Response to referee #2

The referee's comments are in italics, our responses in plain font.
This work reported water activity and surface tension of aqueous AS and D-glucose aerosol particles using DKA. The manuscript fits well to the scope of ACP. I recommend it to be published after the following comments have been adequately addressed.

We thank referee #2 for suggestions for improvement that were taken into account upon manuscript revision. Responses to individual comments are given below.

1. I am worried about the novelty. AS and D-glucose chemicals are not new in hygroscopicity study. Growth factor, water activity and surface tension have been reported in many papers, but not cited in the manuscript. I would suggest the authors demonstrate the new findings (maybe the mixtures and high RH) relative to the previous studies.

      We cannot agree with statement that the manuscript does not include data on previously reported "growth factors, water activity, and surface tension of AS and D-glucose". Sections 3.3 and 4.3 contain references to early modelling and experimental studies. These are used extensively in the following sections when comparing DKA-derived a_w and σ for both pure and mixed particles. Overall, 45 literature sources were used. The third and fourth paragraphs of the "Introduction" briefly describes the results of the comparison.

      Please note, in contrast to the generally accepted approach, where the measured growth factors are compared with thermodynamic model values. In this paper, the inverse problem is solved, i.e., the thermodynamic parameters of the Koehler equation are determined from the measured dependence of the aerosol growth factor $g_b(RH)$. This is the essence and novelty of the work. As a result, we do not compare the measured $g_b(RH)$ dependences with literature, but rather the thermodynamic parameters derived from them.

2. Section 4.1: why is the distinct trend of restructuring of AS/Gl particles ($g_{b,H\&D}$) between mass ratio of 4:1 and 1:1, as shown in Figs.1c and 1d.

      The observed differences between AS:Gl particles with mass ratios of 4:1 and 1:1 are due to the different morphology of the dry particles. The 4:1 particles are porous and irregularly shaped, while the 1:1 particles are compact spheres without voids. In the H&D mode, the porous 4:1 particles are restructured by water absorption into compact particles due to the Ostwald-Freundlich effect, while the dry 1:1 particles, already having a compact structure, are not restructured. In the H&D regime, these particles adsorb water on their surface, which is shown by the increase in particle growth ratio with increasing relative humidity (Figure 1c).
      Line 99, new clarifying text is added:
The description of the particle restructuring mechanism is beyond the scope of this work. It will be considered in detail in the next paper.

3. Line 186: any explanation?
4. Line 190: is there any evidence about the phase state description?

      Both remarks refer to water uptake by pure Gl and AS:Gl=1:1 at low RH. Lines 185-205 explain why the observed dependence of the growth factor on particle size contradicts the Kelvin effect. Two characteristic features allow us to interpret the phase state of Gl and AS:Gl=1:1 nanoparticles at low RH as a semi-solid amorphous state (Mikhailov et al. 2009; Koop et al., 2011). The first future is the absence of stepwise deliquescence and efflorescence phase transitions, which would be characteristic of crystalline substances. Secondly, the moisture-induced phase transition was found to occur at a "glass transition relative humidity"

of $RH_g \approx 48\%$ (Fig.3b, d). Below the RH_g, the particles are mainly in the glassy state. Due to the very low molecular diffusivity of glasses, the uptake of water vapor by glassy aerosol particles is limited to surface adsorption, whereas above the RH_g the particles are in a viscous liquid state and absorb water in the particle bulk. Differences in water sorption mechanisms explain why the growth factors of small particles at low RH are higher than for larger particles.

5. Figure 9b: could you explain why the surface tension decrease firstly, and then increase along with increasing solution concentration?

As mentioned in lines 290-304, the decrease in surface tension of AS:Gl = 1:1 aerosol particles is due the fact the NH_4^+ and SO_4^{2-} ions effectively neutralize the bipolarity of the monomeric D-glucose molecules, facilitating their association into less polar aggregates with reduced σ values, which are more readily accommodated at the air-droplet interface. The subsequent increase of σ at high concentrations ($X_s > 0.5$) is the result of particles solidification.

Line 288, new clarifying text is added:

According to Eq. (10) at $X_s \rightarrow 1$, σ approaches the surface tension of a substance in the molten state, σ_s which for AS and Gl is 185 mNm^{-1} (Dutcher et al., 2010) and 150.9 mNm^{-1} (Docoslis et al., 2000), respectively. Thus, at $X_s=1$ the mole fraction weighted value of σ for AS:Gl = 4:1 and 1:1 is 179 and 170 mNm^{-1}, respectively, which agrees reasonably well with DKA-derived σ at high X_s values (Fig. 9a,b).

6. Is it possible to provide the parametrization for water activity and surface tension of different chemicals using DKA? Then the method could be used more widely.

The DKA method is used to obtain a_w and σ from $g_b(RH)$ dependences. These dependencies can be obtained using various experimental methods, including HTDMA. A feature of DKA is that the $g_b(RH)$ dependence must be obtained for several dry particle diameters in the size range below 100 nm, where the Kelvin term has a significant effect on particle hygroscopicity (Fig. 2). The DKA-derived concentration dependences a_w and σ can then be parameterized and used to validate thermodynamic models of nanoparticles. In this paper, some data were parameterized and the resulting coefficients are presented in Tables S1, S2 and S3 in the Supplement.

References

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