2 data are discussed in the text and simply note that the points from ATom 2 do not closely anti-correlate. We have added the potential temperature data measured on the aircraft, to Figure S4 (see Figure R4 below), and that shows that for most of the low altitude profiles, the potential temperature indicates close to neutral, or stable conditions at about 600m and below.

Abstract: It would be helpful to expand the abstract to briefly explain how the conclusions presented on Lines 26-29 were obtained, including mentioning that box modeling was conducted. Further, it would be helpful to also mention the observed seasonality of the BrCN and its relationship to the other compounds measured that currently missing from the abstract but are key to the study (O<sub>3</sub>, CHBr<sub>3</sub>/CH<sub>2</sub>Br<sub>2</sub>, BrCl, BrO; also mention in the abstract that these were measured).

9. We apologize, the abstract as submitted was left over from a version of the paper submitted to a journal that had a rather restrictive word limit on abstracts. We agree that it will be helpful to expand the abstract in the manner suggested by the Reviewer. We propose to use this more expansive abstract:

“Bromine activation (the production of Br in an elevated oxidation state) promotes ozone destruction and mercury removal in the global troposphere, and commonly occurs in both springtime polar boundary layers, often accompanied by nearly complete ozone destruction. The chemistry and budget of active bromine compounds (e.g. Br<sub>2</sub>, BrCl, BrO, HOBr) reflect the cycling of Br and affect its environmental impact. Cyanogen bromide (BrCN) has recently been

measured by iodide ion high resolution time-of-flight mass spectrometry ( $I^-$  CIMS), and trifluoromethoxide ion time-of-flight mass spectrometry ( $CF_3O^-$  CIMS) during the NASA Atmospheric Tomography mission 2<sup>nd</sup>, 3<sup>rd</sup> and 4<sup>th</sup> deployments (NASA ATom), and could be a previously unquantified participant in active Br chemistry. BrCN mixing ratios ranged from below detection limit (1.5 pptv) up to as high as 36 pptv (10 s avg) and enhancements were almost exclusively confined to the polar boundary layers in the arctic winter and in both polar regions during spring and fall. The coincidence of BrCN with active Br chemistry (often observable BrO, BrCl and  $O_3$  loss) and high  $CHBr_3/CH_2Br_2$  ratios imply that much of the observed BrCN is from atmospheric Br chemistry rather than a biogenic source. Likely BrCN formation pathways involve the heterogeneous reactions of active Br ( $Br_2$ , HOBr) with reduced nitrogen compounds, for example hydrogen cyanide ( $HCN/CN^-$ ), on snow, ice, or particle surfaces. Competitive reaction calculations of HOBr reactions with  $Cl^-/Br^-$  and  $HCN/CN^-$  in solution, as well as box model calculations with bromine chemistry, confirm the viability of this formation channel and show a distinct pH dependence, with BrCN formation favored at higher pHs. Gas phase loss processes of BrCN due to reaction with radical species are likely quite slow and photolysis is known to be relatively slow (BrCN lifetime of ~4 months in mid-latitude summer). These features, and the lack of BrCN enhancements above the polar boundary layer, imply that surface reactions must be the major loss processes. The fate of BrCN determines whether BrCN production fuels or terminates bromine activation. BrCN reactions with other halogens ( $Br^-$ , HOCl, HOBr) may perpetuate the active-Br cycle, however, preliminary laboratory experiments showed that BrCN did not react with aqueous bromide ion (<0.1%) to reform  $Br_2$ . Liquid phase reactions of BrCN are more likely to convert Br to bromide ( $Br^-$ ) or form a C-Br bonded organic species and so would constitute a loss of atmospheric active Br. Thus, accounting for the chemistry of BrCN will be an important aspect of understanding polar Br cycling.”

Both the abstract (Line 29) and conclusions (Lines 615-616) refer to condensed phase loss of BrCN resulting in either  $Br^-$  or C-Br bonds. What is this based on, and where is it shown in the text?

10. The condensed phase (i.e. solution phase) chemistry of BrCN is relatively well studied, since BrCN is used as a diagnostic reagent in biochemistry, and a reagent in organic synthesis. In fact, the reaction of BrCN with tertiary amines is a name reaction in Organic Chemistry (von Braun and Schwarz, 1902). The requested information was covered in Lines 578-584 in the original paper.

Additional comments:

- Lines 32-37, 42-46, 93-96, 99-101, 188, 329-330, 430-432, 560-565, 608-609: Add references to these sentences.

We will add references where requested, however there are three instances (42-46, 99-101, 560-565, where the reference(s) for these statements are given in the sentence immediately before the lines noted. Perhaps this is a style issue? We will clarify so that it is clear that the statements are supported by the preceding reference(s).

- Line 43: Note that Pratt et al. (2013, Nat. Geosc.) showed that Br<sub>2</sub> does not form on sea ice because the surface is buffered (Wren & Donaldson et al 2012, ACP). Also, Wang et al. (2019, PNAS, "Direct detection of atmospheric atomic bromine leading to mercury and ozone depletion") is an appropriate reference to include in this sentence.

We thank the referee for pointing out these features of the Br<sub>2</sub> chemistry. They are consistent with our theory that BrCN formation is favored on substrates that have high pH. We will note these aspects of Br<sub>2</sub> chemistry at this point and include the suggested references.

- Lines 82-92: It would be helpful if this intro discussion could be clarified, as it connects to later discussion of results. The following papers may be useful to consider incorporating: Swanson et al. 2007 (Atmos. Environ., "Are methyl halides produced on all ice surfaces? Observations from snow-laden field sites"), Rhew et al. 2007 (JGR, "Methyl halide and methane fluxes in the northern Alaskan coastal tundra"), Macdonald et al. 2020 (ACP, "Consumption of CH<sub>3</sub>Cl, CH<sub>3</sub>Br, and CH<sub>3</sub>I and emission of CHCl<sub>3</sub>, CHBr<sub>3</sub>, and CH<sub>2</sub>Br<sub>2</sub> from the forefield of a retreating Arctic glacier").

We will clarify the discussion at this point and include the references that the reviewer has pointed out to us. The methyl halides do show some minor production in snow packs that might be associated with photochemistry, or might be biological. The paper on emissions from glacial margins and forefields in particular showed CHBr<sub>3</sub>/CH<sub>2</sub>Br<sub>2</sub> ratios in the range of 1 to 3 (Macdonald et al., 2020). All of these observations are either tangential to, or support our assertion that CH<sub>2</sub>Br<sub>2</sub> and CHBr<sub>3</sub> are the important markers that can distinguish biogenic and abiotic formation.

- Lines 131-132: Please define "MA", "level legs", and other common flight terms here and elsewhere to make the text more readable to the non-aircraft audience.

We will add the following explanations to the text at this point, and continue using the terms 'level leg' and 'MA' in the rest of the manuscript since it has been defined here. The term 'level leg' refers to a flight segment during which the aircraft maintained a constant altitude ( $\pm$  a few percent) for at least several minutes or more. The term 'MA' will be defined here as 'missed approach' during which the aircraft descended into the controlled air space above a designated airfield, almost to the point of touching down, and climbed back out.

- Lines 149-150: Please provide information about the PTR-ToF, including the ion(s) used to detect ClCN.

The Hi-Res PTR-ToF instrument has been described in the literature (Yuan et al., 2016; Yuan et al., 2017), and the compound was detected at ClCN•H<sup>+</sup>, a process greatly aided by the Cl isotope distribution and negative mass defect.

- Lines 153-154: Were the BrCN and ICN in N<sub>2</sub>?

No, the standards were in zero air, which will be noted here.

- Lines 160-161: Please add these data showing the sensitivity as a function of IMR water vapor concentration to the SI. Also, were the analyte ions normalized to I•H<sub>2</sub>O<sup>-</sup> prior to calibration?

We will add figure R5, showing the water vapor dependence of the sensitivity to BrCN to the SI. The signals were normalized to I•H<sub>2</sub>O<sup>-</sup> prior to application of the appropriate sensitivity as will be noted in this section.

- Section 2.3: Describe how the HCN calibration and zero measurements were conducted. How frequent were the “periodic zero measurements”?

The CIT-CIMS instrument was zeroed every 15 minutes in flight by closing instrument to ambient air sampling and adding UHP N<sub>2</sub> gas. A second type of zero was also periodically performed by flowing ambient air through a filter containing NaHCO<sub>3</sub> coated nylon wool and palladium coated alumina pellets. Both types of zero generally agree well, except at very high ambient RH where HCN is not efficiently scrubbed by the NaHCO<sub>3</sub> filter. CIT-CIMS HCN calibration was performed in the lab using a compressed gas mixture of HCN in N<sub>2</sub>, and a dilution system composed of three mass flow controllers, and additional N<sub>2</sub> gas source (LN<sub>2</sub> boiloff). FTIR measurements were used to quantify, and monitor the stability of, the HCN mixing ratio in the compressed gas mixture using HITRAN HCN lines strengths. The laboratory HCN calibrations were proxied to flight data using periodic in-flight calibrations (every ~3 hours) of other species (H<sub>2</sub>O<sub>2</sub>, peroxy acetic acid, and ethylene glycol), after accounting for appropriate water vapor and temperature dependencies. We will include this material in the Experimental Methods section.

Lines 197-198: Define “IR” and “MS” acronyms, and check that other acronyms throughout the manuscript are all defined at their first use.

We will define IR (infrared) and MS (mass spectrometry) here, and check that the manuscript to make sure other acronyms are defined at first use.

- Figure 2b: Given the instrument sensitivity issues on several ATom-4 flights (Line 268), perhaps it would be helpful to distinguish data below the limit of quantiation in the figure, as it is more difficult to distinguish the polar enhancement from this figure due to the data points that appear to be below the LOQ.

We have attempted to modify Figure 2b in the manner the Reviewer has suggested by overlaying in grey the points from May 12, 17 and 21, 2018 that were below detection limit. We hope this makes the polar data points more apparent.



- Lines 275 and 278: Please state the dates of these BRW MAs for improved clarity.

We will add the requested information, Oct 2, 2017 for the ATom-3 MA over BRW and Apr 28, 2018, for the ATom-4 MA over BRW.

- Lines 290-301: It would be helpful to add a map of the ATom-2 flights to the SI for context.

We will add a map of the ATom-2 BrCN observations (Figure R7) to the SI.

- Line 333: Please refer to the section where these atmospheric lifetimes are discussed.

We will refer to Section 4.2 here.

- Line 334, Figure 4 slope, Figure S2 slope: Fix significant figures (e.g., should be 26.5 +/- 0.2 ppbv).

We are not sure we understand this point based on the Reviewer's example, as our number in Figure 4 also has 3 significant figures ( $3.75 \pm 0.17$ ). We think the Reviewer's point is that the uncertainty ( $\pm 0.17$ ) implies that only 2 significant figures are appropriate? We will change Figure 4 to read  $3.8 \pm 0.2$  and Figure S2 to read  $(0.52 \pm 0.01) - 0.34$ .

- Add references for R4-R6, S1-S4, EqS2

R4-R6 are not presented as fundamental reactions but rather to illustrate that active bromine reactions are in competition, the rate constants where available can be found in the literature described in the introduction. We note that we had mis-labeled the reaction that appeared on Line 572, it should have been R7, but it also has no reference as such, since it is being hypothesized here. We think this is clear from the text following R7. The reference for S1-S4 (Gerritsen et al., 1993) will be given in that section. We will add a good general reference for S2, effective Henry's Law constants (Sander, 2015).

- Line 328-329: Discuss altitudes and boundary layer heights here for context since the figure is in the SI.

We will now note that the layer in which the BrCN was observed extended up to 0.8km.

- Lines 335-336: Define the "intermediate polar boundary layer" and "middle of the polar boundary layer".

We would prefer not to try to discuss this here, as it is presented in the Discussion section, (see the paragraph starting on line 508 of the original manuscript). Instead, we will add a sentence "See the Discussion section for a further description of the polar boundary layer."

- Line 349: Clarify whether the Southern Ocean was frozen.

Unfortunately, video from the DC-8 is not available for that flight. The TERRA/MODIS imagery from that date (NASA and Worldview, 2018) and nearby dates indicates that the Weddell Sea was largely frozen, but we cannot be sure there were no open leads. We will add a sentence to this section about this.

- Lines 356-360: Refer to where these data are shown. If not currently included, please add as a figure in the SI.

We will add figures to the SI (figures R8 and R9 below).

- Lines 380-381: The authors are encouraged to also consider the work of Halfacre et al. (2014, ACP, "Temporal and spatial characteristics of ODEs from measurements in the Arctic") here.

We thank the reviewer for suggesting the paper, (Halfacre et al., 2014) which provides further evidence of the widespread nature of ODEs, and we will also add the reference for Ridley et al., who showed widespread ODEs in the Arctic springtime in the year 2000 (Ridley et al., 2003). We will include these in our revised paper at this point.

- Lines 498-499: Note that this statement does not appear to consider aldehydes or ketones.

No, we do not explicitly mention aldehydes and ketones. The point of the text in these lines was to note the condensed phase reactions for which we have rate constants and to examine the implied atmospheric lifetimes in relation to these rates. A search of the literature turns up a number of papers in which BrCN reacts with  $\beta$ -diketones by substitution of Br at the central carbon (ie. forming a C-Br bond) and a paper that describes a general route to epoxides (Li and Gevorgyan, 2011). The key point of this section is that we have not found any reactions in which BrCN reforms active-Br. We feel that any more extensive summary of BrCN condensed phase chemistry is beyond the scope of this work.

- Section 4.1: The modeling of the pH dependence of BrCN production (Figures S17 and S18) is noteworthy, and these results warrant more discussion in the main text.

We will present more discussion of this feature and in the process discuss some of the pH-dependent processes noted by this Reviewer above.

- Line 513: It is stated that "there are often clouds in the vicinity of the inversion", but no reference is provided for this statement.

This is discussed in one of our references (Serreze et al., 1992) and will be noted here.

- Lines 524, 561-62: Fix reference formatting.

These will be fixed.

- Lines 566-568: Describe the details of this laboratory experiment in the SI.

We will put a brief description of this experiment in the SI. The reactor and flow system was similar to that described previously (Roberts et al., 2009) and the instrument responses were calibrated via Br<sub>2</sub> permeation tubes and the BrCN diffusion source.

- Lines 596-597: Provide the date for context.

These will now be added to the text at this point.

- Lines S55-S56: Are both [Cl<sup>-</sup>] and [Br<sup>-</sup>] assumed to be 16 μM? No reference is provided, and previous measurements by Krnavek et al (2012, Atmos. Environ.) suggests that this assumption is not appropriate.

The 16μM refers to the sum of Cl<sup>-</sup> and Br<sup>-</sup>. This is at the low end of the distributions of these ion concentrations measured in snow and ice. We will note that and include the suggested reference (Krnavek et al., 2012), and further references to support this point. Higher halide concentrations would have the effect of shifting the blue dashed line in Figure S17 upwards, and we will note that in the SI text.

- Line 598: Change “ice and particle” to “ice, and/or particle”.

This will be changed.

- Line 605: Note that the sea ice surface is buffered (Wren & Donaldson et al 2012, ACP).

We will include this in the text at this point.

- Figure 1: State in the caption that this was during ATom-3.

We will include this in the Figure caption.

- Figure 3: In the caption, provide the full date ranges for ATom-3 and ATom-4 for context, and also provide the general locations for the highlighted flights. Since a significant fraction of the data are below LOD and LOQ, please state these in the caption for context, or perhaps draw lines in the figure.

We will include the full range of dates of ATom3&4 in the figure caption, and provide the general locations for the dates that are shown explicitly. We will modify the figures to include lines for the LODs.

- The data in many of the figures is below LOD and/or LOQ, it would be helpful to indicate these values in the captions or show more clearly in the figures.

We will repeat the LODs and dates in the figure captions for Figures 2&3, which are really the only ones where all the data are summarized.

- Figures 11, S7, S9, S11, S19: In the captions, state the location where these data were collected.

Much of this requested information is redundant, for example the 22:30 UTC leg on May 19 is shown in several figures and the profiles shown in Figure S11 are at the locations shown in Figure 6a. Nevertheless, we will put the latitudes and longitudes in the figure captions. We also added this information to the caption for Figures 5&6.

- Figure S2: Add error to the slope fit through 0.

The uncertainty ( $\pm 0.008$ ), will be noted in the figure.

- Table S1: Define abbreviations in the caption.

The abbreviations in the Table title will be spelled out.

Table S2: My understanding from Line 298 is that the LOD for BrCN during ATom-4 was higher. Please clarify that here. Also indicate the m/z used for quantitation of each compound.

As we stated in the experimental approach section, there were 3 flights for which the instrument sensitivity was degraded due to the use of higher water vapor in the IMR. We will note that again here.

Figures

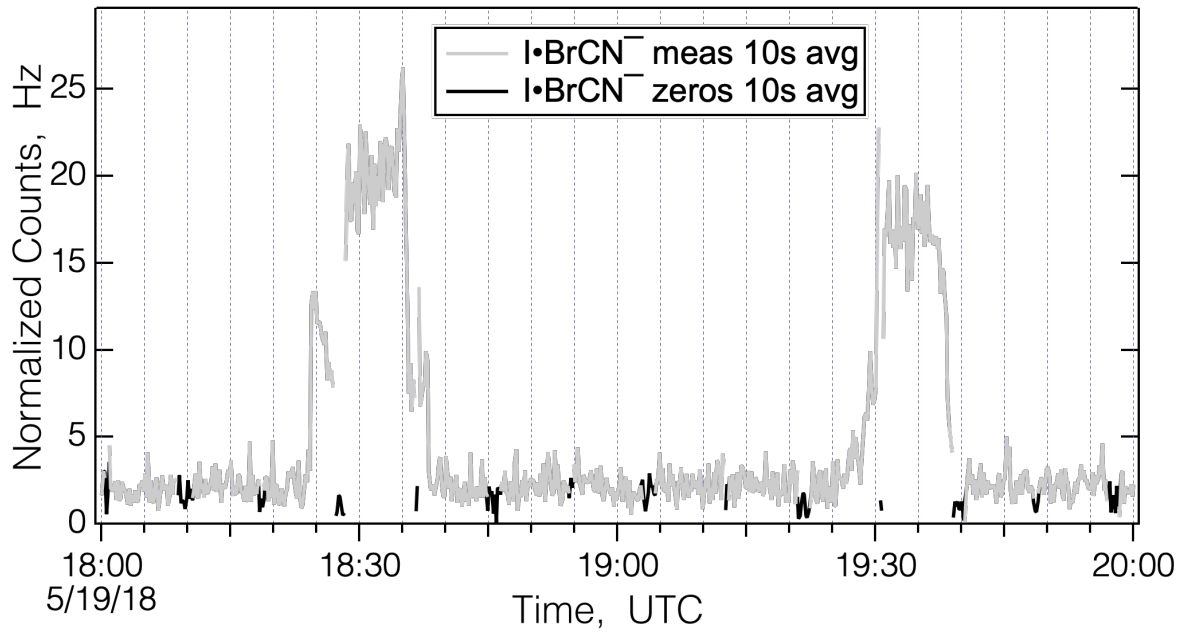


Figure R1. The time series of  $I^{79}\text{BrCN}^-$  signals measured during the 5/19/18 flight over the Arctic. The grey lines are measurements of ambient air and the black lines are measurements during times the inlet was over-flowed with scrubbed air.

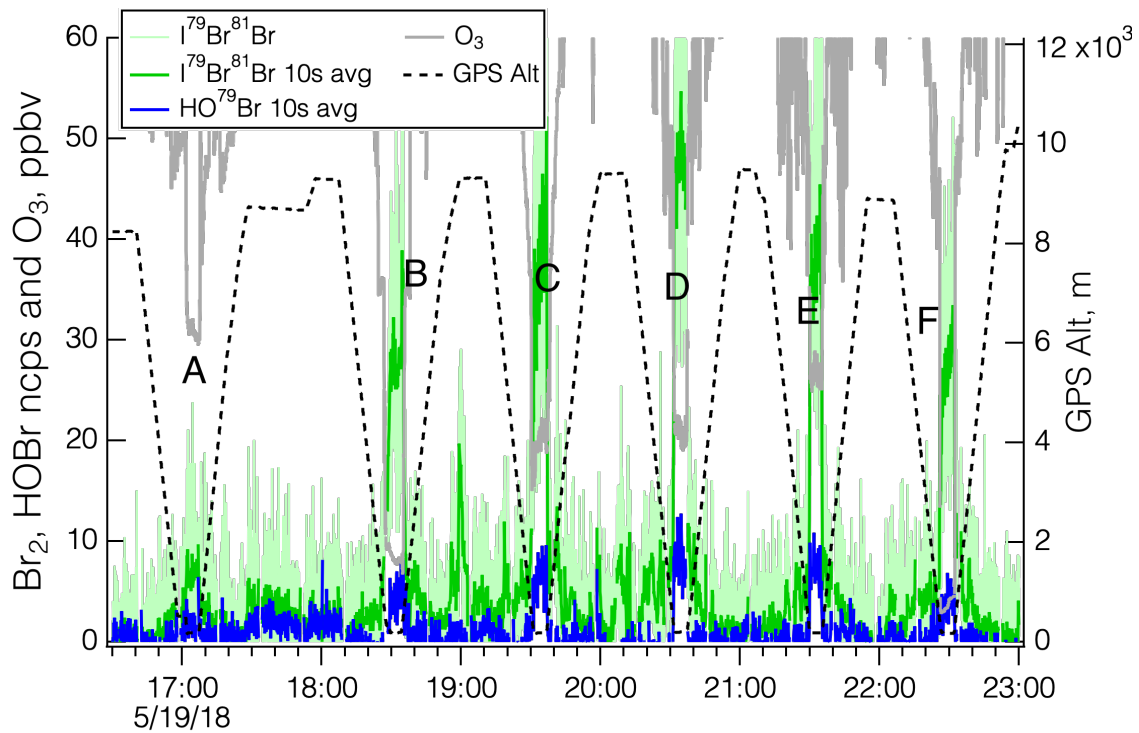


Figure R2. The time series of  $I\text{Br}_2^-$  and  $I\text{HOBr}^-$  signals (in normalized counts per second) along with GPS altitude and  $\text{O}_3$  for the flight on May 19, 2018.

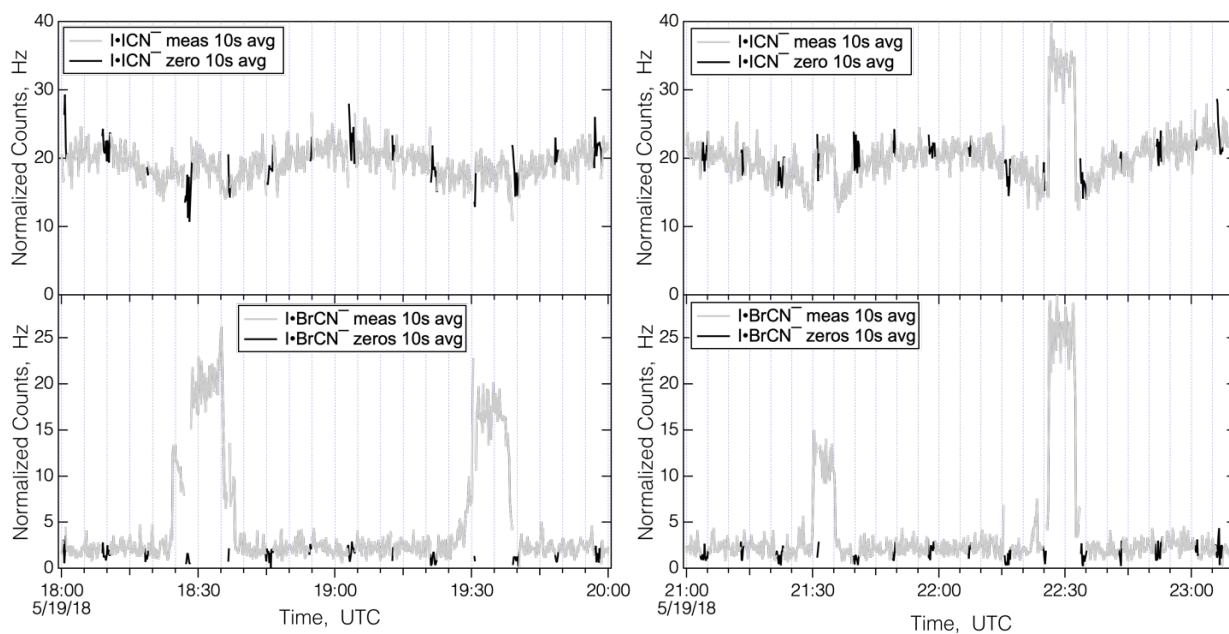


Figure R3. Comparison of the  $I\bullet ICN^-$  signals observed on 5/19/18 with those of  $I\bullet BrCN^-$ . The grey lines are signals from ambient air and the black lines are signals during the periods when the inlet was overflowed with scrubbed air.

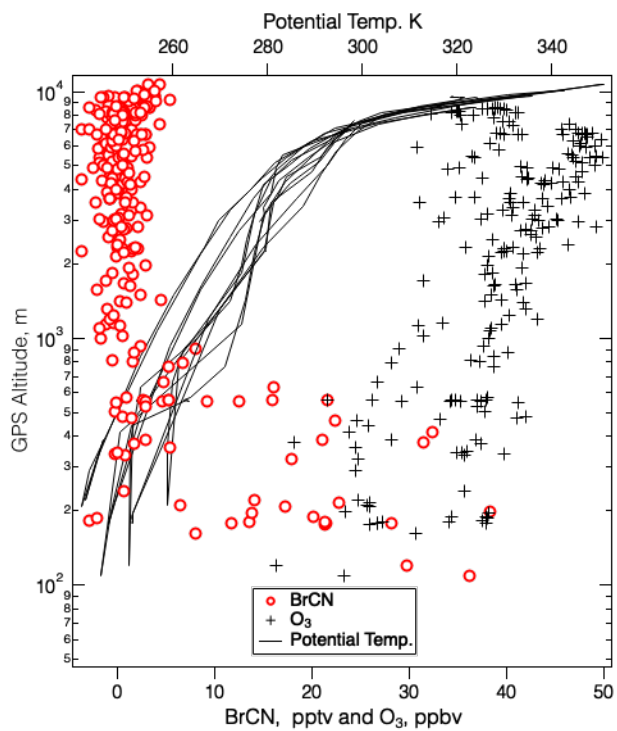


Figure R4. The ATom-2 BrCN and  $O_3$  measurements as a function of altitude for both Arctic flights. The measurements were averaged over the UCI-WAS sample time to be compatible with

the data in Fig. 4 in the main text. The BrCN values are shown as red circles and the O<sub>3</sub> values are shown as black crosses, and the potential temperature is shown as a black line.

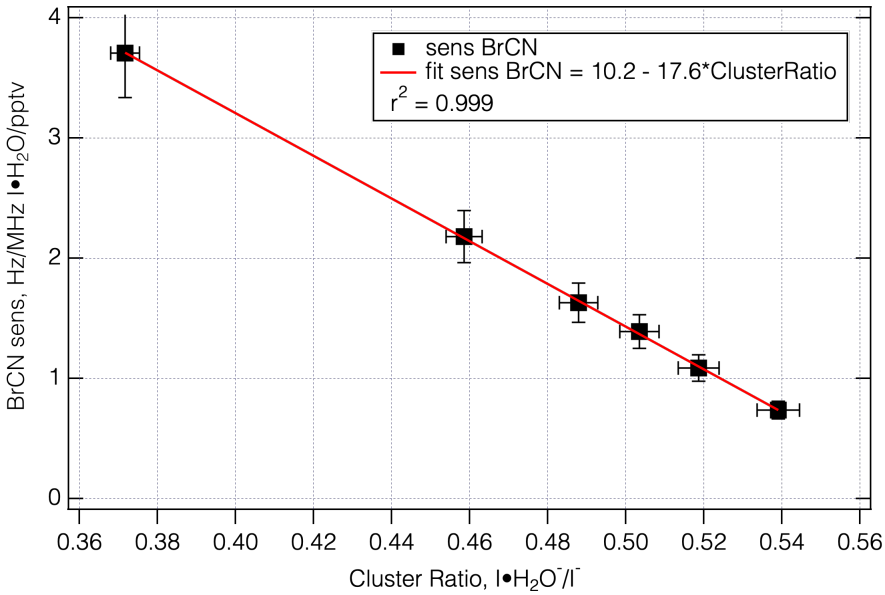


Figure R5. The water vapor dependence of I-CIMS sensitivity to BrCN.

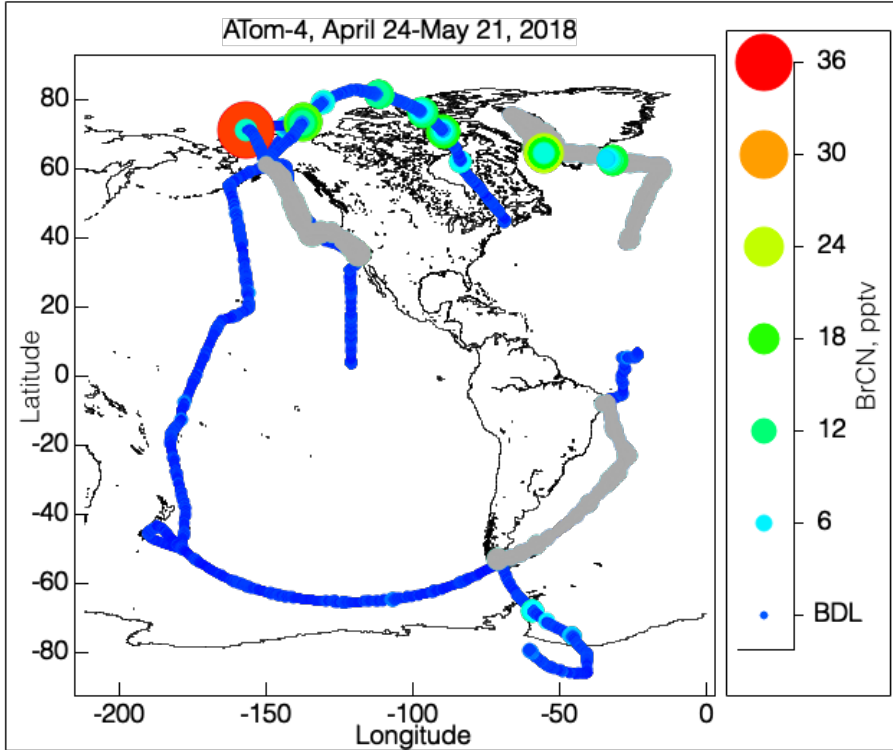


Figure R6. Modified Figure 2b with the observations on May 12, 17 and 21, 2018 that were below detection limit shown in grey.

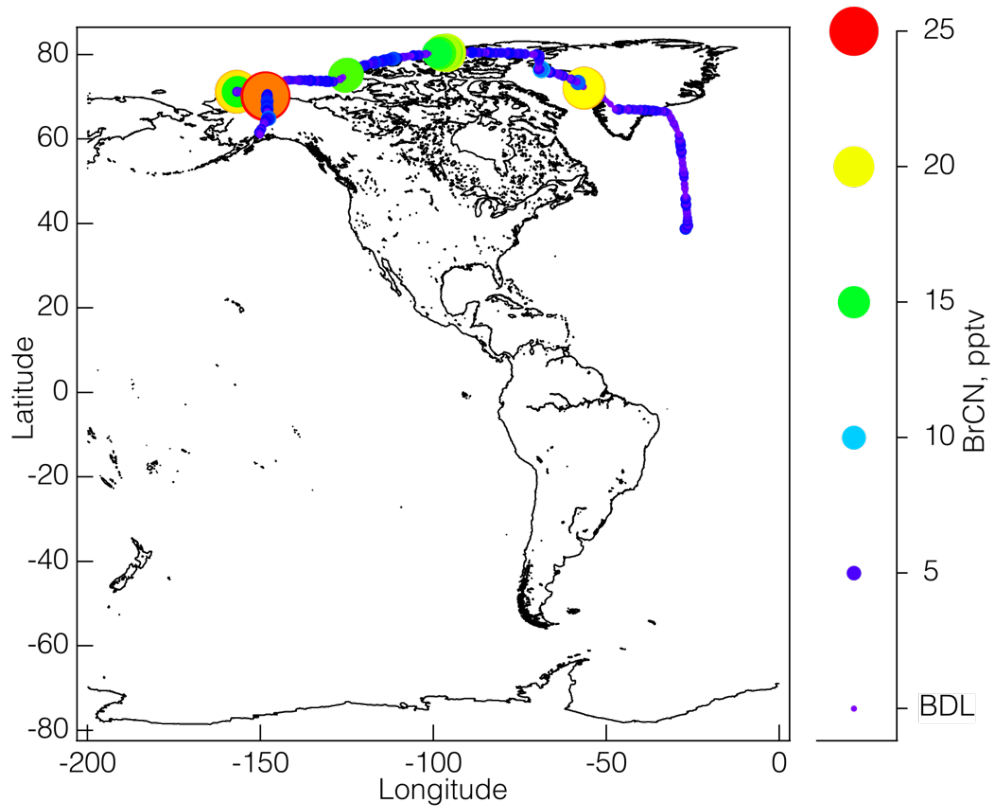


Figure R7. The map of BrCN mixing ratios (60s avg) observed during the Feb 18 and Feb 19, 2017 flights during ATom-2.



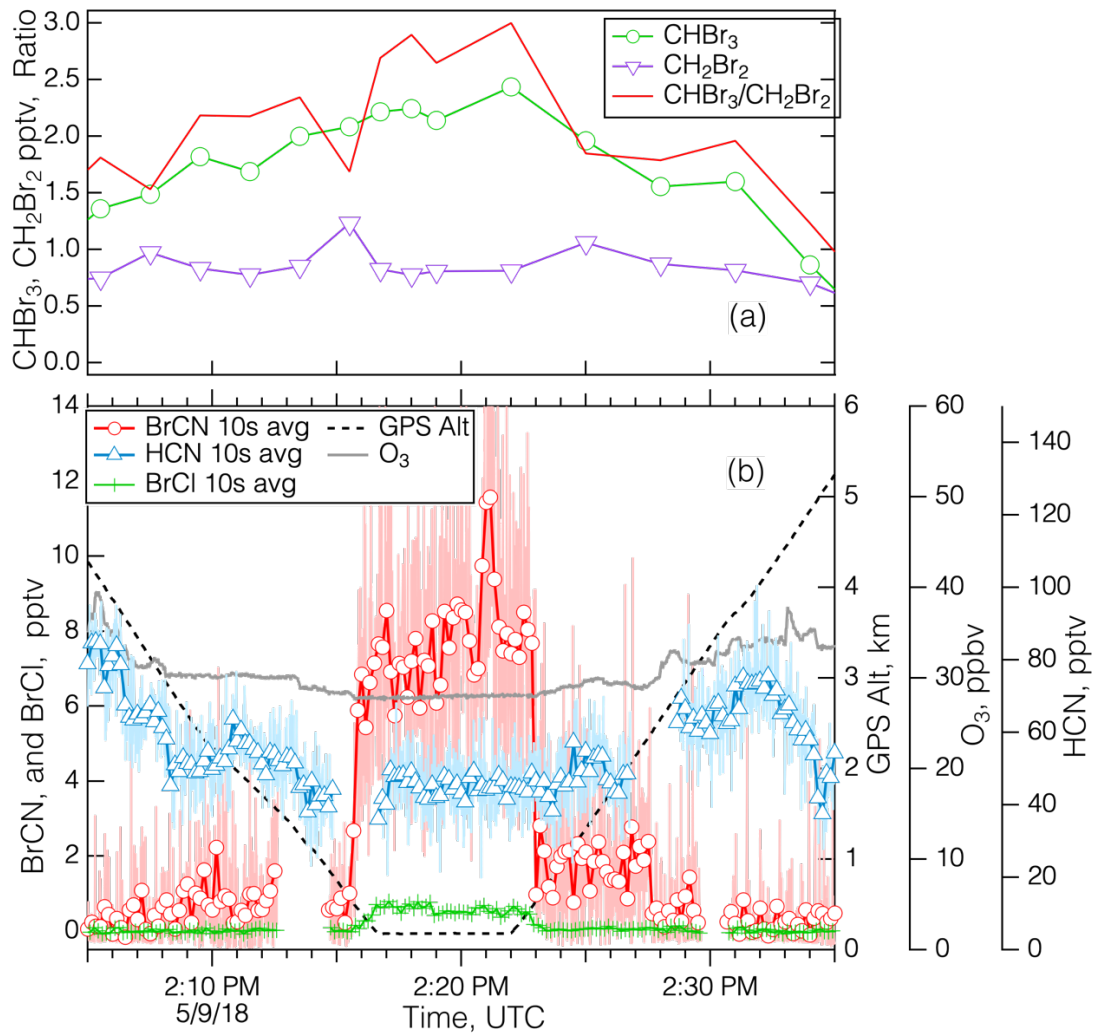


Figure R8. The time series of bromocarbons (Panel a) and  $\text{BrCN}$ ,  $\text{BrCl}$ ,  $\text{HCN}$ ,  $\text{O}_3$  and altitude (Panel b) for the first period in May, 09, 2018 when  $\text{BrCN}$  was encountered.

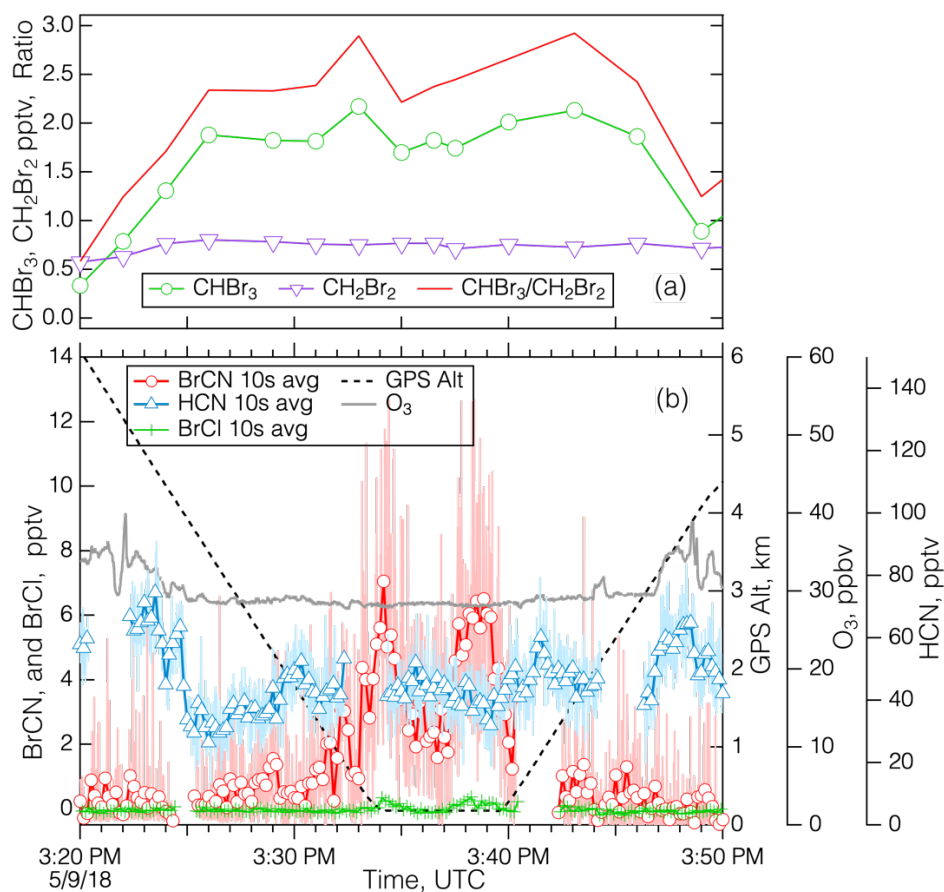


Figure R9. The time series of bromocarbons (Panel a) and BrCN, BrCl, HCN,  $\text{O}_3$  and altitude (Panel b) for the second period in May, 09, 2018 when BrCN was encountered.

#### References:

- Allen, H. M., Crouse, J. D., Kim, M. J., Teng, A. P., E.A., R., McKain, K., Sweeny, C., and Wennberg, P. O.:  $\text{H}_2\text{O}_2$  and  $\text{CH}_3\text{OOH}$  (MHP) in the remote atmosphere: 1. Global distribution and regional influences, *J. Geophys. Res. Atmos.*, 127, e2021JD035701, 10.1029/2021JD035701, 2022.
- Bourgeois, I., Peischl, J., Thompson, C. R., Aikin, K. C., Campos, T., Clark, H., Commane, R., Daube, B., Diskin, G. W., Elkins, J. W., Gao, R. S., Gaudel, A., Hints, E. J., Johnson, B. J., Kivi, R., McKain, K., Moore, F. L., Parrish, D. D., Querel, R., Ray, E., Sánchez, R., Sweeney, C., Tarasick, D. W., Thompson, A. M., Thouret, V., Witte, J. C., Wofsy, S. C., and Ryerson, T. B.: Global-scale distribution of ozone in the remote troposphere from the ATom and HIPPO airborne field missions, *Atmos. Chem. Phys.*, 20, 10611-10635, 10.5194/acp-20-10611-2020, 2020.
- Crouse, J. D., McKinney, K. A., Kwan, A. J., and Wennberg, P. O.: Measurement of gas-phase hydroperoxides by chemical ionization mass spectrometry, *Anal. Chem.*, 78, 6726-6732, 10.1021/ac0604235, 2006.

- Gerritsen, C. M., Gazda, M., and Margerum, D. W.: Non-metal redox kinetics: Hypobromite and hypoiodite reactions with cyanide and the hydrolysis of cyanogen halides, *Inorg. Chem.*, 32, 5739-5748, 10.1021/ic00077a016, 1993.
- Halfacre, J. W., Knepp, T. N., Shepson, P. B., Thompson, C. R., Pratt, K. A., Li, B., Peterson, P. K., Walsh, S. J., Simpson, W. R., Matrai, P. A., Bottenheim, J. W., Natcheva, S., Perovich, D. K., and Richter, A.: Temporal and spatial characteristics of ozone depletion events from measurements in the Arctic, *Atmos. Chem. Phys.*, 14, 4875-4894, 10.5194/acp-14-4875-2014, 2014.
- Krnavek, L., Simpson, W. R., Carlson, D., Domine, F., Douglas, T. A., and Sturm, M.: The chemical composition of surface snow in the Arctic: Examining marine, terrestrial, and atmospheric influences, *Atmos. Environ.*, 50, 349-359, 10.1016/j.atmosenv.2011.11.033, 2012.
- Li, Z., and Gevorgyan, V.: Double duty for cyanogen bromide in a cascade synthesis of cyanoepoxides, *Angew. Chem. Int. Ed. Comm.*, 50, 2808-2810, 10.1002/anie.201006966, 2011.
- Liao, J., Huey, L. G., Tanner, D. J., Flocke, F. M., Orlando, J. J., Neuman, J. A., Nowak, J. B., Weinheimer, A. J., Hall, S. R., Smith, J. N., Fried, A., Staebler, R. M., Wang, Y., Koo, J.-H., Cantrell, C. A., Weibring, P., Walega, J., Knapp, D. J., Shepson, P. B., and Stephens, C. R.: Observations of inorganic bromine (HOBr, BrO, and Br<sub>2</sub>) speciation at Barrow, Alaska, in spring 2009, *J. Geophys. Res. Atmos.*, 117, D00R16, 10.1029/2011JD016641, 2012.
- Liu, X., Deming, B., Pagonis, D., Day, D. A., Palm, B. B., Talukdar, R., Roberts, J. M., Veres, P. R., J., K., Thornton, J. A., de Gouw, J. A., Ziemann, P. J., and Jimenez, J. L.: Effects of gas-wall interactions on measurements of semivolatile compounds and small polar molecules, *Atmos. Meas. Tech.*, 12, 3137-3149, 10.5194/amt-12-3137-2019, 2019.
- Macdonald, M. L., Wadham, J. L., Young, D., Lunder, C. R., Hermansen, O., Lamarche-Gagnon, G., and O'Doherty, S.: Consumption of CH<sub>3</sub>Cl, CH<sub>3</sub>Br, and CH<sub>3</sub>I and emission of CHCl<sub>3</sub>, CHBr<sub>3</sub>, and CH<sub>2</sub>Br<sub>2</sub> from the forefield of a retreating Arctic glacier, *Atmos. Chem. Phys.*, 20, 7243-7258, 10.5194/acp-20-7243-2020, 2020.
- NASA, Worldview: <https://worldview.earthdata.nasa.gov/?v=-70.36366792325461,-81.02284661158711,-12.912908989255904,-41.738455836273374&t=2017-10-14-T21%3A14%3A33Z>, access: September 30,, 2018.
- Nault, B. A., Campuzano-Jost, P., Day, D. A., Guo, H., Jo, D., Handschy, A., Pagonis, D., Schroder, J. C., Schueneman, M., Cubison, M. J., Dibb, J. E., Hodzic, A., Hu, W., Palm, B. B., and Jimenez, J. L.: Interferences with aerosol acidity quantification due to gas-phase ammonia uptake onto acidic sulfate filter samples, *Atmos. Meas. Technol.*, 13, 6193-6213, 10.5194/amt-13-6193-2020, 2020.
- Neuman, J. A., Nowak, J. B., Huey, L. G., Burkholder, J. B., Dibb, J. E., Holloway, J. S., Liao, J., Peischl, J., Roberts, J. M., Ryerson, T. B., Scheuer, E., Stark, H., Stickel, R. E., Tanner, D. J., and Weinheimer, A.: Bromine measurements in ozone depleted air over the Arctic Ocean, *Atmos. Chem. Phys.*, 10, 6503-6514, 10.5194/acp-10-6503-2010, 2010.
- Ridley, B. A., Atlas, E. L., Montzka, D. D., Browell, E. V., Cantrell, C. A., Blake, D. R., Blake, N. J., Cinquini, L., Coffey, M. T., Emmons, L. K., Cohen, R. C., DeYoung, R. J., Dibb, J. E., Eisele, F. L., Flocke, F. M., Fried, A., Grahek, F. E., Grant, W. B., Hair, J. W., Hannigan, J. W., Heikes, B. L., Lefer, B. L., Mauldin, R. L., Moody, J. L., Shetter, R. E., Snow, J. A., Talbot, R. W., Thornton, J. A., Walega, J. G., Weinheimer, A. J., Wert, B. P.,

- and Wimmers, A. J.: Ozone depletion events observed in the high latitude surface layer during the TOPSE aircraft program, *J. Geophys. Res. Atmos.*, 108, 8356, 10.1029/2001JD001507, 2003.
- Roberts, J. M., Osthoff, H. D., Brown, S. S., Ravishankara, A. R., Coffman, D., Quinn, P. K., and Bates, T. S.: Laboratory studies of products of N<sub>2</sub>O<sub>5</sub> Uptake on Cl<sup>-</sup> Containing Substrates, *Geophys. Res. Lett.*, 36, L20808, 10.1029/2009GL040448, 2009.
- Roberts, J. M., and Liu, Y.: Solubility and solution-phase chemistry of isocyanic acid, methyl isocyanate and cyanogen halides, *Atmos. Chem. Phys.*, 19, 4419-4437., 2019.
- Ryerson, T. B., Huey, L. G., Knapp, K., Neuman, J. A., Parrish, D. D., Sueper, D. T., and Fehsenfeld, F. C.: Design and initial characterization of an inlet for gas-phase NOy measurements from aircraft, *J. Geophys. Res.*, 104, 5483-5492, 10.1029/1998JD100087, 1999.
- Sander, R.: Compilation of Henry's law constants (version 4.0) for water as solvent, *Atmos. Chem. Phys.*, 15, 4399-4981, 10.5194/acp-15-4399-2015, 2015.
- Serreze, M. C., Maslanik, J. A., Rehder, M. C., Schnell, R. C., Kahl, J. D., and Andreas, E. L.: Theoretical heights of buoyant convection above open leads in the winter Arctic pack ice cover, *J. Geophys. Res. Oceans*, 97, 9411-9422, 10.1029/92JC00688, 1992.
- Veres, P. R., Neuman, J. A., Bertram, T. H., Assaf, E., Wolfe, G. M., Williamson, C. J., Weinzierl, B., Tilmes, S., Thompson, C. R., Thames, A. B., Schroder, J. C., Saiz-Lopez, A., Rollins, A. W., Roberts, J. M., Price, D., Peischl, J., Nault, B. A., Møller, K. H., Miller, D. O., Meinardi, S., Li, Q., Lamarque, J.-F., Kupc, A., Kjaergaard, H. G., Kinnison, D., Jimenez, J. L., Jernigan, C. M., Hornbrook, R. S., Hills, A., Dollner, M., Day, D. A., Cuevas, C. A., Campuzano-Jost, P., Burkholder, J. B., Bui, T. P., Brune, W. H., Brown, S. S., Brock, C. A., Bourgeois, I., Blake, D. R., Apel, E. C., and Ryerson, T. B.: Global airborne sampling reveals a previously unobserved dimethyl sulfide oxidation mechanism in the marine atmosphere, *Proc. Natl Acad. Sci.*, 117, 4505-4510, 10.1073/pnas.1919344117, 2020.
- von Braun, J., and Schawarz, R.: Die einwirkung von bromcyan auf tertiäre amine, *Chem. Ber.*, 35, 1279-1285, 10.1002/cber.19020350214, 1902.
- Yuan, B., Koss, A., Warneke, C., Gilman, J. B., Lerner, B. M., Stark, H., and de Gouw, J. A.: A high-resolution time-of-flight chemical ionization mass spectrometer utilizing hydronium ions (H<sub>3</sub>O<sup>+</sup> ToF-CIMS) for measurements of volatile organic compounds in the atmosphere, *Atmos. Meas. Tech.*, 9, 2735-2752, 10.5194/amt-9-2735-2016, 2016.
- Yuan, B., Koss, A. R., Warneke, C., Coggon, M. M., Sekimoto, K., and de Gouw, J. A.: Proton-transfer-reaction mass spectrometry: Applications in atmospheric sciences, *Chem. Rev.*, 117, 13187-13229, 10.1021/acs.chemrev.7b00325, 2017.