Combining lime and organic amendments based on titratable alkalinity for efficient amelioration of acidic soils

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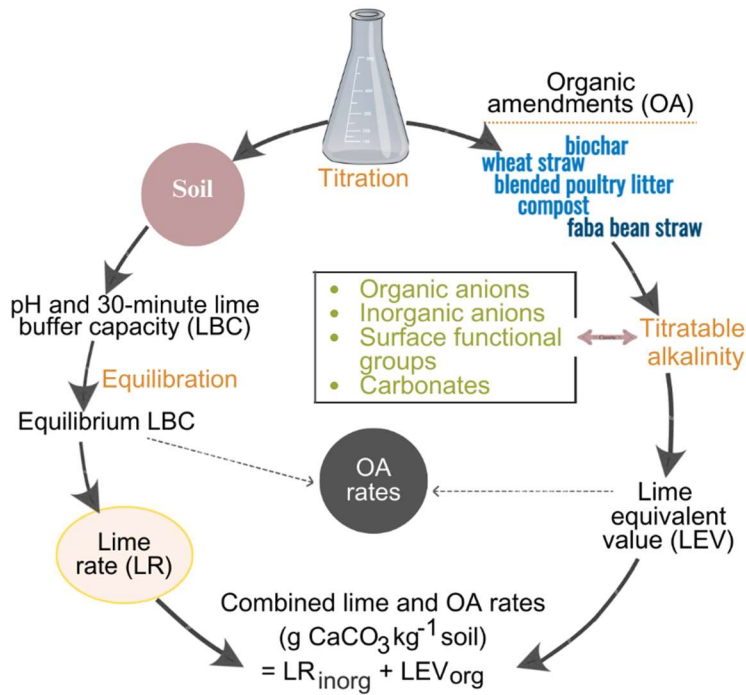
Abstract. Ameliorating soil acidity using a combination of lime and organic amendments (OAs) can be an alternative to lime alone, but determining the appropriate OA rates can be difficult. We developed a new method for calculating the combined application rate of lime and OAs (wheat straw, faba bean straw, blended poultry litter, biochar, and compost), based on the

10 titratable alkalinity of OA and the equilibrium lime buffer capacity (LBC_{eq}) of acidic soils. The effect of calculated soil amendment rates on soil pH was validated at soil water contents of 60, 100, and 150% field capacity (FC). The soil used to develop and validate the method was a sandy loam with pH_W 4.84 and pH_{Ca} 4.21. The LBC_{eq} of the soil was 1657 mg $CaCO_3$ kg^{-1} pH^{-1}. The titratable alkalinity of the OAs ranged from 11.7 cmol H^+_{eq} kg^{-1} for wheat straw to 357 cmol H^+_{eq} kg^{-1} for compost. At 60% FC, faba bean and wheat straw amendment increased soil pH_W to 6.48 and 6.42, respectively, but lower pH

15 values were reached in soil amended with less biodegradable or resistant OAs (ROAs) (i.e. blended poultry litter, biochar, and compost). At 150% FC, the two straws increased soil pH_W to only 5.93 and 5.75, respectively, possibly due to slower decomposition in submerged conditions, resulting in limited alkalinity production, whereas amendment with ROAs produced pH_W values close to 6.5. With increasing lime equivalent value (LEV) of the OA, from 5.8 g $CaCO_3$ kg^{-1} (wheat straw) to 179 g $CaCO_3$ kg^{-1} (compost), the lime requirement to reach pHw 6.5 in lime-OA combinations decreased from 2.72 to 0.09 g

20 $CaCO_3$ kg^{-1}. The developed method was shown to be effective in determining appropriate rates of OAs (with or without additional lime) for management of acidic sandy loam soils in this study and highlights the importance of soil water content for their acid neutralising effect.

Keywords: Lime requirement, soil amendments, alkalinity production, soil pH, lime buffer capacity, soil water content

1 Introduction

Acidic soils with $pH_W \leq 5.5$ are considered harmful to sensitive plants. Soil acidification can be caused by various biogeochemical processes such as oxidation and acid-dissolution reactions, root exudates, solubilisation and hydrolysis of Al^{3+} which releases H^+, and leaching of cations (K^+, Ca^{2+}, Na^+, Mg^{2+}) (Brady and Weil, 2016; Goulding, 2016; Marschner and Noble, 2000; Mosley et al., 2014). Soil acidification in agricultural systems can also be attributed to high inputs of urea and ammonium (NH_4^+)-containing fertilisers where protons (H^+) are generated during nitrification of NH_4^+ to nitrate (NO_3^-) and by NH_4^+ uptake by plants as well as by removal of alkalinity by removing plant material at harvest (Bolan et al., 2003; Hume et al., 2022; Iticha and Takele, 2019; Kunhikrishnan et al., 2016). Even a small decrease in soil pH can have a large impact on nutrient availability, nitrogen fixation by soil microbes, and sustainability of crop production (Kopittke et al., 2019).

Lime and, more recently, OAs can be used to improve the productivity of acidic soils (Garbowski et al., 2023). Agricultural lime ($CaCO_3$) produces alkalinity (OH^- and CO_3^{2-}/HCO_3^- ions) and neutralises soil acidity by reacting with Al^{3+} and H^+ ions to form $Al(OH)_3$ precipitates and H_2O. Organic amendments on the other hand can provide alkalinity via their surface organic functional groups (e.g. carboxylic, $-COO^-$), presence of solid and dissolved carbonates, and by release of organic and inorganic anions, which can neutralise or remove toxic ions such as Al^{3+}, Fe^{3+}, Mn^{2+}, and H^+ from the soil by reacting, complexing, and binding (Brown et al., 2008; Larney and Angers, 2012; McCauley et al., 2009). Alkalinity production by OAs is related to the excess cation content and dissolution or decomposition rate of the amendments, which are in turn influenced by factors such

as the chemical composition of the amendments and soil water content (Anderson et al., 2020; Cai et al., 2020; Védère et al., 2022). As a result, it is important to investigate the effect of soil water content on the timing of OA application.

Determination of the optimum rates of lime and OAs is necessary for efficient management of acidic soils. Lime requirement can be determined using various methods including titration, incubation, field experiments, standard buffer solutions, and predictive equations developed from measured soil properties (Aitken and Moody, 1994; Nelson and Su, 2010). Titration is considered the most effective approach for lime recommendations (Wang et al., 2015). It involves the addition of bases such as $Ca(OH)_2$ to acidic soils in a given soil to liquid ratio (e.g. 1:5 soil to deionized water or 0.01 M $CaCl_2$ solution) followed by measurement of changes in soil pH and pH buffering capacity (pHBC) with time to predict lime requirement of the soils (Barouchas et al., 2013; Liu et al., 2004; Thompson et al., 2010). Titration results are often validated by experiments to determine the equilibrium lime buffer capacity (LBC_{eq}), i.e. measure of the amount of soil acidity that must be neutralized by addition of lime to raise soil pH by one unit in order to calculate lime requirements of acidic soils (Kissel et al., 2012).

The amounts of OAs required to neutralise soil acidity and achieve a desired pH can be calculated using the alkalinity of the amendments. Alkalinity is determined by titrating acidified OAs with a base to neutrality, using either ashed amendments, i.e., potential or 'ash' alkalinity (Slattery et al., 1991), or air-dried amendments, i.e. titratable or available alkalinity (Feizi et al., 2017; Fidel et al., 2017; Singh et al., 2017). The ash alkalinity method overestimates alkalinity or liming potential of OAs due to the loss of anions such as sulfur and chlorine during the ashing process (Noble et al., 1996), leading to underestimation of the amounts of amendments required to neutralise soil acidity. The short-term acid neutralising effect of OAs is better determined by available alkalinity, because potential alkalinity becomes only gradually available over a long period of time (Sakala et al., 2004).

Previous studies have shown that combination of lime and OAs can generate more alkalinity than lime or OA alone (Butterly et al., 2021; Lauricella et al., 2021; Li et al., 2022). However, little effort has been made to develop methods for calculating the application rates of lime combined with OAs needed to neutralise soil acidity and achieve the desired pH for plant growth. Previous research aimed at developing appropriate lime and OA combinations has been conducted in field trials, which involve mixing different rates of lime and OAs into acidic soils and determining the response of acidity and crop yield (Celestina et al., 2018; Khoi et al., 2010; Li et al., 2019). This approach is time consuming and costly, and it is difficult to predict the soil acidity that can be neutralised. In addition, little information exists about the impact of an alkalinity-based mixture of OAs on soil pH. A laboratory method for determining lime and OAs combinations based on titratable alkalinity could shorten the time required and improve the cost-effectiveness of soil acidity amelioration.

Therefore, the first aim of this study was to determine the lime requirement of acidic soil using titration and equilibration methods as well as estimate the application rates of OAs and OAs mixes based on their titratable alkalinity. The second aim was to develop a laboratory method for calculating the rates of lime and OA in combinations required to achieve a desired soil pH, based on titratable alkalinity of the amendments and LBC_{eq} of the acidic soil. Finally, the third aim was to assess the effect of soil water content on amelioration by incubating acidic soil with the amendments at different soil water contents.

2 Materials and Methods

2.1 Collection of soil and amendments

The soil used in this study was collected from the 0-10 cm layer of the non-limed treatment from the acid soil management trial site at Sandilands on the Yorke Peninsula of South Australia. It is located in the North-West of Adelaide at $34^0 33 ` 14 " S$ latitude and $137^0 42 ` 14 " E$ longitude. The long-term mean annual rainfall and temperature of the site were 409 mm and 21.9 ^OC. Penlime Plus™ (Angaston lime, Penrice Quarry & Mineral, South Australia) with a neutralising value of 98% was used as the lime source. The OAs used in this study were wheat straw, faba bean straw, blended poultry litter, biochar, and compost, which differ in their decomposability. The two straws are more decomposable than the other OAs because they haven't undergone decomposition. Due to its lower C/N ratio, faba bean straw is likely to be more readily decomposable than mature wheat straw. The blended poultry litter comprised of equal proportions of poultry manure and sawdust/wood shavings. The biochar was made from pyrolysis (400°C) of Eucalyptus species (sourced from Green Man Char Pty Ltd, Victoria, Australia), the compost was prepared from a mixture of various organic wastes (sourced from Bunnings Pty Ltd, Australia). Blended poultry litter, biochar and compost are partially decomposed, leaving little readily available compounds for decomposition after addition to soil and are therefore referred to as poorly decomposable or resistant amendments.

The soil and OAs were dried at 30 ^0C in a fan-forced oven. Then the oven-dried soil was crushed, sieved to pass through a 2 mm sieve size, the dried OAs were ground and sieved to 0.25-2 mm, and then used in the experiment.

2.2 Analysis of soil and amendments

Soil particle size distribution was determined using the hydrometer method (Bouyoucos, 1962). To measure gravimetric water content at field capacity, air-dried soil was weighed into small container with 10 cm diameter and 10 cm height, watered to saturation, covered at the top with plastic film to prevent evaporation, and drained to a constant mass for 2 to 3 days, and reweighed. Soil pH was measured potentiometrically using a calibrated glass electrode in a soil/water (1:5) suspension (pH_w) and a soil/0.01 M $CaCl_2$ (1:5) suspension (pH_{Ca}) (Rayment and Higginson, 1992). The pH and electrical conductivity (EC) of OAs were measured in a 1:10 amendment to water ratio (Singh et al., 2017). Exchangeable acidity was determined by extracting soil with 1 M KCl solution (Pansu and Gautheyrou, 2006). The cation exchange capacity (CEC) of the soil was determined using a colorimetric method after displacing cations with 1 M ammonium acetate (pH 7) and then extracting the ammonium ions with 1 M KCl (Carter and Gregorich, 2007). Total organic carbon (TOC) in the soil and OAs was determined using the Walkley and Black wet digestion method (Walkley and Black, 1934), whereas total N was determined using the Kjeldahl method (Bremner and Mulvaney, 1982). After digestion with concentrated nitric acid and hydrogen peroxide, the total concentrations of Fe and Al in soil as well as Ca, K, Mg, Na, P, and S in OAs were determined using inductively coupled plasma optical emission spectroscopy (ICP-OES). The excess cation content in the OAs was then calculated by subtracting anions (SO_4^{2-} and $H_2PO_4^-$) from cations (Ca^{2+}, Mg^{2+}, K^+, and Na^+) (Tang and Yu, 1999).

2.3 Properties of soil and amendments

In the Australian Soil Classification (Isbell and NCST, 2021), the soil was categorized as a Chromosol, it is classified as Lixisol in WRB (FAO, 2015). The upper A horizon which has a sandy loam texture was used in this experiment (Table 1). The soil pH_W and pH_{Ca} measured in 1:5 soil to water/0.01 M $CaCl_2$ solution ratios are 4.84 and 4.21, respectively. The exchangeable acidity of the soil was 2.95 $cmol_c$ kg^{-1}. Based on the ratings of Hazelton and Murphy (2016), the soil has a moderate CEC (21.38 $cmol_c$ kg^{-1}), likely due to its total OC content (1.43%) (Table 1).

Table 1 Basic properties of the soil used in this study (mean ±SD).

Parameters	Mean
Clay (%)	$14.41^{\pm1.16}$
Silt (%)	$10.67^{\pm2.31}$
Sand (%)	$74.92^{\pm2}$
pH_W	$4.84^{\pm0.02}$
pH_{Ca}	$4.21^{\pm0.01}$
EA ($cmol_c$ kg^{-1})	$2.95^{\pm0.14}$
CEC ($cmol_c$ kg^{-1})	$21.21^{\pm1.62}$
TOC (%)	$1.43^{\pm0.04}$
Total Al^{3+} (g kg^{-1})	$3.42^{\pm0.44}$
Total Fe^{2+} (g kg^{-1})	$6.78^{\pm0.25}$

SD: standard deviation, pH_W: soil pH in deionised water, pH_{Ca}: soil pH in 0.01 M $CaCl_2$ solution, EA: exchangeable acidity, CEC: cation exchange capacity, TOC: total organic carbon.

The mean pH_W of the OAs varied between 5.50 (wheat straw) and 9.75 (biochar) (Table 2). EC of the amendments ranged from 0.55 dS m^{-1} for faba bean straw to 3.22 dS m^{-1} for compost. Total concentrations of cations was highest in compost (383 cmol kg^{-1}) and lowest in wheat straw (53 cmol kg^{-1}) (Table 2).

Table 2 Chemical properties of organic amendments used in this study ($n = 3$). Different letters indicate significant differences ($p \leq 0.05$).

Organic amendment	pH_W (1:10)	EC (dS m^{-1}) (1:10)	Na	K	Ca	Mg	P	S	TOC (g kg^{-1})	TN (g kg^{-1})	C:N
			cmol kg^{-1}								
Wheat straw	5.50^a	1.87^a	15.62^{ad}	28.89^a	11.28^a	6.68^a	2.65^a	6.37^a	453^a	7.30^a	63.15^a
Faba bean straw	6.96^b	0.55^b	3.12^b	7.17^b	69.46^b	13.52^{ac}	17.27^b	10.49^b	429^b	12.88^b	33.40^b
Blended poultry litter	7.88^c	1.39^c	12.20^a	15.49^c	168.95^c	41.33^b	24^c	27^c	256^c	17.09^c	15.09^{cd}
Biochar	9.75^d	0.91^d	25.03^c	15.74^c	108.82^d	22.16^c	10.82^d	6.98^a	180^d	9.63^a	18.96^c

| Compost | 8.12[e] | 3.22[e] | 17.24[d] | 29.94[a] | 87.67[e] | 327.15[d] | 41.49[e] | 37.12[d] | 160[d] | 20.18[c] | 7.95[d] |

EC: electrical conductivity, TOC: total organic carbon, TN: total nitrogen, C:N: carbon to nitrogen ratio

2.4 Determination of lime requirement of acidic soil

2.4.1 Titration with calcium hydroxide

Titrations and pH measurements in this study were carried out in 1:5 soil to deionised water or 0.01 M $CaCl_2$ solution ratios. First, 5 g of dry soil was weighed to 50 mL polyethylene tubes in triplicate and then 25 mL of deionised water or 25 mL of 0.01 M $CaCl_2$ solution were added into the tubes. The initial soil pH was measured after 30 minutes stirring and calibration of the pH meter with standard pH 7.0 and 4.0 buffers. Then, 0.5 mL aliquots of 0.022 M $Ca(OH)_2$ were added to the suspensions, continuously stirred, and pH measurements were taken at the end of each time interval. Based on Liu et al. (2004), the reaction time allowed between consecutive titrations to obtain constant pH measurements was 30 minutes. The titrations with incremental additions of 0.022 M $Ca(OH)_2$ were carried out while recording the cumulative volume of the 0.022 M $Ca(OH)_2$ added versus the corresponding soil pH. A digital titrator (Burette Digital Titrette® Bottle-top Burette 50 mL Capacity Class A, Australian Scientific, Australia) was used for the titrations.

This data was used to plot a regression curve between incremental rates of equivalent $CaCO_3$ (Mg ha^{-1}) consumed and the corresponding soil pH. The regression curve was used to derive the slope, which was then used to calculate the 30-minute pHBC (pHBC$_{30}$) of the acidic soil. The amount of base needed to neutralise H^+ (independent variable) was on the x-axis and the change in pH (dependent variable) was on the y-axis to generate the slope of the titration curve as:

$$\text{Slope} = \frac{\Delta Y}{\Delta X} = \frac{\Delta pH}{V_{Ca(OH)_2}} = \frac{\Delta pH}{\Delta H^+} \tag{1}$$

where ΔY is the change in pH, ΔX is the amount of base consumed or protons removed to neutralize acids, $V_{Ca(OH)2}$ is the volume of base consumed during titration (mL), ΔH^+ is the mmol kg^{-1} of H^+ removed during titration. The amount of protons (ΔH^+) neutralised during the titration is equivalent to the amount of $Ca(OH)_2$ consumed. From this model, the LR of the acidic soil can be calculated as:

$$\text{LR} = \frac{\Delta pH}{\text{Slope}} = \Delta pH \times pHBC \tag{2}$$

Given that the molecular weight of $CaCO_3$ is 100 g per mol, the amount of $Ca(OH)_2$ consumed by the acidic soil expressed as $CaCO_3$ equivalent (Mg ha^{-1}) was calculated as:

$$\text{CaCO}_3 \text{ (Mg ha}^{-1}) = \frac{V \times M \times 100 \text{ g CaCO}_3 \text{ mol}^{-1} \times 10^{-9} \times \text{Wt. (kg ha}^{-1})}{S \text{ (kg)}} \tag{3}$$

Where V is the volume (mL) of 0.022 M $Ca(OH)_2$ consumed to raise the initial pH of the acidic soil to the target pH, M = 0.022 is the molarity of $Ca(OH)_2$ in mmol mL^{-1}, 10^{-9} is the conversion factor from $CaCO_3$ (mg mmol^{-1}) to $CaCO_3$ (Mg mmol[-]

[1]), S is the mass of soil used for the titration (0.005 kg), and Wt is the weight of soil per hectare which in this case was calculated as 1.5×10^6 kg, assuming a liming depth of 0.1 m and soil bulk density of 1500 kg m^{-3}. By using these known variables in eqn.3, CaCO$_3$ (Mg ha^{-1}) equivalent can be simply calculated as V x 0.66.

The pH buffer capacity (pHBC$_{30}$), expressed in mmol H$^+$ (kg soil)$^{-1}$ pH^{-1}, was calculated from the titration curve as the inverse of the slope of the linear regression between pH and the added base (Mg CaCO$_3$ ha^{-1}) (Shi et al., 2019; Thompson et al., 2010). The unit of pHBC derived from the slope was Mg CaCO$_3$ ha^{-1} pH^{-1}. This unit was converted to mg CaCO$_3$ kg^{-1} pH^{-1} and then to mmol H$^+$ kg^{-1} pH^{-1}.

2.4.2 Equilibration experiment

An additional experiment was conducted to assess whether the 30-minute titration time above was sufficient to complete the equilibrium exchange reaction between H$^+$ and 0.022 M Ca(OH)$_2$. Since there was no significant difference between the slopes of regression lines fitted for titrations in water and 0.01 M CaCl$_2$ (p value of 0.231 at $\alpha = 0.05$), the equilibration experiment was carried out only in a 1:5 soil to deionised water ratio. The experiment consisted of unamended soil (control) and soils amended with the equivalent titration point (ETP) multiplied by 0.5, 0.75, 1, 1.25, and 1.5. The ETP is the final volume of 0.022 M Ca(OH)$_2$ solution consumed over a 30-minute complete titration to achieve the target pH$_W$ of 6.5. These treatments were added in 50 mL polypropylene tubes with replicates of 5 g dried soil and 25 mL deionised water suspension. After stirring the mixture for 30 minutes, the initial 30-minute soil pH was measured, which was used to calculate the lime buffer capacity (LBC$_{30}$) for each rate of 0.022 M Ca(OH)$_2$. Then, three drops of chloroform were added to minimise microbial activity. The tubes were covered with parafilm with only small opening left for air exchange to reduce evaporation, stirred regularly, and incubated at room temperature. The pH was measured every 24 h for 5 days while stirring the suspensions. The lime buffer capacity (mg CaCO$_3$ kg^{-1} pH^{-1}) was calculated for each incubation period (Kissel et al., 2007):

$$\text{LBC} = \frac{V \times M \times M_W}{S \times (pH_e - pH_o)} \tag{4}$$

Where V is the volume of Ca(OH)$_2$ added (mL), M is the molarity of Ca(OH)$_2$ used for the titration (= 0.022 M), M$_W$ is the molecular weight of CaCO$_3$ (100 mg mmol^{-1}), S is the weight of soil titrated (kg), pH$_e$ is the pH of the suspension taken after addition of certain volume of Ca(OH)$_2$ and equilibration for specific incubation time, and pH$_o$ is the pH of a suspension without Ca(OH)$_2$ taken at similar incubation time.

Total acidity expressed as proton (H$^+$) concentrations that was neutralised over the incubation periods to finally attain equilibrium pH was calculated using eqn. 5, which was derived from ΔpH x pHBC, by substituting 1/slope as pHBC in eqn.1.

$$H^+ = (LBC \times \Delta pH \times 2)/100 \tag{5}$$

Where H$^+$ is the proton concentration (mmol H$^+$ kg^{-1} soil) that was neutralised at a given incubation time (t$_i$), LBC is the lime buffer capacity (mg CaCO$_3$ kg^{-1} pH^{-1}) calculated for the incubation time (t$_i$), ΔpH is the difference between initial soil pH taken

after 30-minutes of shaking the soil with base and the final soil pH taken at the incubation time (t_i), 100 is the conversion factor from mg $CaCO_3$ kg^{-1} pH^{-1} to mmol $CaCO_3$ kg^{-1} pH^{-1}, and 2 is the conversion factor from mmol $CaCO_3$ to mmol H^+.

2.4.3 Equilibrium buffer curves and lime rates

Nonlinear regression curves between soil pH and incubation time as well as between LBC and incubation time were plotted for all concentrations of 0.022 M $Ca(OH)_2$ added to evaluate the change in soil pH and LBC over time. The curves were used to determine the equilibrium pH and LBC at equilibrium (LBC_{eq}, i.e. point where the pH did not change significantly with time). In addition, a linear regression was fitted between LBC_{30} and LBC_{eq}, with different rates of 0.022 M $Ca(OH)_2$ as covariate. The regression equation between LBC_{30} and LBC_{eq} was used to calculate equilibrium lime buffer capacity (LBC_{eq}) for the acidic soil based on the pHBC determined during the 30-minute titration, LBC_{30}. After the determination of LBC_{eq}, the lime requirement (LR) of the acidic soil was calculated as:

$$LR \text{ (mg CaCO}_3 \text{ kg}^{-1} \text{ soil)} = (LBC_{eq} \text{ x } (pH_t - pH_i))/ENV \tag{6}$$

Where LBC_{eq} is the lime buffer capacity at equilibrium soil pH (mg $CaCO_3$ kg^{-1} pH^{-1}), pH_i is the initial pH_W before addition of $Ca(OH)_2$, pH_t is the target pH_W (i.e. 6.5 used in this study), and ENV is the effective neutralising value (ENV) of the lime used in percent.

2.5 Titratable alkalinity and application rate of organic amendments

The titratable alkalinity was determined by extracting the dried OAs with acid and then back titrating the suspension to pH 7 with base using the modified methods of Singh et al. (2017) and Yuan and Xu (2011). There was no washing step in this method. The acid-neutralising value of all the compounds released by acid pre-treatment, including any labile compounds, was estimated by back-titrating the solution. Briefly, 0.5 g of dried and sieved (0.25-2 mm) OAs were weighed into 50 mL polyethylene tubes in replicates. Then, 10 mL of 1 M HCl was added, and the mixture was shaken on a reciprocal shaker for 2 h. The suspension was allowed to stand for 24 h and then titrated with 0.5 M NaOH using a digital titrator. The amount of acid neutralized by each amendment, expressed as mmol H^+ per gram OA, was calculated as the difference between the volume of 0.5 M NaOH consumed by the blank and the sample. The lime equivalent value (LEV), expressed as $CaCO_3$ equivalents, of the OAs was then calculated as:

$$LEV_{OA} = \frac{(0.5)*(b-s)*M_w}{2*W*1000} \tag{7}$$

Where LEV is the lime equivalent value (g $CaCO_3$ kg^{-1} OA), 0.5 is the molarity of NaOH used for titration (mole L^{-1}), b is the volume of NaOH consumed by the blank (mL), s is the volume of NaOH consumed by the sample (mL), M_w is the molecular weight of $CaCO_3$ (g $mole^{-1}$), 2 is the moles of H^+ neutralised by one mole of $CaCO_3$, W is the weight of sample (kg), and 1000

is the unit conversion factor. Then, the amounts of OAs required to raise soil pH to a desired level (in this study pH_W 6.5) were calculated by:

$$RAR_{OA} \text{ (g OA kg}^{-1} \text{ soil)} = (LBC_{eq} * (pH_t - pH_i) * 1000)/LEV_{OA} \tag{8}$$

Where RAR_{OA} is the recommended application rate of OAs, LBC_{eq} is the lime buffer capacity at equilibrium pH (mg $CaCO_3$ kg^{-1} soil pH^{-1}), pH_t is the target soil pH, pH_i is the initial soil pH, and 1000 is the unit conversion factor.

2.6 Calculating rates of lime and organic amendment in their combinations

The application rates in the combinations of pure lime and OAs were calculated based on the titratable alkalinity of the OAs and the $CaCO_3$ required to raise soil pH_W to 6.5 that depends on LBC_{eq}. The method used to calculate lime-OA combinations is illustrated in Fig. 1.

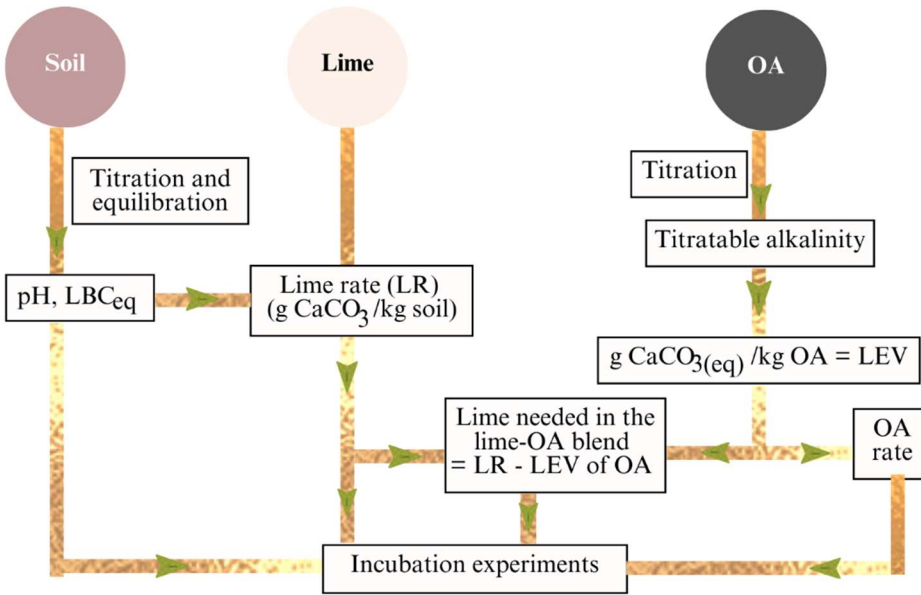



Figure 1. Conceptual method used to calculate the amount of lime, organic amendments (OA), and lime-OA combinations. LBC_{eq}: equilibrium lime buffer capacity, LR: lime rate required to achieve pH_W 6.5, LEV: lime equivalent value.

The amount of pure lime that was combined with each OA to achieve a target pH_W of 6.5 was calculated by subtracting the LEV or $CaCO_3$ equivalent of the OA (i.e. added at 1.5% w/w in this study) from the lime rate using eqn. 9.

$$LOA_{comb} \text{ (g } CaCO_3 \text{ kg}^{-1} \text{ soil)} = LEV_{OA} + LR_{inorg} \tag{9}$$

The LOA$_{comb}$ is the total CaCO$_3$ required in the lime-OA combinations to raise pH$_W$ to 6.5, LEV$_{OA}$ is the lime equivalent value or CaCO$_3$ content of each OA (i.e. added at 15 g OA kg^{-1} in this study), and LR$_{inorg}$ is the amount of lime (i.e. inorganic source of CaCO$_3$) needed in the combinations (g CaCO$_3$ kg^{-1} soil).

2.7 Calculating organic amendment in mixes

A 50:50% alkalinity-based mixture of selected OAs was prepared by combining rapidly decomposable organic material (wheat straw, faba bean straw) with resistant organic materials (compost, biochar) to test the effect of the mixture on soil pH over time. The amount of each amendment in the two OAs mix was calculated in order to contribute 50% of the total alkalinity (i.e. 50% of 2.81 g CaCO$_3$ kg^{-1}) required to neutralise soil acidity and achieve pH$_W$ 6.5.

2.8 Validation experiments at different water contents

Incubation experiments were conducted at 60, 100, and 150% field capacity (FC) to validate whether the amounts of soil amendments calculated based on their titratable alkalinity could raise soil pH to the desired level (pH$_W$ 6.5). The treatments included unamended or control soil, lime, OAs, OAs mix, and lime-OA combinations. The amendment rates calculated based on the method developed in this study are presented in Supplementary Table ST1.

For the incubations at 60 and 100% FC, the amendments were mixed with 50 g soil (< 2 mm) in pots in replicates. The pots were incubated in a dark room with deionized water added on a weight basis. The pH was measured after 14, 30, 60, 90, and 120 days for incubation at 60% FC and after 10, 20, 30, 40, and 50 days for the 100% FC until equilibrium pH was attained. For incubation at 150% FC, 5 g air-dry soil (< 2 mm) was weighed into 50 mL polyethylene plastic tubes and the amendments were mixed with the soil.

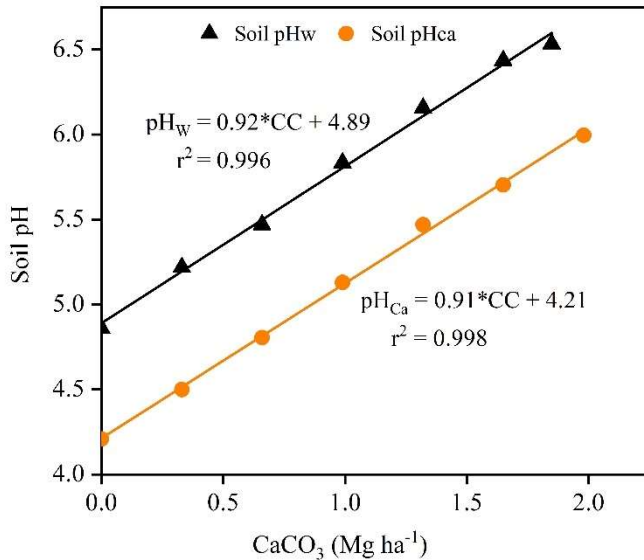
The treatments were incubated in a dark room at 150% field capacity with moisture content maintained on a weight basis. Soil pH was measured at 4, 6, 8, 10, and 12 days by adding the amount of deionized water needed to make a 1:5 soil to water ratio and shaking the suspension for 45-minutes.

2.9 Statistical analysis

Regression curves for the titration points and equilibration experiments were fitted using OriginPro Lab version 2022 (9.95). Mean comparisons of titratable alkalinity and other chemical properties of OAs were carried out using One-Way ANOVA. Mean comparisons of pH values measured over time during titration and equilibration experiments, as well as the rate of H$^+$ neutralisation with incubation time, were undertaken using One-Way repeated measures ANOVA. Changes in pH values of amended soils over time for validation experiments conducted at different soil water contents were also analysed using One-Way repeated measures ANOVA. The mean comparisons, correlation analysis, and tests of significance were conducted at p $\leq$ 0.05 using IBM-SPSS version 28.0.1.0.

3 Results

3.1 Soil pHBC$_{30}$ and LBC$_{30}$

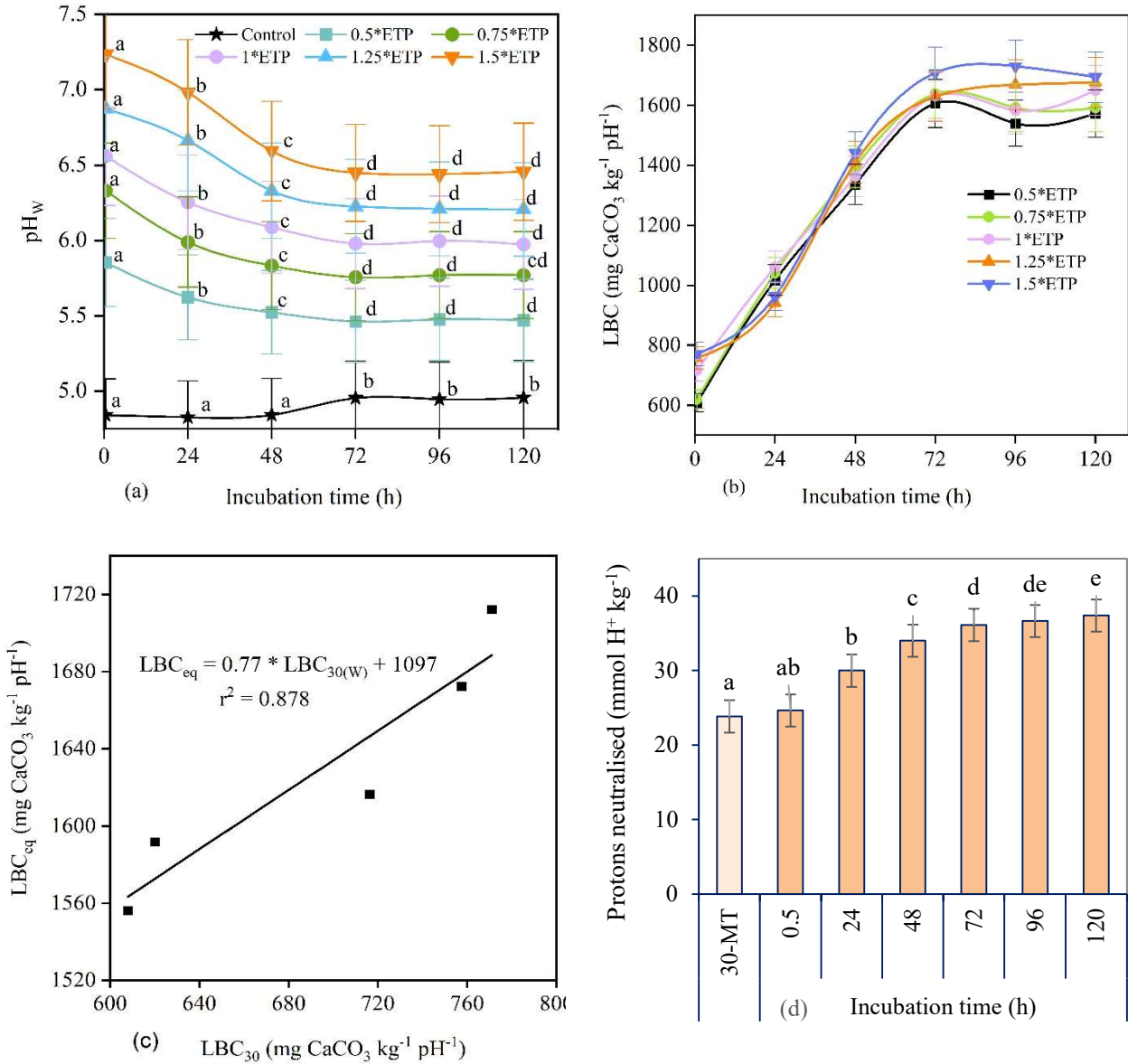
250 The regression lines between soil pH and incremental base additions (expressed as Mg CaCO$_3$ ha^{-1} equivalents, eqn. 3) were linear for titrations in water (r^2 = 0.996) and 0.01 M CaCl$_2$ (r^2 = 0.999) (Fig. 2). The slopes of the regression lines were 0.92 and 0.91 pH ha Mg^{-1} CaCO$_3$ for titrations in water and 0.01 M CaCl$_2$ solution, respectively with no significant difference between them (p = 0.231).

255 **Figure 2.** Fitted linear regression lines for titrations to pH 6.5 in deionised water and pH 6.0 in 0.01 M CaCl$_2$ solution after 30 minutes. In the equations, 'CC' is the CaCO$_3$ (Mg ha^{-1}) consumed during titrations of the acidic soil with 0.022 M Ca(OH)$_2$.

The buffering capacity (pHBC$_{30}$) of the acidic soil was estimated from the inverse slope of the titration curve and was 1.09 Mg CaCO$_3$ ha^{-1} pH^{-1} in water and 1.10 Mg CaCO$_3$ ha^{-1} pH^{-1} in 0.01 M CaCl$_2$, which are equivalent to 14.54 and 14.66 mmol H$^+$ kg^{-1} pH^{-1}, respectively. The pHBC is the amount of H$^+$ that is consumed to raise soil pH by one unit, whereas LBC is the

260 amount of CaCO$_3$ required to raise soil pH by one unit. As a result, pHBC expressed as LBC is more convenient for calculating the amount of lime required to neutralise soil acidity. The average values of LBC$_{30}$ calculated from pHBC$_{30}$ were 727 mg CaCO$_3$ (kg soil)$^{-1}$ pH^{-1} for titration in water and 733 mg CaCO$_3$ (kg soil)$^{-1}$ pH^{-1} for titration in 0.01 M CaCl$_2$. The slightly higher pHBC$_{30}$ in the 0.01 M CaCl$_2$ solution than in water indicates that soil resistance to pH change increases when Ca^{2+} ions in the CaCl$_2$ solution replace Al^{3+} and H$^+$ ions in the soil exchange complex. The pHBC varied slightly across the pH values,

265 indicating that the change in pH due to base additions is not uniform across the range of pH values.

3.2 Equilibrium pH and lime buffer capacity

The pH_W and LBC of soils amended with 0.5, 0.75, 1, 1,25, and 1.5 times the ETP of 0.022 M $Ca(OH)_2$ and incubated for 5 days in a 1:5 soil to deionised water ratio are shown in Figs. 3a and 3b. The pH_W of unamended soil and soils amended with different rates of 0.022 M $Ca(OH)_2$ decreased nonlinearly with incubation time until it reached a relatively constant level. In contrast, compared to its initial pH, pH_W of the unamended soil increased slightly by 0.11 units by 72 h (Fig. 3a). To account for this change, which was not associated with base addition, the pH value for lime-amended treatments was corrected by subtracting the change in pH of the control soil from the observed pH at the respective incubation time.

 Figure 3. (a) pH_W of unamended (control) and amended soils with incubation time, (b) relationship between lime buffer capacity (LBC – mg $CaCO_3$ kg^{-1} pH^{-1}) and incubation time. The LBC was calculated for different rates of 0.022 M $Ca(OH)_2$ using eqn. 4. The amended soils received 0.022 M $Ca(OH)_2$ at rates of 0.5, 0.75, 1, 1.25, and 1.5*ETP where equivalent titration point (ETP) is the volume of 0.022 M $Ca(OH)_2$ consumed during a 30-minute titration in water to raise soil pH_W to 6.5 (i.e. 2.8 mL). (c) relationship between corrected LBC_{eq} and LBC_{30} of soils amended with different rates of 0.022 M $Ca(OH)_2$ converted to $CaCO_3$ equivalent. LBC_{eq}: equilibrium lime buffer capacity (i.e. after 96-120 h), LBC_{30}: lime buffer capacity determined after 30-minute titration in water. (d) proton concentration neutralised by 0.022 M $Ca(OH)_2$ added at a rate of 1*ETP during a 30-minute titration (30-MT) and 5-day incubation in water. 30-MT represents the time required to complete a titration to achieve pH_W 6.5 with 30-minute interval between each base addition, whereas 0.5 h refers to the initial time in the 5-day incubation. Pairwise comparisons using repeated measures ANOVA, different letters indicate significant differences ($p \leq 0.05$) over time.

The soil pH_W taken after 0.5 h for soil treated with 1*ETP was similar to those of the 30-minute titration. However, pH_W values at 0.5, 24, and 48 h incubation were significantly different from each other and the rest of the incubation periods for all rates of $Ca(OH)_2$. In contrast, the pH_W values obtained after 72, 96, and 120 h incubation were similar. This indicates that the acid-base reaction had reached equilibrium after 72 h incubation. Based on this and the trends of regression lines shown in Fig. 3a, as well as the quantity of protons neutralised over time (Fig. 3d), the equilibrium pH was calculated as the average of the adjusted pH values obtained after 96 and 120 h incubation. For example, the equilibrium pH for 1 ETP was 5.99 (i.e. (6+5.97)/2), implying that the 0.022 M $Ca(OH)_2$ equivalent determined during titration in water with 30 minute equilibration time only neutralised 69% of the soil acidity when the target pH_W was 6.5.

Furthermore, LBC was calculated using eqn.4 with incubation time adjusted with respect to the control (unamended) soil, i.e., dividing the observed LBC by the correction factor (CF). The CF was calculated by dividing the difference in pH between a soil amended with 'a' rate of 0.022 M $Ca(OH)_2$ and the control soil 'c' at a specific incubation time 't_i' (e.g., 24, 48, 72 h) by the difference in pH between a soil amended with the same rate of $Ca(OH)_2$ and the control measured at 0.5 h incubation.

The LBC had an increasing non-linear trend with incubation time until equilibrium was reached, and values calculated for 72, 96, and 120 h incubations were not significantly different from each other (Fig. 3b) since LBC_{eq} is obtained when the pH reaches equilibrium. By comparing the LBC values and the trends of regression lines in Fig. 3b, the average LBC values obtained at 96 and 120 h incubations were used as LBC_{eq} for the acidic soil used in this study.

Subsequently, a regression equation was developed between adjusted LBC_{eq} and LBC_{30} (measured after 0.5 h), with different amounts of $Ca(OH)_2$ as covariate (Fig. 3c). The LBC_{eq} for the soils incubated with base in deionised water was then calculated from LBC_{30} value. For example, for LBC_{30} 727 mg $CaCO_3$ kg^{-1} pH^{-1}, the LBC_{eq} was 1657 mg $CaCO_3$ kg^{-1} pH^{-1}. This relationship between LBC_{30} and LBC_{eq} can be used to calculate LBC_{eq} from LBC_{30} data for similar soils.

3.3 Proton concentrations and lime requirements

The cumulative concentrations of protons neutralised by the added base during the 5 days incubation with 1*ETP are shown in Fig. 3d. Thus, 1*ETP incubated for 120 h neutralised 37.39 mmol H$^+$ kg^{-1} soil (eqn. 5) and raised pH_W from 4.84 to 5.99,

thus neutralising 36% more protons than the initial 30-minute titration (30-MT) (Fig. 3d). To raise the soil pH_W to 6.5, 55 mmol H^+ kg^{-1} must be neutralised. This corresponds to the total active acidity because base (OH^-) neutralises not only indigenous H^+ but also H^+ produced by hydrolysis of Al^{3+}. Therefore, 2751 mg $CaCO_3$ kg^{-1} soil were needed to raise pH_W to 6.5. For the lime used in this study (98% effective neutralising value) would be 2807 mg $CaCO_3$ kg^{-1} soil (eqn.6).

3.4 Effect of the alkalinity of organic amendments on soil pH

The titratable alkalinity of OAs ranged from 12 cmol H^+ kg^{-1} for wheat straw to 357 cmol H^+ kg^{-1} for compost (Table 3). Titratable alkalinity of blended poultry litter and biochar was about half of that of compost (Table 3). The titratable alkalinity of OAs was significantly correlated with EC of the amendments ($r^2 = 0.69$**, $p = 0.004$) (Supplementary Table ST2). In addition, alkalinity was moderately correlated with the inherent pH of OAs ($r^2 = 0.60$*, $p = 0.018$). The LEV or $CaCO_3$ content of each OA added at 1.5% (w/w) (i.e. 15 g kg^{-1} soil) (eqn.7) varied with titratable alkalinity of the OA (Table 3). The amounts of OAs required to neutralise soil acidity and achieve pH_W of 6.5, which was calculated based on LEV of the amendments and LBC_{eq} of soil (eqn.8), was inversely related to the alkalinity content of the amendments (Table 3).

Table 3 Titratable alkalinity of organic amendments and the amount of lime combined with the organic amendments added at 1.5% to raise soil pH_W from 4.84 to 6.5. Different letters indicate significant differences ($p \leq 0.05$).

Type of organic amendment	Titratable alkalinity (cmol H^+_{eq} kg^{-1} OA)	LEV_{OA} (g $CaCO_3$ kg^{-1})	Calculated rates of OA (g kg^{-1} soil) to achieve pH_W 6.5	LEV_{OA} (g $CaCO_3$ (15 g OA)$^{-1}$)	Amount of lime (g $CaCO_3$ kg^{-1} soil) required with 1.5% OA to achieve pH_W 6.5
Wheat straw	$11.7^{\pm5.2a}$	$5.8^{\pm2.5}$	$471.1^{\pm7.3}$	$0.09^{\pm0.04a}$	$2.72^{\pm0.04a}$
Faba bean straw	$43.0^{\pm3.6b}$	$21.5^{\pm1.8}$	$127.8^{\pm6.8}$	$0.32^{\pm0.03b}$	$2.49^{\pm0.03b}$
Blended poultry litter	$176.0^{\pm4.6c}$	$88.1^{\pm2.3}$	$31.2^{\pm0.8}$	$1.32^{\pm0.04c}$	$1.49^{\pm0.04c}$
Biochar	$168.7^{\pm1.5c}$	$84.4^{\pm0.8}$	$32.6^{\pm0.3}$	$1.27^{\pm0.02c}$	$1.54^{\pm0.01c}$
Compost	$357.0^{\pm5.5d}$	$178.7^{\pm2.7}$	$15.4^{\pm0.2}$	$2.68^{\pm0.04d}$	$0.13^{\pm0.04d}$

OA: organic amendment, LEV: lime equivalent value. The sum of $CaCO_3$ of OA (column 5) and amount of lime (column 6) is equal to 100% $CaCO_3$ equivalent (2.81 g $CaCO_3$ kg^{-1} soil for the soil used in this study).

Changes in soil pH (ΔpH) caused by all OAs were significantly different at all water contents (60, 100, and 150% FC). Incubation with faba bean and wheat straw at 60% FC increased pH_W of the soil from 4.84 to 6.53 and 6.42, respectively (Fig. 4a). For the less biodegradable or resistant OAs (ROAs) (i.e. blended poultry litter, compost, and biochar), the soil pH at 60% FC was lower than for crop straws (Fig. 4a). Biochar resulted in higher soil pH_W (6.27) than the other ROAs.

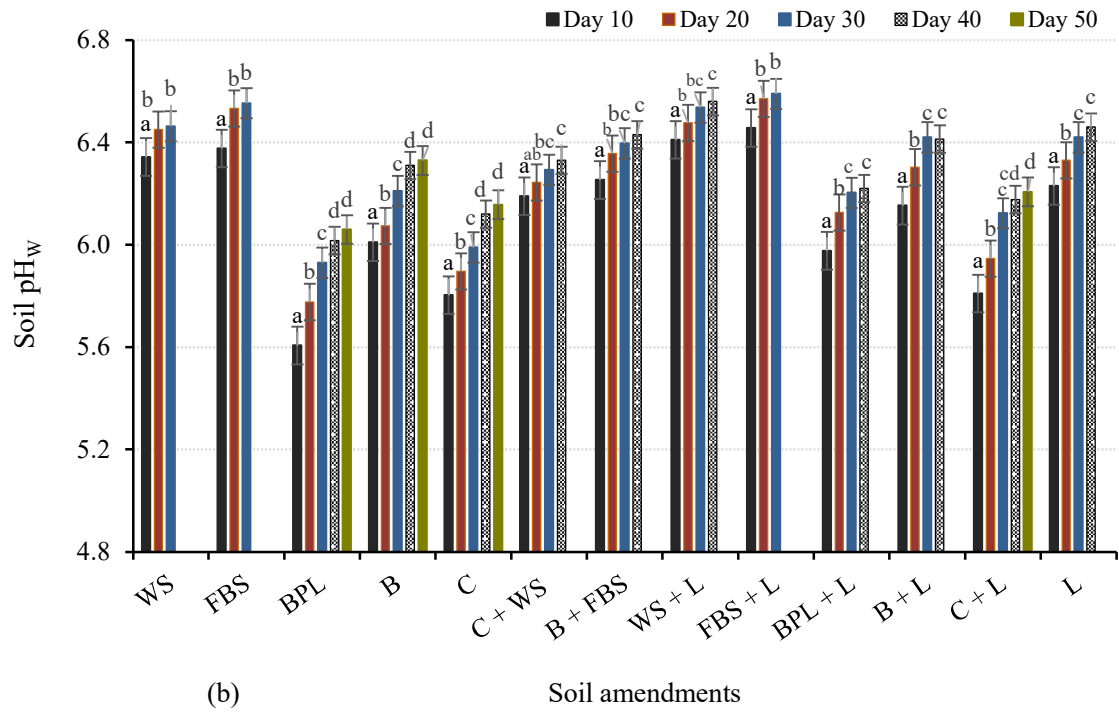
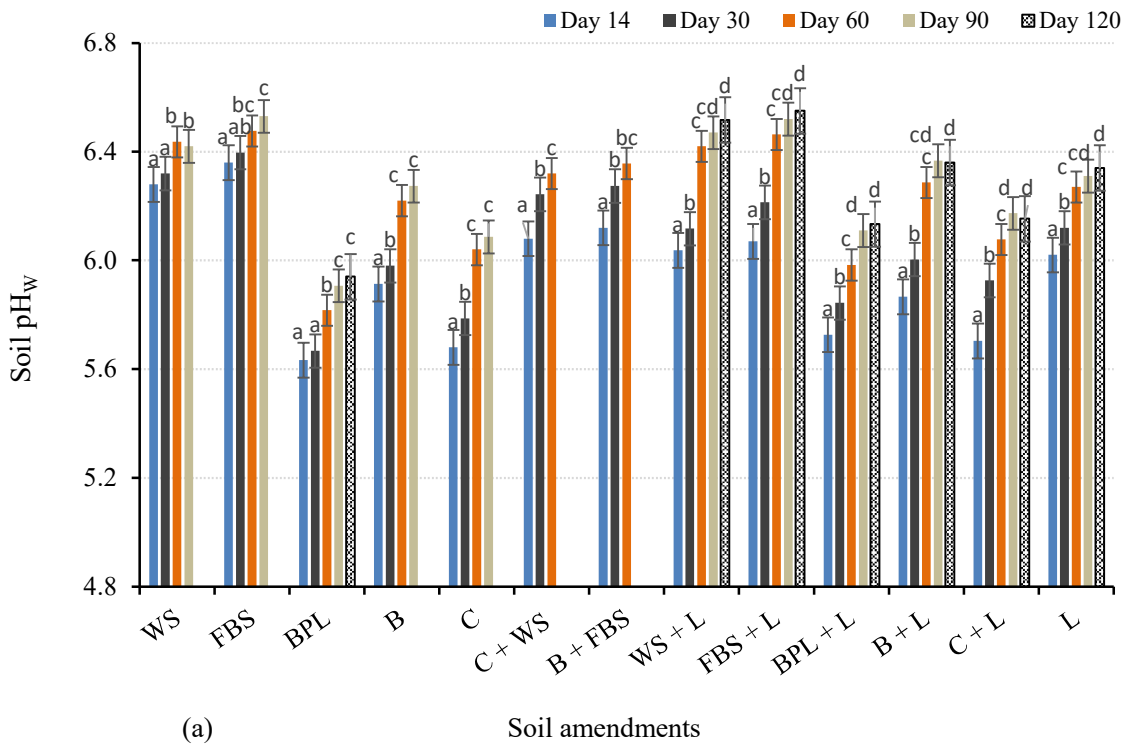
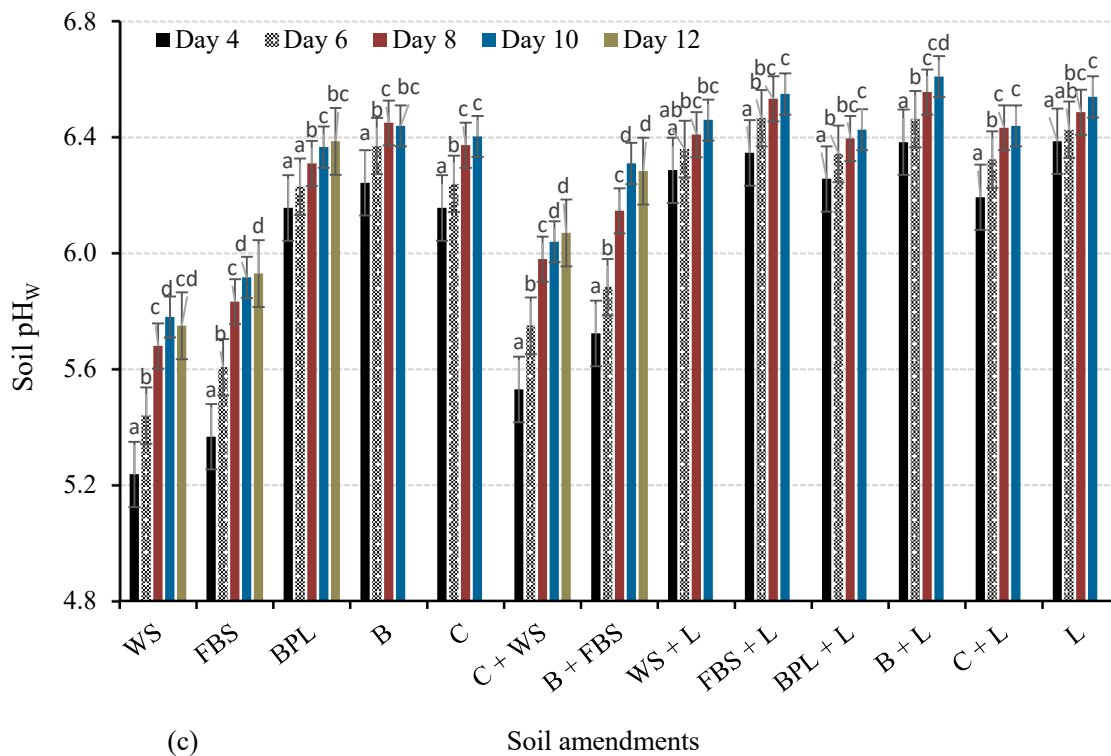
(a)   Soil amendments

(b)   Soil amendments

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Figure 4. Changes in pHw of soils incubated with individual or combinations of different amendments (a) at 60% FC for 120 days, (b) at 100% FC for 50 days, and (c) at 150% FC for 12 days. Different letters indicate significant changes in soil pH with incubation time. The final pH values represent the equilibrium pH of the treatment. WS: wheat straw, FBS: faba bean straw, BPL: blended poultry litter, B: biochar, C: compost, L: lime.

335   At 100% FC, faba bean and wheat straws also resulted in soil pH higher than ROAs, with mean pH_W values of 6.55 and 6.46, respectively (Fig. 4b). At this soil moisture content, ROAs such as biochar and compost increased soil pH_W from 4.84 to 6.33 and 6.16, respectively, thus higher pH values than at 60% FC (Fig. 4a). The rate of change in soil pH over time (ΔpH/time in days) was also rapid and higher than at 60% FC. For instance, the rates of pH change for faba bean straw and biochar-amended soils were 0.22 and 0.13 pH units day^{-1} at 100% FC, whereas 0.11 and 0.07 pH units day^{-1} at 60% FC, respectively.

340   Faba bean straw-amended soil reached equilibrium pH after 30 days of incubation at 100% FC, but after 90 days at 60% FC (Figs. 4a & 4b). The remaining soil amendments had a similar decrease in equilibrium period with increase in soil water content. This could be due to the increased availability of both water and air for biodegradation of OAs at 100% FC, resulting in higher alkalinity generation and neutralisation of acidic soil.

    Soils incubated with ROAs for 12 days at 150% FC (submerged conditions) had a higher soil pH than with faba bean and

345   wheat straws (Fig. 4c). Biochar, compost and blended poultry litter raised soil pH_W to 6.44, 6.40, and 6.39, respectively. On

the other hand, wheat and faba bean straws resulted in lower pH values (5.75 and 5.93) at 150% FC than at the lower water contents. For all soil amendments, the rate of change in soil pH at 150% was greater than at 100 and 60% FC. For example, at 150% FC, faba bean straw and biochar-amended soils reached equilibrium pH after 12 and 10 days (Fig. 4c), and their rates of change in pH were 0.48 and 0.64 pH units day^{-1}, respectively.

350   The magnitude of change in pH or alkalinity production by straws decreased above 100% FC whereas that of lime and less biodegradable OAs linearly increased with soil water content. The correlation between titratable alkalinity of amendments and changes in soil pH increased significantly with soil water content with r^2 values of 0.72**, 0.75**, 0.82** at 60, 100, and 150% FC, respectively (Supplementary Table ST2).

     The alkalinity-based mixture of wheat straw + compost or faba bean straw + biochar aimed to generate 50:50% total
355   alkalinity was calculated as 235.6 g kg^{-1} wheat straw + 7.7 g kg^{-1}compost or 63.9 g kg^{-1} faba bean straw + 16.3 g kg^{-1} biochar. The mixes of resistant and easily biodegradable OAs are based on their alkalinity to provide the required 100% $CaCO_3$ to neutralise soil acidity. The 50:50% alkalinity-based mixture of resistant and easily decomposable OAs produced soil pH values that were intermediate between the individual amendments. At 60% FC, the wheat straw and compost mixture resulted in a soil pH_W of 6.32, which was 0.23 units higher than the pH changes caused by compost but 0.10 units lower than with wheat
360   straw (Fig. 4a). The faba bean straw and biochar mixture increased pH_W to 6.36, resulting in values that were higher than biochar but lower than faba bean straw. A similar trend was observed at 100% FC, but with slightly higher equilibrium pH values of 6.33 (compost – wheat straw mix) and 6.43 (biochar – faba bean straw mix) (Fig. 4b). This shows that the pH changes by alkalinity-based OA mixes can be predicted from pH changes by individual OAs.

4.5 Effect of lime – organic amendment combinations

365   The amount of lime required in the lime-OA combinations was calculated using eqn.9. When the OAs were added at 15 g kg^{-1}, the amount of lime needed to ameliorate acidic soils in combined lime-OAs was inversely related to the titratable alkalinity of the OAs. With increasing alkalinity of the OAs from 5.8 g $CaCO_3$ kg^{-1} (wheat straw) to 178.7 g $CaCO_3$ kg^{-1} (compost), the lime requirement in lime-OA combinations decreased from 2.72 (wheat straw) to 0.09 g $CaCO_3$ kg^{-1} (compost) (Table 3, Fig. 5). Compost addition for example, reduced the lime requirement by 95% compared to lime alone. Wheat straw with very low
370   alkalinity, on the other hand, reduced lime requirement by only 3% compared to lime alone.
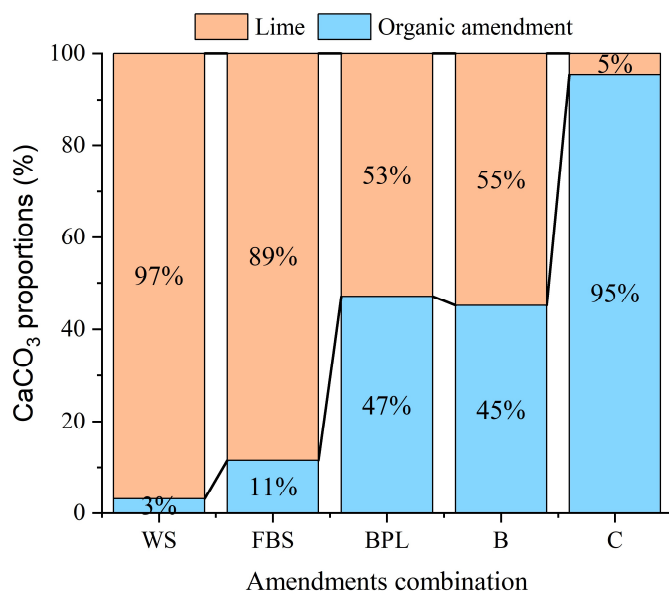
Figure 5. Proportions of alkalinity generated by organic amendments (OAs) added at 15 g kg^{-1} and the lime required to achieve 100% CaCO$_3$ equivalent in combined lime-OA applications. WS: wheat straw, FBS: faba bean straw, BPL: blended poultry litter, B: biochar, C: compost.

Lime added alone increased soil pH by 0.20 units when soil moisture content increased from 60 to 150%. Combinations of lime and OAs resulted in a relatively higher soil pH than OAs alone at all incubation moistures (Figs. 4a, 4b, & 4c). Combined additions of faba bean straw + lime resulted in soil pH$_W$ of 6.55 followed by wheat straw + lime (pH$_W$ of 6.52), after 120 days of incubation at 60% FC (Fig. 4a). At 100% FC, the faba bean + lime and wheat straw + lime combinations generated soil pH$_W$ 6.59 and 6.56, within 30 and 40 days of incubation, respectively (Fig. 4b). At 100% FC, lime combined with biochar, blended poultry litter, and compost increased soil pH$_W$ to 6.41, 6.22, and 6.21, respectively. At 150% FC, the combined lime-ROA treatments increased soil pH$_W$ close to 6.5 (Fig. 4c).

4 Discussion

Our findings show that the laboratory method developed based on titratable alkalinity of OAs and standard soil parameters (pH and LBC) can be used to accurately calculate OA rates and lime-OA combinations for amelioration of acid soils.

4.1 Optimum application rates of soil amendments

The acidity in the soil that had to be neutralised to achieve the target of pH$_W$ 6.5 was estimated by determining an equilibrium LBC. In this study, an equilibrium pH of 5.99 was reached after 96-120 h at the equivalent CaCO$_3$ rate determined during titration, giving an LBC$_{eq}$ of 1657 mg CaCO$_3$ kg^{-1} pH^{-1}. Lime requirement was then calculated by multiplying this LBC$_{eq}$ by

the required change in soil pH. However, it should be noted that the time required to reach equilibrium pH depends on the pHBC of the soil (Jalali and Moradi, 2020; Kissel et al., 2005; Wang et al., 2015). Hence it is important to check this equilibrium time when using this method for different soils. LBC estimates based on 30-minute titration (LBC_{30}) did not represent the total concentrations of soil acidity that has to be neutralised to bring the pH to the desired level, because the acid-base neutralisation reactions do not attain equilibrium in this short time period (Barrow and Cox, 1990; Kissel et al., 2007; Liu et al., 2004). However, we showed a linear regression between LBC_{eq} and LBC_{30} can be used for rapid assessment, once this relationship is developed for a particular soil without the need to conduct a long-term incubation each time. The results showed that the percent by weight (w/w) of OAs needed to achieve a soil pH_W of 6.5 is inversely related to the titratable alkalinity of those amendments. The titratable alkalinity of OAs increased in the order of wheat straw < faba bean straw < biochar < blended poultry litter < compost (Table 3), the optimum application rates of the amendments decreased as wheat straw (47.1%) > faba bean straw (12.8%) > biochar (3.3%) > blended poultry litter (3.1%) > compost (1.5%).

4.2 Effect of different organic amendment properties on soil pH

Alkalinity content has been previously found to be a primary measure of acid neutralizing capacity of OAs (Noble et al., 1996; Noble and Randall, 1998). This was confirmed in the present study by the significant correlation between soil pH and titratable alkalinity of OAs at various soil water contents. Differences in alkalinity content of OAs has previously been found to be due to differences in excess cation content and other physicochemical properties which are influenced by soil, climate, and management practices (Noble and Randall, 1998; Slattery et al., 1991).

Previous studies have also showed that OAs generate alkalinity in the form of hydroxyl (–OH) and carboxyl (–COOH) surface functional groups, carbonates, organic anions (e.g. oxalate, malate), inorganic anions such as sulfate (SO_4^{2-}), phosphate (PO_4^{3-}), silicate (SiO_4^{4-}), and iron hydroxides (FeO–O–) (Cai et al., 2020; Dai et al., 2017; Fidel et al., 2017; Li et al., 2022; Sakala et al., 2004). These groups detoxify H^+, Al^{3+}, and other acid-forming ions through neutralisation and association reactions. Different OAs have different types and amounts of these groups.

In our experiments the magnitude of soil pH changes at the same soil water content was not consistently related to the titratable alkalinity of the OAs added. This suggests other physicochemical properties of OAs (e.g. inherent pH, carbonate content, and C/N ratio) may significantly influence soil pH changes. Biochar, for example, was added at a rate of 100% $CaCO_{3eq}$, which was comparable to the rates for compost and blended poultry litter, but biochar increased soil pH more than these amendments. This might be caused by other properties of biochar such as high inherent pH and solid phase carbonate content following pyrolysis (Mosley et al., 2015) as well as oxygen-rich surface functional groups and a large specific surface area, which leads to high surface adsorption of protons and Al^{3+} (Chintala et al., 2014; Cui et al., 2021).

We found that when applied at similar $CaCO_3$ equivalents, faba bean straw produced higher soil pH than wheat straw at all soil water contents. This could be because faba bean straw has a lower C/N than wheat straw (Table 2) and therefore is decomposed more quickly and results in net ammonification during the mineralization of organic N and a pH increase. Higher

total N concentrations in legumes than in cereals likely lead to a higher N cycling rates and a faster biological decarboxylation (Butterly et al., 2013), resulting in faster and greater soil pH increases.

4.3 Impact of soil water content on alkalinity production by organic amendments

The present study showed that changes in soil pH by wheat and faba bean straws decreased in the order 100 % FC > 60% FC > 150% FC, whereas ROAs soil pH increased with water content as 60% FC < 100% FC < 150% FC. When applied at similar $CaCO_3$ equivalent, wheat and faba bean straws resulted in higher pH values than ROAs at 60% FC. The more rapid soil pH changes with wheat and faba bean straws to pH_W 6.53 and 6.42 compared to the other OAs at 60% FC are likely because crop straws decompose fast because microbial activity is high when oxygen availability is high (Grzyb et al., 2021; Jin et al., 2023), leading to the rapid release of available alkalinity. In contrast, the pH changes with the ROAs were slow at 60% FC. The lignin content of wood chips in blended poultry litter and the resistance of compost and biochar to further decomposition is likely to result in slow decomposition of these amendments, thereby reducing the release of organic anions. In addition, the water content at 60% FC may not be high enough for the dissolution of soluble organic compounds and thus production of alkalinity in ROAs. As a result, soil amended with ROAs reached equilibrium pH after 90-120 days at 60% FC, indicating that ROAs should be added several months before planting to neutralise acidity.

We found that an increase in soil water content from 60% to 100% FC reduced the time required for amended soils to reach equilibrium pH nearly twofold. Thus, soil water content should be considered when deciding on the application time and comparing the effects of amendments on soil pH (e.g. optimal application may be before a significant rainfall event).

At 150% FC, amendment with ROAs resulted in pH_W close to 6.5, likely because there was enough water for the dissolution of soluble organic/inorganic anions and carbonates. Even after the equilibrium pH was reached, undissolved/undecomposed ROA particles were found in the suspension, causing minor deviations in soil pH_W from 6.5. However, it is unlikely that the entire acid-extractable alkalinity of the ROAs is released by water. Adeleke et al. (2017) suggested that organic anions that provide various functional groups (amino, carboxylic, phenolic) for surface adsorption of H^+ and Al^{3+} are weak acids and do not dissolve completely in water. Acid-extractable alkalinity determined by reaction of organic materials with acids (e.g. HCl) over a long time can also lead to increased solubility of inorganic alkali, exposing occluded inorganic alkali and occluded conjugate bases of functional groups (Fidel, 2012), which are not soluble in water. Nevertheless, water-insoluble fractions of acid-extractable alkalinity can increase long-term buffering capacity of acidic soils (Fidel et al., 2017; Yuan et al., 2011).

However, wheat and faba bean straws induced lower pH at 150% FC. The lower pH with wheat and faba bean straws at 150% FC than at the lower water contents may be because undecomposed crop straws were added at high rates which resulted in low oxygen content at this high water content due to low decomposition rates and thus limited production of alkalinity. For example, Chen et al. (2018) showed that anaerobic conditions reduced the rate of straw decomposition by 30% compared with aerobic conditions. Reduction of organic material decomposition is associated with a limited production of organic anions (Cai et al., 2020; Xu et al., 2006b). Furthermore, anaerobic conditions may have favoured the formation of acid-forming products

such as H_2S, and promote the formation of protons from the acidic soil via the hydrolysis of $Al(OH)_3$ and the dissolution of Fe hydroxy oxide clay minerals.

4.4 Effect of mixing soil amendments based on alkalinity production

455 An alkalinity-based mixture of easily decomposable and ROAs produced soil pH changes intermediate between the two amendments. The mixtures led to greater pH increases than ROAs alone, but smaller pH changes than rapidly decomposable alone. Previous studies indicated that the nutrient release and decomposition patterns of a mass-based OAs mixtures resulted in non-additive responses (Gartner and Cardon, 2004; Le and Marschner, 2018). However, there is no information on how alkalinity-based mixtures of OAs affect soil pH. Our findings show that for alkalinity-based OA mixtures, pH changes of the
460 mixes can be predicted from the pH changes of the individual amendments. The pH values can be estimated as the sum of pH changes by individual amendments multiplied by their proportion of alkalinity in the mix. Hence, mixing of OAs based on their alkalinity and C/N ratio could promote an early pH increase due to the rapid production of alkalinity by easily decomposable amendments, as well as a sustained increase later from the resistant amendments. Amendment mixtures may occur where crop residues are left in the field after harvest which are mixed with ROAs to generate the additional alkalinity
465 required to neutralise soil acidity.

The results showed that lime-OA combinations induce higher pH than individual OAs. For instance, the combined lime-ROA treatments increased soil pH_W to about 6.5 at 150% FC, which can likely be attributed to the increased alkalinity production for the reasons explained above. As lime dissolution in acidic soils increases with water content (Anderson et al., 2020; Naorem et al., 2022), combined use of lime and ROAs could be a good option for managing acidic soils forming in
470 coastal areas and wetlands, including acid sulfate soils. Mixtures of lime and wheat and faba bean straws also produced higher pH than lime alone. This could be due to synergetic interactions between lime and OAs, leading to high alkalinity generation. Decomposition of OAs generates CO_2 which leads to dissolution of carbonate; liming on the other hand may increase OA decomposition by increasing the soil pH. However, the relative effectiveness of OAs with lime in ameliorating acidic soil is not simply the sum of their neutralising capacity (Butterly and Tang, 2018), likely due to the adsorption of cations (e.g. Ca^{2+})
475 released from lime on the surfaces of OAs.

5 Conclusion

This study showed that titratable alkalinity of OA and LBC of acidic soils can be used to accurately calculate the amounts of OAs and the lime-OA combination ratio needed to neutralise soil acidity. The alkalinity production potential of soil amendments was significantly affected by soil amendment type and soil water content. Crop straws decomposed faster to
480 release more alkalinity under aerobic than anaerobic conditions. On the other hand, lime and ROAs, such as compost, biochar, and blended poultry litter, lead to higher pH in soils with high water content. Alkalinity-based mixtures of two OAs produced soil pH changes that were intermediate between the individual amendments. Lime-OA combinations on the other had led to

positive interactions, generating more alkalinity than individual amendments. The developed method is novel because it could help to reduce the time and costs associated with conducting field experiments to determine the optimum lime-OA rates for a

485   specific soil. This could be particularly important in situations where lime sources are unavailable or need to be transported long distances. It could also allow flexibility in adjusting the amount of lime or OAs in combined applications based on their availability and cost. The method can be reproducible in different soils as it only requires measurements of standard soil parameters (soil pH and lime buffer capacity, LBC) and titratable alkalinity of OAs. The study also showed that uncertainty in achieving the desired soil pH from amendment calculations based on titratable alkalinity could be due to soil moisture

490   following amendment, fractions of water-insoluble alkalinity, and technical errors in measuring soil LBC_{eq} and OA alkalinity.

Data availability

The data generated in this study are available from the corresponding author upon request.

Author contributions

BI, LMM, and PM conceptualized the study. BI conducted soil sampling, performed the experiments and drafted the

495   manuscript with editorial contributions from the two co-authors.

Competing interests

The authors declare no competing interests.

Acknowledgements

Birhanu Iticha receives a University of Adelaide Research Scholarship and support from the GRDC project on management

500   of subsoil acidity. We would also like to thank Brian Hughes, Ruby Zelda Hume, and Bonnie Armour for assistance in collecting the soil.

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