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Volatile Oxidation Products and Secondary Organosiloxane Aerosol from D_5 + OH at Varying OH Exposures

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Abstract. Siloxanes are composed of silicon, oxygen, and alkyl groups and are emitted from consumer chemicals. Despite being entirely anthropogenic, siloxanes are being detected in remote regions and are ubiquitous in indoor and urban environments. Decamethylcyclopentasiloxane (D_5) is one of the most common cyclic congeners, and smog chamber and oxidation flow reactor (OFR) experiments have found D_5 + OH to form secondary organosiloxane aerosol (SOSiA). However, there is uncertainty about the reaction products, and the reported SOSiA mass yields (Y_{SOSiA}) appear inconsistent. To quantify small volatile oxidation products (VOP) and to consolidate the Y_{SOSiA} in the literature, we performed experiments using a Potential Aerosol Mass OFR while varying D_5 concentration, humidity, and OH exposure (OH_{exp}). We use a proton transfer reaction time-of-flight mass spectrometer to quantify D_5, HCHO, and HCOOH, and detect other VOP, which we tentatively identify as siloxanols and siloxanyl formates. We determine molar yields of HCHO and HCOOH between 52 – 211 % and 45 – 127 %, respectively. With particle size distributions measured with a scanning mobility particle sizer, we find Y_{SOSiA} to be < 10 % at $OH_{exp} < 1.3 \times 10^{11}$ s cm^{-3} and ~20 % at OH_{exp} corresponding to that of the lifetime of D_5 at atmospheric OH concentrations. We also find that Y_{SOSiA} is dependent on both organic aerosol mass loading and OH_{exp}. We use a kinetic box model of SOSiA formation and aging (aging-VBS model) to reconcile the Y_{SOSiA} values found in this study and the literature. The model uses a volatility basis set (VBS) of the primary oxidation products as well as an aging rate coefficient in the gas phase, $k_{age,gas}$, of 2.17×10^{-11} cm^3 s^{-1}, and an aging rate coefficient in the particle phase, $k_{age,particle}$, which is ten times smaller. The combination of primary VBS and OH-dependent oxidative aging predicts SOSiA formation much better than a standard-VBS parameterization that does not consider aging ($R^2 = 0.970$ vs. 0.847). The need for an ageing-dependent parameterization to accurately model SOSiA formation shows that concepts developed for secondary organic aerosol precursors, which are able to form low-volatile products at low OH_{exp}, do not necessarily apply to D_5 + OH. The resulting yields of HCHO and HCOOH and the parameterization of Y_{SOSiA} may be used in larger scale models to assess the implications of siloxanes on air quality.

Keywords: D_5 siloxane, organic aerosol, proton transfer reaction mass spectrometer, oxidation flow reactor, chemical kinetics

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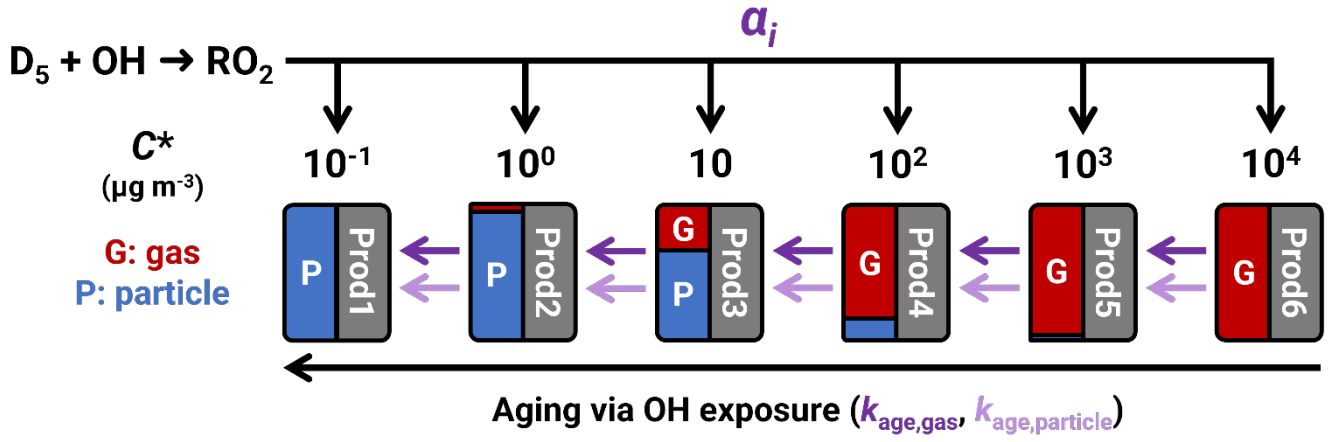

Graphical Abstract: Schematic of the kinetic box model.

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1 Introduction

Organosiloxanes are molecules composed of silicon-oxygen bonds with alkyl groups on the silicons and encompass linear and cyclic species, some of which have vapor pressures on par with volatile organic compounds (VOC). Siloxanes are entirely anthropogenic pollutants (Rücker and Kümmerer, 2015) commonly used in consumer and industrial chemical products (Seltzer et al., 2021a; Gkatzelis et al., 2021) and their emissions are projected to increase in the coming decades (Tansel and Surita, 2017). Decamethylcyclopentasiloxane (D_5, $C_{10}H_{30}O_5Si_5$), where "D" refers to units of $(CH_3)_2SiO$, is a ubiquitous cyclosiloxane in the ambient environment.

Siloxanes can be detected in the indoor environment (Tang et al., 2015; Tran and Kannan, 2015; Arata et al., 2021; Katz et al., 2021; Kaikiti et al., 2022; Wang et al., 2022), near landfills (Schweigkofler and Niessner, 1999), and sewage treatment sites (Lee et al., 2014; Horii et al., 2019). Siloxanes are also found in outdoor urban air (Xiang et al., 2021), and organosilicon compounds have been found in varying amounts in ambient particulates in China (Lu et al., 2019; Cheng et al., 2021; Meng et al., 2021; Song et al., 2022; Xu et al., 2022) and the United States (Milani et al., 2021).

Siloxanes are suspected to be environmentally persistent or "pseudo persistent" (Howard and Muir, 2010; Xiang et al., 2021), but this long-lifetime assessment is disputed (Graiver et al., 2003; Whelan and Kim, 2021). Reaction rate coefficients of D_5 with atmospheric oxidants have been reported, and Atkinson (1991) found D_5 to be effectively unreactive with atmospheric concentrations of O_3 ($k_{D5+O3} < 3 \times 10^{-20}$ cm^3 s^{-1}) and NO_3 radicals ($k_{D5+NO3} < 3 \times 10^{-16}$ cm^3 s^{-1}) at ~298 K. While D_5 is reactive with OH and Cl, Alton and Browne (2020) calculated that the removal of D_5 by Cl radicals would only be a few percent of that by OH radicals at typical ambient oxidant concentrations.

Atkinson (1991), Safron et al. (2015), Xiao et al. (2015), Kim and Xu (2017), and Alton and Browne (2020) have measured k_{D5+OH} at ~298 K to be 1.55×10^{-12}, 2.6×10^{-12}, 2.46×10^{-12}, 1.46×10^{-12}, and 2.1×10^{-12} cm^3 s^{-1}, respectively. These measurements are summarized in Table S1. Xiao et al. (2015) derived k_{D5+OH} computationally as 2.90×10^{-12} cm^3 s^{-1}. In this paper, we use $k_{D5+OH} = 2.0 \times 10^{-12}$ cm^3 s^{-1}, which is a rounded average of the empirically determined rate coefficients. This k_{D5+OH} corresponds to a D_5 atmospheric lifetime of ~4 days via removal by OH, assuming a daily average OH concentration ($[OH]_{avg}$) of 1.5×10^6 cm^{-3}.

D_5 is expected to suppress O_3 formation in urban environments. Carter et al. (1993) performed a series of chamber experiments mimicking urban air conditions and found that D5 siloxane would inhibit ozone formation by suppressing the OH radical. In contrast, formaldehyde (HCHO) is known to contribute to O_3 formation (Derwent et al., 1996). Fu et al. (2020) predicted the formation of HCHO as a product of D_5 + OH at low NO/HO_2 conditions using quantum chemical calculations and kinetics modelling, but an experimental yield of HCHO from D_5 + OH has not been reported. Atkinson (1991) proposed HCHO as a

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product of the siloxane alkoxyl radical (RO) pathway, assuming an analogous mechanism to that of VOC. Sommerlade et al. (1993) suggested that HCHO may arise from siloxane RO decomposition and from ROOH rearrangement in the presence of acids and H_2O. Alton and Browne (2022) predicted HCHO as a product of RO_2 rearrangement in the case of D_3 siloxane. Because HCHO is a secondary product, the O_3 formation potential of D_5 may differ between source and downwind locations.

Formic acid (HCOOH) is a common acid catalyst in the atmosphere (Hazra et al., 2014) and a particle-nucleating species (Yu, 2000). Studies have identified some HCOOH sources in the atmosphere (Millet et al., 2015; Franco et al., 2021), however, HCOOH is suspected to have unidentified anthropogenic sources in the troposphere (Millet et al., 2015; Chen et al., 2021) as some urban sources remain unaccounted for (le Breton et al., 2012; Yuan et al., 2015). Chandramouli and Kamens (2001) proposed that the RO_2 initially formed from D_5 + OH makes a siloxanyl formate (D_4T-OCHO, where "T" refers to CH_3SiO) that reacts with H_2O to a siloxanol (D_4T-OH) and HCOOH. However, we are unaware of experimental HCOOH yields reported for D_5 + OH.

Whelan et al. (2004) used known siloxane chemistry in a partitioning model to assess the atmospheric fate of siloxanes and found that silanols are the predominant oxidation products. These silanols are generally water soluble and either removed from the atmosphere via wet deposition, or undergo a pH-dependent process of hydrolysis, forming smaller and smaller silanols (Whelan et al., 2004). Eventually, the small silanols are converted to SiO_2, H_2O, and CO_2 through photolytic reactions in water or biological processes in soil (Spivack et al., 1997; Stevens, 1998; Graiver et al., 2003).

The intermediate products between D_5 and those small silanols are less well studied. The intermediates may be composed of a variety of alcohols, aldehydes, and hydroperoxides if the reaction mechanisms of D_5 behave in a similar manner to that of organics, with oligomers in the condensed phase (Chen et al., 2023). However, there is evidence that siloxanes do not necessarily follow such reaction mechanisms (Sommerlade et al., 1993; Alton and Browne, 2020, 2022), so there is a need to understand the formation of volatile oxidation products (VOP) and secondary aerosol.

Secondary aerosol mass yield (Y, Eq. 1) is defined as the ratio of produced aerosol mass (Δm(SOSiA)) to reacted precursor mass (Δm(D_5)), which we adopt here for secondary organosiloxane aerosol (SOSiA). Reports about secondary aerosol formation from D_5 siloxane are conflicting, with some experiments reporting much higher Y_{SOSiA} than others. For instance, Wu and Johnston (2017) and Janechek et al. (2019) saw maximum Y_{SOSiA} of 23 % and 50 %, respectively, in their photo-oxidation chamber and oxidation flow reactor (OFR) experiments, albeit at different OH exposures (OH_{exp}). Charan et al. (2022) found a Y_{SOSiA} of 158 % with their OFR at an OH_{exp} of 3.2×10^{12} s cm^{-3}. Avery et al. (2023) reported a wide range of Y_{SOSiA} (2 – 146 %) from their PAM-OFR experiments.

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$$Y_{\mathrm{SOSiA}} = \frac{\Delta m(\mathrm{SOSiA})}{\Delta m(\mathrm{D_5})} \quad (1)$$

In contrast, Charan et al. (2022) reported almost negligible Y_{SOSiA} (< 5 %) from their chamber studies where [OH] was on the order of ~10^6 cm^{-3}, which is closer to [OH] found in ambient conditions (Peng and Jimenez, 2020). Han et al. (2022) conducted OFR experiments and found that Y_{SOSiA} would be 2 % at [OH] of 4.6 × 10^8 cm^{-3} or OH_{exp} of 5.5 × 10^{10} s cm^{-3}. The variation of Y_{SOSiA} reported in the literature suggests that oxidation conditions need to be considered to accurately parameterize Y_{SOSiA}, especially given that D_5 is being considered in air quality models as a part of volatile chemical product inventories (Pennington et al., 2021; Seltzer et al., 2021a, b). In this study, we aim to develop parameterizations that reconcile the reported Y_{SOSiA} for use in such air quality models.

2 Method and Materials

2.1 Experiments

The Aerodyne Research (Billerica, MA, USA) PAM-OFR (Kang et al., 2007) has a volume of 13.3 L and is made of chromated aluminum (Xu and Collins, 2021). We operated the PAM-OFR in "OFR185" mode (Peng and Jimenez, 2020), where 185 nm lamps that also emit 254 nm light (GPH436T5VH, LightSources, Orange, CT, USA) generate OH and O_3 with injected H_2O vapor from a Nafion humidifier (FC-100-80-6MKK, Perma Pure, Lakewood, NJ, USA). There were two of these 185 nm lamps placed across from each other in clear fused quartz sleeves. The 185 nm lamps were wrapped with covers at even intervals to reduce the UV intensity so that 90 % of the lamp surface was covered. We operated the PAM-OFR at residence times (τ_{res}) of 120 and 180 s with flow rates of 6.65 and 4.43 L min^{-1}, respectively. Additional details about the experiment setup are summarized in Fig. S1 and Sect. S1.

We used the D_5 siloxane trace measured from the proton transfer reaction mass spectrometer (PTR-MS) to calculate OH_{exp} with Eq. (2), where k_{D5+OH} = 2.0 × 10^{-12} cm^3 s^{-1}. $[D_5]_0$ and $[D_5]_{final}$ are the D_5 concentrations before and after the exposure to OH.

$$\mathrm{OH_{exp}} = -\frac{1}{k_{\mathrm{D5+OH}}} \times ln\left(\frac{[\mathrm{D_5}]_{\mathrm{final}}}{[\mathrm{D_5}]_0}\right) \quad (2)$$

Prior to experiments, we checked the background particle and D_5 concentrations with the scanning mobility particle sizer (SMPS) and PTR-MS. In all experiments, the background particle number concentrations were < 10 cm^{-3}, and the background $[D_5]$ were below the limit of detection (3σ = 80 ppt). Then, we injected D_5 with a syringe pump while monitoring the PTR-MS, with major ions at *m/z* 371 and *m/z* 355. We performed the experiments with target $[D_5]_0$ of 50, 100, or 200 ppb. With these target $[D_5]_0$, we get external OH reactivities (OHR_{ext}) of 2.5 – 9.8 s^{-1} at 298 K and 1 atm (Peng and Jimenez, 2020).

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When the D_5 trace stabilized near the target $[D_5]_0$, we began the experiment by turning on the UV lamps in the PAM-OFR to either 2.4 or 8.0 V. We waited 30 minutes for the UV lamps to stabilize and for the PAM-OFR walls to equilibrate with gaseous species. The Y_{SOSiA} (Eq. (1)) were calculated using the average SOSiA mass concentration from four SMPS cycles following those 30 minutes. We obtained $\Delta m(D_5)$ as the difference between $[D_5]_0$ and $[D_5]_{final}$. At the end of an experiment, we turned off the UV lamps to check the D_5 trace return.

To clean the PAM-OFR between experiments, we stopped the syringe pump and removed the syringe from the glass bulb while keeping the humid air flow through into the PAM-OFR. We turned on the PAM-OFR UV lamps and connected the outlet directly to the exhaust, until D_5 and particle number concentrations we below the limit of detection. We used Igor Pro 9 (Wavemetrics, Portland, OR, USA) for data post-processing and visualization.

2.2 Instrumentation

2.2.1 PTR-MS

To measure D_5 and VOP, we used a PTR-MS (PTR-TOF 1000, Ionicon Analytik, Innsbruck, Austria) equipped with the extended volatility range (EVR) option (Piel et al., 2021), where the wetted inlet components and the drift tube are passivated with a silicon coating. The PTR-MS also had ion transfer lens between the drift tube and time-of-flight mass spectrometer (Jordan et al., 2009). An internal permeation source (PerMaSCal) emitted a steady stream of 1,3-diiodobenzene into the mass spectrometer for mass calibration scale adjustments. Additional PTR-MS details are in Sect. S1.

To reduce H_2O clusters at high humidities, we operated the PTR-MS at 137 Td (U_{drift} = 600 V, Td = Townsend, 1 Td = 10^{-17} V cm^2) for quantification. The drift tube pressure and temperatures were set to 2.30 mbar and 80 °C. For the reagent ion source, we set the U_s, U_{so}, and the H_2O flow rate to 150 V, 80 V, and 6.00 sccm respectively. The ion source hollow cathode discharge current was set to 5.0 mA. The PTR-MS drift tube was 9.6 cm long, and at 137 Td, the $(H_2O)H^+$ reaction time (Δt) was 94 µs (de Gouw et al., 2003). We calculated the primary reagent ion signal, $(H_2O)H^+$, by multiplying the signal of its isotope, $(H_2{}^{18}O)H^+$, by 500.

We used the PTR-MS data for the quantification of D_5 (*m/z* 371), HCHO (*m/z* 31), and HCOOH (*m/z* 47), where the primary reagent ion counts were normalized to 10^6 counts per second (ncps). For D_5, we used a calibration gas cylinder (Apel-Riemer Environmental, Miami, FL, USA) containing D_5 to calibrate the PTR-MS. We also calculated the normalized measurement sensitivity (ncps ppb^{-1}) of D_5, HCHO, and HCOOH using Eq. (3) adapted from de Gouw and Warneke (2007). $I_{(VOC)H+}$ and $I_{(H2O)H+}$ are the ion counts of the protonated VOC and the reagent ion respectively. Additional details on the mass spectra interpretation and quantification are in Sect. S1.5 and S3.

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$$Sensitivity = \frac{\frac{I_{(VOC)H^+}}{I_{(H_2O)H^+}} \times 10^6}{[VOC]} \quad (3)$$

We tested the instrument sensitivity response with humidity by keeping the species concentrations constant while changing the sample air humidity. The sensitivity of D_5 at *m/z* 371 was not heavily affected by humidity at 137 Td, and we did not correct for humidity in the D_5 quantification (Fig. S5). On the other hand, HCHO and HCOOH sensitivities varied with humidity, and we corrected their sensitivities as detailed in Sect. S3. Prior to experiments, we tuned the micro channel plate (MCP) to prevent signal bias against higher mass ions (Müller et al., 2014). We adjusted the MCP voltage in steps to increase the signal strength at *m/z* 331, a PerMaSCal ion, until the relative signal increase was < 20 %.

2.2.2 Scanning Mobility Particle Sizer

An SMPS (Model 3938, TSI, Shoreview, MN, USA) equipped with an impactor (0.0508 cm) measured the particle mobility diameter size distribution between diameters of 14.3 to 723.4 nm. The SMPS consisted of a Model 3082 Electrostatic Classifier, a Model 3081A Differential Mobility Analyzer (DMA), a Model 3088 Soft X-ray Neutralizer, and a Model 3756 Ultrafine Condensation Particle Counter. We set the SMPS sheath flow at 3.0 L min^{-1} and the aerosol flow rate at 0.3 L min^{-1}, and the DMA voltage ranged from 10.6 to 9921.4 V. The SMPS scanned for 150 s, followed by a 5 s retrace and 10 s purge while recording on a 3 min cycle. We referred to the manufacturer's recommendations when deciding the above SMPS settings (TSI Inc., 2012), and a sample particle size distribution from experiment 12 (Table 1) is shown in Fig. S4.

For the Y_{SOSiA} calculations, we converted the SMPS integrated particle volumes into mass using a SOSiA mass density (ρ_{SOSiA}) of 1.07 g cm^{-3} for all experiments. We obtained this ρ_{SOSiA} from PAM-OFR experiments separate from the ones described here, where we weighed the masses of SOSiA collected on filters and particle volumes with the SMPS. Additional details on ρ_{SOSiA} are available in Sect. S2.

2.3 Volatility Distribution Parameterization

Janechek et al. (2019) and Charan et al. (2022) fitted their Y_{SOSiA} data to the Odum two-product model (Odum et al., 1996) and we follow the same methodology for comparison with the literature (Sect. S4). Similarly, we fit the standard volatility basis set (VBS) parameters α (Donahue et al., 2006) in Eq. (4) to the measured Δm(SOSiA) using the measured $\Delta m(D_5)$, where α_i is the product mass yield for volatility bin *i*.

$$\Delta m(\text{SOSiA}) = \Delta m(D_5) \times \sum_{i=1}^{n} \frac{\alpha_i}{1+\frac{C_i^*}{C_{OA}}} \quad (4)$$

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In the experiments, the organosiloxane aerosol mass loading (C_{OA}) was equivalent to the SOSiA mass concentrations. As the produced aerosol mass in the experiments ranged from 3.7 to ~1000 μg m^{-3}, we use six logarithmically spaced effective saturation mass concentration (C^*) bins ranging from 0.1 to 10000 μg m^{-3} at 298 K to cover the low and high-volatility products. For reference, D_5 liquid has a vapor pressure of 20.4 Pa at 298 K or C* = 3.05 × 10^6 μg m^{-3} (Lei et al., 2010).

As the experiments were performed for a range of OH_{exp}, the products between experiments may have varied due to multigenerational aging (Zhao et al., 2015). To account for aging and parameterize Y_{SOSiA} as a function of OH exposure, we also analyse the yield data using a kinetic box model with four chemical reactions (R1–R3) written in MATLAB (MathWorks, Natick, MA, USA).

$$D_5 + OH \rightarrow \sum \alpha_i \times prod(i) \quad (k_{D5+OH} = 2.0\times10^{-12}\ cm^3\ s^{-1}) \quad (R1)$$

$$(1\text{-}f_i) \times prod(i) + OH \rightarrow (1\text{-}f_i) \times prod(i\text{-}1) \quad (k_{age,gas},\ i = 2, \ldots, 6) \quad (R2)$$

$$f_i \times prod(i) + OH \rightarrow f_i \times prod(i\text{-}1) \quad (0.1 \times k_{age,gas},\ i = 2, \ldots, 6) \quad (R3)$$

Eq. (R1) describes the initial oxidation of D_5 and formation of RO_2, which immediately forms products of varying volatility (Eq. (R2)). Here, prod(i) refers to the sum of products (gas + particle) in volatility bin i, which are formed with a molar branching ratio α_i. We assume that prod(i) have the same molecular weights (g mol^{-1}) as D_5, and so the α_i are equivalent to the product mass yields at $OH_{exp} \rightarrow 0$. In the model, a fraction f_i of each oxidation product partitions instantaneously from the gas phase to the particle phase according to absorptive partitioning theory (Donahue et al., 2006) (Eq. 5).

$$f_i = \left(\frac{1}{1+\frac{C_i^*}{C_{OA}}} \right) \quad (5)$$

Eqs. (R2) and (R3) describe how OH_{exp} causes volatility to decrease (Robinson et al., 2007). This decrease in volatility via "bin-hopping" (Sommers et al., 2022) occurs at a rate proportional to the chemical aging rate coefficient for gaseous species ($k_{age,gas}$, $cm^3\ s^{-1}$), with the oxidation of particle-phase products being ten times slower than that of the gas. Note that we assume that products in the lowest-volatility bin cannot be removed from that bin and that the highest-volatility bin does not receive product with aging (i = 2, …, 6). The [OH] are set by dividing the experimental OH_{exp} from Eq. 2 by the PAM-OFR residence times.

We use $k_{age,gas}$ and $k_{age,particle}$ as aggregate chemical aging rate coefficients, not specific to any species or volatility bin. Studies on chamber experiments (Robinson et al., 2007) and ambient measurements (Sommers et al., 2022) applied chemical aging only to the gas phase as heterogeneous aging is relatively slower. However, studies have found that the high oxidant

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concentrations in OFRs would appreciably oxidize OA within experiment timescales (Kessler et al., 2012; Kroll et al., 2015). To accommodate OH uptake to the bulk phase, we follow the approach used by Zhao et al. (2015) and assume that the effective particle-phase aging rate coefficient ($k_{age,particle}$) is equivalent to 10 % of the gas-phase aging rate coefficient ($k_{age,gas}$). The timescales and atmospheric relevance of heterogeneous oxidation in OFRs are areas of ongoing research (Zhao et al., 2019; Peng and Jimenez, 2020), but for now we opt to fit a single chemical aging rate coefficient to reduce dimensionality. We use the Monte Carlo genetic algorithm (Berkemeier et al., 2017) to fit $k_{age,gas}$ and the six coefficients α_i.

3 Results and Discussion

3.1 Volatile Organic Products (VOP)

3.1.1 Siloxanol and Formate Ester Trends

Fig. 1 shows the PTR-MS mass spectra for experiment 12 (Table 1), where $[D_5]_0$ and OH_{exp} were high. The PTR-MS signals before and after D_5 is oxidized are displayed relative to the protonated D_5 ion at *m/z* 371 to identify changes more easily in the mass spectra. Using the mass spectra and species reported by Alton and Browne (2022), we attribute the indicated ions in Fig. 1 to siloxanol (D_4T-OH), siloxanediol (D_3T_2-$(OH)_2$), siloxanyl formate (D_4T-OCHO), and siloxanolyl formate (D_3T_2-OH-OCHO). Here, "D" and "T" refers to units of $(CH_3)_2SiO$ and CH_3SiO respectively. The multifunctional VOP are reported to arise from multiple steps of oxidation (Alton and Browne, 2022).

D_5 siloxane loses a methyl group during the PTR, which forms a large signal at *m/z* 355. The isotopologues of the -CH_3 fragment of D_5 overlap with fragments of VOP, which complicates the VOP identification. To separate the signal of the VOP, we use the ratios of the D_5 signal and its -CH_3 signals prior to oxidation. The red and pink shaded areas in the inset of Fig. 1 refer to the enhancement in signal over that of the -CH_3 fragment of D_5, which we attribute to the -OH fragments of D_4T-OH and D_3T_2-$(OH)_2$, respectively. We use the masses of the -OH fragments of the siloxanols as large alcohols dissociate during the PTR (Brown et al., 2010).

As we did not have calibration standards to quantify these VOP, we calculate the relative molar yields of the VOP to that of protonated D_5 siloxane at *m/z* 371 to study the trends of siloxane VOP (Fig. 2). The y-axes in Fig. 2 are the relative molar yields (ncps/ncps), which refers to the change in signal attributed to a VOP over that of *m/z* 371. $\Delta m371$ refers to the change in the signal at *m/z* 371 before and after OH oxidation. In the right-side panels for each VOP in Fig. 2, the relative molar yield of VOP decreases with increasing OH_{exp} (x-axes). This decrease in VOP signal is consistent with these gaseous products undergoing further oxidation or increased gas-particle partitioning due to higher C_{OA} at higher OH_{exp}.

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In the left-side panels for each VOP in Fig. 2, the relative signals of the VOP (y-axes) decreases with increasing OH_{exp} (color scale). Then, assuming [OH] is consistent throughout the PAM-OFR, that D_5 + OH is the rate-limiting step in VOP formation, and that removal via gas-particle partitioning is negligible, we can consider a simplified D_5 + OH chemical mechanism, (R4) and (R5).

$$D_5 + OH \rightarrow \sum \gamma_i VOP_i \qquad (k_{D5+OH} = 2.0 \times 10^{-12}\ cm^3\ s^{-1}) \qquad (R4)$$

$$VOP_i + OH \rightarrow \qquad (k_{VOPi+OH}) \qquad (R5)$$

In Eq. (R4), γ_i is the relative molar yield of a given VOP_i found by extrapolating $\Delta m(VOP_i)/\Delta m371$ (y-axes in Fig. 2) to $OH_{exp} \rightarrow 0$. With ordinary differential equations (ODE) from these reactions (Eq. (7) and (8)) and experimental inputs, we fit γ_i and the VOP_i + OH rate coefficient ($k_{VOPi+OH}$, $cm^3\ s^{-1}$). The fits are shown as black lines in the right-side panels of each VOP in Fig. 2.

$$\frac{d[D_5]}{dt} = -k_{D5+OH}[D_5][OH] \qquad (7)$$

$$\frac{d[VOP_i]}{dt} = \gamma_i k_{D5+OH}[D_5][OH] - k_{VOPi+OH}[VOP_i][OH] \qquad (8)$$

The fitted $k_{VOPi+OH}$ for each VOP are on the order of ~10^{-12} $cm^3\ s^{-1}$ (Table S7), but faster than k_{D5+OH}, which suggests that these VOP have atmospheric lifetimes shorter than that of D_5. Alton and Browne (2022) have estimated these VOP to be volatile with quantitative structure activity relationship models. However, there are uncertainties in those models, and the VOP may have lower saturation mass concentrations than expected. Moreover, the chemical mechanism might be more complex than the one outlined with the simple reactions (R4) and (R5). Consequently, we present these $k_{VOPi+OH}$ as estimates for secondary chemistry in this simplified reaction scheme, and future work using quantitative measurements should improve the calculated lifetimes of these intermediate D_5 + OH products in the atmosphere.

3.1.2 Formaldehyde (HCHO) Yields

As shown in Table S8 and Fig. 3, the experimental molar yields of HCHO (Y_{HCHO}, $\Delta HCHO/\Delta D_5$ in ppb/ppb) exceed 100 % at low OH_{exp} and decrease with higher OH_{exp}. We attribute the decreasing Y_{HCHO} with increasing OH_{exp} to HCHO removal by OH in the PAM-OFR. HCHO has a lifetime of 0.91 days at $[OH]_{avg} = 1.5 \times 10^6$ cm^{-3} (Atkinson et al., 2006) or 78 s at [OH] = 1.5×10^9 cm^{-3}. In such high [OH] conditions, some HCHO is oxidized while being produced, which is consistent with the decreasing Y_{HCHO} with increasing OH_{exp} (Fig. 3a1). Thus, we use the ODE from Eqs. (7) and (8) and a fixed $k_{HCHO+OH} = 8.5 \times 10^{-12}$ $cm^3\ s^{-1}$ at 298 K to obtain the molar yield of HCHO as $OH_{exp} \rightarrow 0$, which we denote as γ_{HCHO}.

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We fit γ_{HCHO} to be 269 % (black line in Fig. 3a2), assuming a constant [OH] in the PAM-OFR, that HCHO is rapidly formed from D_5 + OH, and that HCHO removal via partitioning or reactive uptake is negligible. This γ_{HCHO} is consistent with the modeled yields of those for VOC used by Millet et al. (2006), who used γ_{HCHO} from chemical models ranging from 60 – 230 % for a variety of VOC. Thus, D_5 has a comparable γ_{HCHO} to that of isoprene or aromatic VOC.

Fu et al. (2020) proposed a mechanism for D_3 siloxane, where high Y_{HCHO} is produced under low NO/HO_2 conditions. In that mechanism, RO_2 rearrangement and RO H-shift rate coefficients become progressively faster as the D_3 siloxane backbone is oxidized, and HCHO is produced at each rearrangement step. The γ_{HCHO} exceeding 100 % in these D_5 experiments is consistent with HCHO production over multiple rapid oxidation steps. The results we report suggests that a similar HCHO production mechanism exists for D_5.

Mao et al. (2009) found that models under-predicted tropospheric HCHO during their aircraft campaign studying Asian pollution outflows into the Pacific ocean. This discrepancy between the measurements and calculations was pronounced near the surface and up to 2 km. They proposed that there is some missing OH reactivity, and that the unaccounted species would be reactive with OH and yield HCHO when oxidized. Based on the D_5 experiments present here, the inclusion of siloxane species may reduce the HCHO formation gap; Coggon et al. (2021) already noted that including volatile chemical products in their model would increase HCHO production.

The large formation of HCHO may entail that D_5 siloxane could contribute to O_3 formation, albeit indirectly. We were unable to observe O_3 enhancement due to the high concentrations of O_3 produced from the PAM-OFR internal chemistry itself and the lack of NO_x. Given that k_{D5+OH} is relatively slow compared to that of other common anthropogenic VOC, we suspect that the oxidation of D_5 will occur downwind of urban sources in low-NO_x conditions or in cases of air stagnation. Whether D_5 has a net positive or negative effect on O_3 formation in these VOC/NO_x scenarios needs to be assessed with models. To get a rough estimate of O_3 production, we consider a case where 20 ppt of D_5 react with OH to form 40 ppt of HCHO, which also fully react. This D_5 concentration is within the range reported by Coggon et al. (2018) in ambient urban air. The molar maximum incremental reactivity (MIR) of HCHO under high-NO_x conditions is ~20 % (Carter et al., 1995), which makes HCHO a prominent precursor for tropospheric O_3. By multiplying the MIR with the HCHO reacted with OH, we can estimate an O_3 formation potential of 8 ppt from D_5 in urban air.

3.1.3 Formic Acid (HCOOH) Yields

We find molar yields of HCOOH (Y_{HCOOH}, $\Delta HCOOH/\Delta D_5$ ppb/ppb) between 45 – 127 %, as shown in Fig. 3b, although a trend with OH_{exp} is not obvious (Fig. 3b).We assume HCOOH loss via OH oxidation to be minor given the rate coefficient of $k_{HCOOH+OH} = 4.5 \times 10^{-13}$ cm^3 s^{-1} at 298 K (Atkinson et al., 2006), which corresponds to 17 days of OH_{exp} at $[OH]_{avg} = 1.5 \times 10^6$ cm^{-3} or an OH-oxidation lifetime of 440 s in our highest OH_{exp} experiment. In addition to D_4T-OCHO hydrolysis, HCOOH

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may have been produced by heterogeneous reactions of HCHO at the surface of the SOSiA or the OFR walls in these humid experiments. In the atmosphere, HCOOH is presumed to form heterogeneously from HCHO and methanediol ($HOCH_2OH$) in the presence of wet particles (Franco et al., 2021).

The Y_{HCOOH} from D_5 + OH we report are higher than the values from isoprene + OH (Link et al., 2020) or monoterpene + OH reported by Friedman and Farmer (2018), who quantified the Y_{HCOOH} of 7 monoterpenes at varying OH_{exp} without NO_x. The range of Y_{HCOOH} from these references is shown as shaded areas in Fig. 3b2. The Y_{HCOOH} from D_5 is on par with the humid isoprene ozonolysis cases reported by Link et al. (2020). Friedman and Farmer (2018) also used a PAM-OFR, but with 254 nm UV lamps in dry conditions (~1 % RH), and Link et al. (2020) used a reaction chamber, which limits a direct comparison with our results. Nevertheless, Friedman and Farmer (2018) found Y_{HCOOH} of 0.64 – 8.5 % at $OH_{exp} = 2.0 \times 10^{11}$ s cm^{-3}. Aside from the different precursor VOC and mechanism, Friedman and Farmer (2018) may have encountered less heterogenous production of HCOOH due to the dry OFR conditions. Our laboratory findings suggest that D_5 siloxane should be considered as an atmospheric HCOOH source.

3.2 SOSiA Mass Yields

3.2.1 Volatility Basis Set Parameterization

The Odum two-product model does not reconcile the Y_{SOSiA} in the literature in the high C_{OA} range (Sect. S4), so we apply a VBS model. Fig. 4a shows the fitted aerosol mass yield curve (black line) using a standard-VBS model (Eq. (4)), but the Y_{SOSiA} (y-axis) appears to depend on both C_{OA} (x-axis) and OH_{exp} (color scale). To address whether accounting for the varying OH_{exp} in these experiments would improve the VBS model outputs, we fit the produced SOSiA mass using a standard-VBS model (Eq. (4)) and a kinetic model with VBS and chemical aging rate coefficients ("aging-VBS model", Eqs. (R1) – (R3)) based on OH_{exp} and $[D5]_0$ (Table 1). We fit $k_{age,gas}$ in Eq. (R2) to be 2.17×10^{-11} cm^3 s^{-1}. The fitted VBS parameters are summarized in Table S11.

In both the standard and aging-VBS model fits (blue and red, respectively in Fig. 4b), ~95 % of the D_5 + OH product mass is in the gas phase at a C_{OA} of 10 µg m^{-3}. The high fraction of gaseous products is consistent with low Y_{SOSiA} in the lower OH_{exp} experiments, whereas additional oxidation in the higher OH_{exp} experiments leads to a shift towards products that partition into the particle phase, thus increasing Y_{SOSiA}. Secondary organic aerosol (SOA) often exhibits a maximum yield as a function of OH_{exp}, after which the yield decreases due to fragmentation becoming dominant at high OH_{exp} (Isaacman-VanWertz et al., 2018). We do not find such a maximum in the range of OH_{exp} studied, which suggests that an even higher Y_{SOSiA} could have been found at higher OH_{exp}. Moreover, SOSiA is reported to be non-hygroscopic compared to SOA (Janechek et al., 2019), and we do not see an obvious relationship between the experiment humidity conditions and aerosol formation.

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Figs. 4c and 4d show comparisons of the standard and aging-VBS with experimental SOSiA mass concentrations and Y_{SOSiA}. We see an improvement in the R^2 with the aging-VBS over the standard-VBS model, suggesting that incorporating OH_{exp} into the yield parameterization improves model outcomes. Fig. 5 illustrates compares the standard-VBS model with the aging VBS for a range of OH_{exp}, showing that product volatility gradually decreases with increasing OH_{exp} in the ageing VBS model. The high volatility of the initial products is consistent with the lack of the rapid formation of low-volatile species, like highly oxygenated molecules, known to form SOA (Isaacman-VanWertz et al., 2018).

3.2.2 Reconciling Literature Y_{SOSiA}

To address the variation in the literature Y_{SOSiA} and to generate parameters for air quality models, we fit the parameters in the aging-VBS model with all available data in the literature and those from our experiments. Given that the literature used differing ρ_{SOSiA} to calculate Y_{SOSiA} from SMPS data, we adjust the Y_{SOSiA} and C_{OA} reported in the literature to that of the ρ_{SOSiA} used here (ρ_{SOSiA} = 1.07 g cm^{-3}). Similarly, we re-calculate the OH_{exp} in the literature using Eq. (2) and the $[D_5]_0$ and $[D_5]_{final}$ values.

Fig. 6 shows experimental values (markers) and model outputs (contours) of Y_{SOSiA} (panels a1 and a2) and SOSiA mass concentrations (panels b1 and b2) as a function of $[D_5]_0$ and OH_{exp}. Figs. 6a1 and 6b1 are generated using the aging-VBS model fit using only data from experiments presented in this study, while Figs. 6a2 and 6b2 show a fit including data from the literature. The aging-VBS model is able to capture the increasing Y_{SOSiA} with increasing $[D_5]_0$ and OH_{exp}. At a given $[D_5]_0$, Y_{SOSiA} and the SOSiA mass concentration increase with higher OH_{exp}. Fig. 6a2 shows that the relatively high Y_{SOSiA} (> 50 %) is feasible at $OH_{exp} > 10^{12}$ s cm^3. Moreover, the aging-VBS model predicts that Y_{SOSiA} is almost negligible (< 5 %) under atmospheric concentrations of D_5 and OH_{exp}.

Fig. S8 shows that the aging-VBS model used here leads to a much higher correlation between modelled and experimental values for SOSiA mass concentration compared to the same analysis with a standard-VBS model (R^2 = 0.956 vs. R^2 = 0.745). The better correlation suggests that the volatility distribution evolves with OH_{exp} and that chemical aging should be considered when evaluating the volatility distribution of SOSiA from D_5 + OH.

We note that, in reality, bulk-phase chemistry is more complex than logarithmic shifts in volatility with OH_{exp} and not fully captured in the above aging-VBS parameterization. For example, Wu and Johnston (2017), Avery et al. (2023), and Chen et al. (2023) characterized D_5 + OH SOSiA with mass spectrometry and found spectra indicative of oligomers. The formation of oligomers may reduce the bulk volatility by more than one bin and change the gas-particle equilibrium timescales (Berkemeier et al., 2020). Here, we incorporated $k_{age,gas}$ and a simple "bin-hopping" approach to illustrate that a change in the volatility distribution with OH_{exp} can adequately reconcile the Y_{SOSiA} variation in the literature. Future work with more sophisticated chemical models should close that gap further.

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4 Conclusions and Atmospheric Implications

With a PAM-OFR, PTR-MS, and SMPS, we studied the formation of VOP and SOSiA under various OH_{exp} conditions. Using a simplified VOP oxidation scheme (Eqs. (R5) and (R6)), we find that the VOP of tentatively identified siloxanols and formate esters have shorter OH-oxidation lifetimes than their precursor D_5 (Table S7). In addition, we find the mass yield of HCHO of D_5 comparable to that of isoprene or aromatics (Millet et al., 2006), suggesting that D_5 siloxane is a potential O_3-contributing species in downwind scenarios. We find the mass yield of HCOOH ranging from 45 – 127 %, which suggests that D_5 + OH is a source of atmospheric HCOOH.

An aging-VBS model incorporating OH_{exp} and chemical aging adequately describes gas-particle partitioning at atmospheric OH_{exp} and C_{OA}. Based on these experiments, low-NO_x Y_{SOSiA} should be < 10 % under commonly observed atmospheric OH_{exp} < 5×10^{11} s cm^{-3} (Fig. 6a1). The first-generation products of D_5 + OH are likely volatile, but their volatility decreases with increasing OH_{exp} (Fig. 5). This shift in volatility suggests that further oxidation of secondary products would reduce the volatility enough to form SOSiA. Unlike α-pinene (Isaacman-VanWertz et al., 2018) or other precursors for secondary organic aerosol (SOA), D_5 + OH does not appear to produce low-volatile species within a single oxidation step. Instead, additional OH_{exp} is needed to form aerosol, which suggests that multiple oxidation steps lead to gradual decrease of product volatility. Hence, concepts that can be successfully applied to SOA formation may not accurately capture SOSiA formation, for which models must consider chemical aging. In the atmosphere, SOSiA from D_5 + OH may be easier to detect downwind of urban sources due to the higher OH_{exp} and dilution/removal of competing OH-reactive species.

Based on KinSim calculations (Sect. S5), we expect that the RO_2 fate is dominated by RO_2 + HO_2 and RO_2 + OH, which is consistent with the calculations performed by Avery et al. (2023). However, we note that the reaction rate coefficients of RO_2 and its subsequent products are uncertain for D_5, and we cannot directly address the atmospheric relevance of these calculated RO_2 fates at this time. To improve Y_{SOSiA} parameterizations for the atmosphere, there is a need to study the impact NO_x has on siloxane RO_2 chemistry, given that siloxanes are likely emitted from urban sources where [NO_x] is high. In such scenarios, RO_2 + NO_x is likely an important fate (Peng et al., 2019; Newland et al., 2021). Han et al. (2022) found that the addition of N_2O into their OFR would reduce Y_{SOSiA}, although the cause is unclear. However, Charan et al. (2022) did not find Y_{SOSiA} to change with NO_x in their chamber experiments, which is consistent with rapid RO formation across RO_2 fates. Quantifying secondary species across RO_2 fates and identifying their subsequent oxidation reactions may also be useful to adapt the D_5 oxidation mechanism into chemical kinetics models.

The high [OH] used in OFRs may induce faster radical reactions and dimerization near the particle surface (Zhao et al., 2019), which affects particle composition and equilibrium timescales. Dimers and oligomers have been found in SOSiA (Wu and Johnston, 2017; Avery et al., 2023; Chen et al., 2023), and how oligomerization in the D_5 + OH SOSiA system evolves the

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volatility distribution and particle properties is currently not considered in the aging-VBS model. Moreover, high degrees of oxidation should lead to fragmentation and increasing volatility (Isaacman-VanWertz et al., 2018), which is also not considered in the aging-VBS model. Hence, multiphase modeling to evaluate SOSiA chemistry and translate experimental findings to atmospheric conditions remains a direction for future research.

Appendix A Abbreviations

C_{OA}: organic aerosol mass loading
C*: effective saturation mass concentration
D_5: decamethylcyclopentasiloxane
EVR: extended volatility range
ID: inner diameter of tubing
I_{254}, I_{185}: flux of 254 and 185 nm photons
OA: organic aerosol
OD: outer diameter of tubing
OFR: oxidation flow reactor
OH: hydroxyl radical
$[OH]_{avg}$: 24 hour average daily hydroxyl radical concentration
OH_{exp}: hydroxyl radical exposure
OHR_{ext}: external hydroxyl radical reactivity
O_3: ozone
ncps: normalized counts per second
NO_x: nitric oxide and nitrogen dioxide
PAM: potential aerosol mass
PTR: proton transfer reaction
PTR-MS: proton transfer reaction mass spectrometer
RH: relative humidity
RO: alkoxyl radical
RO_2: peroxyl radical
SMPS: scanning mobility particle sizer
SOA: secondary organic aerosol
SOSiA: secondary organosiloxane aerosol
UV: ultraviolet radiation
VBS: volatility basis set

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VOP: volatile oxidation products

Y_{HCHO}: formaldehyde molar yield from D_5

Y_{HCOOH}: formic acid molar yield from D_5

Y_{SOSiA}: SOSiA mass yield from D_5

γ: molar yields extrapolated to when $OH_{exp} \rightarrow 0$

ρ_{SOSiA}: SOSiA aerosol mass density

τ_{res}: residence time

Data Availability

Summary data are available in the supplementary. Additional data will be provided upon reasonable request.

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Author Contribution

HGK, YC, JJ, and YP conducted the experiments. YP performed the offline calibrations of OH exposure on the PAM-OFR. HGK analysed the data. HGK and TB developed the kinetic model. HGK, HK and TB wrote the manuscript with contributions from all co-authors. HK supervised the project.

Competing Interests

TB is a member of the editorial board of Atmospheric Chemistry and Physics, but the peer-review process was guided by an independent editor. The authors declare that they have no other personal nor financial conflicts of interest. Instruments and products used in the research are listed for reference and not as endorsements.

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Table 1. Summary of SOSiA mass yields (Y_{SOSiA}) with aerosol sampling line corrections assuming ρ_{SOSiA} = 1.07 g cm^{-3} for all experiments. [H_2O] is the molar mixing ratio of H_2O in air. For C_{OA} and [D_5], the errors are the standard deviation of the data points averaged, while for Y_{SOSiA}, they are calculated with error propagation. For reference, at 25 °C and 1 atm, 1 ppb of D_5 is ~15 µg m^{-3} and one day equivalent of OH_{exp} is ~1.3 × 10^{11} s cm^{-3} at a daily $[OH]_{avg}$ of 1.5 × 10^6 cm^{-3}.

Experiment	Y_{SOSiA} (%)	[H_2O] (%)	C_{OA} (µg m^{-3})	OH_{exp} (s cm^{-3})	[OH] (cm^{-3})	$[D_5]_0$ (ppb)	$1 - [D_5]_{final}/[D_5]_0$
1	5.9 ± 0.9	0.892	10.5 ± 0.7	1.73 × 10^{11}	9.59 × 10^8	43.4 ± 1.3	0.292
2	4.9 ± 0.6	0.828	19.0 ± 0.6	1.90 × 10^{11}	1.06 × 10^9	85.7 ± 2.5	0.316
3	3.3 ± 0.6	0.742	17.7 ± 0.5	1.26 × 10^{11}	6.99 × 10^8	165.8 ± 4.5	0.222
4	19.5 ± 1.5	1.95	75.2 ± 1.9	4.66 × 10^{11}	2.59 × 10^9	44.0 ± 1.7	0.606
5	29.3 ± 2.7	2.06	179.2 ± 3.1	3.80 × 10^{11}	2.11 × 10^9	78.3 ± 3.2	0.532
6	26.5 ± 1.8	2.09	286.2 ± 7.1	3.12 × 10^{11}	1.73 × 10^9	157.8 ± 3.6	0.464
7	8.6 ± 0.5	0.733	36.8 ± 1.3	5.76 × 10^{11}	3.20 × 10^9	43.8 ± 1.3	0.684
8	18.6 ± 1.7	0.736	118.6 ± 5.6	4.00 × 10^{11}	2.22 × 10^9	78.9 ± 3.2	0.550
9	21.8 ± 1.1	0.797	304.5 ± 2.8	4.19 × 10^{11}	2.33 × 10^9	166.8 ± 4.1	0.567
10	39.8 ± 2.2	1.93	212.9 ± 8.1	9.01 × 10^{11}	5.00 × 10^9	43.8 ± 1.4	0.835
11	47.4 ± 1.9	2.08	420.2 ± 3.0	7.78 × 10^{11}	4.32 × 10^9	76.5 ± 2.2	0.789
12	54.0 ± 2.4	2.15	965.7 ± 25	7.39 × 10^{11}	4.10 × 10^9	156.9 ± 3.9	0.772
13	4.7 ± 1.7	0.712	3.9 ± 0.3	8.70 × 10^{10}	7.25 × 10^8	37.9 ± 1.6	0.160
14	1.9 ± 0.4	0.718	4.1 ± 0.3	1.09 × 10^{11}	9.10 × 10^8	80.8 ± 2.3	0.196
15	1.1 ± 0.3	0.704	3.7 ± 0.7	8.29 × 10^{10}	6.91 × 10^8	162.8 ± 4.9	0.153

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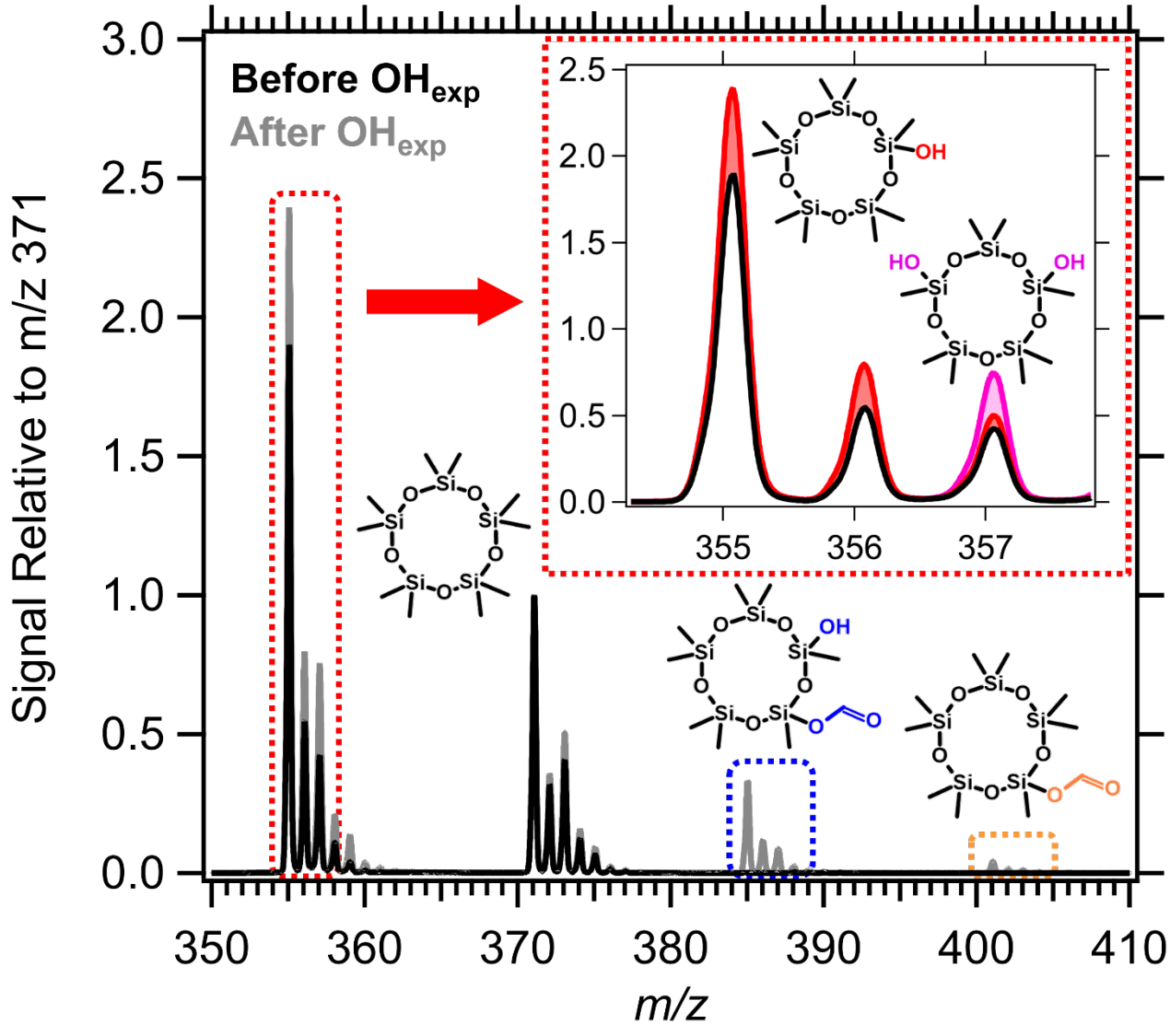

867

868 **Figure 1. Example PTR-MS mass spectra from experiment 12 and proposed VOP ions.** The signal intensities, before (black) and after
869 (grey) oxidation, are each normalized to the signal intensity of the D_5 ion at *m/z* 371, which is set to 1. The multifunctional species (blue,
870 pink) are expected to be formed through multiple steps of OH-oxidation. The red and pink areas in the inset each refer to the enhancement
871 in signal attributed to D_4T-OH and D_3T_2-$(OH)_2$ over that of the -CH_3 fragment of D_5 and isotope signals, respectively.

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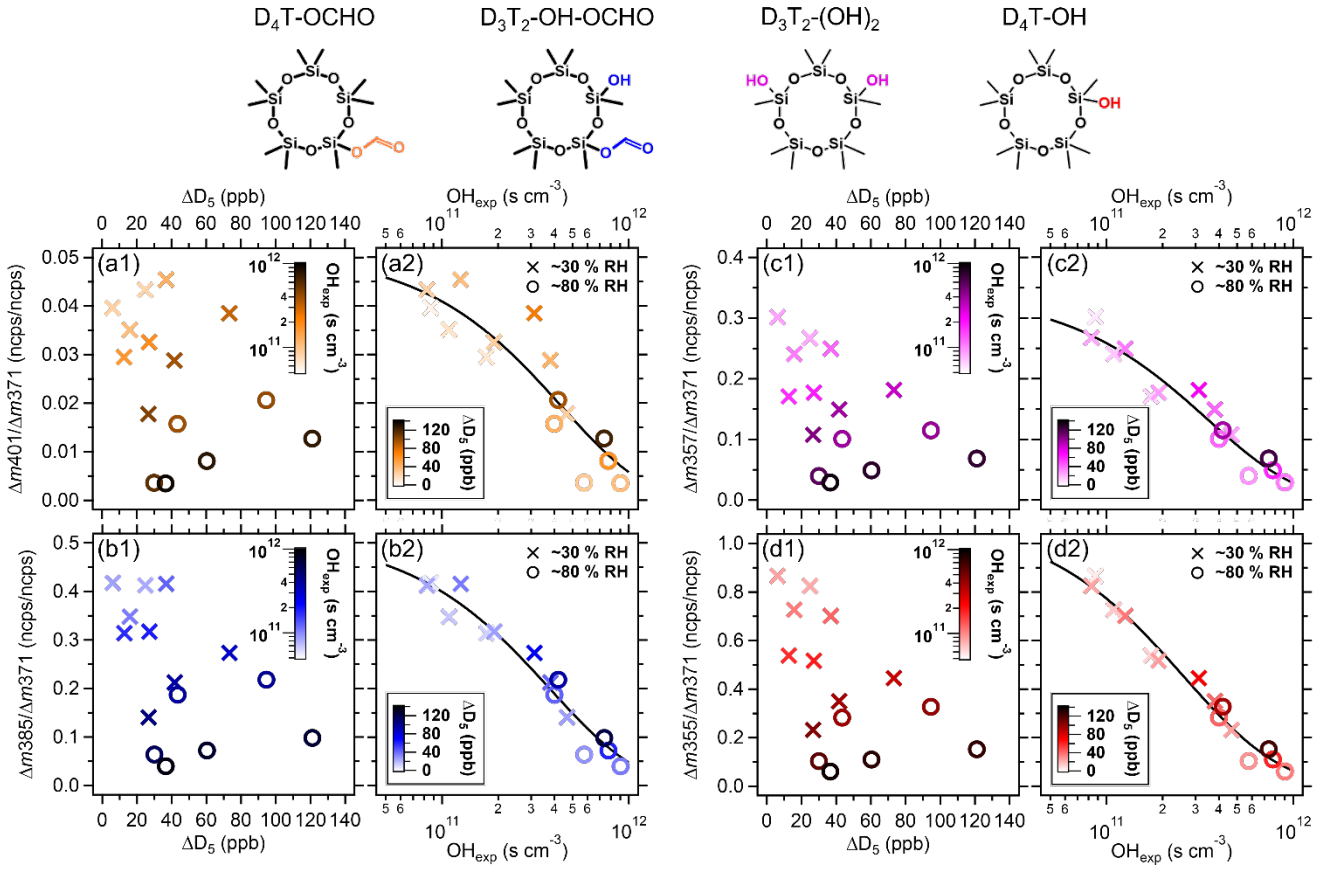

Figure 2. Relative molar yields of selected VOP. Molar yields as a function of OH_{exp} and D_5 consumed in experiments for (a1, a2) D_4T-OCHO, (b1, b2) D_3T_2-OH-OCHO, (c1, c2) D_3T_2-$(OH)_2$, and (d1, d2) D_4T-OH. We did not have a calibration for the suspected VOP, so the y-axes are relative molar yields (ncps/ncps) calculated with the change in signal attributed to each VOP and that of D_5 at *m/z* 371. The relative molar yields decrease with OH_{exp}, which is used to fit their OH-oxidation rate coefficients and γ_i (black lines).

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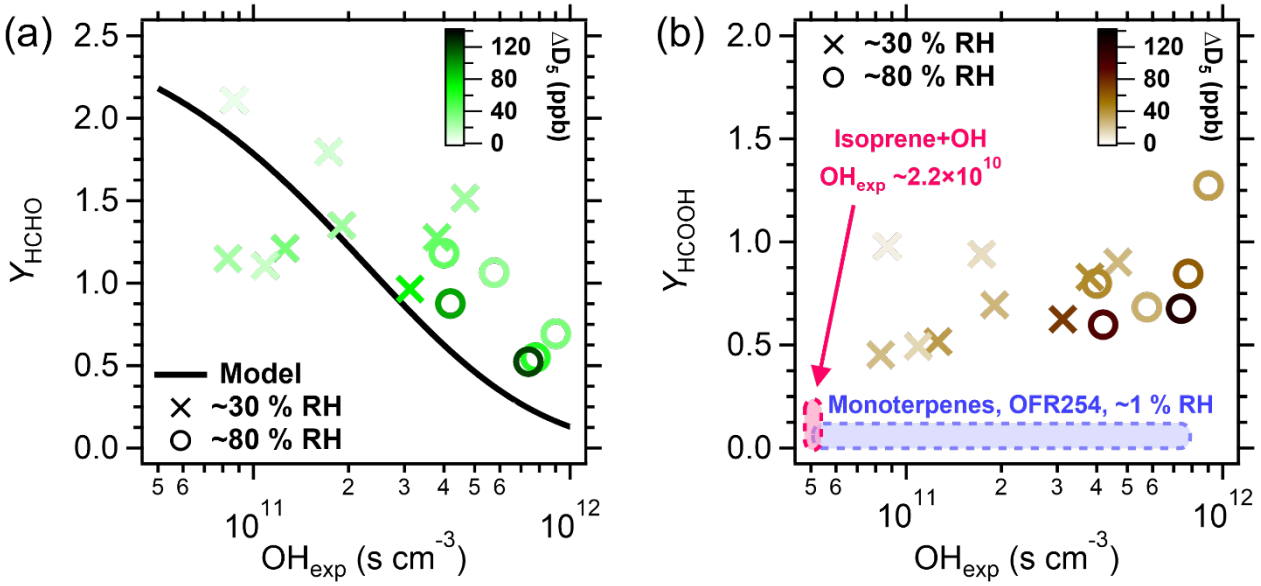

Figure 3. Experimental molar yields of selected VOP: (a) HCHO and (b) HCOOH as functions of OH_{exp}. The blue shaded area in (b) is the range of Y_{HCOOH} (< 10 %) measured by Friedman and Farmer (2018) with monoterpenes under low RH and low NO_x conditions. The pink shaded area refers to Y_{HCOOH} from isoprene + OH chamber experiments (Link et al., 2020) at lower OH_{exp}.

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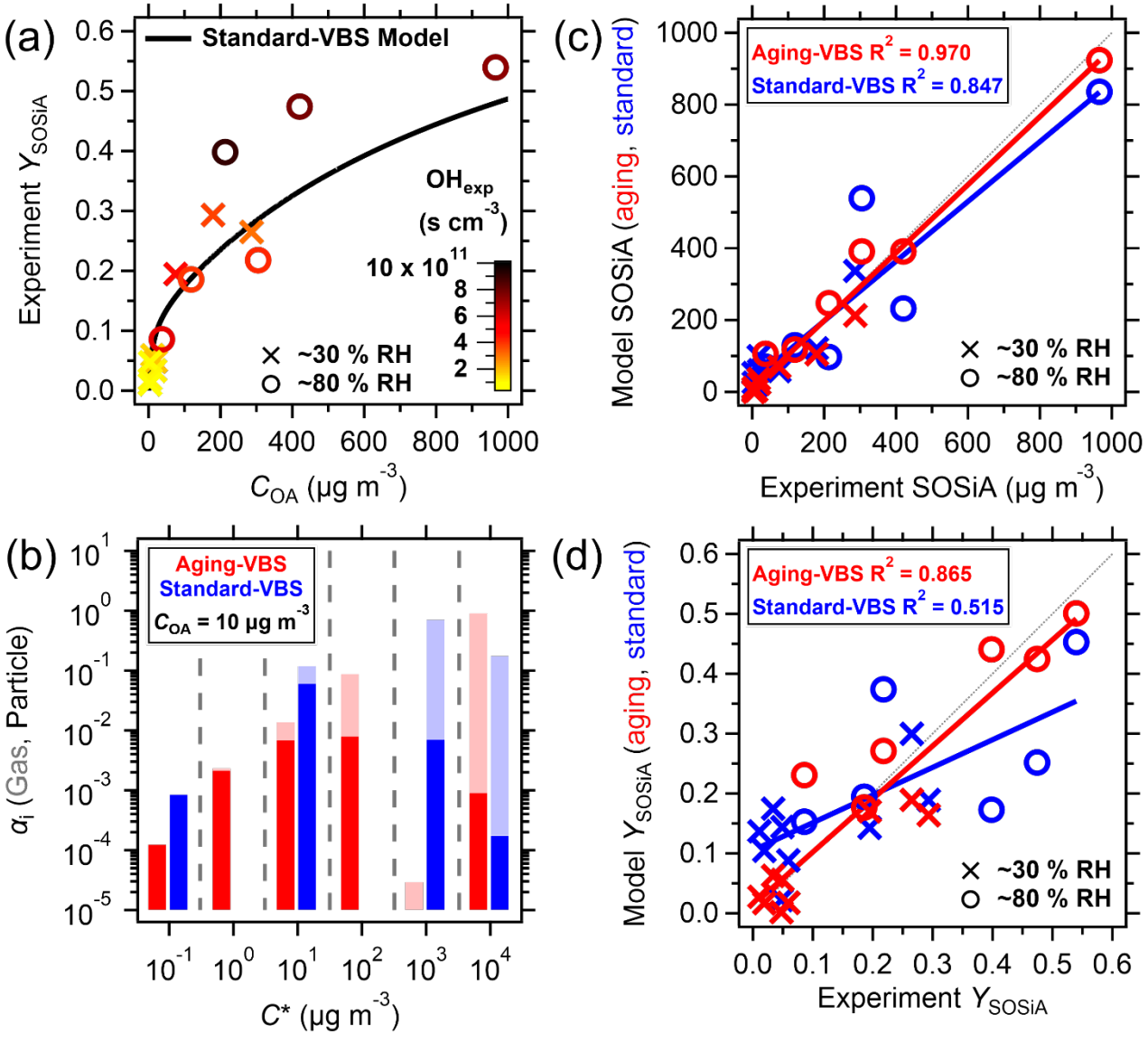

Figure 4. Application of standard-VBS and aging-VBS models to experimental data. (a) Y_{SOSiA} as a function of C_{OA}, where the Y_{SOSiA} appears to be correlated with OH_{exp}. (b) VBS product mass yields for each volatility bin. For the aging-VBS, the values are those of the first-generation products. (c) Comparison of SOSiA mass concentrations from the aging-VBS and standard-VBS models against measurements. (d) Comparison of Y_{SOSiA} from the aging-VBS and standard-VBS models against measurements, where the aging-VBS model has a higher R^2.

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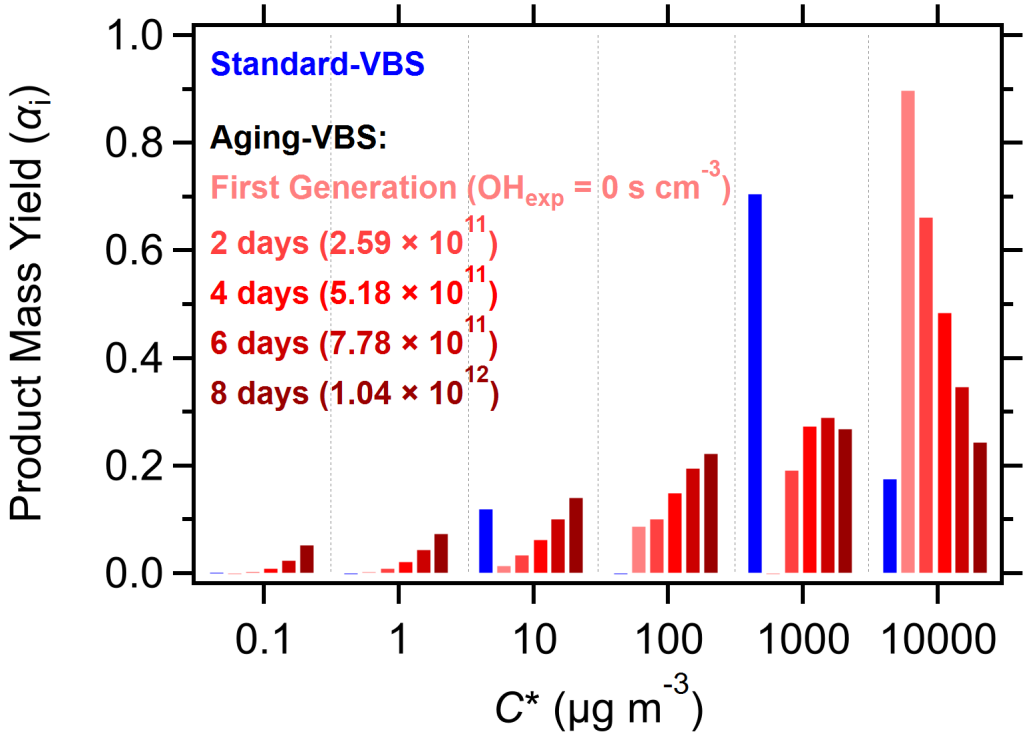

Figure 5. Evolution of the volatility distribution with OH_{exp}. The standard-VBS model parameterization (blue bars) is dominated by the C* = 1 000 μg m^{-3} volatility bin. In the aging-VBS model, the first-generation volatility distribution is dominated by the highest volatility bin (C* = 10 000 μg m^{-3}) but decreases with increasing OH_{exp} (red bars).

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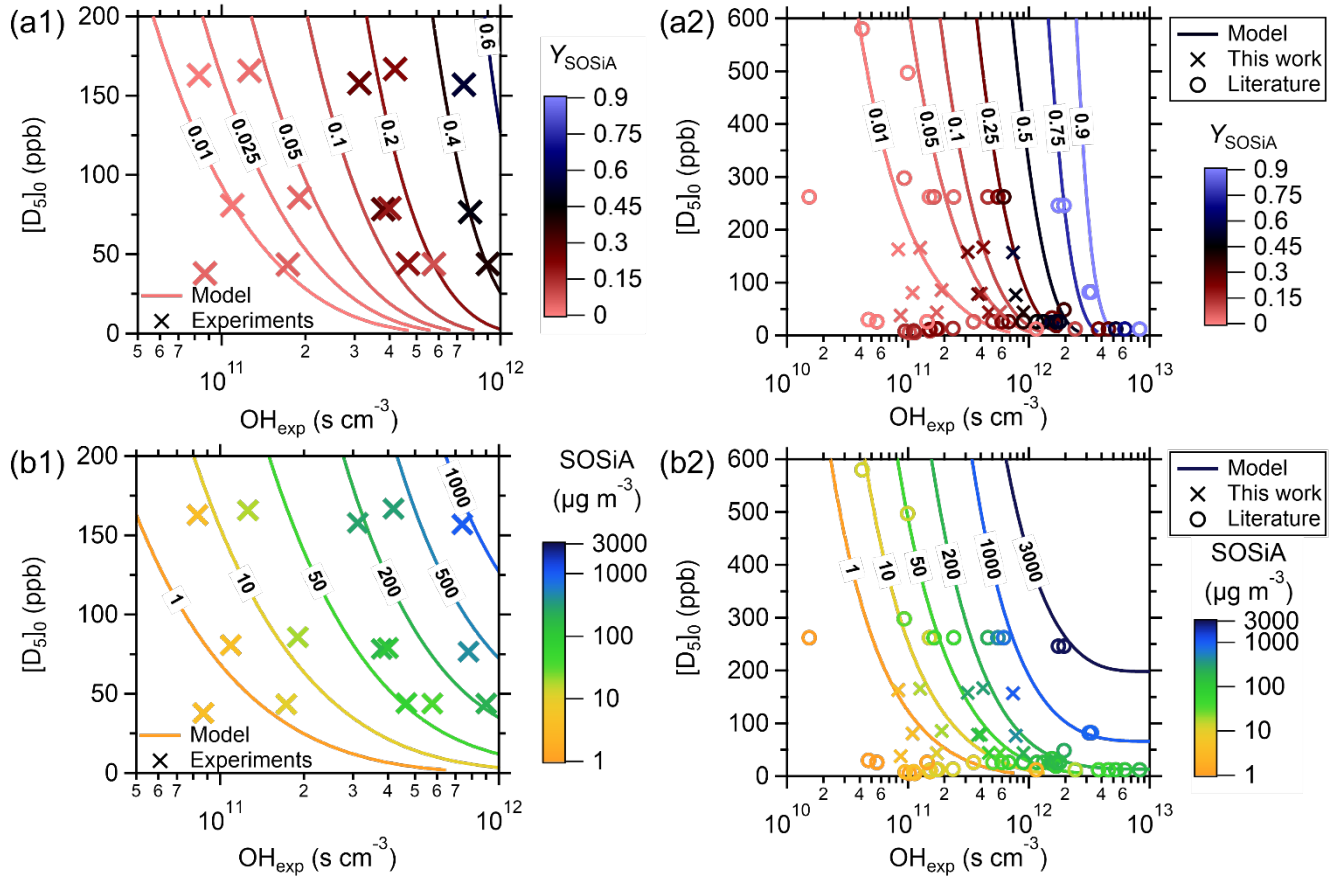

Figure 6. Comparison of experiments, model results, and literature values. (a) Y_{SOSiA} and (b) SOSiA mass concentrations as a function of $[D_5]_0$ and OH_{exp}. The aging VBS-model is fit using experimental data from (1) this study and (2) including those in the literature. SOSiA formation generally increases with $[D_5]_0$ and OH_{exp}. The aging-VBS can capture the broad range of Y_{SOSiA} reported in the literature.