

Comprehensive multiphase chlorine chemistry in the box model CAABA/MECCA: Implications to atmospheric oxidative capacity

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Abstract. Tropospheric chlorine chemistry can strongly impact the atmospheric oxidation capacity and composition, especially in urban environments. To account for these reactions, the gas- and aqueous-phase Cl chemistry of the community atmospheric chemistry box model CAABA/MECCA has been extended. In particular, an explicit mechanism for ClNO₂ formation following N₂O₅ uptake to aerosols has been developed. The updated model has been applied to two urban environments with different concentrations of NO_x (NO + NO₂): New Delhi (India) and Leicester (United Kingdom). The model shows a sharp build-up of Cl at sunrise through Cl₂ photolysis in both the urban environments. Besides Cl₂ photolysis, ClO+NO reaction, and photolysis of ClNO₂ and ClONO are also prominent sources of Cl in Leicester. High-NO_x conditions in Delhi tend to suppress the night-time build-up of N₂O₅ due to titration of O₃ and thus lead to lower ClNO₂, in contrast to Leicester. Major loss of ClNO₂ is through its uptake on chloride, producing Cl₂, which consequently leads to the formation of Cl through photolysis. The reactivities of Cl and OH are much higher in Delhi, however, the Cl/OH reactivity ratio is up to ≈9 times greater in Leicester. The contribution of Cl to the atmospheric oxidation capacity is significant and even exceeds (by ≈2.9 times) that of OH during the morning hours in Leicester. Sensitivity simulations suggest that the additional consumption of VOCs due to active gas- and aqueous-phase chlorine chemistry enhances OH, HO₂, and RO₂ near sunrise. The simulation results of the updated model have important implications for future studies on atmospheric chemistry and urban air quality.

15 1 Introduction

Chlorine (Cl) radicals are one of the most important players in the tropospheric chemistry (Seinfeld and Pandis, 2016; Ravishankara, 2009). Cl impacts the oxidative capacity of the atmosphere and radical cycling, and, therefore, can significantly alter the atmospheric composition (Seinfeld and Pandis, 2016; Faxon and Allen, 2013). In comparison with hydroxyl (OH) radicals,

the so-called atmospheric detergent, the much faster reaction rates of Cl with volatile organic compounds (VOCs), enhance the
20 peroxy radicals (RO_2) formation and, thereby, the production of ozone (O_3) and secondary organic aerosols (SOA) (Qiu et al.,
2019a; Choi et al., 2020). In addition, Cl radicals can also enhance the oxidation of climate-driving gases (such as methane
and dimethyl sulphide) (Saiz-Lopez and von Glasow, 2012). Cl radicals are produced in the atmosphere through photochem-
istry involving heterogeneous reactions of Cl-containing gases and aerosols (Qiu et al., 2019a; Faxon and Allen, 2013). The
major sources of Cl-containing species are anthropogenic activities in continental regions and sea salt aerosols in marine and
25 coastal environments (von Glasow and Crutzen, 2007; Osthoff et al., 2008; Liao et al., 2014; Liu et al., 2017; Thornton et al.,
2010; Gunthe et al., 2021; Zhang et al., 2022). The photolysis of reactive Cl-containing species, such as chlorine gas (Cl_2),
hypochlorous acid (HOCl), nitryl chloride (ClNO_2), and chlorine nitrite (ClONO) and the reaction of hydrochloric acid (HCl)
with OH are known to produce Cl radicals in the lower troposphere (Atkinson et al., 2007; Riedel et al., 2014). With the rise in
anthropogenic activities, emissions of Cl-containing species have increased significantly across the globe (Lobert et al., 1999;
30 Zhang et al., 2022), and hence the importance of Cl in local as well as regional atmospheric chemistry has become prominent.

Despite the aforementioned importance, Cl chemistry and associated mechanisms, especially heterogeneous reactions in the
lower troposphere, are not yet fully understood, and the effects of Cl on atmospheric composition, air quality and oxidation
capacity remain uncertain. Field measurements have revealed high concentrations of Cl species over inland regions in addition
35 to coastal and polar regions (von Glasow and Crutzen, 2007; Osthoff et al., 2008; Liao et al., 2014; Liu et al., 2017; Thornton
et al., 2010), however, quantitative understanding of continental sources remains poorly understood. This is due to lack of the
relevant heterogeneous and gas-phase chemistry in atmospheric photochemical models despite the range of chemical mecha-
nisms complexity used in 3-D chemistry transport models (Xue et al., 2015; Pawar et al., 2023; Pozzer et al., 2022). In addition,
the chemistry of Cl compounds has been less studied using the laboratory/chamber experiments. Qiu et al. (2019b) showed
40 that due to inadequate representation of heterogeneous Cl chemistry, the Community Multiscale Air Quality (CMAQ) model
underestimated nitrate concentrations during daytime but overestimated during night-time in Beijing, China. In addition, the
uncertainties associated with emission inventories of Cl species, can lead to inaccurate estimation of air composition (Zhang
et al., 2022; Sharma et al., 2019). For example, Pawar et al. (2023) noticed that even after the inclusion of HCl emissions from
trash burning the levels of nitrate, sulphate, nitrous acid (HONO) etc., still deviated from the observations in Delhi, India,
45 highlighting the need to include emissions from other sectors, such as industries. A few recent studies assessed the impacts
of the gas phase Cl chemistry by including gas phase ClNO_2 reactions, for example, Xue et al. (2015) reported about 25 %
enhancement in the daytime oxidation of carbon monoxide and VOCs at a coastal site in East Asia. In the same region, the
model predicted a 5-16 % enhancement in peak ozone with ClNO_2 ($\approx 50\text{--}200$ pmol/mol) at a mountain top in Hong Kong,
China (Wang et al., 2016). The measurements of Cl_2 (up to ≈ 450 pmol/mol) and ClNO_2 (up to ≈ 3.5 nmol/mol) were reported
50 from a rural site in the North China Plain and Cl chemistry was showed to enhance the formation of peroxy radicals (by 15 %)
and O_3 production rate (by 19 %) (Liu et al., 2017).

Nevertheless, the heterogeneous chemistry of Cl species remains poorly represented in models, and often neglected in large scale numerical simulations. For example, in several models, the heterogeneous uptake of N₂O₅ on aqueous aerosols yielded
55 nitric acid (HNO₃) via reaction HET1:



However, N₂O₅ uptake on chloride-containing particles can produce ClNO₂ (Behnke et al., 1997; Thornton et al., 2010) especially in urban environments with strong NO_x emissions (Osthoff et al., 2008; Young et al., 2012). Incorporating heterogeneous mechanism of ClNO₂ into the regional models led to 3–12 % increase in O₃ over Northern China (Sarwar et al., 2014;
60 Zhang et al., 2017; Liu et al., 2017). In addition, heterogeneous reactions of Cl-containing species including particulate chloride (pCl⁻), Cl₂, ClNO₂, chlorine nitrate (ClNO₃), and hypochlorous acid (HOCl) are suggested to result in the formation of Cl radicals as well as in recycling of NO_x, and HO_x (OH, and HO₂) (Ravishankara, 2009; Qiu et al., 2019a; Hossaini et al., 2016; Faxon and Allen, 2013). Very recent measurements suggest a reduction in ClNO₂ formation due to the competition of N₂O₅ uptake among chloride, sulphate and acetate aerosols (Staudt et al., 2019). These heterogeneous reactions can be of paramount
65 significance in the Cl budget, however, to the best of our knowledge, these are not yet considered in model simulations.

The main goal of the present study is to investigate the role of chlorine chemistry in chemically contrasting urban environments. In this regard, we incorporate comprehensive gas-phase and heterogeneous Cl chemistry into a state of the art box model. Section 2 provides a detailed description of the Cl chemistry mechanism with gas-phase and heterogeneous reactions.
70 Section 3 describes the model setup and Section 4 shows the simulation results which include a detailed investigation on (i) the sensitivity of air composition to chlorine chemistry, (ii) the production and loss of Cl and ClNO₂, (iii) the role of Cl in the Atmospheric Oxidative Capacity (AOC), and (iv) the sensitivity to ClNO₂ + Cl⁻ reaction.

2 Mechanism Development

The community box model “Chemistry As A Boxmodel Application/Module Efficiently Calculating the Chemistry of the
75 Atmosphere” (CAABA/MECCA, Sander et al., 2019), has been used in this work. A comprehensive gas- and aqueous-phase mechanism of chlorine chemistry has been added to MECCA, here used within the box model CAABA. The gas-phase and heterogeneous chemistry implemented in MECCA is described in the following subsections and the full mechanism is shown in the supplementary section.

2.1 Gas-phase chlorine chemistry

80 A total of 36 inorganic, organic and photolysis reactions which are key contributors of Cl radicals were added to the mechanism (Table 1). The mechanism includes the inorganic reactions of Cl with NO_x, NO₃ (G1–G4), the reactions of Cl-containing species with OH and NO (G5–G7), and the reactions between Cl-containing species (G8–G9) (Qiu et al., 2019a; Burkholder et al., 2015; Atkinson et al., 2007). ClONO is formed through reaction of Cl with NO₂ (G2), and exists as a metastable

intermediate (Janowski et al., 1977; Niki et al., 1978; Golden, 2007). This intermediate subsequently transforms into ClNO₂ (G10), with an average conversion time of ≈ 12 h (ranging from 4 to 20 h), and the corresponding rate constant is $2.3 \times 10^{-5} \text{ s}^{-1}$ (Janowski et al., 1977). The Cl-initiated oxidation of organic species i.e. alkanes (C₃H₈, C₄H₁₀), aromatics (benzene (C₆H₆), toluene (C₇H₈) and xylene (C₈H₁₀)), alcohols (CH₃OH, C₂H₅OH), ketones (CH₃COCH₃, MEK), isoprene (C₅H₈), and other organic compounds (C₂H₅CHO, HOCH₂CHO, BENZAL, GLYOX, MGLYOX) have also been included (G11–G31). The corresponding kinetic data are based on the International Union of Pure and Applied Chemistry and NASA Jet Propulsion Laboratory data evaluations (Atkinson et al., 2006, 2007; Burkholder et al., 2015), and from the literature (Niki et al., 1985, 1987; Green et al., 1990; Shi and Bernhard, 1997; Sokolov et al., 1999; Thiault et al., 2002; Wang et al., 2005; Rickard, 2009; Wennberg et al., 2018). In addition, photolysis reactions (G32–G36) resulting in production of Cl are also added to the module (Atkinson et al., 2007). The abbreviations of species mentioned in Table 1 are kept similar to that in the Master Chemical Mechanism (MCM) nomenclature (Rickard, 2009).

Table 1: Gas-phase chlorine reactions and corresponding rate constants added to MECCA. The rate constants are expressed in units of $\text{cm}^3 \text{ molecule}^{-1} \text{ s}^{-1}$ unless otherwise specified. Model-simulated maximum noontime J -values for Delhi are provided.

Reaction	Rate constant	Reference
Inorganic reactions		
(G1) Cl + NO + M \rightarrow ClNO	$7.6\text{E}(-32) * (\text{T}/300)^{-1.8}$	Qiu et al. (2019a)
(G2) Cl + NO ₂ + M \rightarrow ClONO	1.6E-11	Burkholder et al. (2015)
(G3) Cl + NO ₂ + M \rightarrow ClNO ₂	3.6E-12	Burkholder et al. (2015)
(G4) Cl + NO ₃ \rightarrow ClO + NO ₂	2.40E-11	Qiu et al. (2019a)
(G5) Cl ₂ + OH \rightarrow HOCl + Cl	$3.6\text{E}-12 * \exp(-1200/\text{T})$	Atkinson et al. (2007)
(G6) ClNO ₂ + OH \rightarrow HOCl + NO ₂	$2.4\text{E}-12 * \exp(-1250/\text{T})$	Atkinson et al. (2007)
(G7) OClO + NO \rightarrow NO ₂ + ClO	$1.1\text{E}-13 * \exp(350/\text{T})$	Atkinson et al. (2007)
(G8) Cl + Cl ₂ O \rightarrow Cl ₂ + ClO	$6.2\text{E}-11 * \exp(130/\text{T})$	Atkinson et al. (2007)
(G9) ClO + OClO + M \rightarrow Cl ₂ O ₃	1.2E-12	Atkinson et al. (2007)
(G10) ClONO \rightarrow ClNO ₂	$2.3\text{E}-5 \text{ s}^{-1}$	Janowski et al. (1977)
Organic reactions		
(G11) Cl + C ₃ H ₈ \rightarrow iso-C ₃ H ₇ O ₂ + HCl	$1.4\text{E}-10 * 0.43 * \exp(75/\text{T})$	Rickard (2009)
(G12) Cl + C ₃ H ₈ \rightarrow n-C ₃ H ₇ O ₂ + HCl	$1.4\text{E}-10 * 0.59 * \exp(-90/\text{T})$	Rickard (2009)
(G13) Cl + iso-C ₄ H ₁₀ \rightarrow iso-C ₄ H ₉ O ₂ + HCl	$1.43\text{E}-10 * 0.564$	Rickard (2009)
(G14) Cl + iso-C ₄ H ₁₀ \rightarrow tert-C ₄ H ₉ O ₂ + HCl	$1.43\text{E}-10 * 0.436$	Rickard (2009)
(G15) Cl + n-C ₄ H ₁₀ \rightarrow LC ₄ H ₉ O ₂ + HCl	2.05E-10	Atkinson et al. (2006), Rickard (2009)
(G16) Cl + benzene \rightarrow C ₆ H ₅ O ₂ + HCl	1.3E-16	Sokolov et al. (1999)

(G17)	Cl + toluene → C ₆ H ₅ CH ₂ O ₂ + HCl	6.20E-11	Wang et al. (2005)
(G18)	Cl + isoprene → .63 LISOPAB + .30 LISOPCD + .07 LISOPEFO2 + HCl	7.6E-11*exp(500/T)*1.1*exp(-595/T)	Wennberg et al. (2018)
(G19)	Cl + isoprene → .63 LISOPAB + .30 LISOPCD + .07 LISOPEFO2 + LCHLORINE	7.6E-11*exp(500/T)*(1-1.1*exp(-595/T))	Wennberg et al. (2018)
(G20)	Cl + xylene → C ₆ H ₅ CH ₂ O ₂ + LCARBON + HCl	1.50E-10	Shi and Bernhard (1997)
(G21)	Cl + CH ₃ OH → HOCH ₂ O ₂ + HCl	7.1E-11*0.59*exp(-75/T)	Atkinson et al. (2006)
(G22)	Cl + C ₂ H ₅ OH → HOCH ₂ CH ₂ O ₂ + HCl	6.0E-11*exp(155/T)*0.28*exp(-350/T)	Atkinson et al. (2006)
(G23)	Cl + C ₂ H ₅ OH → C ₂ H ₅ O ₂ + HCl	6.0E-11*exp(155/T)*(1-0.28*exp(-350/T))	Atkinson et al. (2006)
(G24)	Cl + HOCH ₂ CHO → HOCHCHO + HCl	8.0E-12/0.9*0.35	Atkinson et al. (2006), Niki et al. (1987)
(G25)	Cl + HOCH ₂ CHO → HOCH ₂ CO + HCl	8.0E-12/0.9*(1-0.35)	Atkinson et al. (2006), Niki et al. (1987)
(G26)	Cl + GLYOX → HCOCO + HCl	3.8E-11	Niki et al. (1985)
(G27)	Cl + MGLYOX → CH ₃ CO + CO + HCl	4.8E-11	Green et al. (1990)
(G28)	Cl + C ₂ H ₅ CHO → C ₂ H ₅ CO ₃ + HCl	1.3E-10	Atkinson et al. (2006)
(G29)	Cl + CH ₃ COCH ₃ → CH ₃ COCH ₂ O ₂ + HCl	1.5E-11*exp(-590/T)	Atkinson et al. (2006)
(G30)	Cl + MEK → LMEKO ₂ + HCl	3.05E-11*exp(80/T)	Atkinson et al. (2006)
(G31)	Cl + BENZAL → C ₆ H ₅ CO ₃ + HCl	1.0E-10	Thiault et al. (2002)
Photolysis reactions		J-value (s⁻¹)	
(G32)	ClO → Cl + O3P	1.45E-4	Atkinson et al. (2007)
(G33)	Cl ₂ O → Cl + ClO	9.20E-4	Atkinson et al. (2007)
(G34)	Cl ₂ O ₃ → ClO + ClO ₂	5.50E-4	Atkinson et al. (2007)
(G35)	ClNO → Cl + NO	2.89E-3	Atkinson et al. (2007)
(G36)	ClONO → Cl + NO ₂	3.81E-3	Atkinson et al. (2007)

95 2.2 Heterogeneous chemistry

The aqueous-phase and heterogeneous chemistry of Cl compounds added to the MECCA is described in Table 2. In the present study, we assume that N₂O₅ is in equilibrium between the gas- and aqueous-phase (H2) according to Henry's law and the dissociation of N₂O₅(aq) to nitronium ion (NO₂⁺) and nitrate (NO₃⁻), occurs according to reaction (A1). The rate constant for the recombination reaction of NO₂⁺ and NO₃⁻ is 2.7×10⁸ mol⁻¹ L s⁻¹, calculated based on Bertram and Thornton (2009);

100 Staudt et al. (2019). The acid dissociation of nitric acid (HNO₃) in aqueous phase (A3) also results in formation of NO₂⁺ (Sapoli et al., 1985).

Table 2: Aqueous-phase and heterogeneous chlorine reactions added to MECCA

Reaction	Rate constant	Reference
Aqueous-phase reactions		
(A1) N ₂ O ₅ (aq) → NO ₂ ⁺ (aq) + NO ₃ ⁻ (aq)	1.5 × 10 ⁵ s ⁻¹	Staudt et al. (2019)
(A2) NO ₂ ⁺ (aq) + NO ₃ ⁻ (aq) → N ₂ O ₅ (aq)	2.7 × 10 ⁸ mol ⁻¹ L s ⁻¹	Bertram and Thornton (2009); Staudt et al. (2019)
(A3) HNO ₃ (aq) + H ⁺ (aq) → NO ₂ ⁺ (aq) + H ₂ O(aq)	1.6 × 10 ⁹ mol ⁻¹ L s ⁻¹	Sapoli et al. (1985)
(A4) NO ₂ ⁺ (aq) + Cl ⁻ (aq) → ClNO ₂ (aq)	7.5 × 10 ⁹ mol ⁻¹ L s ⁻¹	Staudt et al. (2019)
(A5) ClNO ₂ (aq) → NO ₂ ⁺ (aq) + Cl ⁻ (aq)	2.70 × 10 ² s ⁻¹	Behnke et al. (1997)
(A6) ClNO ₂ (aq) + Cl ⁻ (aq) → Cl ₂ (aq) + NO ₂ ⁻ (aq)	10 ⁷ mol ⁻¹ L s ⁻¹	Roberts et al. (2008)
(A7) OH·Cl ⁻ (aq) + OH·Cl ⁻ (aq) → Cl ₂ (aq) + 2 OH ⁻ (aq)	1.8 × 10 ⁹ mol ⁻¹ L s ⁻¹	Knipping et al. (2000)
(A8) OH·Cl ⁻ (aq) + Cl ⁻ (aq) → Cl ₂ ⁻ (aq) + 2 OH ⁻ (aq)	10 ⁴ mol ⁻¹ L s ⁻¹	Grigorev et al. (1987)
(A9) Cl ₂ ⁻ (aq) + 2 OH ⁻ (aq) → OH·Cl ⁻ (aq) + Cl ⁻ (aq)	4.5 × 10 ⁷ mol ⁻¹ L s ⁻¹	Grigorev et al. (1987)
(A10) NO ₂ ⁺ (aq) + H ₂ O(aq) → HNO ₃ (aq) + H ⁺ (aq)	1.6 × 10 ⁷ mol ⁻¹ L s ⁻¹	Staudt et al. (2019)
(A11) NO ₂ ⁺ (aq) + SO ₄ ²⁻ (aq) → SO ₄ ²⁻ (aq) + NO ₃ ⁻ (aq) + 2 H ⁺ (aq)	7.5 × 10 ⁹ mol ⁻¹ L s ⁻¹	Staudt et al. (2019)
(A12) NO ₂ ⁺ (aq) + HCOO ⁻ (aq) → HCOO ⁻ (aq) + NO ₃ ⁻ (aq) + 2 H ⁺ (aq)	7.5 × 10 ⁹ mol ⁻¹ L s ⁻¹	Staudt et al. (2019)
(A13) NO ₂ ⁺ (aq) + CH ₃ COO ⁻ (aq) → CH ₃ COO ⁻ (aq) + NO ₃ ⁻ (aq) + 2 H ⁺ (aq)	7.5 × 10 ⁹ mol ⁻¹ L s ⁻¹	Staudt et al. (2019)
(A14) NO ₂ ⁺ (aq) + phenol(aq) → HOC ₆ H ₄ NO ₂ (aq) + H ⁺ (aq)	7.5 × 10 ⁹ mol ⁻¹ L s ⁻¹	Ryder et al. (2015); Heal et al. (2007)
(A15) NO ₂ ⁺ (aq) + CH ₃ OH(aq) → CH ₃ NO ₃ (aq) + H ⁺ (aq)	4.5 × 10 ⁸ mol ⁻¹ L s ⁻¹	Iraci et al. (2007)
(A16) NO ₂ ⁺ (aq) + CRESOL(aq) → TOL1OHNO2(aq) + H ⁺ (aq)	7.5 × 10 ⁹ mol ⁻¹ L s ⁻¹	Coombes et al. (1979)
Heterogeneous reactions		
(H2) N ₂ O ₅ (g) ⇌ N ₂ O ₅ (aq)	8.8E-2	Fried et al. (1994)
(H3) ClNO ₂ (g) ⇌ ClNO ₂ (aq)	4.5E-2	Frenzel et al. (1998)

(H4)	$\text{HOC}_6\text{H}_4\text{NO}_2(\text{aq}) \rightleftharpoons \text{HOC}_6\text{H}_4\text{NO}_2(\text{g})$	8.9E1	Müller and Heal (2001)
(H5)	$\text{CH}_3\text{NO}_3(\text{aq}) \rightleftharpoons \text{CH}_3\text{NO}_3(\text{g})$	2.0E0	Sander (2015)

Thus produced nitronium ion (NO_2^+) reacts reversibly with chloride (Cl^-) yielding ClNO_2 (A4, A5) (Staudt et al., 2019; Behnke et al., 1997). After outgassing according to Henry's law (H3), ClNO_2 is photolyzed in the gas phase, producing Cl and NO_2 (Ghosh et al., 2012; Sander et al., 2014). ClNO_2 uptake on chloride containing aerosols results in formation of Cl_2 and nitrite ion (NO_2^-), as shown by the reaction (A6) (Roberts et al., 2008). Chamber experiments suggest the formation of Cl_2 from the self reaction of $\text{OH}\cdot\text{Cl}^-$ (A7), which gets formed via the reaction of OH with Cl^- (Knipping et al., 2000). Through other channel of reversible reactions (A8, A9), $\text{OH}\cdot\text{Cl}^-$ reacts with aqueous chloride and produces Cl_2^- , which can yield Cl_2 through subsequent reactions (Grigorev et al., 1987). The NO_2^+ uptake on aqueous chloride to form ClNO_2 (A4) is ≈ 500 times faster than NO_2^+ reaction with H_2O (A10) (Staudt et al., 2019). At the same time, experimental studies revealed a strong competition of NO_2^+ to react with Cl^- and with other nucleophiles (e.g. SO_4^{2-}) and aqueous organic compounds e.g. phenol, methanol, cresol (A11–A16) (Staudt et al., 2019; Ryder et al., 2015; Heal et al., 2007; Iraci et al., 2007; Coombes et al., 1979). These reactions could suppress the formation of ClNO_2 and also the corresponding rate constants for reactions A11–A14 are similar to the $\text{NO}_2^+ + \text{Cl}^-$ reaction yielding ClNO_2 i.e. $7.5 \times 10^9 \text{ mol}^{-1} \text{ L s}^{-1}$ (Staudt et al., 2019; Ryder et al., 2015; Heal et al., 2007). Methanol reacts with NO_2^+ (A15) and forms aqueous methyl nitrate (CH_3NO_3) (Iraci et al., 2007). Phase exchange for CH_3NO_3 and nitrophenol ($\text{HOC}_6\text{H}_4\text{NO}_2$) is shown by reactions H4 and H5, respectively. The heterogeneous chemistry just discussed is implemented in MECCA and is summarized in Fig. 1. The rate constant for NO_2^+ reaction with methoxyphenol is about ≈ 10000 times smaller than $\text{NO}_2^+ + \text{phenol}$ reaction (Kroflíč et al., 2015), so it is not considered in this study. In addition, nitration reactions of other alcohols (e.g. catechol and polyphenols) could be potentially important, however due to unavailability of corresponding rate constants, these reactions are not considered in this study, nonetheless future studies calculating the kinetics of these reactions are recommended.

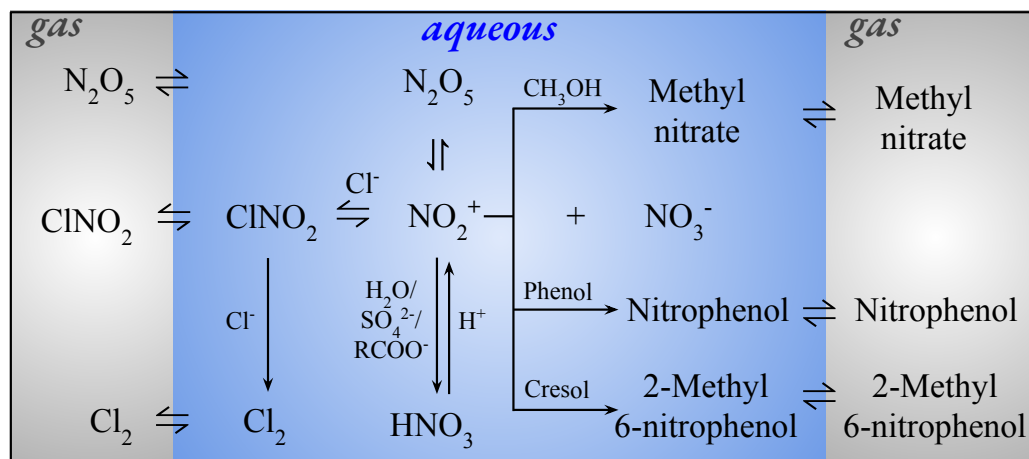


Figure 1. Aqueous-phase and heterogeneous chemistry added to MECCA.

120 In this study, Cl reacts with hydrocarbons and acetone via H-abstraction, and hence does not lead to the formation of any
Cl-containing molecules, such as chloroacetone. This means that there are no such reactions in MECCA in which the Cl atom
becomes part of the organic molecule. The reaction of Cl atoms with isoprene proceeds mainly via addition, and it produces
chlorine-containing organics (Ragains and Finlayson-Pitts, 1997; Fan and Zhang, 2004). However, here we have simplified the
mechanism by not considering the fate of organohalogens. Therefore, for future research, it would be valuable to investigate
125 the chemical kinetics of such reactions kinetics and their importance in the formation of organohalogen compounds.

3 Box model setup

The chemistry described in section 2 has been added into community box model CAABA/MECCA v4.4.2 (Sander et al., 2019).
A comprehensive gas and aqueous phase tropospheric chemistry involving total 3330 reactions was utilized for the simulations,
and the full set of reactions are presented in the electronic supplement. The gas-phase chemistry of organics like terpenes and
130 aromatics is treated by the Mainz Organic Mechanism (MOM) (Taraborrelli et al., 2012; Nölscher et al., 2014; Hens et al., 2014;
Taraborrelli et al., 2021). The aqueous-phase chemistry of oxygenated VOCs is treated by the Jülich Atmospheric Mechanism
of Organic Chemistry (JAMOC) (Rosanka et al., 2021). The numerical integration of the chemical mechanism is performed by
the kinetic preprocessor v2.1 (KPP) (Sandu and Sander, 2006). The photolysis rate constants (J values) are calculated by the
submodel JVAL, based on the method by Landgraf and Crutzen (1998). The Cl chemistry is expected to be more prominent
135 during winter conditions due to higher concentration of Cl-containing species in the boundary layer (Thornton et al., 2010;
Gunthe et al., 2021; Sommariva et al., 2021), and therefore, simulations are performed for the winter season. Hence, the model
is set-up for typical winter conditions of two different urban environments: Delhi (28.6° N, 77.2° E), India and Leicester (52.4°
N, 01.1° W), United Kingdom. Simulations are performed for a 5-day period (17–21 February 2018) and output of 5th day has
been considered for the analysis; by then, radicals had achieved almost a steady state. The typical environmental conditions
140 used in the simulations for Delhi (Tripathi et al., 2022) and Leicester (Sommariva et al., 2021) are summarized in Tab. 3 and
Tab. S1.

VOC emissions are taken from the CAMS inventory (Sindelarova et al., 2014; Granier et al., 2019) and are adjusted iter-
atively in magnitude for better agreement with observations. CAMS-GLOB-ANT v5.3 (0.1° × 0.1°) (Granier et al., 2019)
provides emissions of anthropogenic VOCs (e.g., benzene, toluene etc.), while emissions of biogenic VOCs (e.g., isoprene)
145 are from CAMS-GLOB-BIO v3.1 (0.25° × 0.25°) (Sindelarova et al., 2014). Emission of HCl and particulate chloride are in-
cluded from Zhang et al. (2022) and adjusted iteratively towards reported levels of Cl-containing species (Gunthe et al., 2021;
Sommariva et al., 2021). The Mainz Organic Mechanism (MOM) dry deposition scenario (Sander et al., 2019) is activated in
the model. Ground-based lidar measurements of boundary layer height (BLH) during winter-time, performed as a part of the
European Integrated project on Aerosol Cloud Climate and Air Quality Interactions (EUCAARI) project, are utilized for the
simulations at Delhi (Nakoudi et al., 2018). The diurnal variation in BLH in Leicester is extracted from the European centre
150 for medium-range weather forecast's (ECMWF) fifth-generation reanalysis dataset ERA5 (Hersbach et al., 2020). Air compo-

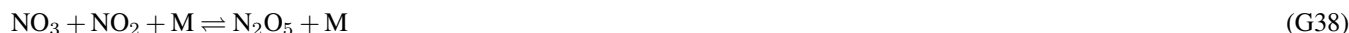
Table 3. Environmental conditions of Delhi and Leicester in the model simulations.

Parameter	Delhi	Leicester
Latitude	28.58° N	52.38° N
Longitude	77.22° E	01.08° W
Time-zone	GMT+5:30	GMT+0:00
Temperature (K)	292	278.1
Pressure (mbar)	1010	1004
Air number density (molecules cm ⁻³)	2.5 × 10 ¹⁹	2.61 × 10 ¹⁹
Relative Humidity	67 %	90 %

sition in the model has been initialized based on previous studies (Tab. S1; Zhang et al. (2007); Lanz et al. (2010); Lawler et al. (2011); Sommariva et al. (2018, 2021); Gunthe et al. (2021); Tripathi et al. (2022)). The values of aerosol properties (e.g., radius, liquid water content, and chemical composition) incorporated in the simulations for both Delhi and Leicester as provided in Table. S1. We constrained the model with the parameterized function best representing the observed diurnal variations of NO_x (Fig. 2) (Tripathi et al. (2022); Sommariva et al. (2018, 2021), <https://uk-air.defra.gov.uk/data/>) which helped in better reproducing the diurnal variations of some VOCs (e.g. isoprene) and ozone. Diurnal observations of HONO from Sommariva et al. (2021) are used for Leicester. For Delhi, however, HONO couldn't be constrained due to lack of observations.

4 Results and Discussion

The model captures the patterns in O₃ variability at both locations (Sommariva et al., 2018; Nelson et al., 2021; Chen et al., 2021; Sommariva et al., 2021; Nelson et al., 2023) to an extent, as shown in Fig. 2. O₃ is underestimated after ≈16:00 h LT in Leicester mainly due to titration by high NO and lack of adequate dynamics/transport of O₃ in the model. Entrainment seems to improve O₃ after mid-night, towards the observed values (Fig. 2i). Simulated isoprene is in agreement with diurnal observations in Delhi (Tripathi et al., 2022) and in accordance with observed mean level in Leicester (Sommariva et al., 2021). The nitrate radical (NO₃), which is a nighttime oxidant, is formed through reaction between NO₂ and O₃ (G37). NO₃ can react with NO₂ forming N₂O₅, which can again produce NO₃ and NO₂ through thermal dissociation (G38).



As seen in Fig. 2e, NO₃ remains negligible during the night-time (≈18:00–07:30 h LT) in Delhi due to unavailability of O₃ under high-NO conditions (up to 200 nmol/mol). Interestingly, despite its very short lifetime (≈5 s), about ≈0.1 pmol/mol of NO₃ sustains during daytime. This is primarily due to prevailing levels of NO₂ (≈30 nmol/mol) and O₃ (≈40 nmol/mol).

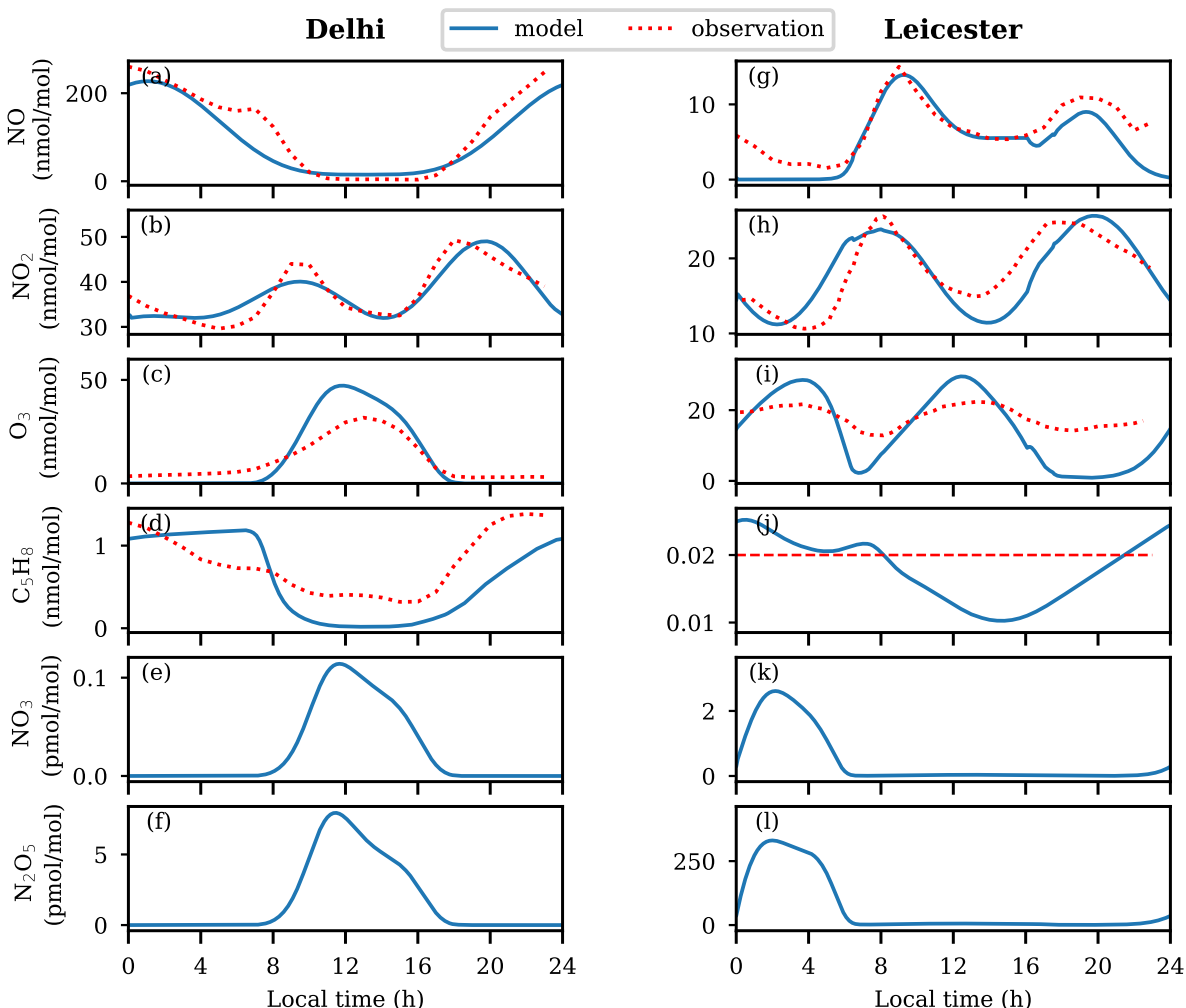


Figure 2. Diurnal variations of NO, NO₂, O₃, C₅H₈, NO₃, and N₂O₅ mixing ratios in Delhi (left) and Leicester (right). The unusual and negligible nighttime NO₃ in Delhi is attributed to the nearly non-existent O₃, due to titration by higher concentrations of NO. This leads to the negligible nighttime N₂O₅ in this region. Although mixing ratios of NO₃ and N₂O₅ peak during the daytime, their levels remain quite low. Mean value of observed C₅H₈ in Leicester is shown by red colored long dashed line.

Such unusual daytime enhanced NO₃ have been reported in recent studies, for example, 5-31 pmol/mol of NO₃ in Texas, USA (Geyer et al., 2003). Aircraft measurements during the New England Air Quality Study showed ≈ 0.5 pmol/mol of NO₃ within boundary layer (≤ 1 km) during noon time (Brown et al., 2005). The calculated NO₃ levels using steady state approximation showed 0.01-0.06 pmol/mol of NO₃ for the 1997-2012 period at urban sites in the UK (Marylebone Road London, London Eltham, and Harwell) (Khan et al., 2015a). Horowitz et al. (2007) suggested that NO₃ in tenths of pmol/mol during daytime over the eastern United States results in formation of $\approx 50\%$ isoprene nitrates through oxidation of isoprene, which could

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further affect the formation of O₃ and SOA significantly (Horowitz et al., 2007). Following to higher NO₃, up to 8 pmol/mol of N₂O₅ is simulated during daytime in Delhi (Fig. 2f). Similar unusual daytime high levels of N₂O₅ ($\approx 21.9 \pm 29.3$ pptv) during wintertime were recently measured at Delhi using a high-resolution iodide adduct chemical ionization mass spectrometer (Haslett et al., 2023).

Enhanced NO₃ ≈ 2.6 pmol/mol and N₂O₅ ≈ 330 pmol/mol are simulated after mid-night in Leicester (Fig. 2k, 2l). In contrast to Delhi, the daytime simulated levels of NO₃ are negligible as it gets removed rapidly during the daytime by photolysis and through its reactions with NO, HO₂, RO₂, and VOCs (Khan et al., 2015b). In conjunction with high NO from $\approx 16:00$ h LT to near midnight that titrates O₃, the corresponding NO₃ and N₂O₅ is negligible (following reactions G37 and G38). Night-time high and negligible day-time levels of NO₃ and N₂O₅ are their typical features which are generally reported in the literature (Brown et al., 2001; Seinfeld and Pandis, 2016).

190 4.1 Sensitivity of air composition to chlorine chemistry

To investigate the effects of Cl chemistry on air composition, other than comprehensive chemistry simulation discussed in previous section (simulation: NEW i.e. chemistry already present in the model + newly added gas and aqueous phase chlorine chemistry), two additional simulations have been performed, which are: (1) OLD – this includes default chemistry already present in the model, and (2) NOCL – OLD minus chlorine chemistry (i.e. without Cl chemistry). OLD simulation also encompassed some basic chlorine chemistry that was part of the model prior to its update (full mechanism is also shown in supplement). Figure 3 shows the comparison of Cl, ClNO₂, ClONO, OH, HO₂, and RO₂ variations among the three simulations in Delhi and Leicester. Figure S5 shows the differences in diurnal variations of Cl, ClONO + ClNO₂, OH, HO₂, and RO₂ in NEW simulation with: NOCL and OLD simulations.

200 The Delhi environment is mainly characterized by two peaks in Cl, a predominant sharp peak just after sunrise followed by a broad shallow peak during noontime, corresponding to different mechanisms as discussed in the next section. With newly added chemistry (NEW simulation), a sharp peak in Cl is seen near sunrise, with the maximum values attained is ≈ 3.5 fmol/mol (8.75×10^4 molec cm⁻³) in Delhi (Fig. 3a). A broad smaller peak with magnitude of ≈ 0.8 fmol/mol maximizing around noontime is seen, which is ≈ 4 times smaller than the first morning peak. OLD simulation also show a sharp peak in Cl near sunrise in Delhi, with a maximum of ≈ 11 fmol/mol (2.75×10^5 molec cm⁻³). Cl get suppressed by up to ≈ 0.01 pmol/mol of maximum value in the OLD simulation, in the presence of added chlorine chemistry (NEW) as shown in Fig. S5. Similar to Cl, a peak is seen in ClONO + ClNO₂ of ≈ 100 pmol/mol with sunrise, which gradually decreases and attain ≈ 7 pmol/mol from nearly 11:00–16:00 h LT. Afterwards it increases to ≈ 20 pmol/mol from late evening as shown by Fig. 3b,c. The pathways for the formation of ClNO₂ and ClONO were absent in earlier version of the model (OLD). Simulated OH, HO₂, and RO₂ show a prominent peak just after sunrise in the presence of Cl chemistry for both the OLD and NEW simulations (Fig. 3d,e,f). As a consequence of greater oxidation of VOCs by Cl, enhanced levels of OH by 0.05 pmol/mol (up to a factor of ≈ 1.8), HO₂

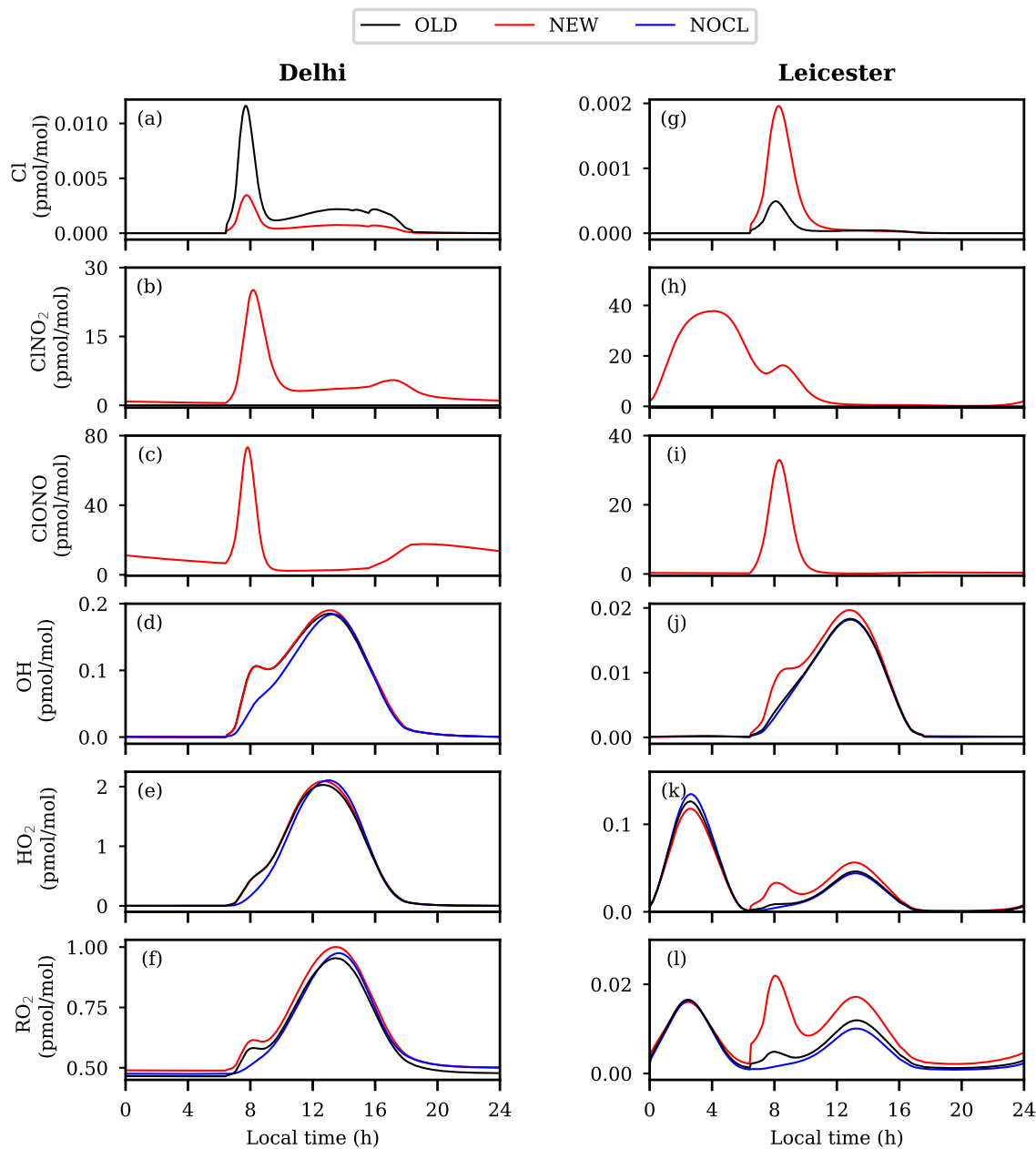


Figure 3. Model simulated diurnal variations of Cl, ClNO₂, ClONO, OH, HO₂, and RO₂ at Delhi (left panel) and Leicester (right panel).

by 0.21 pmol/mol and RO₂ by 0.1 pmol/mol are noted with added Cl chemistry compared to NOCL case (see Fig. S5). No significant changes are seen in noon-time levels of OH and HO₂, whereas ≈ 1.1 times more RO₂ is produced with added Cl

chemistry (NEW) compared to the OLD simulation.

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The model-predicted Cl peaks at ≈ 2 fmol/mol (5.2×10^4 molec cm^{-3}) during sunrise in Leicester (Fig. 3g). In contrast to Delhi, suppressed Cl (up to ≈ 3.2 times) with a narrow peak is simulated by OLD simulation in comparison with NEW simulation containing newly added Cl chemistry, at Leicester. In contrast to negligible night-time ClONO + ClNO₂ in Delhi, it shows a strong build-up over Leicester during 0-4 hours with a maximum of ≈ 40 pmol/mol, with higher levels (up to 220 50 pmol/mol) prevailing until about sunrise. ClONO + ClNO₂ is negligible during mid-day until mid-night, in accordance with N₂O₅ in Leicester as shown in Fig. 2l. Previous studies have demonstrated that the formation of ClNO₂ occurs within the nocturnal residual layer, which contains lower levels of NO compared to the surface layer. Subsequently, ClNO₂ mixes downward during the morning when the convective mixed layer develops (Bannan et al., 2015; Tham et al., 2016). However, the present study does not account the the effect of transport processes due to the limitations of the box model. The effects 225 of added Cl chemistry on OH, HO₂, and RO₂ are more prominent in Leicester compared to Delhi. NEW simulation show strong enhancements in OH (up to ≈ 2 times), HO₂ (up to ≈ 5 times), and RO₂ (up to ≈ 8 times) after sunrise which is gradually progressive, resulting in higher levels during noon-time as well (Fig. 3, Fig. S5). Remarkably elevated levels of RO₂ (by a factor of ≈ 2) are prominent during the noon hours. Such elevated levels of RO₂ could favour enhanced levels of secondary organic aerosols in Leicester. The impact of Cl chemistry on aerosols (NO₂⁺, NO₃⁻, and oxalic acid) is discussed 230 in Supplementary section 2.2 (Fig. S6). Though significant differences in NO₂⁺, NO₃⁻, and oxalic acid are seen due to Cl chemistry, further measurements are required for validation. In the next sections, we have analysed the observed behaviour of Cl and ClNO₂ in the NEW simulation over both the locations in more detail.

4.2 Production and loss of Cl and ClNO₂

The sources and sinks of Cl in Leicester and Delhi are presented in Fig. 4. The left-upper panel (a) delineates the sources 235 and sinks of Cl radical on diurnal scale in Delhi. The morning sharp peak in Cl radical is caused mainly by the photolysis of Cl₂ with a maximum rate of 1.2×10^7 molec $\text{cm}^{-3} \text{ s}^{-1}$. The shallow secondary peak is due to the reaction HCl + OH with a noon time rate of $\approx 0.4 \times 10^7$ molec $\text{cm}^{-3} \text{ s}^{-1}$. However, there is a smaller contribution from other reactions (photolysis of ClNO₂, ClONO and reaction of ClO with NO) to the morning peak, which have negligible contributions during the daytime. Interestingly, there is a strong consumption of Cl to oxidize VOCs (peak rate $\approx 2.4 \times 10^7$ molec $\text{cm}^{-3} \text{ s}^{-1}$) during sunrise, and 240 a lesser consumption during the rest of the day. Cl + NO₂ is also a Cl sink during the morning time in Delhi. The Cl-initiated oxidation of VOCs in the morning hours in Delhi may lead to formation of secondary organic aerosols and new particle formation, which opens up pathways of future research in this direction. In addition to Cl₂ photolysis ($\approx 1.0 \times 10^6$ molec $\text{cm}^{-3} \text{ s}^{-1}$), photolysis of ClNO₂ and ClONO, and ClO + NO reaction (total rate $\approx 0.8 \times 10^6$ molec $\text{cm}^{-3} \text{ s}^{-1}$) are other prominent sources of Cl in Leicester. VOCs are the major sink for Cl (rate $\approx 1.3 \times 10^6$ molec $\text{cm}^{-3} \text{ s}^{-1}$), followed by NO₂ (rate $\approx 0.6 \times$ 245 10^6 molec $\text{cm}^{-3} \text{ s}^{-1}$).

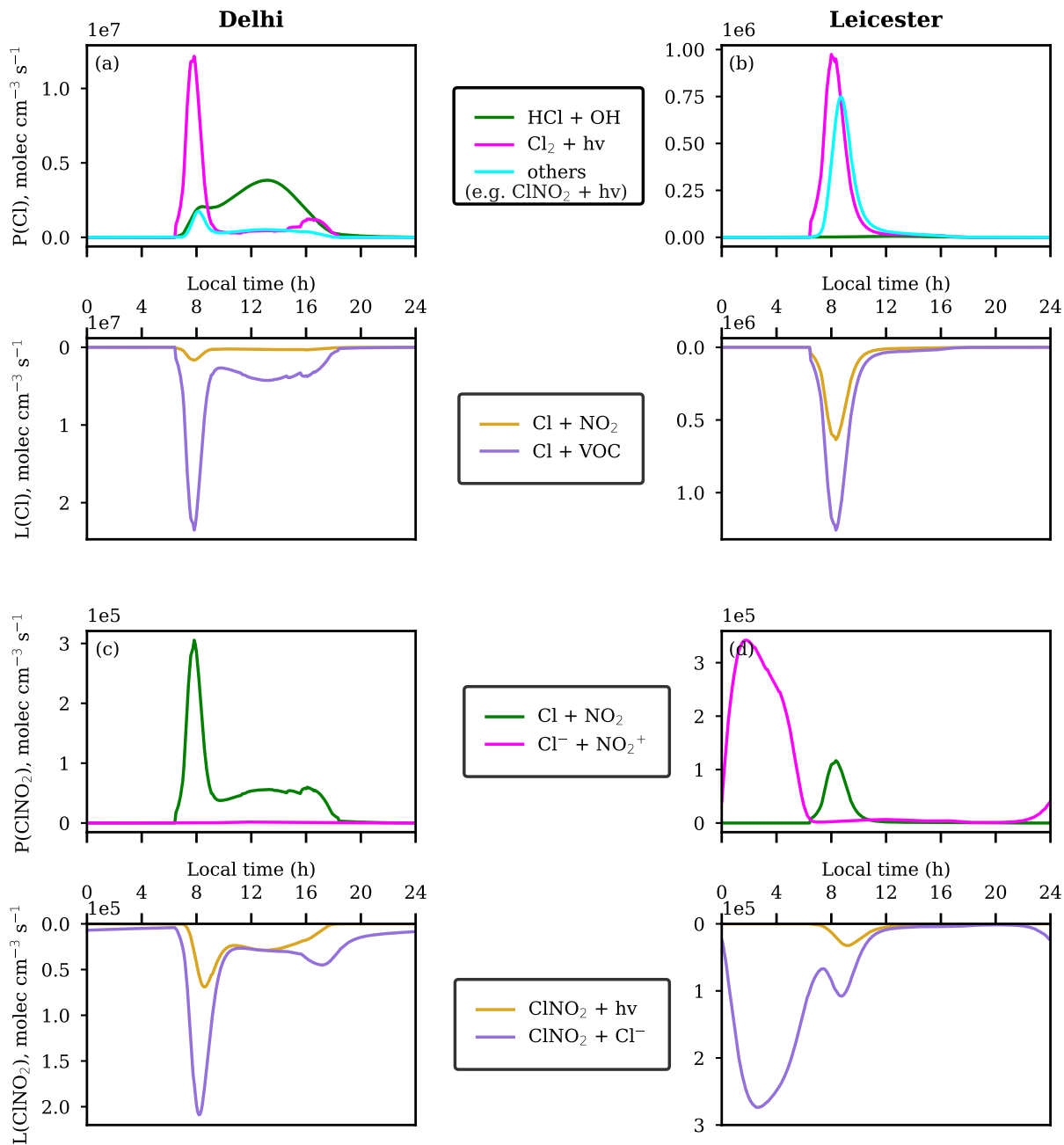


Figure 4. Production and loss rates of (a, b) Cl and (c, d) ClNO₂ in Delhi (left panel) and Leicester (right panel).

We further analyzed the production and loss pathways of ClNO₂, as shown in Fig. 4c,d. While the major source of ClNO₂ is through the Cl + NO₂ reaction with a reaction rate $\approx 3 \times 10^5 \text{ molec cm}^{-3} \text{ s}^{-1}$ in Delhi, the aqueous phase reaction $\text{Cl}^- + \text{NO}_2^+$

($\approx 3.4 \times 10^5$ molec $\text{cm}^{-3} \text{s}^{-1}$) is the prominent source in Leicester corresponding to the peak ClNO_2 (Fig. 2h,p). Though gas-phase reaction $\text{Cl} + \text{NO}_2$ is discussed in the literature (Burkholder et al., 2015; Qiu et al., 2019a), however, to the best of our knowledge, such an unusually higher contribution of this reaction (seen in Delhi) as compared to the aqueous-phase reaction of $\text{Cl}^- + \text{NO}_2^+$ has not been reported in any study. The reaction of Cl with NO_2 ($\approx 1.1 \times 10^5$ molec $\text{cm}^{-3} \text{s}^{-1}$) is the major ClNO_2 source during sunrise in Leicester. In contrast, there is lesser contribution of $\text{Cl}^- + \text{NO}_2^+$ reaction (rate $\approx 1 \times 10^3$ molec $\text{cm}^{-3} \text{s}^{-1}$) in ClNO_2 production in Delhi. The prominent sink for ClNO_2 is through its heterogeneous reaction with Cl^- ($\approx 1.8 \times 10^5$ molec $\text{cm}^{-3} \text{s}^{-1}$ or 7.2×10^{-15} mol $\text{mol}^{-1} \text{s}^{-1}$) in Delhi almost throughout the day, while its loss through the photolysis ($\approx 0.5 \times 10^5$ molec $\text{cm}^{-3} \text{s}^{-1}$ or 2×10^{-15} mol $\text{mol}^{-1} \text{s}^{-1}$) is also an important sink during the daytime. We are using ClNO_2 uptake coefficient, $\gamma = 9\text{E-}3$ from Fickert et al. (1998) in the simulation. Sensitivity simulation with $\gamma = 1\text{E-}5$ (Haskins et al., 2019) results in considerably slower (by a factor of ≈ 270 and ≈ 17 , near sunrise and during mid-day, respectively) loss rate of ClNO_2 with Cl^- than in the NEW simulation over Delhi. ClNO_2 loss through the reaction $\text{ClNO}_2 + \text{Cl}^-$ ($\approx 2.7 \times 10^5$ molec $\text{cm}^{-3} \text{s}^{-1}$ or 1.0×10^{-14} mol $\text{mol}^{-1} \text{s}^{-1}$) is its major sink in Leicester from mid-night to mid-day, while photolysis ($\approx 0.3 \times 10^5$ molec $\text{cm}^{-3} \text{s}^{-1}$ or 1.1×10^{-15} mol $\text{mol}^{-1} \text{s}^{-1}$) is smaller sink from sunrise to mid-day here. The diurnal variation in Cl_2 , and its production and loss mechanisms over Delhi and Leicester are shown by Fig. S1 and Fig. S2. In conjunction with major loss of ClNO_2 , $\text{ClNO}_2 + \text{Cl}^-$ reaction is the major contributor to Cl_2 formation over Delhi and Leicester.

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We also calculated ClNO_2 yield from NO_2^+ (Fig. S3), which is the ratio of $P_{\text{ClNO}_2}/L_{\text{total}}$, where P_{ClNO_2} is the rate of ClNO_2 production through $\text{Cl}^- + \text{NO}_2^+$ reaction and L_{total} denotes the loss rate of NO_2^+ through its reactions with Cl^- , H_2O , SO_4^{2-} , HCOO^- , CH_3COO^- , phenol, CH_3OH , and cresol (A4, A10–A16). ClNO_2 yield is ≈ 0.9 over Delhi, indicating the strongest loss of NO_2^+ is through its reaction with Cl^- , which is also mimicked in Fig. S4a showing the same concentrations of ClNO_2 as in NEW simulation and when other NO_2^+ reactions (A10–A16) are turned off (simulation: without other NO_2^+ reactions). ClNO_2 yield over Leicester is between ≈ 0.40 – 0.55 , which is about half the yield in Delhi. Stronger ClNO_2 yield in Delhi could be attributed to ≈ 2 times higher Cl^- than Leicester. Lesser ClNO_2 yield in Leicester portrays the importance of NO_2^+ loss reactions (A10–A15) other than with Cl^- , which could be seen through Fig. S4b where ClNO_2 is increased by more than twice during early morning hours when A10–A15 reactions are kept inactive in the model. The determination of ClNO_2 yield using cavity ring-down spectroscopy and chemical ionization mass spectrometry, shows yield ranging between 0.2 to 0.8 for Cl^- concentrations of 0.02 to 0.5 mol/L (Roberts et al., 2009). The measurements of ClNO_2 yield for coastal and open ocean waters were found to be between 0.16–0.30 which is suppressed by up to 5 times than equivalent salt containing solutions, due to the addition of aromatic organic compounds (e.g., phenol and humic acid) to synthetic seawater matrices (Ryder et al., 2015).

280 4.3 Role of Cl in Atmospheric Oxidative Capacity (AOC)

In order to understand the role of Cl as oxidising agent with respect to the OH radical, we calculated the reactivity of Cl and OH as $\sum X_i (k_{\text{radical}+X_i} \times [X_i])$, where radical is Cl or OH , and $[X_i]$ is the concentration of specie X_i (here X_i includes CO ,

CH₄, primary VOCs and NMHCs which are initialized in the model) (Fig. 5). The corresponding rate constants for Cl + X and OH + X reactions are taken from the MECCA. The reactivity of both Cl and OH decreases rapidly nearly from sunrise to noon time and afterwards increases gradually at both locations. In comparison to Leicester, the magnitudes of Cl and OH reactivity in Delhi are higher by up to ≈ 1.4 and ≈ 12 times, respectively. However, the Cl/OH reactivity ratio in Leicester is up to ≈ 9 times higher than that in Delhi. Cl reactivity is lower (Delhi: ≈ 685 s⁻¹, Leicester: ≈ 553 s⁻¹) during noontime and higher (Delhi: ≈ 750 s⁻¹, Leicester: ≈ 554 s⁻¹) during nighttime and early morning hours at both locations. The OH reactivity follows a similar pattern as that of Cl in Delhi and Leicester. The ratio of Cl to OH reactivity starts increasing after sunrise, reaching a maximum value of ≈ 42 at nearly 16:00 h LT and then decreases further in Delhi. As mentioned above, Cl/OH reactivity ratio in Leicester shows a double peak pattern, with one peak (≈ 270) during early morning $\approx 04:00$ h LT and other peak (≈ 276) at about 16:00 h LT.

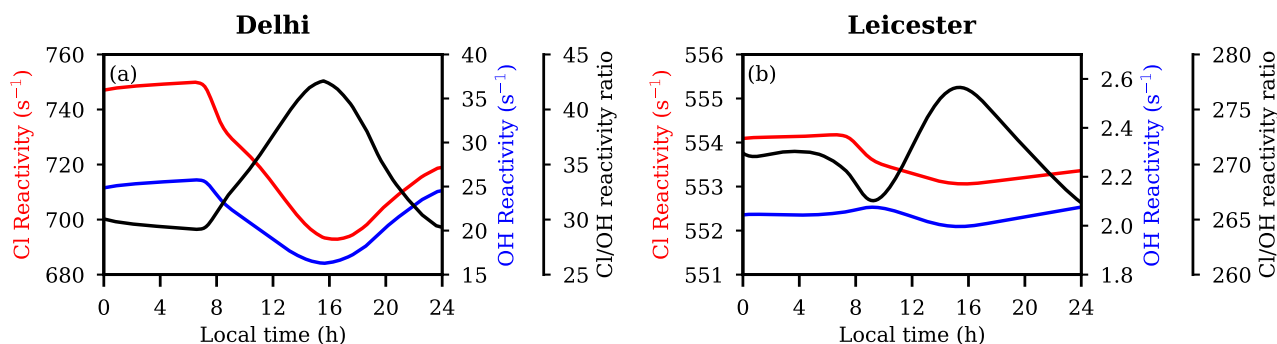


Figure 5. Reactivity of Cl and OH with CO, CH₄, and VOCs, and Cl/OH reactivity ratio during the simulation period in (a) Delhi and (b) Leicester.

We quantified the relative contribution of Cl in atmospheric oxidative capacity (AOC) using the model. AOC represents the sum of oxidation rates of specie X_i by oxidants Y (OH, Cl, and other radicals: NO₃ and O₃) (Elshorbany et al., 2009):

$$\text{AOC} = \sum k_{X_i} [X_i][Y] \quad (1)$$

where, k_{X_i} is the corresponding rate constant for X_i + Y reaction. Accordingly, the magnitude of AOC depends upon the concentration and reactivity of Cl. Figure 6 shows the contribution of individual oxidants in AOC at both locations. Besides OH, Cl is the second most important oxidant in Delhi, with a significant contribution of 23.4 % during morning (averaged over 07:00-09:00 h LT), and 8.2 % throughout the day (06:00-16:00 h LT). In Leicester, Cl is the highest contributor (74.0 %) towards AOC during morning. In fact, with 34.1 % contribution, Cl is major oxidant after OH, during the daytime. Besides the abundance of Cl, higher reactivity enhances the contribution of Cl in AOC, which is further substantiated by the ratio of Cl reactivity to OH reactivity (Fig. 5b). This ratio indicates that Cl reactivity exceeds OH reactivity by a significant margin,

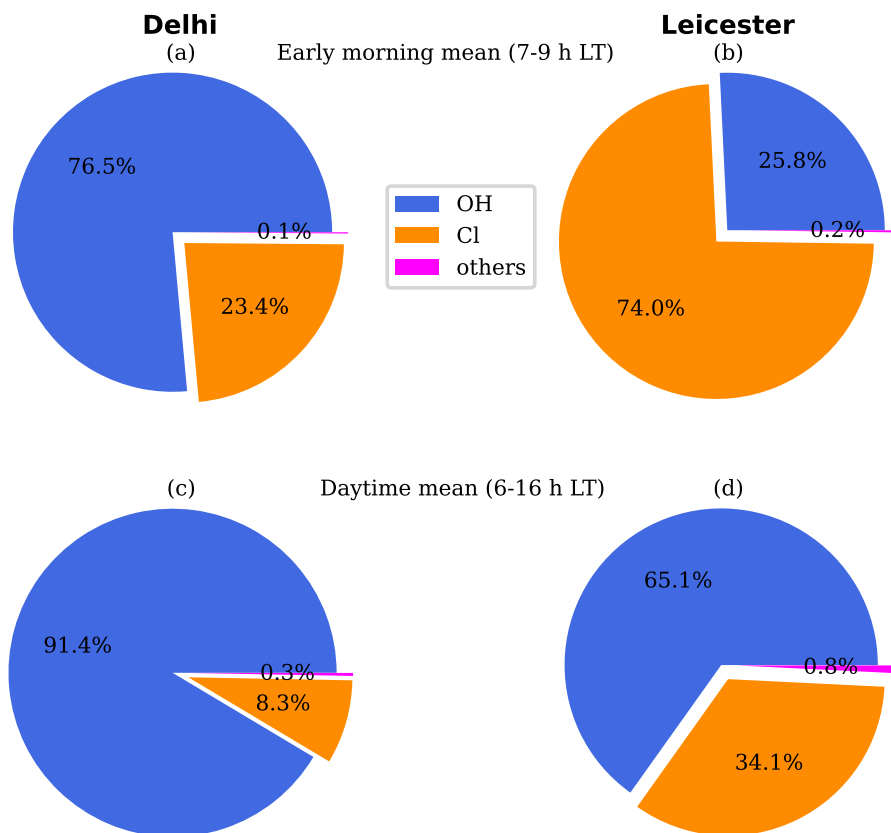


Figure 6. Atmospheric oxidative capacity (AOC) of radicals during (a, b) early morning mean (7-9 h LT) and (c, d) daytime mean (6-16 h LT) in Delhi (left panel) and Leicester (right panel).

ranging from 265 to 276 times greater throughout the day in Leicester. Such a substantial contribution of Cl in AOC leads to enhancements of RO_2 as seen in Fig. 3(f,i). Especially, a prominent peak in RO_2 during early morning (07:00-09:00 h LT) is imparted to strong participation of Cl in atmospheric oxidation during this time. Notably strongest contribution of Cl in AOC during early morning in Leicester, strengthens RO_2 peak by up to a factor of 8 (Fig. 3i). The role of Cl is predominant in Leicester as well as in Delhi during early morning, compared to a polluted environment of Hong Kong, China where Cl contribution was estimated to be 21.5 % (Xue et al., 2015). NO_3 and O_3 were found to play a relatively minor role in AOC at both urban environments.

4.4 Sensitivity to $ClNO_2 + Cl^-$ reaction

In a study conducted by Haskins et al. (2019), using the reacto-diffusive length-scale framework, it was demonstrated that field and laboratory observations could be reconciled by considering an aqueous-phase reaction rate constant for the $ClNO_2 + Cl^-$ reaction on the order of $\approx 10^4 \text{ s}^{-1}$. This reaction rate constant is considerably lower (by ≈ 179 times) than reported in

315 Roberts et al. (2008). In this context, sensitivity simulation (NEWrate) is performed using a reaction rate coefficient of $5.6 \times 10^4 \text{ mol}^{-1} \text{ L s}^{-1}$ (Haskins et al., 2019) for the $\text{ClNO}_2 + \text{Cl}^-$ reaction, for both Delhi and Leicester. As depicted in Figure S7a, the concentration of Cl remains nearly the same in the NEWrate simulation compared to the NEW simulation over Delhi. However, there are significant changes in the concentration of ClNO_2 , as shown in Fig. S7b. The simulated ClNO_2 exhibits a broader peak and is approximately 30 pmol/mol higher near sunrise in the NEWrate simulation when compared to the NEW
320 simulation. During the nighttime, approximately 20 pmol/mol of ClNO_2 is simulated in the NEWrate simulation, whereas it is negligible in the NEW simulation (see Fig. 3b). Since the Cl concentration is almost similar in both the NEW and NEWrate simulations, the differences in the simulated concentrations of OH, HO_2 , and RO_2 remain consistent between the NEWrate or NEW simulations and the OLD and NOCL simulations (refer to Fig. S7d, e, f, and Fig. 3d, e, f). The production and loss mechanisms of Cl are similar in both the NEW and NEWrate simulations (see Fig. S8a and Fig. 4a). The contributions
325 from ClNO_2 formation reactions are also similar. However, in contrast to the NEW simulation, the loss of ClNO_2 through photolysis becomes dominant and is ≈ 6 times greater than its loss through $\text{ClNO}_2 + \text{Cl}^-$ reaction, in NEWrate simulation. The contribution of radicals to AOC is also similar between the NEW and NEWrate simulation, as depicted in Fig. 6a,c and Fig. S9a,c respectively, over Delhi.

330 In contrast to Delhi, significant differences are seen in atmospheric composition in Leicester when the rate coefficient of the $\text{ClNO}_2 + \text{Cl}^-$ reaction is altered (as shown in Fig. S7). The peak concentration of Cl becomes ≈ 0.6 fmol/mol during the morning hours of NEWrate simulation (Fig. S7g), which is about 4 times lower than the concentration of Cl in NEW simulation (Fig. 3g). However, due to slower rate of ClNO_2 consumption with Cl^- , the simulated ClNO_2 using the NEWrate is significantly enhanced (by ≈ 5 times) compared to NEW simulation, reaching a maximum of about 210 pmol/mol around
335 sunrise (see Fig. S7h). Due to lower Cl concentrations, the levels of ClONO also decrease by 3.5 times in NEWrate simulation (as shown in Fig. S7i) compared to NEW simulation (Fig. 3i). The dominant peak seen at sunrise in the NEW simulation for OH, HO_2 , and RO_2 is significantly reduced with the lower rate of the $\text{ClNO}_2 + \text{Cl}^-$ reaction, as illustrated in Fig S7j,k,l. Significant changes in the production and loss mechanisms of Cl and ClNO_2 are seen in Leicester when the reaction rate of A6 is changed, as shown in Fig. S8 and Fig. 4b. For example, in the NEWrate simulation, other reactions, including the
340 photolysis of ClNO_2 and ClONO, and $\text{ClO} + \text{NO}$ reaction, become prominent sources of Cl (with a rate of approximately $6.0 \times 10^5 \text{ molec cm}^{-3} \text{ s}^{-1}$), whereas in the NEW simulation, the major source for Cl is photolysis of Cl_2 . The primary source for ClNO_2 production remains the $\text{Cl}^- + \text{NO}_2^+$ reaction in both the NEW and NEWrate simulations. However, in the NEWrate simulation, ClNO_2 loss from photolysis becomes the major sink, whereas in the NEW simulation, loss from the $\text{ClNO}_2 + \text{Cl}^-$ reaction is prominent. In addition, remarkable changes in AOC are seen between the NEWrate (Fig. S9b, d) and the
345 NEW simulation (Fig. 6b,d). In the NEWrate simulation, even though Cl remains the major oxidant its contribution is notably reduced from 74% (in NEW simulation) to 58.1% during the early morning hours.

5 Summary and Conclusions

Extended gas- and aqueous-phase chemistry of chlorine compounds has been added to the MECCA mechanism. It consists of 36 gas-phase reactions (inorganic, organic, and photolysis reactions). A total of 24 aqueous-phase and heterogeneous reactions have been added, containing detailed chemistry of N_2O_5 uptake on aerosols to yield ClNO_2 and various other competing reactions. The updated model is applied to two different urban environments: Delhi (India) and Leicester (United Kingdom) during winter time. The major conclusions are:

1. The model predicts up to 0.1 pmol/mol of NO_3 and up to 8 pmol/mol of N_2O_5 during daytime in Delhi. However, night-time production of NO_3 and N_2O_5 is seen to be negligible primarily due of the unavailability of O_3 . In contrast to Delhi, NO_3 and N_2O_5 after mid-night in Leicester is ≈ 2.6 pmol/mol and ≈ 330 pmol/mol, respectively. N_2O_5 uptake on aerosols yields ClNO_2 , which produces Cl via photolysis.
2. A sharp build-up of Cl with sunrise is mainly through Cl_2 photolysis in Delhi. Besides Cl_2 , photolysis of ClNO_2 and ClONO and the reaction of ClO with NO are prominent Cl sources in Leicester. VOCs are the main sink for Cl at both locations, whereas NO_2 is also an important sink for Cl in Leicester. The latter results in the formation of ClNO_2 with a major contribution in Delhi, while $\text{Cl}^- + \text{NO}_2^+$ is a stronger source in Leicester. Photolysis is the major sink for ClNO_2 in Delhi, however, its uptake on chloride aerosols is a prominent sink in Leicester.
3. The magnitude of Cl ($\approx 750 \text{ s}^{-1}$) and OH ($\approx 25 \text{ s}^{-1}$) reactivities are significantly greater in Delhi, particularly during the morning hours, when compared to Leicester. However, Cl to OH reactivity ratio (≈ 270) is pronounced in Leicester coinciding with higher contribution of Cl in AOC.
4. Sensitivity simulations reveal substantial post-sunrise enhancements of in OH, HO_2 , and RO_2 radicals, with a prominent secondary peak due to Cl chemistry. Up to 8 times higher RO_2 is simulated in Leicester primarily because of leading role of Cl in AOC potential.

It is important to note that box models, despite their general limitation of neglecting transport phenomena and assuming species to be well mixed, do include highly detailed chemical mechanisms. Furthermore, because the model is initialized with measurements of chemical species at both locations and the modeled levels align with observed data, significant discrepancies in model estimates would be unexpected. Future studies focussing on modeling vertical gradients, in particular for radical reservoir species such as HONO, and ClNO_2 (Young et al., 2012) are recommended.

This study highlights the vital role of Cl chemistry in governing the oxidation capacity of the atmosphere and air quality, and therefore it is important to account for it in detailed photochemical as well as in 3-D chemical transport models. This will lead to better quantification of the importance of radicals in atmospheric oxidation and hence, the formation of ozone as well as secondary aerosols, over regional to global scale. Future studies focusing on secondary aerosol formation and new particle formation from heterogeneous reactions are needed to deepen the understanding of transformation of trace gases to aerosols.

Code and data availability. CAABA/MECCA is a community box model published under the GNU General Public Licence, available from
380 the Gitlab repository (<https://gitlab.com/RolfSander/caaba-mecca>). The version of CAABA/MECCA updated in this study is currently available in the 'delhi' branch of the repository. The new chlorine mechanism will be included in the next release of CAABA/MECCA. All the model outputs associated with this study are archived at zenodo (<https://zenodo.org/record/8332131>; Soni et al. (2023)).

Author contributions. M. Soni, R. Sander, and D. Taraborrelli designed the study with inputs from S. S. Gunthe, P. Liu, and N. Ojha. M. Soni, R. Sander, and D. Taraborrelli developed and analyzed the chemical mechanism and M. Soni performed the simulations. A. Pozzer,
385 R. Sander, L. K. Sahu, D. Taraborrelli, I. A. Girach, and N. Ojha helped M. Soni in the analyses and interpretations of the results. A. Patel assisted M. Soni in compiling literature and some input dataset. M. Soni wrote the manuscript and all the co-authors contributed to the review and editing.

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