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Influence of acidity on liquid–liquid phase transitions of mixed SOA proxy–inorganic aerosol droplets

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Abstract. Phase state and morphology of aerosol particles play a critical role in determining their effect on climate. While aerosol acidity has been identified as a key factor affecting the multiphase chemistry and phase transitions, the impact of acidity on phase transition of multicomponent aerosol particles has not been extensively studied in situ. In this work, we employ an aerosol optical tweezer (AOT) to probe the impact of acidity on the phase transition behavior of levitated aerosol particles. Our results reveal that higher acidity decreases the separation relative humidity (SRH) of aerosol droplets mixed with ammonium sulfate (AS) and secondary organic aerosol (SOA) proxy, such as 3-methylglutaric acid (3-MGA), 1,2,6-hexanetriol (HEXT) and 2,5-hexanediol (HEXD) across aerosol pH in atmospheric condition. Phase separation of organic acids was more sensitive to acidity compared to organic alcohols. We found the mixing relative humidity (MRH) was consistently higher than the SRH in several systems. Phase-separating systems, including 3-MGA/AS, HEXT/AS, and HEXD/AS, exhibited oxygen-to-carbon ratios (O:C) of 0.67, 0.50, and 0.33, respectively. In contrast, liquid-liquid phase separation (LLPS) did not occur in the high O:C system of glycerol/AS, which had an O:C of 1.00. Additionally, the morphology of 38 out of the 40 aerosol particles that underwent LLPS was observed to be a core-shell. Our findings provide a comprehensive understanding of the pH-dependent LLPS in individual suspended aerosol droplets and pave the way for future research on phase separation of atmospheric aerosol particles.

1 Introduction

Atmospheric aerosol particles can directly and indirectly impact climate by absorbing and scattering light and acting as cloud condensation nuclei (Rosenfeld et al., 2014). Particle morphology is a critical factor influencing the physiochemical properties

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of aerosols such as their optical properties, chemistry, and nucleation processes (Freedman et al., 2009; Corral Arroyo et al., 2022; Cosman et al., 2008; Lam et al., 2021; Petters and Kreidenweis, 2007; Mikhailov et al., 2021). Morphology can be broadly categorized into single-phase homogeneous morphology and phase separation morphology (Gorkowski et al., 2020), based on the phase state of the particle. For droplets with a phase separation morphology, the two main equilibrium morphologies are a fully engulfed (core-shell) structure and a partially engulfed structure (Freedman, 2020). Droplets can undergo phase transition processes and thus the morphology would be changed. The composition and mass of inorganic and organic components impact the phase transition characteristics of a particle. With a decrease of particle water content, a transition occurs from single homogenous liquid phase to two separated liquid phases, which is known as liquid-liquid phase separation (LLPS). The relative humidity (RH) when the LLPS occurs is defined as separation relative humidity (SRH). The reverse process, in which two liquid phases mix into a single homogenous liquid phase, is referred to as liquid-liquid phase mixing and the corresponding RH is the mixing RH (MRH; Gorkowski et al., 2017).

The phenomenon of LLPS has garnered considerable attention from the atmospheric research community due to its potential role in affecting the physiochemical properties of atmospheric aerosols. Song et al. (2012) using optical microscopy studied the relationship between LLPS and the oxygen-to-carbon ratio (O:C) and discovered that LLPS was consistently observed when O:C < 0.56, while it was never observed when O:C > 0.80. For O:C between 0.56 and 0.80, the occurrence of LLPS was influenced by the types of organic functional groups. Gorkowski et al. (2020) utilized experimental results of previous studies on LLPS and morphology, observing a general trend in morphology from partially engulfed to core shell and finally homogeneous as oxidation increased. More recently, Kucinski et al. (2021) found that submicrometer-sized aerosol particles had a lower SRH compared to micrometer-sized droplets. Meanwhile, Stewart et al. (2015) employed aerosol optical tweezer (AOT) to investigate the morphologies of aqueous droplets. They found in the polyethylene glycol (PEG)/ammonium sulfate (AS) system, droplets formed predominately core-shell particles when the AS content was high and partially engulfed when the PEG content was high.

One factor that could influence the phase transitions of aerosol particles is the aerosol pH. The pH values for misty cloud and fog droplets generally range between 2 and 7, whereas continental and marine aerosol particles exhibit a wider range of pH values, from -1 to 5 and 0 to 8, respectively (Pye et al., 2020; Angle et al., 2021; Weber et al., 2016; Tilgner et al., 2021; Zheng et al., 2020). Meanwhile, aerosol pH is size-dependent, with the fine mode showing lower 1–4 pH units than the coarse mode (Fang et al., 2017; Young et al., 2013; Guo et al., 2017). Losey et al. (2018) measured the RH of phase transitions using optical microscopy and discovered that for low-pH aerosol particles (≤0.35), phase separation may be hindered by the addition of sulfuric acid. However, it should be noted that the study utilized substrate-deposited droplets, which means the effect of the contact of the coverslip on droplet morphology cannot be disregarded. More recently, Tong et al. (2022) investigated the effect of acidity on phase separation in single suspended microdroplets using AOT. Their results showed that the pH can affect the miscibility of the mixture and high acidity results in a reduced SRH of 1,2,6-hexanetriol. Nevertheless, parallel experiments in this study were not conducted to accurately determine the uncertainty of the measurements.

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Our aim with this work is to gain a comprehensive understanding of the influence of pH on phase transitions in suspended droplets. To that end, we investigate pH-dependent SRH and MRH, as well as morphologies of aqueous droplets using AOT, meanwhile discussed the effect of O:C on phase separation behavior. Compared to substrate-based measurement techniques, AOT can suspend droplets without any substrate contact, providing a more realistic simulation of the behavior of aerosols in the atmosphere. (Wang et al., 2021; Cui et al., 2021; Redding et al., 2015; Gong et al., 2018; Rafferty et al., 2023). We measured droplets containing AS and a range of organic compounds with varying O:C. We discuss how our findings provide insight into the mechanisms behind pH-dependent phase transitions in levitated droplets, along with the implications for fields such as climate science. Overall, our study highlights the importance of considering pH as a key factor in the phase transition behavior of micron-sized droplets and underscores the need for further research to fully understand the complex interactions between pH and phase transitions in these atmospherically relevant systems.

2 Methods

2.1 Aerosol generation

Four organics components: glycerol (GL), 3-methylglutaric acid (3-MGA), 1,2,6-hexanetriol (HEXT), and 2,5-hexanediol (HEXD), were chosen because they are commonly-used secondary organic aerosol (SOA) proxies (Lam et al., 2021; Gorkowski et al., 2020). O:C of the selected chemicals varied from 1 to 0.33 (**Table 1**), which is similar to the real atmospheric SOA (Canagaratna et al., 2015; Mahrt et al., 2021). AS was chosen as the inorganic salt component due to its widespread occurrence in the atmospheric environment. All concentrations of organics and AS in the mother solutions were 50 g/L. The pure organic and inorganic components were dissolved in ultrapure water (Millipore, resistivity of 18.2MΩ) to create solutions with OIR of 1:1. The pH of each solution was measured using a pH meter (Mettler Toledo Instruments Co., Ltd., Shanghai, China), and adjusted as necessary using either concentrated sulfuric acid (SA) or sodium hydroxide (NaOH) solution (5.29 mol/L). Four to six solutions were prepared for each system with pH values ranging from 0.48 to 6.53. The purity and supplier of the compounds used in this study were summarized in **Table S1**.

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Table 1. Information of the solutions used to generate aerosol droplets.

Solution ID	Organic component	O:C ratio	pH
GL	glycerol	1.00	5.24±0.01
3-MGA-I	3-methylglutaric acid	0.67	0.48±0.01
3-MGA-II			1.19±0.01
3-MGA-III			2.70±0.01
3-MGA-IV			3.70±0.01
3-MGA-V			5.21±0.02
3-MGA-VI			6.53±0.02
HEXT-I	1,2,6-hexanetriol	0.50	0.92±0.01
HEXT-II			2.02±0.01
HEXT-III			3.14±0.01
HEXT-IV			5.11±0.02
HEXD-I	2,5-hexanediol	0.33	1.39±0.01
HEXD-II			2.03±0.01
HEXD-III			2.71±0.01
HEXD-IV			3.13±0.01
HEXD-V			5.01±0.01

2.2 Experimental setup

A schematic illustration of the experimental setup is presented in **Fig. S1**. The aerosol optical tweezer system consists of a custom-made levitation chamber that integrates the optical trapping system, the illumination and imaging system, and the aerosol generation system. A 532 nm (Opus 532-2W) laser was used to create an optical trap with a 100x oil immersion objective (Olympus, UPLFLN100XO, NA 1.30) pressed against a glass coverslip (Nest, thickness 160-190 μm). The illumination and imaging system includes a 450 nm LED (Daheng Optics, GCI060404) and a camera (Thorlabs, CS165CU/M) to illuminate and image the particle. Two low pass filters (Andover, 500FL07-25) were used in front of the camera lens to remove the influence of back scattered light of the 532 nm laser to photograph clear image of the particle. The Raman scattered light passed through two 50:50 beam splitters (CVI Laser Optics, BTF-VIS-50-2501M-C) and a notch filter (Semrock, NFD01-532-25x36) and was focused into the Raman spectrograph. A spectrograph (ZOLIX, Omni-λ5004i) is used to measure the Stokes shifted Raman spectrum. A 20 μm entrance slit width and 1200 groove/mm diffraction grating with blaze wavelength of 500 nm were used to achieve a spectral resolution of 0.021 nm. The wavelength position of spectrograph was calibrated with Hg-laser. The Raman scattered light was recorded every 4 second with range of 624.24-665.40 nm.

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As droplets are introduced continuously into the chamber from a medical nebulizer (LANDWIND, PN100), smaller droplets undergo a process of collision and coalescence, leading to the formation of larger droplets that can be readily trapped near the focal point of the laser. In most cases, droplets can be successfully captured within 30 s after the introduction of an aerosol plume into the cell. Air with relative humidity (RH) of 100% and 0% were mixed to produce wet air with a desired RH. The flow rates of the humidified and dry air streams were regulated by mass flow controllers (MFCs, Tianjin Gastool Instruments Co., Ltd., Tianjin, China, GT130D), with a combined flow rate of 0.3 L/min in total. Two humidity sensors (Sensirion, SHT85) were utilized, with a precision of ±1.5%. Since the sensor located behind the chamber was positioned in close proximity (~80 mm) to the droplet, its observed values were used as a surrogate for measuring the RH inside the chamber. The RH values were reduced in increments of 5% every 30 minutes (Tong et al., 2022; Stewart et al., 2015) until droplet phase separation occurred. The measured values of RH given by the sensors were used as the phase separation RH. Subsequently, the RH level was set to 100%, to investigate the phase mixing of the droplets. The entire experiment was repeated 1~4 times for each system.

2.3 Determination of phase transitions

When a transparent or weakly absorbing spherical particle is trapped, it can behave as a high-quality factor optical cavity that supports sharp optical resonances, resulting in cavity-enhanced Raman scattering. These resonances can be observed as peaks in the Raman spectrum of a particle and are often referred to as whispering gallery modes (WGMs). In principle, particle morphology can be deduced from the WGMs, as inhomogeneities in the refractive index can disrupt the circulation of the WGMs (Lin et al., 1992; Mitchem et al., 2006). Raman spectra measurements of single droplets in various morphological states are presented in **Figure 1**. When the droplet was in a homogeneous phase morphology, the droplet acted as a high-quality microcavity and sharp WGM peaks overlapped with the spontaneous Raman spectrum (**Fig. 1a**). When the droplet was in a state of a core-shell structure, observed WGMs were clearly diminished in measured spectra (**Fig. 1b**). The origin of the damping of the WGMs is the radial homogeneity that is present when the particle is separated into a hydrophilic core and a hydrophobic shell. As a result, when fitting the Raman spectra with the Mie scattering model for homogeneous droplets, the error in the best-fits greatly increase. Examination of the retrieved radius and refractive index reveals a clear break with fits for that of a homogeneous sphere. Therefore, the point at which a significant break in particle size and refractive index occurred can be used as the point at which core-shell phase separation occurs. As illustrated in **Fig. 1c**, when the droplet was partially-engulfed and non-spherical, WGM peaks in the spectrum are absent (Reid et al., 2011). Overall, the results of this analysis demonstrate the dynamic changes in the Raman spectra of single droplets as they undergo morphological transitions (Sullivan et al., 2020; Stewart et al., 2015; Tong et al., 2022).

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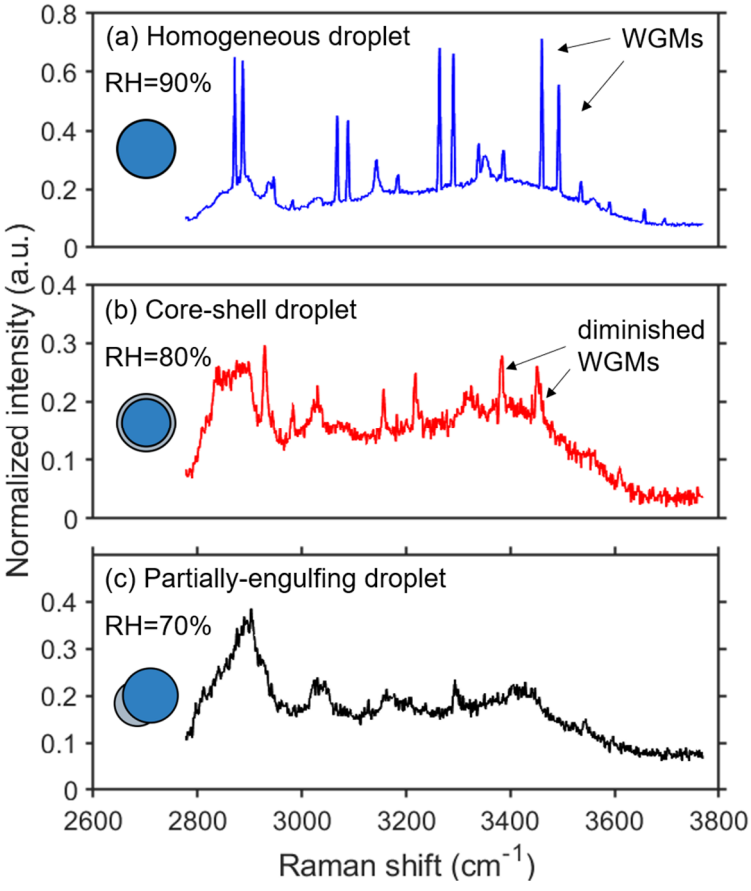

Figure 1. Raman spectra of 3-MGA-II microdroplets: (a) a homogenous droplet (RH = 90%); (b) a core-shell droplet (RH = 80%); (c) a partially-engulfed droplet (RH=70%). The WGMs are marked by black arrows.

The peak finding method used in this study is based on the ipeak code developed by O'Haver (2022). In short, the code first smooths the first derivative of the signal and identified downward-going zero-crossings that met a certain predetermined minimum slope and amplitude threshold. By adjusting the corresponding parameters, it is possible to accurately detect the desired peaks. The algorithm used to fit WGM peaks in spectra from homogenous spheres in this study was proposed by Preston and Reid (2013) and Preston and Reid (2015). The algorithm compares observed peak positions to expected positions calculated using a resonance condition from Mie theory. Error is minimized by varying particle size and refractive index (i.e. the parameters of best-fit). The method has been demonstrated to provide a rapid determination of the fitted radius and refractive index with an accuracy of ±2 nm and ±0.0005, respectively. All of the Raman spectra used in this study were normalized by the area below the spontaneous Raman signals.

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3 Results and discussion

3.1 Phase behaviors of droplets mixed SOA proxy with AS

Figure 2 presents the results of time-resolved Raman spectra of aerosol droplets produced from a 3-MGA-II solution under continuously varying RH, as well as the corresponding particle size and refractive index values. To enable temperature and RH to stabilize, the chamber was conditioned with airflow for 50 minutes after trapping a particle. During the dehumidification process, the particle diameter decreased from 11.85 μm to 9.03 μm and the refractive index increased from 1.379 to 1.475 when RH decreased from 93.0% to 70.0%. The particle size and water content decreased with RH due to the equilibrium partitioning of water molecules between vapor and droplets. Meanwhile, the refractive index of the droplets gradually increased as the water content decreases. When LLPS occurred, the droplets changed from a symmetrical homogeneous phase to an asymmetrical partially engulfed structure which led to the disappearance of the WGMs, or the formation of a core-shell structure. As RH in the reaction chamber was reduced, the LLPS was initiated, marked by the variations of the WGM signal (See **Fig. 1b**). This was achieved by reducing setting RH (setting values) by 5% at 30-minute intervals until the organic phase separated from the water-rich phase and then continuing decreasing RH by 10%-15%. **Fig. 2a** illustrates how the fitting of the droplet diameter and the refractive index deteriorated as the shell develops, indicating phase separation. The refractive index's shift results from a significant change in the radial profile due to the formation of a core-shell structure. Additionally, the persistence of strong WGMs indicates that the morphology of the droplet remains spherical following LLPS and is core−shell. During the RH increased from 70% to 95%, the reappearance of the continuously shifting WGM signal is observed, suggesting that the inorganic phase has mixed with the organic phase, and droplet returned to a homogeneous phase. During the humidification process, there is an opposite trend observed in the particle size and refractive index of the droplet compared to the dehumidification process. In conclusion, the variations of the WGM signal can serve as a reliable indicator of the occurrence of liquid-liquid phase separation or mixing, and the RH at these points can be considered as the SRH or MRH, respectively. The observed phase transitions of droplets produced from HEXT-IV and HEXD-V solutions were shown in **Fig. S2** and **Fig. S3**, respectively.

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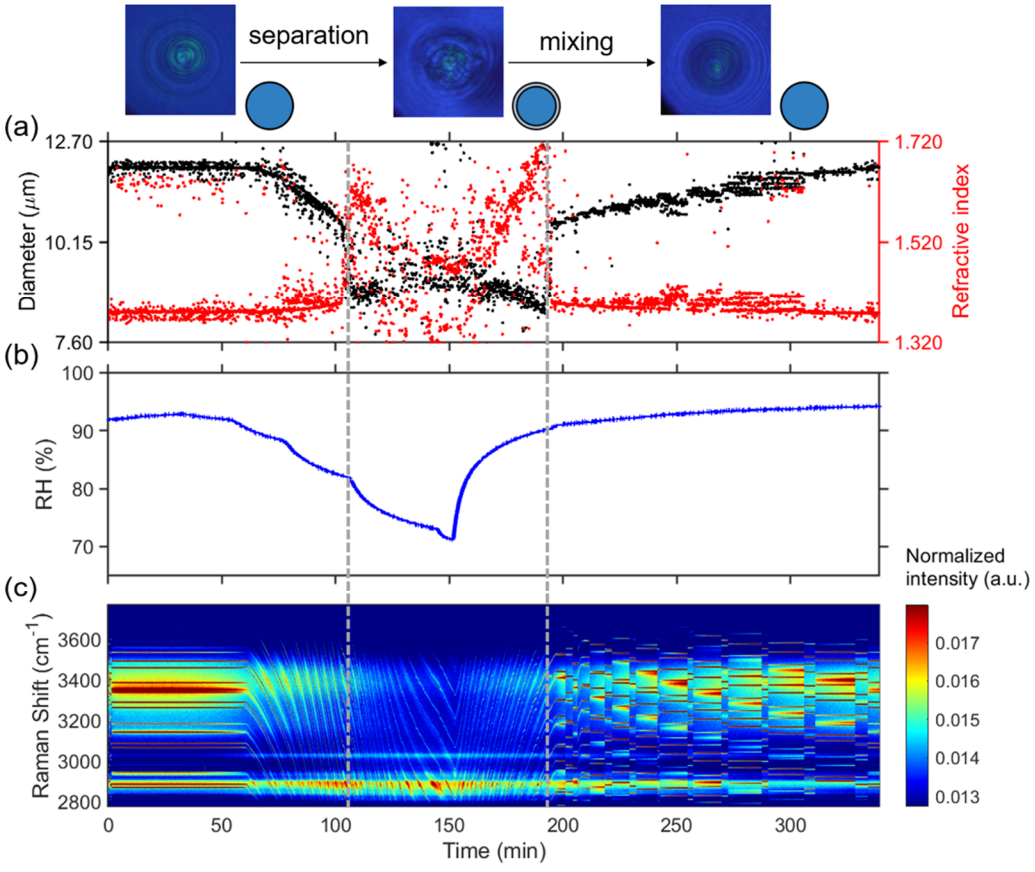

Figure 2. Liquid-liquid phase separation and mixing of aqueous 3-MGA-II. Schematic diagram of phase states is on the top of the figure. (a) Timescale of changes in droplet size and refractive index, determined from fitting the Raman shift positions of the WGMs. (b) RH variation after the trapping chamber during the humidity changing process. (c) Time-resolved Raman spectra. The cessation of the random motion of inclusions within the droplet and the resultant formation of a core-shell structure are indicated by the grey dashed line on the left. The grey dashed line on the right serves as an indication of the point at which the droplet morphology transited from a state of separated phases to a homogeneous phase. The Raman spectra at 53 min, 113 min, 130 min are shown in **Fig. 1(a), (b), (c)**, respectively. Fitting errors of the WGMs was presented in **Fig. S5**.

Figure S2 presents the results of time-resolved Raman spectra of aerosol droplets produced from GL/AS solution under continuously varying RH, as well as the corresponding particle diameter and refractive index values. At the start of the experiment, the chamber RH was held at 93% for approximately 75 minutes. The spectrum during this period reveals a clear bright trend, indicative of the presence of many WGMs in the newly captured droplets. As the chamber RH dropped to a minimum value of 71.5% at around 200 minutes, the position of the WGMs in each spectral snapshot shifted continuously, following the same trend as the chamber RH. This observation suggests that the droplet was homogeneous and that no phase

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separation occurred in the experimental RH range. The phenomenon regarding the GL/AS system is consistent with the conclusion by Song et al. (2013) and Gorkowski et al. (2020).

3.2 Effect of pH on SRH and MRH of different systems

The SRH and MRH of aerosol droplets produced from 3-MGA-I~VI solution are shown in **Fig. 3a**. The SRH values were 92.7%, 89.5%, 80.6%, 79.7%, 76.2% and 69.7% at pH of 6.53, 5.21, 3.70, 2.70, 1.19 and 0.48, respectively. This decrease in SRH was attributed to the salting out ability of ammonium sulfate that is weakened at lower pH, and thus hinders the ability of organic matter to precipitate out of the solution (Losey et al., 2018). The MRH values at pH 6.53, 5.21, 2.70, 1.19 and 0.48 were 87.6%, 89.5%, 87.3%, 83.9% and 83.5%, respectively, and are generally higher than corresponding SRH, especially in the low pH range (<5.00). The hysteresis between SRH and MRH existed because the SRH process has an activation barrier while the MRH process does not, and lower RH is needed for the aerosol droplet to overcome the activation barrier to form two phases (Freedman, 2020). Similar results were also observed in HEXT/AS and HEXD/AS systems. Additionally, the pH-dependent SRHs obtained in this study were compared to those reported by Losey et al. (2018), as depicted in **Fig. 3a**. It is worth mentioned that the solute concentration used in our study (50g/L) is comparable to Losey et al. (2018) (5.0 wt%), allowing for meaningful comparison of results. Overall, the SRHs of 3-MGA obtained in this study was higher than the results of Losey et al. (2018). When the pH was lower than 3.70, in 3-MGA system, the present study followed a similar trend as the results of Losey et al. (2018), with the SRH decreasing as the pH decreased. However, when the pH was greater than 3.70, our study showed an opposite trend compared to the results of Losey et al. (2018). The observed discrepancy may be attributed to the different methodologies employed. The present study utilized an optical tweezer system to capture droplets of ~10 μm diameter and the results were obtained using a spectrometer, while Losey et al. (2018) allowed ~100 μm size droplets to settle on a hydrophobic surface and observed the samples using an optical microscope. Kucinski et al. (2019) reported that the particle size has a significant impact on the phase separation of micro-droplets. The spherical morphology and smaller size in this study enhances the Kelvin effect, resulting in higher possibilities of phase separation occurring within droplets. The difference in droplet size is significant and the droplets used in this study are more representative of real atmospheric particulate matter, whereas the droplet properties in Losey et al. (2018) may be influenced by the surface they came into contact with. Thus, the utilization of an optical tweezer system provides a more realistic simulation of the evolution of particulate matter in the atmosphere.

In addition to 3-MGA, we also studied two organic/AS systems to investigate how acidity affects SRH and MRH of aerosols of differing composition. These results are shown in **Fig. 3** and tabulated in **Table 2**. The data suggests that acidity did not have a noticeable effect on the MRH of the various systems. The pH of the HEXT/AS solution without the addition of any acid was 5.11, and sulfuric acid was utilized to adjust the pH to lower levels (3.14, 2.02 and 0.92). The SRH values of HEXT/AS system (O:C=0.50) decreased as the pH decreased, with values of 78.3%, 76.6%, 76.4% and 75.7% at pH values of 5.11, 3.14, 2.02 and 0.92, respectively. The trend is similar to the 3-MGA (O:C=0.67) system, and the reason why SRH

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decreased may be due to the acid enhancing the miscibility of organic alcohols and inorganic substances, resulting in a greater difficulty in separating the hydrophobic phase from the water-rich phase (Tong et al., 2022). Still, we observed SRH was not strongly dependent on pH for HEXT/AS, compared to 3-MGA/AS system. This is likely due to the fact that organic alcohols have a large pK_a (e.g. the pK_a of HEXT is 14.3) and therefore exhibit minimal ionization in the pH range studied here (Wade and Simek, 2020). Additionally, the relative molecular interactions between alcohols and water are weaker than those of acids, leading to a weaker dependence of salting out ability of AS in the HEXT/AS system. The results of Losey et al. (2018) and Tong et al. (2022) were also depicted in **Fig. 3b**. Our results were higher than those of Losey et al. (2018), but the trend was similar. We attribute this discrepancy to a similar reason as that of the 3-MGA/AS system, which was previously discussed in this article. In contrast to the findings of Tong et al. (2022), our study observed a less pronounced trend in the values of SRH, and a narrower range in the distribution of SRH compared to literature values. The difference in OIR between this study (1:1) and Tong et al. (2022) (2:1) may account for the discrepancy in SRH. Previous studies (Ma et al., 2021; Stewart et al., 2015; Song et al., 2012) indicated that OIR differences could affect SRH, but SRH was not significantly dependent on OIR. For HEXD/AS (O:C=0.33) system, SRH decreases significantly when the pH is less than 2.00, while acidity had no significant effect on SRH when pH is greater than 2.00, with values of 79.4%, 81.0%, 77.7%, 82.8% and 70.9% at pH values of 5.01, 3.13, 2.71, 2.03 and 1.39, respectively. This phenomenon was attributed to a mechanism similar to that observed in HEXT/AS. To our knowledge, this is the first investigation on the pH-dependent phase transition of HEXD/AS at the single particle level in a contact-free environment.

Our research suggests that in real atmospheric conditions where many ambient aerosol particles exhibit high acidity (Angle et al., 2021; Song et al., 2018; Liu et al., 2019; Li et al., 2022), droplets encounter heightened impediments to phase separation and tend to display a homogeneous structure. Although we used bulk solution pH as an indicator of pH at droplet phase transition, previous studies (Craig et al., 2018; Coddens et al., 2019; Li et al., 2023) have shown that the pH deviation between single microdroplet measurements and bulk solution measurements is small. Therefore, we consider our results valid. However, this study focused on volatile organics and was conducted over a relatively long period, which may have affected our results. Nevertheless, since the droplets studied here were relatively large (diameter 6-14 μm), we believe that any such influence can be neglected.

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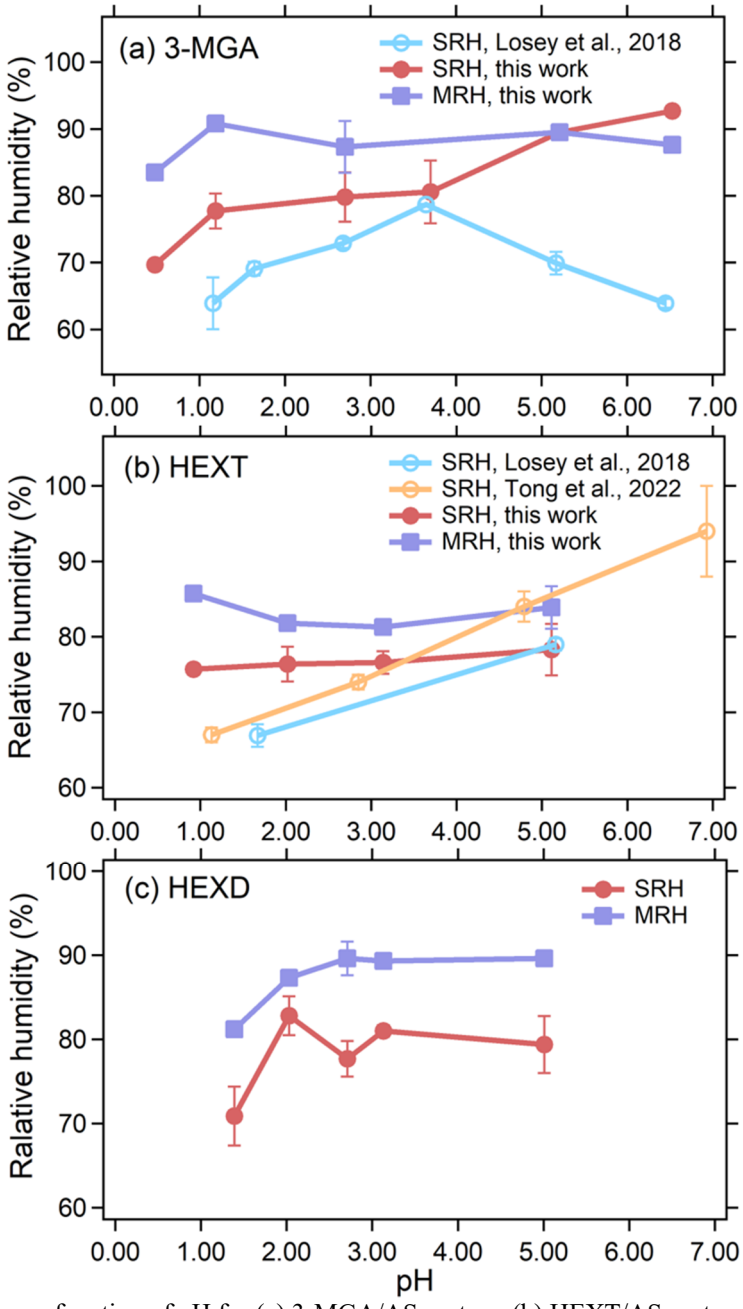

Figure 3. SRHs and MRHs as a function of pH for (a) 3-MGA/AS system, (b) HEXT/AS system, (c) HEXD/AS system. Hollow circles represent data from Losey et al. (2018) and Tong et al. (2022). The error bars of SRHs and MRHs are derived from multiple measurements.

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Table 2. SRH information for each pH studied as well as initial diameter, separation diameter (SD), separation relative index (SRI), MRH, mixing diameter (MD), and mixing relative index (MRI) data.

Initial pH	Initial Diameter(nm)	SRH (%)	SD (nm)	SRI (λ=650nm)	MRH (%)	MD (nm)	MRI (λ=650nm)
3-MGA/AS system (O:C=0.67)							
0.48	10.97±1.57	69.7±0.2	7.23±1.72	1.515±0.086	83.5	6.82	1.540
1.19	11.23±1.20	77.7±2.6	8.68±2.38	1.454±0.100	90.8±0.2	9.08±1.64	1.394±0.009
2.70	10.60±2.28	79.8±3.7	7.83±1.71	1.480±0.112	87.3±3.9	7.86±1.63	1.483±0.118
3.70	10.87±1.87	80.6±4.7	7.24±1.00	1.491±0.088			
5.21	11.65±1.80	89.5±0.4	8.88±0.20	1.369±0.007	89.5	7.89	1.381
6.53	13.79	92.7	10.10	1.262	87.6	7.85	1.387
HEXT/AS system (O:C=0.50)							
0.92	14.04	75.7	10.58	1.438	85.7	10.83	1.420
2.02	12.88±1.0	76.4±2.3	9.09±0.46	1.409±0.007	81.8	9.34	1.410
3.14	12.31±0.8	76.6±1.5	9.01±0.47	1.408±0.002	81.3	9.04	1.409
5.11	13.53±0.4	78.3±3.4	9.15±0.35	1.396±0.014	83.9±2.8	9.04±0.73	1.412
HEXD/AS system (O:C=0.33)							
1.39	11.48±0.78	70.9±3.5	7.45±0.77	1.406±0.008	81.2	7.93	1.406
2.03	10.54±0.57	82.8±2.3	7.90±0.99	1.382±0.007	87.3	8.83	1.392
2.71	14.55±1.36	77.7±2.1	8.30±0.28	1.391±0.009	89.6±2.0	8.53±0.32	1.388±0.010
3.13	11.02±0.62	81.0±0.7	8.97±0.22	1.384±0.016	89.3	9.14	1.384
5.01	12.22±2.73	79.4±3.4	8.33±0.40	1.384±0.019	89.6±0.1	8.38±0.54	1.390±0.004

3.3 Effect of O:C on phase separation behavior in different systems

The phase separation behavior of organic aerosol particles is strongly influenced by their O:C. As shown in **Fig. 4**, our findings, as well as those from previous studies (You et al., 2013; O'Brien et al., 2015), indicated that there is no correlation between the occurrence of LLPS and the hydrogen-to-carbon (H:C) ratios of the organics, which is consistent with results in previous findings (Bertram et al., 2011; Song et al., 2012). However, a clear trend was observed between LLPS occurrence and the O:C of the organic components. We observed that droplets of 3-MGA/AS, HEXT/AS and HEXD/AS systems with O:C between 0.33 and 0.67 undergo LLPS. With the decrease of water content in the droplets, two distinct phases were formed: an organic-rich phase and a salt-rich aqueous phase, under both acidic and neutral conditions. By contrast, no LLPS occurred in the GL/AS system, as shown in **Fig. S2**. In general, particles with low O:C are more prone to undergo LLPS. This observation is consistent with the findings of Song et al. (2012) who reported that LLPS was never observed when O:C > 0.80 and always observed when O:C < 0.56.

As shown in **Fig. 2** and **Table S2**, for most spectra, WGMs remained after LLPS occurred for droplets of 3-MGA/AS. This phenomenon indicated that the droplets undergo LLPS with a core-shell morphology in most conditions, which is consistent

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with the prediction of Gorkowski et al. (2020). Meanwhile, morphology of phase-separated droplets containing either HEXT or HEXD were also core-shell shape mostly, as depicted in **Figure S3/S4** and **Table S3/S4**. It is attributed to the lower interfacial tension observed at higher O:C, leading to higher possibility condition for forming core-shell shaped droplets (Gorkowski et al., 2020). These findings support the idea that the O:C plays a crucial role in determining the morphology of phase-separated particles in organic/inorganic mixed aerosols.

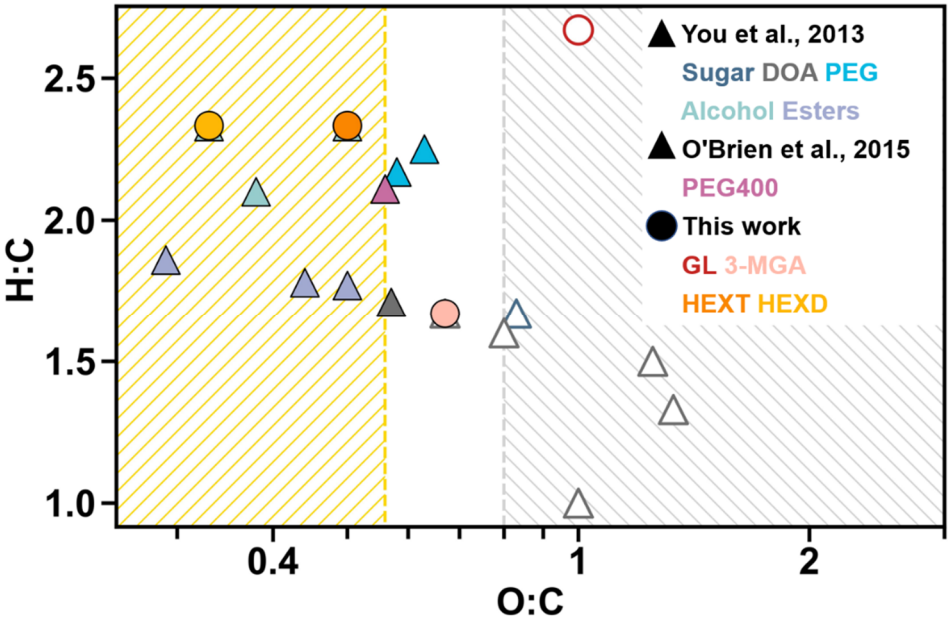

Figure 4. Van Krevelen Diagram for the mixed organic/AS particles: Solid symbols indicate that LLPS was observed, while hollow symbols indicate that LLPS was not observed. Solid triangles represent dicarboxylic acids (DOA, including malonic acid, malic acid, maleic acid, glutaric acid and diethylmalonic acid), sugars (levoglucosan), esters (including diethyl sebacate, suberic acid monomethyl ester and poly diacrylate), alcohols (including 2,5-hexanediol, propylene glycol and 1,2,6-hexanetriol), PEG (including PEG200 and PEG300) obtained from You et al. (2013), and AS-PEG400 obtained from O'Brien et al. (2015).

4 Conclusion

The aim of this study was to investigate the effect of pH and O:C on phase transition behavior of levitated particles using the AOT. Our results show that across aerosol pH in atmospheric condition, the presence of sulfuric acid inhibited the LLPS of aerosol droplets that contained organics (3-MGA, HEXT, HEXD) and AS. Additionally, the MRHs were found to be higher than the SRHs. The O:C of phase-separating systems is 0.67, 0.50, 0.33, and by contrast, LLPS of the high O:C system (GL,

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O:C=1.00) did not occur. Meanwhile, the morphology of levitated aerosol particles was studied and we found that 38 out of 40 droplets that underwent LLPS for a core-shell structure. The results presented here provide new insights into the behavior of different types of aerosol droplets, and the findings have important implications for our understanding of physical and chemical processes that occur in the atmosphere. It is anticipated that future studies will be carried out to examine the OIR-dependent phase separation in real acidified ambient aerosols. Such research will provide insights into the morphological characteristics of real aerosols and the ways in which these characteristics influence important properties such as hygroscopicity and homogenous chemistry. Such information will be helpful in furthering our understanding of the impacts of ambient aerosols on the environment and human health.

Additionally, in-situ measurement or pH estimation methods, such as the real-time AOT analysis in microdroplets reported by Boyer et al. (2020) could be combined with SRH measurements for a more accurate and comprehensive analysis. Furthermore, our study used a surrogate for SOA instead of in situ measurements of real SOA, which can be addressed in future work using SOA generated from a smog chamber or real SOA precursors and oxidized species.

Data availability. The data used in this paper can be obtained from the corresponding author upon request.

Author contributions. YC built the instrument, performed the experiments, analyzed the data, plotted the figures, and wrote the original draft. XP conceptualized the study, contributed to instrumentation, data analysis, discussion, and reviewed the manuscript. HL and CX contributed to the instrumentation and discussion. YM contributed to the experiments and discussion. ZX, FZ contributed to the discussion and manuscript review. TCP contributed to data analysis and manuscript review. ZW administrated the project, conceptualized the study, reviewed the manuscript, and contributed to funding acquisition.

Competing interests. The contact author has declared that none of the authors has any competing interests.

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